ISOTACHOPHORES FOR THE ANALYSIS OF URINARY TRACT STONE: A PRELIMINARY REPORT

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ISOTACHOPHORESIS FOR THE ANALYSIS OF URINARY TRACT STONE: A PRELIMINARY REPORT

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ABSTRACT

Newly devised isotachophoretic apparatus was used for the analysis of urinary tract stones and was found to be useful for rapid and simple quantitative analysis.

INTRODUCTION

It is important to determine the composition of urinary tract stone for understanding the mechanism of stone formation and preventing recurrence. Infrared spectroscopy, X-ray defraction, optical crystallographic analysis and other methods have been used for this purpose. These methods are essentially qualitative or semi-quantitative. Chemical analysis of stones is not easy because it required somewhat larger specimen and the sample must be separately processed to determine each component such as calcium, magnesium, oxalate, phosphate, urate, and so on. A sensitive, rapid, simultaneous, easy and inexpensive method is desired for the quantitative analysis of stone. In order to accomplish this purpose a newly developed isotachophoresis was applied for stone analysis.

MATERIALS AND METHODS

The isotachophoretic analyser used (Fig. 1) is the Shimazu IP-1B, equipped with a PGD-1 potential gradient detector and a recorder (Shimazu Seisakusho, Ltd., Kyoto, Japan). The principle of isotachophoresis is shown on Fig. 2. This is an electrolytic analysis performed in the fluid phase. Leading electrolyte and terminal electrolyte are filled in reservoirs. The reservoirs contain respectively an anode and a cathode.
An electrophoretic capillary tube is connected between the reservoirs. The tube is 0.5 mm in diameter and 20 cm in length and controlled at 20°C of the temperature. Microanalysis is one of the benefits of this method. Usually 10 μl of the sample is introduced into the sample tap with a microsyringe.

In our study 0.01 M histidine and 0.01M potassium acetate (pH 5.4) was used as the leading electrolyte and 0.01M tris acetate (pH 5.0) was used as the terminal electrolyte for the separation of cations. For the anions a mixture of 0.01M histidine and 0.01N histidine-HCl (pH 6.02) was used as the leading electrolyte and 0.01M caproic acid as the terminal electrolyte. Chemical agents for leading and terminal electrolytes, and all the standard materials were used in the highly purified form commercially available.

Stones were washed with water to remove blood and other adhering material. Dried and sectioned with a blade and powdered. One mg of the powder was dissolved with 10 ml of 0.01M HCl and 20–50 μl of the sample solution were applied to the analyser.

The electric current was stabilized at 75 μA for both cation analysis and anion analysis.

Ions were detected at the end of the electrophoretic capillary. They were identified by the ratio of potentials.

\[
\text{potential ratio} = \frac{\text{terminal potential} - \text{potential of the sample ion}}{\text{terminal potential} - \text{leading potential}}
\]

Zone length of ion correlates with the amount of the ion. As the ion moves in the same speed, the passing time of ion on the detector correlates with the amount of the ion.

**RESULTS**

Calcium, magnesium, oxalic acid, citric acid and phosphatic acid were satisfactorily separated and calculated. Potential gradient curves and standard curves were shown in Fig. 3 and Fig. 4 respectively. Satisfactory reproducibility was obtained.

Linear relationship is observed between the passing time and the amount of ions introduced. Ammonium and carbonic ions, which are suggested as occasional components of urinary stones, escaped under the condition of the study. Uric acid moved more slowly than the terminal caproic acid and formed a round curve in the terminal ion.

The mobility of cystine and xanthine, very rare components of urinary stones, was smaller than that of uric acid and they could not be detected with the histidine-HCl and caproic acid system.

To identify and calculate the latter three substances 0.01M HCl adjusted to pH 8.5 with 2-amino-2-methyl-1, 3-proamediol was used as the leading electrolyte and 0.01M β-alanine adjusted to pH 10.8 with Ba (OH)₂ was, after filtration, used as the terminal electrolyte. Although the three

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**Fig. 3.** Potential gradient curves of main urinary stone components. Integral and differential curves are shown.

*Left: cation analysis*
- leading electrolyte: 0.01 M histidine and 0.01 M potassium acetate (pH 5.4)
- terminal electrolyte: 0.01 M Tris acetate (pH 5.0)

*Right: anion analysis*
- leading electrolyte: 0.01 M histidine and 0.01 M histidine-HCl (pH 6.02)
- terminal electrolyte: 0.01 M caproic acid
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Fig. 4. Standard curves. The length of the zone, passing time in turn, well correlate with the amounts of ion.

Fig. 5. Summary of analysis on 58 urinary stones. Mean value obtained from 58 urinary stones is shown.

**DISCUSSION**

The principle of isothacophoresis was anologued to displacement chromatogram. In anion analysis, for example, all anions move towards the anode. In moving to the anode, the anions arrange themselves in order of mobility and once the arrange has been completed all the anions move in the same concentration and in the same speed. The passage of each zone can be detected at the end of the capillary tube by their potential gradient, heat production or ultraviolet absorbance.

On the basis of our observation this method is useful for stone analysis because it is simple and rapid. The time required for a separation is about 20 minutes. It is well known that the composition of urinary stone differs in portion to portion. This method can deal with small specimens taken from any desired portion of the stone. When the stone is throughly powdered and mixed, mean value of the components can be easily obtained.

The electrolyte solution which completely satisfy our purpose has not been found. According to Beckers and associates the step-height of uric acid is lower than that of cacodylic acid in using histidine-HCl as leading electrolyte and a thermometer as detector. The use of cacodylic acid for terminal electrolyte, however, failed to separate uric acid. Conversely, it was easily identified as a sloped peak in the terminal caproic acid. Caproic acid was therefore used as the terminal electrolyte in the anion analysis. Although the β-alanine-Ba(OH)$_2$ system can separate uric acid, cystine and xanthine between the boundry of the leading and terminal electrolyte, difficulties arise because of interactions of fast-moving ions. The system is therefore can not be recommended for the analysis of oxalic acid and phosphatic acid.

Chemical analysis of urinary tract stones is not easy because each component must be estimated separately. An easy, simultaneous quantititative analytical method is desirable. It is the reason we tried to apply isothacopheretic method for urinary stone analysis. The method could not satisfy our purpose completely. All the possible components of urinary stones could not be detected. Ammonium and carbonate ions escaped from the detection. Stone constituents were not detected in their crystalline form. From these points isothacophoresis does not replace the existing methods.

However, it should be mentioned that
the analysis is performed in the fluid phase. It may be useful in experimental studies on the relationship between urinary constituents and precipitated stones such as crystallization from synthetic urine. Important constitutions of urinary stones can be simultaneously analysed.

It is found that the addition of a non-ionic detergent increases the sharpness of the the zone boundaries\(^1\). Further improvement of the present equipment would bring down an excellent tool for the study on urinary stone.

REFERENCES


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