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Translational and internal energy distributions of methyl and hydroxyl radicals produced by 157 nm photodissociation of amorphous solid methanol

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Methanol is typically observed within water-rich interstellar ices and is a source of interstellar organic species. Following the 157 nm photoexcitation of solid methanol at 90 K, desorbed CH3(υ=0) and OH(υ=0,1) radicals have been observed in situ, near the solid surface, using resonance-enhanced multiphoton ionization (REMPI) detection methods. Time-of-flight and rotationally resolved REMPI spectra of the desorbed species were measured, and the respective fragment internal energy and kinetic energy distributions were obtained. Photoproduction mechanisms for CH3 and OH radicals from solid methanol are discussed. The formation of O(1D) and H2O was investigated, but the yield of these species was found to be negligible. CH3 products arising following the photoexcitation of water-methanol mixed ice showed similar kinetic and internal energy distributions to those from neat methanol ice. © 2009 American Institute of Physics.

1. INTRODUCTION

Methanol is the smallest closed shell alcohol and has long served as a model for the dissociation of organic molecules. There is a comparably large body of literature on the vacuum ultraviolet (VUV) photolysis of methanol in the gas phase.1–6 The results of the gas phase methanol photolysis studies at 157 nm showed that the relative contribution of the atomic hydrogen elimination process is larger than that of the molecular hydrogen elimination process, i.e., the ratio is 1:0.2.1,2 C–O bond cleavage [reaction (1)] is another primary photoprocess following excitation of the 21A−X1A transition in the 151–163 nm region, where it competes with the atomic and molecular hydrogen elimination processes,

\[
\text{CH}_3\text{OH} + h\nu \rightarrow \text{CH}_3 + \text{OH}.
\]

(1)
The relative contribution of the C–O bond cleavage channel in the gas phase photolysis at 157 nm has yet to be determined.2,3 Reaction (1) was also observed following the photoexcitation of CH3OH in various rare-gas matrices at 4 K and the photodissociation threshold energy for reaction (1) identified at λ ∼ 175 nm.7 Since methanol is known to be a component of the icy mantle of interstellar grains, comets, and other solar system bodies,8–11 the effects of irradiating pure methanol ice by electrons, ions, and UV photons have each received much experimental study.12–16

Compared to the gas phase, however, the photochemical processes occurring in condensed phase samples of CH3OH are more complicated because of the many possible secondary reactions on/in solid CH3OH.15,16 UV absorption of methanol in the condensed phase is observed at wavelengths shorter than the ∼185 nm threshold.17 Gerakines et al. reported VUV (λ > 110 nm) photolysis studies of solid CH3OH at 10 K. Hydrogen (H2), carbon monoxide (CO), carbon dioxide (CO2), formyl radical (HCO), formaldehyde (H2CO), methane (CH4), and methyl formate (CH3OCHO) were detected as products by in situ infrared (IR) absorption spectroscopy, but CH3 and OH radicals were not.16 CH4 production from VUV photolysis of solid methanol was reported to scale linearly with irradiation time.15,16 Unimolecular photodissociation,

\[
\text{CH}_3\text{OH} + h\nu \rightarrow \text{CH}_3(X^1A) + \text{O}(1D),
\]

(2)
is thus one plausible CH4 formation mechanism. As Gerakines et al. noted, however, further research is necessary since gas phase photolysis studies have failed to reveal any evidence of reaction (2).2,3,16

To obtain a more complete mechanism understanding of amorphous solid methanol (ASM) photolysis—including possible secondary reactions on/in ASM—it would be helpful to determine the kinetic and internal energy distributions of the photoproducts generated from the primary photodissociation of ASM. Hama et al. previously used sensitive resonance-enhanced multiphoton ionization (REMPI) methods to determine the translational and internal energy distributions of, and thus formation mechanisms for, photodesorbed H atoms and H2 molecules following VUV photolysis of ASM.18

In this paper, we investigate the mechanisms, the dynamics of production and the possible reactions of CH3 and OH...
products following 157 nm photodissociation of ASM at 90 K, using REMPI methods to determine their respective translational energy distributions and internal state population distributions. Possible secondary reactions on/in ASM are discussed in light of the obtained results. An attempt was also made to detect O($^1D$) and O($^3P$) atoms and H2O products.

