Title

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Translational and internal energy distributions of methyl and hydroxyl radicals produced by 157 nm photodissociation of amorphous solid methanol

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Methanol is typically observed within water-rich interstellar ices and is a source of interstellar organic species. Following the 157 nm photoexcitation of solid methanol at 90 K, desorbed CH₃(v=0) and OH(v=0,1) radicals have been observed in situ, near the solid surface, using resonance-enhanced multiphoton ionization (REMPI) detection methods. Time-of-flight and rotationally resolved REMPI spectra of the desorbed species were measured, and the respective fragment internal energy and kinetic energy distributions were obtained. Photoproduction mechanisms for CH₃ and OH radicals from solid methanol are discussed. The formation of O(^1D) and ^3P) atoms and H₂O was investigated, but the yield of these species was found to be negligible. CH₃ products arising following the photoexcitation of water-methanol mixed ice showed similar kinetic and internal energy distributions to those from neat methanol ice. © 2009 American Institute of Physics. [doi:10.1063/1.3259877]

I. INTRODUCTION

Methanol is the smallest closed shell alcohol and has long served as a model for the dissociation of organic molecules. There is a comparably large body of literature on the vacuum ultraviolet (VUV) photolysis of methanol in the gas phase. The results of the gas phase methanol photolysis studies at 157 nm showed that the relative contribution of the atomic hydrogen elimination process is larger than that of the molecular hydrogen elimination process, i.e., the ratio is 1:0.21. C–O bond cleavage [reaction (1)] is another primary photoprocess following excitation of the 2^1A^n–X^1A^n transition in the 151–163 nm region, where it competes with the atomic and molecular hydrogen elimination processes,

\[ \text{CH}_3\text{OH} + h\nu \rightarrow \text{CH}_3 + \text{OH}. \]  

The relative contribution of the C–O bond cleavage channel in the gas phase photolysis at 157 nm has yet to be determined. Reaction (1) was also observed following the photolysis of CH₃OH in various rare-gas matrices at 4 K and the photodissociation threshold energy for reaction (1) identified at \( \lambda \approx 175 \text{ nm} \). Since methanol is known to be a component of the icy mantle of interstellar grains, comets, and other solar system bodies, the effects of irradiating pure methanol ice by electrons, ions, and UV photons have each received much experimental study. Compared to the gas phase, however, the photochemical processes occurring in condensed phase samples of CH₃OH are more complicated because of the many possible second-
products following 157 nm photodissociation of ASM at 90 K, using REMPI methods to determine their respective translational energy distributions and internal state population distributions. Possible secondary reactions on/in ASM are discussed in light of the obtained results. An attempt was also made to detect O(^1D) and O(^3P) atoms and H2O products.

II. EXPERIMENTAL

A. Apparatus and preparation of ice films

The apparatus has been described before, and only details relevant to the present experiments are given here.21 As shown in the schematic diagram of the experimental arrangement (Fig. 1), photodesorbed products were ionized at the focal point of the REMPI probe laser light, which was separated from the ice surface by a vertical distance of r. ASM was prepared by backfilling deposition of methanol vapor onto the gold substrate at 90 K for 60 min. The exposure was typically 1500 L (1 L=1×10⁻⁶ Torr s) to exclude the influence of reactions at the ASM/substrate interface.21,22

Since crystallization of solid methanol occurs at temperatures above 103.4 K, the morphology of the prepared solid methanol is amorphous.22 Unfocused 157 nm laser (Lambda Physik, OPTexPro) radiation with a pulse duration of 10 ns [full width at half maximum (FWHM)] was incident on the ice surface at an angle of ~80° to the surface normal and at a fluence, F<0.1 mJ (~10¹⁴ photons) cm⁻² pulse⁻¹, which roughly corresponds to that in molecular clouds over ~10⁴ yr.23 CH₃ photoproducts were ionized at a distance of r=2 mm from the substrate surface (4 mm for the case of OH products). CH₃ products were detected by (2+1) REMPI via the 4P_g 2A⁰_2(v' =0) ← X 2Π_2(v'=0) transition in the wavelength range of 285.0–287.0 nm and collected with a small mass spectrometer aligned perpendicular to the ice surface.24 OH products were probed by (2+1) REMPI via the D 2Σ⁺(v' =0) ← X 2Π_1(v'=0) transition at 243.5–245.0 nm and via the D 2Σ⁺(v'=1) ← X 2Π_1(v'=0) and 3 2Σ⁺(v'=0) ← X 2Π_1(v'=1) transitions in the wavelength range of 237.5–237.7 nm.25 PGOPHER, a program for simulating rotational structure,26 was used to simulate the measured spectra and thereby establish the product rotational temperatures. Detection of O(^1D₂), O(^3P_{j=2,1,0}) atoms and H₂O molecules was attempted also by (2+1) REMPI via the O(^1F₃ ← ^1D₂) transition at 203.8 nm, the O(^3P_j←^3P_j) transition at 225.6–226.4 nm, and the C(000)-X(000) transition at 247.3–248.6 nm, respectively.27–29 but no signal was ob-served. Photodesorption of O(^1D₂), O(^3P_{j=2,1,0}) atoms, and H₂O molecules has been previously observed with the same REMPI setup following the 157 nm photolysis of H₂O ice at 90 K.29–31

