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<td>Hama, Tetsuya; Yokoyama, Masaaki; Yabushita, Akihiro; Kawasaki, Masahiro</td>
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Kyoto University
Role of OH radicals in the formation of oxygen molecules following vacuum ultraviolet photodissociation of amorphous solid water

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Photodesorption of O2(X 3Σg−) and O2(a 1Δg) from amorphous solid water at 90 K has been studied following photoexcitation within the first absorption band at 157 nm. Time-of-flight and rotational spectra of O2 reveal the translational and internal energy distributions, from which production mechanisms are deduced. Exothermic and endothermic reactions of OH+O(3P) are proposed as plausible formation mechanisms for O2(X 3Σg−) and a 1Δg). To examine the contribution of the O(3P)+O(3P) recombination reaction to the O2 formation following 157 nm photolysis of amorphous solid water, O2 products following 193 nm photodissociation of SO2 adsorbed on amorphous solid water were also investigated. © 2010 American Institute of Physics.

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I. INTRODUCTION

The effect of radiation on water ice has intrigued many scientists in the fields of interstellar chemistry and planetary ice science as well as reaction dynamics since water is the predominant component of interstellar icy grain mantles in dense molecular clouds and small solar system bodies such as comets.1−12 Oxygen molecule is known to be a product when water ice is irradiated with photons, electrons, or with energetic ions.3−6 Westley et al.7,8 observed desorption of H2, O2, and H2O by a quadrupole mass spectrometer (QMS) during vacuum ultraviolet (VUV) irradiation of water ice at 35−100 K with mainly Lyman−α photons (λ=121 nm). Öberg et al.9 used a VUV lamp (118 nm ≤ λ=177 nm) to irradiate water ice at 18−100 K, and detected OH, H2, O2, and H2O as desorbing species by QMS. These experimental studies indicate that VUV photodissociation of amorphous solid water (ASW) leads to secondary reactions that result in molecular oxygen formation on or in ASW.

Various primary processes are energetically possible following photoexcitation of water ice at λ ≤ 130 nm,10−12 whereas photodissociation of H2O in the first absorption band of water ice (130−165 nm) involves mainly two primary processes,10,13,14

\[ \text{H}_2\text{O} + h\nu \rightarrow \text{H} + \text{OH}, \quad (1) \]
\[ \text{H}_2\text{O} + h\nu \rightarrow \text{H}_2 + \text{O}(1\text{D}). \quad (2) \]

Measurements of the translational and internal energy distributions of the photoproducts generated from photodissociation of ASW allow the assessment of possible secondary reactions on/in ASW from reaction dynamics point of view. The translational and internal energy distributions of OH(v=0,1) radicals and O(1D) and (3P) atoms were previously measured following 157 nm photodissociation of ASW.15−17 For the O(3P) production, two different formation mechanisms were proposed: the exothermic recombination reaction of OH and the photodissociation of OH on the ASW surface.17

\[ \text{OH} + \text{OH} \rightarrow \text{H}_2\text{O} + \text{O}(3\text{P}), \quad \Delta H = -67 \text{ kJ mol}^{-1}, \quad (3) \]
\[ \text{OH(ads)} + h\nu \rightarrow \text{H} + \text{O}(3\text{P}). \quad (4) \]

Thermodynamic data are calculated using solid phase data for the condensed or adsorbed species (“ads”) and the gas phase data for other species.18−20 From the fact that OH(v=0 and 1) and O(3P) are formed with large excess energy via reactions (1), (3), and (4),15,17 O2(X 3Σg−) and O2(a 1Δg) can be produced via subsequent reactions (5) and (6) in the 157 nm photolysis of ASW,

\[ \text{OH}(v=1) \text{ has a sufficient vibration energy of 43 KJ mol}^{-1} \text{ to proceed in reaction (6).22 In fact, Lunt et al.23 observed the formation of O}_2(a 1\Delta_g) \text{ via the reaction of O}(3\text{P}) \text{ and vibrationally excited OH(v=1)} \text{ in the gas phase by monitoring the near-infrared emission at 1.27 \mu m from the product O}_2(a 1\Delta_v). \]