II. EXPERIMENTAL

A. Apparatus and preparation of ice films

The apparatus has been described before, and only details relevant to the present experiments are given here.19 As shown in the schematic diagram of the experimental arrangement (Fig. 1), photodesorbed products were ionized at the focal point of the REMPI probe laser light, which was separated from the ice surface by a vertical distance of v. ASM was prepared by backfilling deposition of methanol vapor onto the gold substrate at 90 K for 60 min. The exposure was typically 1500 L (1 L=1×10⁶ Torr s) to exclude the influence of reactions at the ASM/substrate interface.20,21

Since crystallization of solid methanol occurs at temperatures above 103.4 K, the morphology of the prepared solid methanol is amorphous.22 Unfocused 157 nm laser (Lambda Physik, OPTExPro) radiation with a pulse duration of 10 ns [full width at half maximum (FWHM)] was incident on the ice surface at an angle of 80° to the surface normal and at a fluence, $F < 0.1$ mJ ($\sim 10^{14}$ photons cm⁻² pulse⁻¹), which roughly corresponds to that in molecular clouds over $\sim 10^{4}$ yr.23 CH₃ photoproducts were ionized at a distance of r=2 mm from the substrate surface (4 mm for the case of OH products). CH₃ products were detected by (2+1) REMPI via the $4p_x \rightarrow 2A^1_g(v'=0) \rightarrow X^2A^1_g(v''=0)$ transition in the wavelength range of 285.0–287.0 nm and collected with a small mass spectrometer aligned perpendicular to the ice surface.24 OH products were probed by (2+1) REMPI via the $D^2Σ^+ (v'=0) \rightarrow X^2Π (v''=0)$ transition at 243.5–245.0 nm and via the $D^2Σ^+ (v'=1) \rightarrow X^2Π (v''=0)$ and $3\,^2Σ^+ (v'=0) \rightarrow X^2Π (v'=1)$ transitions in the wavelength range of 237.5–237.7 nm.25 PGOPHER, a program for simulating rotational structure,26 was used to simulate the measured spectra and thereby establish the product rotational temperatures. Detection of O($^1D_2$), O($^3P_{J=2,1,0}$) atoms and H₂O molecules was attempted also by (2+1) REMPI via the $O((^1F_3 \leftarrow ^1D_2))$ transition at 203.8 nm, the $O(^3P_{J=3} \leftarrow ^3P_2)$ transition at 225.6–226.4 nm, and the C(000)-X(000) transition at 247.3–248.6 nm, respectively,27–29 but no signal was observed.

Photodesorption of O($^1D_2$), O($^3P_{J=2,1,0}$) atoms, and H₂O molecules has been previously observed with the same REMPI setup following the 157 nm photolysis of H₂O ice at 90 K.29–31

Recognizing that photoproducts might accumulate on the ASM surface after prolonged irradiation by 157 nm photolysis pulses, and that the photochemistry of these products might influence the measured product (state) distributions, the ASM was continuously refreshed by intermissive exposure to CH₃OH vapor. In the present experiments, the pulsed valve was opened after each laser shot so as to deposit a fresh layer of ASM.

For the CH₃OH/H₂O codeposited ice photolysis experiments, CH₃OH/H₂O (1:1 mixture) vapor was deposited on the gold substrate. As before, the exposure was typically 1500 L and, again, fresh surfaces of cocondensed CH₃OH/H₂O were prepared as described above. All of the present photolysis experiments were performed at a sample temperature of 90 K, and the chamber pressure was 5×10⁻⁷ Torr (due to the intermissive CH₃OH or CH₃OH/H₂O vapor injection into the chamber).