Recognizing that photoproducts might accumulate on the ASM surface after prolonged irradiation by 157 nm photolysis pulses, and that the photochemistry of these products might influence the measured product (state) distributions, the ASM surface was continually refreshed by intermissive exposure to CH₃OH vapor. In the present experiments, the pulsed valve was opened after each laser shot so as to deposit a fresh layer of ASM.

For the CH₃OH/H₂O codeposited ice photolysis experiments, CH₃OH/H₂O (1:1 mixture) vapor was deposited on the gold substrate. As before, the exposure was typically 1500 L and, again, fresh surfaces of cocondensed CH₃OH/H₂O were prepared as described above. All of the present photolysis experiments were performed at a sample temperature of 90 K, and the chamber pressure was 5×10⁻⁷ Torr (due to the intermissive CH₃OH or CH₃OH/H₂O vapor injection into the chamber).

B. Simulation of (2+1) REMPI spectra

CH₃ products were monitored by (2+1) REMPI via the 4P_g 2A⁰_2(v' =0) ← X 2Π_2(v'=0) transition. Spectral simulation employed the program PGOPHER (Ref. 26) and the spectroscopic constants reported by Black and Powis.24 The 4P_g 2A⁰_2(v' =0) state predissociates with a level dependent efficiency.24 This affects the measured REMPI line intensities and linewidths such that the latter vary as [0.4+0.08[(N(N+1)−K²)] (in cm⁻¹), while the former declines by the square of this term. The measured spectra show evidence of power broadening also, which we accommodate by further broadening each transition with a Lorentzian function (3 cm⁻¹ FWHM) and by reducing the relative weighting of the zero rank component of this two-photon transition. The two-photon transition probability is carried by two (a zero and a second rank) components. The former contributes only to the intense central Q branch (which is noticeably saturated under the present conditions), while the latter gives rise to O, P, Q, R, and S branches.

The REMPI features associated with OH products were also assigned and simulated using literature constants for the relevant electronic states.25,32 The two-photon absorption cross sections reported by Greenslade et al.25 allow estimation of the (v''=1)/(v''=0) branching ratio in the OH products.

C. Simulation of time-of-flight spectra of photoproducts

Time-of-flight (TOF) spectra of CH₃ and OH photoproducts were taken as a function of time t between the photolysis and REMPI laser pulses using a delay generator (Stanford Research) in order to investigate the flight times and thus translational energies of the desorbing photoprod-
B. Kinetic energy and rotational distributions of the OH radical