Reaction (5) is a barrierless process in the gas phase,21

\[ \text{OH}(v=1) \text{ has a sufficient vibration energy of 43 KJ mol}^{-1}. \quad (5) \]
\[ \text{OH} + \text{O}(3\text{P}) \rightarrow \text{O}_2(X 3\Sigma_g^-, v=0) + \text{H}, \quad -68. \quad (6) \]

In the present work, we have investigated the kinetic and internal energy distributions of O2(X 3Σg−, v=0) and O2(a 1Δg, v=0) following 157 nm photodissociation of ASW at 90 K using the resonance-enhanced multiphoton ionization (REMPI) method. H2O2 photolysis experiments have also been performed at 157 nm to elucidate the role of OH in the O2 formation. In addition, we have also studied desorption of O2 following 193 nm photolysis of SO2 adsorbed on ASW to measure the kinetic and internal energy

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b)Author to whom correspondence should be addressed. Electronic mail: kawasaki@moleng.kyoto-u.ac.jp. FAX: +81-75-383-2572.
TABLE I. Spectroscopic constants for the upper \(d^3\Pi_u\) and \(C^3\Pi_u\) Rydberg states, and the lower \(X^3\Sigma_g^-\) states of \(O_2\). [The states of origin for \(d^3\Pi_u(v'=2)\) and \(C^3\Pi_u(v'=3, 4, \text{ and } 5)\) were taken from Ref. 35. The values of state origins for \(X^3\Sigma_g^-(v''=2 \text{ and } 3)\) were taken from Refs. 36 and 37.]

<table>
<thead>
<tr>
<th>State origin ((v', \Omega))</th>
<th>Rotational constant (B) ((\text{cm}^{-1}))</th>
<th>Centrifugal distortion constant (D) ((10^{-6} \text{ cm}^{-1}))</th>
<th>Reference</th>
</tr>
</thead>
<tbody>
<tr>
<td>(C^3\Pi_u(v=2), F_1(\Omega=0))</td>
<td>69 353</td>
<td>1.633</td>
<td>1.5</td>
</tr>
<tr>
<td>(C^3\Pi_u(v=2), F_3(\Omega=1))</td>
<td>69 437</td>
<td>1.665</td>
<td>1.0</td>
</tr>
<tr>
<td>(C^3\Pi_u(v=2), F_5(\Omega=2))</td>
<td>69 548</td>
<td>1.685</td>
<td>1.3</td>
</tr>
<tr>
<td>(d^3\Pi_u(v=2))</td>
<td>70 011</td>
<td>1.29</td>
<td></td>
</tr>
<tr>
<td>(X^3\Sigma_g^-(v=1))</td>
<td>70 142</td>
<td>1.12</td>
<td></td>
</tr>
<tr>
<td>(X^3\Sigma_g^-(v=0))</td>
<td>70 156.3</td>
<td>1.422</td>
<td>4.840</td>
</tr>
</tbody>
</table>

Δ\(H\) (kJ mol\(^{-1}\))

\[
O(3P) + O(3P) + M \rightarrow O_2(X \, 3\Sigma_g^-) + M, \quad -498. \quad (7)
\]