B. Simulation of (2+1) REMPI spectra

CH₃ products were monitored by (2+1) REMPI via the $4p_x \rightarrow 2A^1_g(v'=0) \rightarrow X^2A^1_g(v''=0)$ transition. Spectral simulation employed the program PGOPHER (Ref. 26) and the spectroscopic constants reported by Black and Powis.24 The $4p_x \rightarrow 2A^1_g(v'=0)$ state predissociates with a level dependent efficiency.24 This affects the measured REMPI line intensities and linewidths such that the latter vary as [0.4+0.08(N(N+1)−K²)] (in cm⁻¹), while the former declines by the square of this term. The measured spectra show evidence of power broadening also, which we accommodate by further broadening each transition with a Lorentzian function (3 cm⁻¹ FWHM) and by reducing the relative weighting of the zero rank component of this two-photon transition. The two-photon transition probability is carried by two (a zero and a second rank) components. The former contributes only to the intense central Q branch (which is noticeably saturated under the present conditions), while the latter gives rise to O, P, Q, R, and S branches.

The REMPI features associated with OH products were also assigned and simulated using literature constants for the relevant electronic states.25,32 The two-photon absorption cross sections reported by Greenslade et al.25 allow estimation of the $(v''=1)/(v''=0)$ branching ratio in the OH products.

C. Simulation of time-of-flight spectra of photoproducts

Time-of-flight (TOF) spectra of the CH₃ and OH photoproducts were taken as a function of time t between the photolysis and REMPI laser pulses using a delay generator (Stanford Research) in order to investigate the flight times and thus translational energies of the desorbing photoprod-
A. Kinetic energy and rotational energy distributions of the CH₃ radical

Figure 2(a) shows a rotationally resolved REMPI spectrum of CH₃(v₀=0) products formed following the 157 nm photolysis of freshly deposited ASM recorded at a fixed delay of t=6.0 μs. The spectrum is readily assignable to the Q branch of the CH₃ products monitored at a REMPI wavelength of 286.0 nm (i.e., in the congested Q branch), which is reproduced well by a sum of two MB distributions with $T_{trans}=3000$ K (50% $E_{trans}=49.9 \pm 16.6$ kJ mol⁻¹) and 90 K ($E_{trans}=1.5 \pm 0.3$ kJ mol⁻¹). Table I summarizes these results. Changing the REMPI probe wavelength, e.g., to the S(2) line at 285.45 nm, led to no appreciable change in the TOF profile. We were unable to characterize the rotational resolved REMPI spectrum for the $T_{trans}=3000$ K component in terms of a specific $T_{rot}$ because of the weak signal intensity.

B. Kinetic energy and rotational distributions of the OH radical

Figure 4(a) shows a rotationally resolved REMPI spectrum of the OH(v₀=0) products recorded at t=2.0 μs, i.e., at the maximum of the single peak in their TOF spectrum. Spectral simulation [Fig. 4(b)] returns a best-fit rotational temperature $T_{rot}(v₀=0)=300$ K, and the rotational temperature $T_{rot}(v₀=0)$ is estimated to be 300±100 K by spectral simulation ($E_{rot}=2.5+0.8$ kJ mol⁻¹). Figures 5(a) and 5(b) show typical TOF spectra of the OH(v₀=0) and OH(v₀=1) products.

<table>
<thead>
<tr>
<th>Time-of-flight component, contributions</th>
<th>Translational energy $E_T$ (kJ mol⁻¹)</th>
<th>Rotational temperature $T_{rot}$ (K)</th>
<th>Rotational energy $E_R$ (kJ mol⁻¹)</th>
</tr>
</thead>
<tbody>
<tr>
<td>CH₃($T_{trans}=3000$ K), 10%</td>
<td>49.9±16.6</td>
<td>*</td>
<td>*</td>
</tr>
<tr>
<td>CH₃($T_{trans}=90$ K), 90%</td>
<td>1.5±0.3</td>
<td>150±50</td>
<td>1.2±0.4</td>
</tr>
</tbody>
</table>

*The rotational spectrum could not be characterized by a specific $T_{rot}$ because of the weakness of the signal intensity.