Figure 4(a) shows a rotationally resolved REMPI spectrum of the OH(\(v=0\)) products recorded at \(t=2.0 \mu s\), i.e., at the maximum of the single peak in their TOF spectrum. Spectral simulation [Fig. 4(b)] returns a best-fit rotational temperature \(T_{\text{rot}}(v=0)=300\) K, and the rotational temperature \(T_{\text{rot}}(v=0)\) is estimated to be \(300 \pm 100\) K by spectral simulation \((\langle E_{\text{rot}} \rangle=2.5 \pm 0.8 \text{ kJ mol}^{-1})\). Figures 5(a) and 5(b) show typical TOF spectra of the OH(\(v=0\)) and OH(\(v=1\)) photolysis of freshly deposited ASM recorded at a fixed delay of \(t=6.0 \mu s\). The spectrum is readily assignable to the \(4p_u^2A_2^e(v'=0) \leftrightarrow X^2A_1^e(v'=0)\) transition. The best-fit simulation is shown in Fig. 2(b). The rotational temperature, \(T_{\text{rot}}\), is estimated to be \(150 \pm 50\) K \((\langle E_{\text{rot}} \rangle=1.2 \pm 0.4 \text{ kJ mol}^{-1})\) by comparison with the spectral simulation. Figure 3 shows a typical TOF spectrum of the CH\(_3\) products monitored at a REMPI wavelength of 286.0 nm (i.e., in the congested \(Q\) branch), which is reproduced well by a sum of two MB distributions with \(T_{\text{trans}}=3000 \pm 1000\) K \((\langle E_{\text{trans}} \rangle=49.9 \pm 16.6 \text{ kJ mol}^{-1})\) and \(90 \pm 20\) K \((\langle E_{\text{trans}} \rangle=1.5 \pm 0.3 \text{ kJ mol}^{-1})\). Table I summarizes these results. Changing the REMPI probe wavelength, e.g., to the (2) line at 285.45 nm, led to no appreciable change in the TOF profile. We were unable to characterize the rotational resolved REMPI spectrum for the \(T_{\text{trans}}=3000\) K component in terms of a specific \(T_{\text{rot}}\) because of the weak signal intensity.

### III. RESULTS

#### A. Kinetic energy and rotational energy distributions of the CH\(_3\) radical

Figure 2(a) shows a rotationally resolved REMPI spectrum of CH\(_3\)(\(v'^*=0\)) products formed following the 157 nm photodissociation of a fresh ASM sample, monitoring on the \(Q\) branch of the \((2+1)\) REMPI spectrum shown in Fig. 2. The solid curves are fits to the data derived assuming two MB distributions with \(T_{\text{trans}}=3000\) K (10\%) and 90 K (90\%). The vertical flight distance used in these experiments is 2 mm.

**FIG. 2.** (a) \((2+1)\) REMPI excitation spectrum of CH\(_3\) radicals from the 157 nm photolysis of a fresh ASM sample at 90 K, recorded at \(t=6.0 \mu s\). (b) Simulated REMPI excitation spectrum of CH\(_3\) radicals, resonance enhanced at the two photon energy by the \(4p_u^2A_2^e(v'=0)\) level, assuming a Boltzmann distribution rotational state population distribution with \(T_{\text{rot}}=150\) K.

**FIG. 3.** TOF spectrum of CH\(_3\) radicals from the 157 nm photolysis of a fresh ASM sample, monitoring on the \(Q\) branch of the \((2+1)\) REMPI spectrum shown in Fig. 2. The solid curves are fits to the data derived assuming two MB distributions with \(T_{\text{trans}}=3000\) K (10\%) and 90 K (90\%). The vertical flight distance used in these experiments is 2 mm.

**TABLE I.** Translational and rotational temperatures and energies of CH\(_3\)(\(v=0\)) products.

<table>
<thead>
<tr>
<th>Time-of-flight component</th>
<th>Translational energy</th>
<th>Rotational temperature</th>
<th>Rotational energy</th>
</tr>
</thead>
<tbody>
<tr>
<td>CH(<em>3)((T</em>{\text{trans}}=3000) K)</td>
<td>(49.9 \pm 16.6)</td>
<td>*</td>
<td>*</td>
</tr>
<tr>
<td>CH(<em>3)((T</em>{\text{trans}}=90) K)</td>
<td>(1.5 \pm 0.3) (150 \pm 50)</td>
<td>(1.2 \pm 0.4)</td>
<td></td>
</tr>
</tbody>
</table>

*The rotational spectrum could not be characterized by a specific \(T_{\text{rot}}\) because of the weakness of the signal intensity.

*From spectra recorded at \(t=6.0 \mu s\).
products monitoring, respectively, the $R_t(1) + R_t(5)$ line and the $R_t(2)$ line in the REMPI spectrum. Both are reproduced well by a single MB distribution with $T_{\text{trans}} = 3000 \pm 500$ K ($\langle E_{\text{trans}} \rangle = 49.9 \pm 8.3$ kcal mol$^{-1}$). No component with $T_{\text{trans}} = 90$ K was detected, in contrast to CH$_3$ products. Figure 6(a) shows another portion of the OH REMPI spectrum which contains overlapping contributions from both OH($v=0$) and OH($v=1$) products, recorded at $t = 2.0$ $\mu$s; the accompanying simulation [Fig. 6(b)] employs $T_{\text{rot}}(v=0) = 300$ K and $T_{\text{rot}}(v=1) = 200$ K. The rotational temperature $T_{\text{rot}}(v=1)$ is estimated to be $200 \pm 50$ K by spectral simulation ($\langle E_{\text{rot}} \rangle = 1.7 \pm 0.4$ kJ mol$^{-1}$). The OH($v=1$)/OH($v=0$) population ratio is determined to be $0.2 \pm 0.1$. Table II summarizes these results.