II. EXPERIMENTAL

A. Apparatus and preparation of ice

A photodissociation study of ASW at 90 K was carried out in an ultrahigh vacuum chamber, which was equipped with two turbomolecular pumps in tandem, a pulsed molecular beam, an excimer laser, and a dye laser. Experimental details are described elsewhere. For the concentrated \(H_2O_2\) photolysis experiments, a commercially available \(H_2O_2\) solution was deposited on the ASW substrate. For the concentrated \(H_2O_2\) photolysis experiments, a commercially available \(H_2O_2\) solution was used for the laser irradiation area to have the disk substrate with a radius of 6 mm. The incident fluence \(F\) was typically \(<0.1\) mJ cm\(^{-2}\) pulse\(^{-1}\) at 157 nm in 15 ns pulse duration. The \(F\) value at 193 nm was \(<1.0\) mJ cm\(^{-2}\) pulse\(^{-1}\). 193 nm irradiation of ASW produced no measurable REMPI signals of \(O_2(X \, 3\Sigma_g^-)\) and \(OH\) at such low incident intensities, although \(H\) atom photofragments from dimerlike water molecules on ASW surface were observed following 193 nm photodissociation of ASW. Photodesorbed \(O_2(X \, 3\Sigma_g^-)\) were ionized at a vertical distance of 2 mm from the ice surface by 2+1 REMPI process via the \(e^3\Pi_u(v'=2) \rightarrow X \, 3\Sigma_g^-(v''=0)\) at 283–289 nm produced by frequency doubling the output of a Nd\(^3+:YAG\) (yttrium aluminum garnet) pumped dye laser (Lambda Physik, Göttingen, SCANmate), and collected with a small mass spectrometer aligned perpendicular to the ice surface. Electronically excited \(O_2\) \((a \, 1\Delta_g, v=0)\) photodesorbed from ASW has also been observed with the same experimental setup via the \(d^3\Pi_u(v=1) \rightarrow a \, 1\Delta_g(v'=0)\) REMPI transitions at 329–332 nm. Johnson III \(et\, al.\) \(31\) suggested that the REMPI cross section for \(O_2(a \, 1\Delta_g)\) is at least an order of magnitude greater than that for \(O_2(X \, 3\Sigma_g^-)\). Photodesorbed \(O(3P_{1/2,1/2})\) atoms were ionized by the 2+1 REMPI transition via the \(O(3P_{3/2}) \rightarrow \lambda\) transition at 225.6–226.4 nm. \(17\) We tried to search for REMPI signals of oxygen molecules in the higher electronically excited states, i.e., \(b \, \Sigma_g^+, c \, \Sigma_g^-, \text{ and } 3\Pi_g\), using the REMPI transitions reported by Morrill \(et\, al.\) \(32\) and Slanger and Copeland, \(33\) but no evidence for these products was obtained.

III. SPECTRAL ANALYSIS

A. Simulation of 2+1 REMPI spectra of oxygen molecules

1. \(O_2(X \, 3\Sigma_g^-, v=0, 1, 2, 3)\)

The spectra were simulated using the program PGOPHER, \(34\) a program for simulating the rotational structure accompanying multiphoton electronic transitions. The simulated spectra were used to estimate the values of the rotational temperatures of the oxygen products. Spectroscopic parameters for the \(X \, 3\Sigma_g^-(v''=0, 1)\) and \(C^3\Pi_u(v'=2)\) states of \(O_2\) are summarized in Table I. \(30,35\)–\(39\) For spectral simulations of 2+1 REMPI via the sequence bands of the distributions of \(O_2\) formed via highly exothermic recombination reaction of two \(O(3P)\), i.e., reaction (7), which was proposed as a source for \(O_2(X \, 3\Sigma_g^-)\) formation in VUV photon and electron irradiations of ASW.\(^{8,24}\)
O_{2}(a^1\Delta_g, v=0) were fitted to a sum of flux-weighted Maxwell–Boltzmann (MB) distributions, $S_{\text{MB}}$, with coefficients $a_i$, which are defined by a translational temperature, $T_{\text{trans}}$. Details regarding the simulation of such TOF spectra have been reported previously.\(^{25}\)

$$S(t, T_{\text{trans}}) = \sum a_i S_{\text{MB}}(t, T_{\text{trans}}),$$ \hspace{1cm} (8)

$$S_{\text{MB}}(t, T_{\text{trans}}) = r^3 t^{-4} \exp[-mr^2/2k_B T_{\text{trans}}t^2],$$ \hspace{1cm} (9)

where $r$ is the flight distance to the REMPI probe region. In these calculations we assume that signals come from a disk area (VUV photolysis area) with a radius of 6 mm. The MB distribution, $P_{\text{MB}}(E_{\text{trans}})$, as a function of translational energy, $E_{\text{trans}}$, is characterized by the averaged translational energy, $\langle E_{\text{trans}} \rangle = 2k_B T_{\text{trans}}$.\(^{40}\)

$$P_{\text{MB}}(E_t) = (k_B T_{\text{trans}})^{-2}E_t \exp[-E_t/k_B T_{\text{trans}}].$$ \hspace{1cm} (10)

Conversion from the $E_t$ distribution to the TOF distribution was performed using the Jacobian given by Zimmerman and Ho.\(^{41}\) The angular distribution of the molecules photodesorbed from the ice surface was assumed to follow a cos$^n \theta$ function, where $\theta$ is the surface polar coordinate. $n=0$ was adopted in the present best-fit procedures.\(^{25,42}\)