*From spectra recorded at t=6.0 μs.
products monitored, respectively, the $R_t(1)+R_t(5)$ line and the $R_t(2)$ line in the REMPI spectrum. Both are reproduced well by a single MB distribution with $T_{\text{trans}}=3000 \pm 500$ K ($\langle E_{\text{trans}}\rangle=49.9 \pm 8.3$ kcal mol$^{-1}$). No component with $T_{\text{trans}}=90$ K was detected, in contrast to CH$_3$ products. Figure 6(a) shows another portion of the OH REMPI spectrum which contains overlapping contributions from both OH($v=0$) and OH($v=1$) products, recorded at $t=2.0 \mu$s; the accompanying simulation [Fig. 6(b)] employs $T_{\text{rot}}(v=0)=300$ K and $T_{\text{rot}}(v=1)=200$ K. The rotational temperature $T_{\text{rot}}(v=1)$ is estimated to be $200 \pm 50$ K by spectral simulation ($\langle E_{\text{rot}}\rangle=1.7 \pm 0.4$ K mol$^{-1}$). The OH($v=1$)/OH($v=0$) population ratio is determined to be 0.2 $\pm$ 0.1. Table II summarizes these results.

C. Additional 157 nm photolysis experiments involving ASM

To assess possible contributions from secondary processes on the ASM surface and in the bulk, additional experiments were performed wherein TOF spectra of CH$_3$ and OH products formed by 157 nm photolysis of ASM were measured after 30 min photoradiation without the intermissive dosing of CH$_3$OH vapor. No discernible differences were found in either the TOF spectra or the rotationally resolved REMPI spectra of the CH$_3$ or OH products recorded with the fresh or photoradiated ASM samples. These results support the views that (a) the observed CH$_3$ and OH radicals are produced directly via the C–O cleavage reaction (1) and (b) secondary photoprocesses on/in ASM make no contribution to the formation of these desorbing species.

REMPI signals attributable to O($1^D$) and O($3^P$) atoms were sought in the TOF range 0.5 $\mu$s $\leq t \leq 30$ $\mu$s, but not observed. The primary formation of methane and O($3^P$) products is a spin-forbidden process, i.e.,

$$\text{CH}_3\text{OH} + h\nu \rightarrow \text{CH}_3(3\text{A}_1) + \text{O}(3^P).$$  

O($1^D$) and O($3^P$) atoms have been successfully detected with the same REMPI setup following the 157 nm photolysis of H$_2$O or H$_2$O$_2$ ices films at 90 K. The lack of detectable O($1^D$) and O($3^P$) atoms in the present methanol experiments is in accord with the results of gas phase methanol photolysis studies at 157 nm (Refs. 2 and 3) and suggests that the primary formation of methane and O($1^D$) or O($3^P$) products via reactions (2) and (6) plays at most a minor role.

To verify that the measured CH$_3$ product state distributions are not affected by possible contamination from back-

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**FIG. 4.** (a) (2+1) REMPI excitation spectrum of OH radicals from the 157 nm photolysis of a fresh ASM sample at 90 K, recorded at $t=2.0 \mu$s. (b) Simulation of the $D^2\Sigma^+ \rightarrow X^1\Pi(v'=0,v''=0)$ two-photon excitation spectrum of OH assuming a Boltzmann rotational state population distribution with $T_{\text{rot}}=300$ K. The arrow indicates the $R_t(1)+R_t(5)$ line used when measuring the TOF spectrum.

**FIG. 5.** TOF spectra of OH radicals from the 157 nm photolysis of a fresh ASM sample, obtained by monitoring (a) the $R_t(1)+R_t(5)$ line in the OH $D^2\Sigma^+ \rightarrow X^1\Pi(v'=0,v''=0)$ two-photon transition and (b) the $R_t(2)$ line in the OH $3^2\Sigma^- \rightarrow X^1\Pi(v'=0,v''=1)$ two-photon transition. The solid curves are fits to the data derived assuming a MB translational energy distribution with $T_{\text{trans}}=3000$ K in both cases. The vertical flight distance used in these experiments is 4 mm.