C. Additional 157 nm photolysis experiments involving ASM

To assess possible contributions from secondary processes on the ASM surface and in the bulk, additional experiments were performed wherein TOF spectra of CH$_3$ and OH products formed by 157 nm photolysis of ASM were measured after 30 min photolrradiation without the intermssive dosing of CH$_3$OH vapor. No discernible differences were found in either the TOF spectra or the rotationally resolved REMPI spectra of the CH$_3$ or OH products recorded with the fresh or photoirradiated ASM samples. These results support the views that (a) the observed CH$_3$ and OH radicals are produced directly via the C–O cleavage reaction (1) and (b) secondary photoprocesses on/in ASM make no contribution to the formation of these desorbing species.

REMPI signals attributable to O($^1D$) and O($^3P$) atoms were sought in the TOF range 0.5 $\mu$s $\leq t \leq 30$ $\mu$s, but not observed. The primary formation of methane and O($^3P$) products is a spin-forbidden process, i.e.,

$$\text{CH}_3\text{OH} + h\nu \rightarrow \text{CH}_4(1^1\text{A}_1) + \text{O}(^3\text{P}).$$

(6)

O($^1D$) and O($^3P$) atoms have been successfully detected with the same REMPI setup following the 157 nm photolysis of H$_2$O or H$_2$O$_2$ ices films at 90 K. The lack of detectable O($^1D$) and O($^3P$) atoms in the present methanol experiments is in accord with the results of gas phase methanol photolysis studies at 157 nm (Refs. 2 and 3) and suggests that the primary formation of methane and O($^1D$) or O($^3P$) products via reactions (2) and (6) plays at most a minor role.

To verify that the measured CH$_3$ product state distributions are not affected by possible contamination from back-

![FIG. 4.](image1)  
**FIG. 4.** (a) (2+1) REMPI excitation spectrum of OH radicals from the 157 nm photolysis of a fresh ASM sample at 90 K, recorded at $t = 2.0$ $\mu$s. (b) Simulation of the $D^3\Sigma^{-} \rightarrow X^1\Pi(v'=0,v''=0)$ two-photon excitation spectrum of OH assuming a Boltzmann rotational state population distribution with $T_{\text{rot}} = 300$ K. The arrow indicates the $R_t(1)+R_t(5)$ line used when measuring the TOF spectrum.

![FIG. 5.](image2)  
**FIG. 5.** TOF spectra of OH radicals from the 157 nm photolysis of a fresh ASM sample, obtained by monitoring (a) the $R_t(1)+R_t(5)$ line in the OH $D^3\Sigma^{-} \rightarrow X^1\Pi(v'=0,v''=0)$ two-photon transition and (b) the $R_t(2)$ line in the OH $3^3\Sigma^{-} \rightarrow X^1\Pi(v'=0,v''=1)$ two-photon transition. The solid curves are fits to the data derived assuming a MB translational energy distribution with $T_{\text{trans}} = 3000$ K in both cases. The vertical flight distance used in these experiments is 4 mm.

![FIG. 6.](image3)  
**FIG. 6.** (a) (2+1) REMPI excitation spectrum of OH radicals from the 157 nm photolysis of a fresh ASM sample at 90 K recorded at $t = 2.0$ $\mu$s. (b) Simulation of relevant parts of the overlapping $D^3\Sigma^{-} \rightarrow X^1\Pi(v'=1,v''=0)$ and $3^3\Sigma^{-} \rightarrow X^1\Pi(v'=0,v''=1)$ two-photon transitions of OH assuming Boltzmann rotational state population distribution with $T_{\text{rot}}(v'=0) = 300$ K (blue line) and $T_{\text{rot}}(v'=1) = 200$ K (red line). The black line is the sum of the spectral simulations. The arrow indicates the $R_t(2)$ line used when measuring the TOF spectrum.