**FIG. 6.** (a) (2+1) REMPI excitation spectrum of OH radicals from the 157 nm photolysis of a fresh ASM sample at 90 K recorded at $t=2.0 \mu$s. (b) Simulation of relevant parts of the overlapping $D^2\Sigma^+ \rightarrow X^1\Pi(v'=1,v''=0)$ and $3^2\Sigma^- \rightarrow X^1\Pi(v'=0,v''=1)$ two-photon transitions of OH assuming Boltzmann rotational state population distribution with $T_{\text{rot}}(v'=0)=300$ K (blue line) and $T_{\text{rot}}(v'=1)=200$ K (red line). The black line is the sum of the spectral simulations. The arrow indicates the $R_t(2)$ line used when measuring the TOF spectrum.

**TABLE II.** Translational and rotational temperatures and energies of OH($v=0$) and 1) products.

<table>
<thead>
<tr>
<th></th>
<th>OH($v=0$)</th>
<th>OH($v=1$)</th>
</tr>
</thead>
<tbody>
<tr>
<td><strong>Temperature (K)</strong></td>
<td>3000 ± 500</td>
<td>3000 ± 500</td>
</tr>
<tr>
<td><strong>Energy (kJ mol$^{-1}$)</strong></td>
<td>49.9 ± 8.3</td>
<td>49.9 ± 8.3</td>
</tr>
</tbody>
</table>

$^a$Population ratio OH($v=1$)/OH($v=0$) $=0.2 \pm 0.1$.

$^b$From spectra recorded at $t=2.0 \mu$s.
ground water vapor, the ASM sample was changed to a 1:1 mixture of CH₃OH/H₂O ice. The intensities of the T_{trans}=3000 and 90 K components of the CH₃ radical TOF signal from the mixed sample were both reduced, and the majority of the CH₃ products are accommodated to the substrate temperature (90 K) which was similar to that from pure, fresh ASM samples.

A similar check is not possible for the OH products because of the contribution from OH products formed by simultaneous photolysis of H₂O.⁷,²⁸

### IV. DISCUSSION

The C–O bond cleavage has been observed as a primary process following photolysis of CH₃OH both in the gas phase and in solid rare-gas matrices,²³,⁷

\[
\text{CH}_3\text{OH(ads)} + h\nu \rightarrow \text{CH}_3(i) + \text{OH}(i),
\]

\[\Delta H = 422.6 \text{ kJ mol}^{-1}.\] (7)

The available energy for reaction (7) at 157 nm (E_{phot} = 757.3 kJ mol⁻¹) is E_{avail}(7) = 334.7 kJ mol⁻¹, calculated using liquid phase thermodynamic data for the condensed or adsorbed species (“ads”) and gas phase data for species on the ASM/vacuum interface (“i”).³⁹,⁴⁰ The center of mass kinetic energy distribution of the CH₃ and OH products resulting from 157 nm photolysis of gas phase CH₃OH shows maximum probability at 217.6 kJ mol⁻¹ and extends to 372.4 kJ mol⁻¹, which is consistent with the gas phase value for E_{avail}(7).²³ This translational energy is partitioned between the CH₃ and OH fragments according to momentum conservation, such that the most probable fragment translational energies are, respectively, 117.2 and 100.4 kJ mol⁻¹. These values are much higher than those found in the present condensed phase study with E_{avail}(7) = 334.7 kJ mol⁻¹, wherein we determine an average center of mass translational energy of the photodesorbed CH₃ and OH products of only 99.6 kJ mol⁻¹. This reduced partitioning of E_{avail}(7) into product translational energy could be due to cage effects and the dissipation of available energy into the ASM bulk.⁷,¹⁷

Hama et al. performed direct measurements of photodesorbed hydrogen atoms and molecules following VUV photolysis of amorphous solid CH₃OH and CD₂OH.¹⁸ The results showed that H and D atoms are produced via reactions (8) and (9),

\[
\text{CD}_3\text{OH} + h\nu \rightarrow \text{CD}_2\text{O} + \text{H},
\]

\[\Delta H = 422.6 \text{ kJ mol}^{-1}.\] (8)

\[
\text{CD}_2\text{OH} + h\nu \rightarrow \text{CD}_2\text{OH} + \text{D}.
\]

(9)

Highly vibrationally excited D₂ and HD(v = 2–5) molecules were also produced as primary products by unimolecular elimination from methanol, i.e., reactions (10) and (11),

\[
\text{CD}_3\text{OH} + h\nu \rightarrow \text{CD}_2\text{O} + \text{D}_2(v = 2–5),
\]

(10)

\[
\text{CD}_2\text{OH} + h\nu \rightarrow \text{CD}_2\text{O} + \text{HD}(v = 2–5).
\]

(11)

Reactions (8)–(11) and the C–O bond cleavage (7) are the primary photodissociation processes of ASM. However, the branching ratio of the primary photoprocesses could not be quantified either in our previous or the present studies.

No IR absorption signal attributable to CH₃ and OH radicals was observed in earlier VUV photolysis studies of solid methanol at 10 K.¹⁵,¹⁶ These results imply that (a) the quantum yield of the C–O bond cleavage reaction (7) on the surface under VUV irradiation is small or, (b) CH₃ and OH photoproducts, if formed, are removed by photodesorption and reactions on/in ASM and thus remain undetected. In the following sections, possible secondary reactions of CH₃ and OH products on/in ASM are proposed.

### A. CH₃ radical formation in the 157 nm photolysis of fresh ASM

Two MB components, characterized by T_{trans}=3000 and 90 K and with respective relative yields of 10% and 90%, are required to fit the TOF spectrum of CH₃ products measured at 157 nm. The CH₃ rotational temperatures for the T_{trans}=90 K component are almost thermally equilibrated with the substrate temperature (90 K), as shown in Table I. Since 90% of the CH₃ fragments are accommodated to the substrate temperature, and the translational and rotational energies are much lower than E_{avail}(7), we propose that the majority of the nascent CH₃ products are formed by absorption in the ASM bulk and are subsequently relaxed by collisions within the ASM en route to the vacuum. Only a small fraction (10%) of the CH₃ products are produced at the exposed ASM surface and retain a high translational temperature, i.e.,

\[
\text{CH}_3\text{OH} + h\nu \rightarrow \text{CH}_3(T_{trans} = 3000 \text{ K}) + \text{OH},
\]

\[\text{collision}\]

\[
\text{CH}_3(3000 \text{ K}) \rightarrow \text{CH}_3(90 \text{ K}).
\]

(13)

CH₃ photoproducts may also be removed by reactions within the bulk and thereby remain undetected by IR absorption spectroscopy.¹⁵,¹⁶ For example, hydrogen abstraction from methanol by CH₃ radicals, reactions (14) and (15), are known in the gas phase,⁴¹

\[
\text{CH}_3(i) + \text{CH}_3\text{OH(ads)} \rightarrow \text{CH}_4(i) + \text{CH}_2\text{OH(i)},
\]

\[\Delta H = 4.2, \quad E_a = 58.6 \text{ kJ mol}^{-1},\]

\[\text{CH}_3(i) + \text{CH}_3\text{OH(ads)} \rightarrow \text{CH}_4(i) + \text{CH}_3\text{O(i)},
\]

\[\Delta H = 33.5, \quad E_a = 56.9 \text{ kJ mol}^{-1}.\]

The quoted activation energies, E_a, are ab initio estimates from MP2 and MP4 calculations reported by Jodkowski et al.⁴⁹ Some of the CH₃ products formed at the ASM/vacuum interface with (E_{trans}) = 49.8 kJ mol⁻¹ and considerable internal energy may have sufficient energy to induce the H abstraction reactions (14) and (15) but, as Table I shows, most of the CH₃ photoproducts relax to T_{trans} = 90 ± 20 K and T_{rot} = 150 ± 50 K. At this stage, their average energy (∼ 2.7 kJ mol⁻¹) is far too small to overcome such activation barriers. The thermally equilibrated CH₃ photoproducts then diffuse through the bulk ASM until they desorb from the solid/vacuum interface.
Other possible removal processes for CH₃ include (generally barrierless) reaction with other radical photoproducts, e.g., reaction with H-atom photoproducts to form methane,

\[ \text{CH}_3(i) + H(i) + M \rightarrow \text{CH}_4(i) + M, \]

\[ \Delta H = -439.3 \text{ kJ mol}^{-1}. \]  