### Table II. Translational and rotational temperatures and energies of OH($v=0$) and 1) products.

<table>
<thead>
<tr>
<th></th>
<th>OH($v=0$)</th>
<th>OH($v=1$)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Temperature (K)</td>
<td>3000 ± 500</td>
<td>3000 ± 500</td>
</tr>
<tr>
<td>Energy (kJ mol$^{-1}$)</td>
<td>49.9 ± 8.3</td>
<td>49.9 ± 8.3</td>
</tr>
</tbody>
</table>

$^f$Population ratio OH($v=1$)/OH($v=0$) = 0.2 ± 0.1.

$^g$From spectra recorded at $t = 2.0$ $\mu$s.
ground water vapor, the ASM sample was changed to a 1:1 mixture of CH₃OH/H₂O ice. The intensities of the $T_{\text{trans}} = 3000$ and 90 K components of the CH₃ radical TOF signal from the mixed sample were both reduced, and the majority of the CH₃ products are accommodated to the substrate temperature (90 K) which was similar to that from pure, fresh ASM samples.

A similar check is not possible for the OH products because of the contribution from OH products formed by simultaneous photolysis of H₂O.²⁷,³⁸

IV. DISCUSSION

The C–O bond cleavage has been observed as a primary process following photolysis of CH₃OH both in the gas phase and in solid rare-gas matrices,²³,⁷

$$\text{CH₃OH(ads) + hν → CH₃(i) + OH(i),}$$

$$\Delta H = 422.6 \text{ kJ mol}^{-1}.$$ (7)

The available energy for reaction (7) at 157 nm ($E_{\text{phot}} = 757.3 \text{ kJ mol}^{-1}$) is $E_{\text{avail}}(7) = 334.7 \text{ kJ mol}^{-1}$, calculated using liquid phase thermodynamic data for the condensed or adsorbed species (“ads”) and gas phase data for species on the ASM/vacuum interface (“i”).²⁰,⁴⁰ The center of mass kinetic energy distribution of the CH₃ and OH products resulting from 157 nm photolysis of gas phase CH₃OH shows maximum probability at 217.6 kJ mol⁻¹ and extends to 372.4 kJ mol⁻¹, which is consistent with the gas phase value for $E_{\text{avail}}(7)$.²³ This translational energy is partitioned between the CH₃ and OH fragments according to momentum conservation, such that the most probable fragment translational energies are, respectively, 117.2 and 100.4 kJ mol⁻¹. These values are much higher than those found in the present condensed phase study with $E_{\text{avail}}(7) = 334.7 \text{ kJ mol}^{-1}$, wherein we determine an average center of mass translational energy for the photodesorbed CH₃ and OH products of only 99.6 kJ mol⁻¹. This reduced partitioning of $E_{\text{avail}}(7)$ into product translational energy could be due to cage effects and the dissipation of available energy into the ASM bulk.⁷,¹⁵

Hama et al. performed direct measurements of photodesorbed hydrogen atoms and molecules following VUV photolysis of amorphous solid CH₃OH and CD₂OH.¹⁸ The results showed that H and D atoms are produced via reactions (8) and (9),

$$\text{CD₃OH + hν → CD₃O + H,}$$ (8)

$$\text{CD₂OH + hν → CD₂O + D.}$$ (9)

Highly vibrationally excited D₂ and HD($v=2–5$) molecules were also produced as primary products by unimolecular elimination from methanol, i.e., reactions (10) and (11),

$$\text{CD₃OH + hν → CDOH + D₂(v = 2–5),}$$ (10)

$$\text{CD₂OH + hν → CD₂O + HD(v = 2–5).}$$ (11)

Reactions (8)–(11) and the C–O bond cleavage (7) are the primary photodissociation processes of ASM. However, the branching ratio of the primary photoproducts could not be quantified either in our previous or the present studies.

No IR absorption signal attributable to CH₃ and OH radicals was observed in earlier VUV photolysis studies of solid methanol at 10 K.¹⁵,¹⁶ These results imply that (a) the quantum yield of the C–O bond cleavage reaction (7) on the surface under VUV irradiation is small or, (b) CH₃ and OH photoproducts, if formed, are removed by photodesorption and reactions on/in ASM and thus remain undetected. In the following sections, possible secondary reactions of CH₃ and OH products on/in ASM are proposed.

A. CH₃ radical formation in the 157 nm photolysis of fresh ASM

Two MB components, characterized by $T_{\text{trans}} = 3000$ and 90 K and with respective relative yields of 10% and 90%, are required to fit the TOF spectrum of CH₃ products measured at 157 nm. The CH₃ rotational temperatures for the $T_{\text{trans}} = 90$ K component are almost thermally equilibrated with the substrate temperature (90 K), as shown in Table I. Since 90% of the CH₃ fragments are accommodated to the substrate temperature, and the translational and rotational energies are much lower than $E_{\text{avail}}(7)$, we propose that the majority of the nascent CH₃ products are formed by absorption in the ASM bulk and are subsequently relaxed by collisions within the ASM en route to the vacuum. Only a small fraction (10%) of the CH₃ products are produced at the exposed ASM surface and retain a high translational temperature, i.e.,

$$\text{CH₃OH + hν → CH₃(T_{\text{trans}} = 3000 \text{ K}) + OH,}$$ (12)

$$\text{CH₃(3000 K) → CH₃(90 K).}$$ (13)

CH₃ photoproducts may also be removed by reactions within the bulk and thereby remain undetected by IR absorption spectroscopy.¹⁵,¹⁶ For example, hydrogen abstraction from methanol by CH₃ radicals, reactions (14) and (15), are known in the gas phase,⁴¹

$$\text{CH₃(i) + CH₃OH(ads) → CH₄(i) + CH₂OH(i),}$$

$$\Delta H = 4.2, \quad E_a = 58.6 \text{ kJ mol}^{-1},$$ (14)

$$\text{CH₃(i) + CH₂OH(ads) → CH₄(i) + CH₃O(i),}$$

$$\Delta H = 33.5, \quad E_a = 56.9 \text{ kJ mol}^{-1}.$$ (15)

The quoted activation energies, $E_a$, are $ab initio$ estimates from MP2 and MP4 calculations reported by Jodkowski et al.⁴¹ Some of the CH₃ products formed at the ASM/vacuum interface with ($E_{\text{trans}} = 49.8 \text{ kJ mol}^{-1}$ and considerable internal energy may have sufficient energy to induce the H abstraction reactions (14) and (15) but, as Table I shows, most of the CH₃ photoproducts relax to $T_{\text{trans}} = 90 \pm 20 \text{ K}$ and $T_{\text{rot}} = 150 \pm 50 \text{ K}$. At this stage, their average energy ($\sim 2.7 \text{ kJ mol}^{-1}$) is far too small to overcome such activation barriers. The thermally equilibrated CH₃ photoproducts then diffuse through the bulk ASM until they desorb from the solid/vacuum interface.
Other possible removal processes for \( \text{CH}_3 \) include (generally barrierless) reaction with other radical photoproducts, e.g., reaction with H-atom photoproducts to form methane,

\[
\text{CH}_3(i) + \text{H}(i) + \text{M} \rightarrow \text{CH}_4(i) + \text{M},
\]

\( \Delta H = -439.3 \text{ kJ mol}^{-1} \). \( \text{(16)} \)

Compared to the gas phase, we might anticipate a significant enhancement in the rate of reaction \( \text{(16)} \) in ASM since H atoms are very mobile, even at low temperature, and reaction \( \text{(16)} \) is a highly exothermic, three-body reaction. H atoms are known to be present in high density following pulsed laser excitation of ASM at 90 K, and \( \text{H}_2 \) molecule formation from H atoms has also been reported,\(^{15} \) i.e.,

\[
\text{CH}_3\text{OH} + h\nu \rightarrow \text{CH}_2\text{OH/CH}_3\text{O} + \text{H},
\]

\( \text{(17)} \)

\[
\text{H} + \text{CH}_3\text{OH} \rightarrow \text{CH}_2\text{OH/CH}_3\text{O} + \text{H}_2(v = 0 \text{ and } 1),
\]

\( \text{(18)} \)

\[
\text{H} + \text{H} \rightarrow \text{H}_2(v \geq 2).
\]

\( \text{(19)} \)

Earlier VUV photolysis studies of solid methanol at 10 K found the yield of \( \text{CH}_4 \) to depend linearly on the irradiation time, but no IR absorption signal attributable to \( \text{CH}_3 \) was observed,\(^{15,16} \) probably because these radicals were efficiently consumed by reaction with H atoms, i.e., via reaction \( \text{(16)} \).