Compared to the gas phase, we might anticipate a significant enhancement in the rate of reaction (16) in ASM since H atoms are very mobile, even at low temperature, and reaction (16) is a highly exothermic, three-body reaction. H atoms are known to be present in high density following pulsed 157 nm laser excitation of ASM at 90 K, and H₂ molecule formation from H atoms has also been reported, i.e.,

\[ \text{CH}_3OH + h \nu \rightarrow \text{CH}_3OH/\text{CH}_3O + H, \]

\[ \text{H} + \text{CH}_3OH \rightarrow \text{CH}_3OH/\text{CH}_3O + \text{H}_2, \]

\[ \text{H} + \text{H} \rightarrow \text{H}_2 (v \geq 2). \]

Earlier VUV photolysis studies of solid methanol at 10 K found the yield of CH₃ to depend linearly on the irradiation time, but no IR absorption signal attributable to CH₃ was observed, probably because these radicals were efficiently consumed by reaction with H atoms, i.e., via reaction (16).

In the mixed CH₃OH/H₂O ice experiment, the TOF signal intensity of the CH₃ radicals was reduced. This may be due to (a) the reduction in CH₃OH density on the surface, or (b) the reaction of CH₃ products with H atoms or OH radicals formed from photolysis of H₂O. The reaction of CH₃ with water molecules is unlikely: the TOF profile of CH₃ radicals produced from the mixed ice showed a similar behavior to that from the fresh ASM samples, i.e., the majority of the CH₃ products are accommodated to the substrate temperature (90 K). This implies that the CH₃ products desorb without reaction with solid water, as would be expected given the large activation energy of reaction (20),

\[ \text{CH}_3(i) + \text{H}_2\text{O(ads)} \rightarrow \text{CH}_4(i) + \text{OH}(i), \]

\[ \Delta H = 104.6, \quad E_a = 138.1 \text{ kJ mol}^{-1}. \]

The activation energy or reaction (20) was estimated from theoretical calculations for the reverse process, CH₄+OH → CH₃+H₂O, by Bravo-Pérez et al. The translational temperatures of the photodesorbed H or OH products following the 157 nm photolysis of water ice have been previously observed with the same REMPI setup. The previous results showed that the majority of the photodesorbed H atoms were relaxed by collisions within the water ice and the translational temperature was accommodated to the substrate temperature. The photodesorbed OH radicals, on the other hand, were not thermally equilibrated with the substrate temperature but were likely formed near the water ice surface, whereas collisionally cooled OH products were either efficiently trapped in the bulk or reacted with water.

### B. OH radical formation in the 157 nm photolysis of fresh ASM

The translational energies of OH(\(v = 0 \) and 1) products from the 157 nm photolysis of fresh ASM are described by a single temperature \( T_{\text{trans}} = 3000 \pm 500 \text{ K} \) \((\langle E_{\text{trans}} \rangle = 49.9 \pm 8.3 \text{ kJ mol}^{-1}). \)

The OH rotational temperatures, \( T_{\text{rot}}(v = 0) = 300 \pm 100 \text{ K} \) and \( T_{\text{rot}}(v = 1) = 200 \pm 50 \text{ K}, \) are not thermally equilibrated with the substrate temperature of 90 K, as shown in Table III. These results suggest that (a) the observed OH(\(v = 0 \) and 1) fragments originate from the ASM surface, not the bulk, and (b) OH(\(v = 0 \) and 1) radicals formed in the ASM bulk are either efficiently trapped or react in the bulk as described before in the 157 nm photolysis of H₂O, i.e.,

\[ \text{CH}_3OH + h \nu \rightarrow \text{CH}_3 + \text{OH}(T_{\text{trans}} = 3000 \text{ K}), \]

\[ \text{collison} \]

\[ \text{OH}(3000 \text{ K}) \rightarrow \text{trapped or react with ASM}. \]