In the mixed \( \text{CH}_3\text{OH}/\text{H}_2\text{O} \) ice experiment, the TOF signal intensity of the \( \text{CH}_3 \) radicals was reduced. This may be due to (a) the reduction in \( \text{CH}_3\text{OH} \) density on the surface, or (b) the reaction of \( \text{CH}_3 \) products with H atoms or OH radicals from photolysis of H\(_2\)O. The reaction of \( \text{CH}_3 \) with water molecules is unlikely: the TOF profile of \( \text{CH}_3 \) radicals produced from the mixed ice showed a similar behavior to that from the fresh ASM samples, i.e., the majority of the \( \text{CH}_3 \) products are accommodated to the substrate temperature (90 K). This implies that the \( \text{CH}_3 \) products desorb without reaction with solid water, as would be expected given the large activation energy of reaction \( \text{(20)} \),

\[
\text{CH}_3(i) + \text{H}_2\text{O(ads)} \rightarrow \text{CH}_4(i) + \text{OH}(i),
\]

\( \Delta H = 104.6, \quad E_a = 138.1 \text{ kJ mol}^{-1} \). \( \text{(20)} \)

The activation energy or reaction \( \text{(20)} \) was estimated from theoretical calculations for the reverse process, \( \text{CH}_4+\text{OH} \rightarrow \text{CH}_3+\text{H}_2\text{O} \), by Bravo-Pérez et al.\(^{42} \)

The translational temperatures of the photodesorbed H or OH products following the 157 nm photolysis of water ice have been previously observed with the same REMPI setup.\(^{38,43,44} \) The previous results showed that the majority of the photodesorbed H atoms were relaxed by collisions within the water ice and the translational temperature was accommodated to the substrate temperature. The photodesorbed OH radicals, on the other hand, were not thermally equilibrated with the substrate temperature but were likely formed near the water ice surface, whereas collisionally cooled OH products were either efficiently trapped in the bulk or reacted with water.\(^{38,43,44} \)

**B. OH radical formation in the 157 nm photolysis of fresh ASM**

The translational energies of OH\((v=0 \text{ and } 1)\) products from the 157 nm photolysis of fresh ASM are described by a single temperature \( T_{\text{trans}} = 3000 \pm 500 \text{ K} \) \( (E_{\text{trans}} = 49.9 \pm 8.3 \text{ kJ mol}^{-1}) \). The OH rotational temperatures, \( T_{\text{rot}}(v=0) = 300 \pm 100 \text{ K} \) and \( T_{\text{rot}}(v=1) = 200 \pm 50 \text{ K} \), are not thermally equilibrated with the substrate temperature of 90 K, as shown in Table II. These results suggest that (a) the observed OH\((v=0 \text{ and } 1)\) fragments originate from the ASM surface, not the bulk, and (b) OH\((v=0 \text{ and } 1)\) radicals formed in the ASM bulk are either efficiently trapped or react in the bulk as described before in the 157 nm photolysis of H\(_2\)O.\(^{37,38} \) i.e.,

\[
\text{CH}_3\text{OH} + h\nu \rightarrow \text{CH}_3 + \text{OH}(T_{\text{trans}} = 3000 \text{ K}),
\]

\( \text{(21)} \)

\[
\text{x}_{\text{OH}(3000 \text{ K})} \rightarrow \text{trapped or react with ASM.} \quad \text{(22)}
\]

Possible H abstraction reactions from methanol by OH radicals are

\[
\text{OH}(i) + \text{CH}_3\text{OH(ads)} \rightarrow \text{H}_2\text{O}(i) + \text{CH}_3\text{O}(i),
\]

\( \Delta H = -50.2, \quad E_a = 3.8 \text{ kJ mol}^{-1} \). \( \text{(23)} \)

\[
\text{OH}(i) + \text{CH}_3\text{OH(ads)} \rightarrow \text{H}_2\text{O}(i) + \text{CH}_3\text{O}(i),
\]

\( \Delta H = -25.1, \quad E_a = 14.6 \text{ kJ mol}^{-1} \). \( \text{(24)} \)

The activation energies quoted are for the corresponding gas phase reactions, estimated by Jodkowski et al.\(^{41} \) Reaction \( \text{(23)} \) has been proposed previously in methanol-water condensed phase systems.\(^{45} \) Recalling Table II, the OH products from the 157 nm photolysis of ASM have ample energy to overcome the barriers for reactions \( \text{(23)} \) and \( \text{(24)} \), and thus OH radical initiated hydrogen abstraction is likely to occur on/in ASM.