Possible H abstraction reactions from methanol by OH radicals are

\[ \text{OH}(i) + \text{CH}_3\text{OH(ads)} \rightarrow \text{H}_2\text{O}(i) + \text{CH}_3\text{OH}(i), \]

\[ \Delta H = -50.2, \quad E_a = 3.8 \text{ kJ mol}^{-1}. \]

\[ \text{OH}(i) + \text{CH}_3\text{OH(ads)} \rightarrow \text{H}_2\text{O}(i) + \text{CH}_3\text{O}(i), \]

\[ \Delta H = -25.1, \quad E_a = 14.6 \text{ kJ mol}^{-1}. \]

The activation energies quoted are for the corresponding gas phase reactions, estimated by Jodkowski et al. Reaction (23) has been proposed previously in methanol-water condensed phase systems. Recalling Table II, the OH products from the 157 nm photolysis of ASM have ample energy to overcome the barriers for reactions (23) and (24), and thus OH radical initiated hydrogen abstraction is likely to occur on/in ASM.

However, neither photodesorption nor production of H₂O was observed in either the present or earlier VUV photolysis studies on ASM. These results imply that the yield of OH production via the C–O bond cleavage reaction (7) on the surface is small and that the secondary H₂O formation via reactions (23) and (24) does not proceed efficiently.

### C. Other possible secondary reactions

After photolysis of a CH₃OH molecule in bulk ASM, the probability of reaction (25) is expected to be high since the photogenerated CH₃ and OH radicals are effectively caged by the ASM matrix, i.e.,

\[ \text{CH}_3(i) + \text{OH}(i) + M \rightarrow \text{CH}_3\text{OH(ads)} + M, \]

\[ \Delta H = -422.6 \text{ kJ mol}^{-1}, \]

the bimolecular reaction between gas phase CH₃ and OH radicals has been studied extensively, both theoretically and experimentally. The calculations of Jasper et al. show a number of reactive radicals and molecular products from this
reaction, e.g., CH₃OH, HCOOH+H₂, and CH₂+H₂O.⁴⁸ In the condensed phase, Hodıyss et al. reported a UV (130–335 nm) photolysis study of CH₄–H₂O ice mixtures at 20 K.⁴⁹ Their analysis assumes that OH radicals formed by the photodissociation of water react with methane to form CH₃ radicals, which then recombine with OH radicals to form methanol.

Ethane formation, from the recombination reaction of two CH₃ radicals [reaction (26)], was also reported in the photolysis of CH₄–H₂O ice mixtures,⁴⁹

\[ \text{CH}_3(i) + \text{CH}_3(i) + M \rightarrow \text{C}_2\text{H}_6(i) + M, \]  
\[ \Delta H = -376.6 \text{ kJ mol}^{-1}. \]  

(26)

No IR absorption attributable to C₂H₆ was observed in the earlier VUV photolysis studies of solid CH₂OH at 10 K, however.¹⁵,¹⁶ This finding is understandable if CH₃ photo-products react with other species (e.g., H, OH, or CH₂OH) much more frequently than they encounter another CH₃ radical.

Hodıyss et al.⁴⁹ proposed that the presence of CH₄ in CH₄–H₂O ice mixtures inhibited the formation of H₂O₂ (from the combination of two OH radicals) by serving as a trap for OH radicals, since the C–H bonds are also attacked by OH radicals. In fact, no IR absorption features attributable to H₂O₂ or OH were observed in the VUV photolysis of solid CH₂OH at 10 K.¹⁵,¹⁶ Our previous experimental work has confirmed that OH radicals are formed following the 157 nm photolysis of surface H₂O, and that the translational energy of OH radicals from H₂O₂ was much larger than that from H₂O.³⁸ There are no discernible differences in the TOF spectra, or the rotational resolved REMPI spectra, of the OH radicals from the 157 nm photolysis of fresh or photolysed ASM samples, i.e., the OH radicals detected in the present study are produced directly, via reaction (7), and not from secondary photolysis of photogenerated H₂O₂.

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26. C. M. Western, PGOPHER, a program for simulating rotational structure, University of Bristol, available at http://ppgopher.chem.bris.ac.uk.