However, neither photodesorption nor production of \( \text{H}_2\text{O} \) was observed in either the present or earlier VUV photolysis studies on ASM.\(^{15,16} \) These results imply that the yield of OH production via the C–O bond cleavage reaction \( \text{(7)} \) on the surface is small and that the secondary \( \text{H}_2\text{O} \) formation via reactions \( \text{(23)} \) and \( \text{(24)} \) does not proceed efficiently.

**C. Other possible secondary reactions**

After photolysis of a \( \text{CH}_3\text{OH} \) molecule in bulk ASM, the probability of reaction \( \text{(25)} \) is expected to be high since the photogenerated \( \text{CH}_3 \) and OH radicals are effectively caged by the ASM matrix.\(^{17,46,47} \) i.e.,

\[
\text{CH}_3(i) + \text{OH}(i) + \text{M} \rightarrow \text{CH}_3\text{OH(ads)} + \text{M},
\]

\( \Delta H = -422.6 \text{ kJ mol}^{-1} \), \( \text{(25)} \)

the bimolecular reaction between gas phase \( \text{CH}_3 \) and OH radicals has been studied extensively, both theoretically and experimentally.\(^{48} \) The calculations of Jasper et al. show a number of reactive radicals and molecular products from this
reaction, e.g., CH₃OH, HCOH+H₂, and CH₂+H₂O.⁴⁸ In the condensed phase, Hodysy et al. reported a UV (130–335 nm) photolysis study of CH₄–H₂O ice mixtures at 20 K.⁴⁹ Their analysis assumes that OH radicals formed by the photodissociation of water react with methane to form CH₃ radicals, which then recombine with OH radicals to form methanol.

Ethane formation, from the recombination reaction of two CH₃ radicals [reaction (26)], was also reported in the photolysis of CH₄–H₂O ice mixtures,⁴⁹

\[ \text{CH}_3(i) + \text{CH}_3(i) + M \rightarrow \text{C}_2\text{H}_6(i) + M, \]
\[ \Delta H = -376.6 \text{ kJ mol}^{-1}. \]  
(26)

No IR absorption attributable to C₂H₆ was observed in the earlier VUV photolysis studies of solid CH₃OH at 10 K, however.¹⁵,¹⁶ This finding is understandable if CH₃ photoproducts react with other species (e.g., H, OH, or CH₃OH) much more frequently than they encounter another CH₃ radical.

Hodysy et al.⁴⁹ proposed that the presence of CH₄ in CH₄–H₂O ice mixtures inhibited the formation of H₂O₂ (from the combination of two OH radicals) by serving as a trap for OH radicals, since the C–H bonds are also attacked by OH radicals. In fact, no IR absorption features attributable to CH₃OH, HCOH+H₂, and CH₂+H₂O.⁴⁸ In the photolysis of CH₄–H₂O ice mixtures,⁴⁹ proposed that the presence of CH₄ in CH₄–H₂O ice mixtures inhibited the formation of H₂O₂ (from the combination of two OH radicals) by serving as a trap for OH radicals, since the C–H bonds are also attacked by OH radicals. In fact, no IR absorption features attributable to CH₃ radicals have low translational and internal energ-

V. CONCLUSION

The translational and rotational temperatures of CH₃(ν =0) and OH(ν=0 and 1) radicals (i.e., the products from the primary C–O bond cleavage) have been measured following pulsed 157 nm irradiation of ASM at 90 K. Most of the detected CH₃ species have low translational and internal energies, implying that these photoproducts accommodate to the ASM bulk temperature and desorb from the ASM surface without reaction in the bulk near the solid/vacuum interface. The detected OH fragments, in contrast, are deduced to originate exclusively from the ASM surface; any OH radicals formed in the ASM bulk phase are readily trapped or react in the bulk phase. The equivalence of the TOF spectra of CH₃ fragments from ASM and from a mixed CH₃OH–H₂O ice sample indicates that the reactivity of solid water with CH₃ products is small.

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²⁶ C. M. Western, robmap, a program for simulating rotational structure, University of Bristol, available at http://pppophchem.bris.ac.uk.