IV. RESULTS

A. Kinetic and rotational energy distributions of O$_2$ following 157 nm photodissociation of ASW

Figure 1(a) shows a 2+1 REMPI spectrum following 157 nm photolysis of ASW at a fixed delay $t=1.5$ $\mu$s that corresponds to the peak TOF, as shown in Fig. 2. Highly perturbed rotational structures were observed in Fig. 1(a). By comparison with spectral simulation for O$_2$(X $^3\Sigma_g^-$, $v=0$) shown in Fig. 1(c), $T_{\text{rot}}$ was estimated to be $\sim 2000$ K at $t = 1.5$ $\mu$s. Figure 2 shows a typical TOF spectrum at 286.127 nm, which is characterized by $T_{\text{trans}}=4500 \pm 1000$ K and $500 \pm 200$ K. These results are summarized in Table II. Changing the REMPI probe wavelength to other wavelengths, no discernible change in the TOF profiles was ob-
TABLE II. Translational and rotational temperatures of oxygen molecules following 157 nm photolysis of ASW and 193 nm photolysis of SO2 on ASW at 90 K.

<table>
<thead>
<tr>
<th>Molecular State</th>
<th>Translational Temperature $T_{\text{trans}}$ (K)</th>
<th>Averaged Translational Energy $\langle E_{\text{trans}} \rangle$ (kJ mol$^{-1}$)</th>
<th>Rotational Temperature $T_{\text{rot}}$ (K)</th>
<th>Averaged Rotational Energy $\langle E_{\text{rot}} \rangle$ (kJ mol$^{-1}$)</th>
<th>Reaction Mechanism</th>
</tr>
</thead>
<tbody>
<tr>
<td>$O_2(X^3\Sigma_g^-, v=0)$ from ASW</td>
<td>4500 ± 1000 (70%)</td>
<td>75 ± 17</td>
<td>~2000$^b$</td>
<td>~17</td>
<td>OH + O($^3P$)</td>
</tr>
<tr>
<td></td>
<td>500 ± 200 (30%)</td>
<td>8 ± 3</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>$O_2(a^1 \Delta_g, v=0)$ from ASW</td>
<td>2500 ± 500$^c$</td>
<td>42 ± 8</td>
<td></td>
<td></td>
<td>OH + O($^3P$)</td>
</tr>
<tr>
<td></td>
<td>250 ± 100$^d$</td>
<td>4 ± 2</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>$O_2(X^3 \Sigma_g^-, v=1–3)$ from SO2 on ASW$^e$</td>
<td>3000 ± 1000 (55%)</td>
<td>50 ± 17</td>
<td>~8000$^b$</td>
<td>~67</td>
<td>O($^3P$) + O($^3P$)</td>
</tr>
<tr>
<td></td>
<td>500 ± 200 (45%)</td>
<td>8 ± 3</td>
<td></td>
<td></td>
<td></td>
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</tbody>
</table>

$^a$Percentage in the parenthesis is a contribution of each temperature component.
$^b$Measured at time-of-flight=1.5 $\mu$s.
$^c$Each contribution depends on REMPI wavelength. See text for details.
$^d$Not estimated. See text for details.
$^e$With the relative ratios of 0.2: 1: 1 for $v=1: 2: 3$.

We were unable to estimate a rotational temperature for the $T_{\text{trans}}=500$ K component because of weak REMPI spectrum intensity.

2. $O_2(a^1 \Delta_g, v=0)$

Figures 3(a) and 3(b) show REMPI spectra following 157 nm photolysis of ASW, which were recorded at (a) $t=2.0$ $\mu$s and (b) 10.0 $\mu$s. Figure 4 shows TOF spectra at the two different REMPI wavelengths, (a) 329.965 nm for a high rotational level, and (b) 331.215 nm for a low rotational level. The TOF spectra are reproduced by two MB distributions with $T_{\text{trans}}=2500 \pm 500$ K and 250 $\pm 100$ K with different contributions: (a) 80% and 20% at 329.965 nm and (b) 15% and 85% at 331.215 nm, respectively.

B. Evolution of $O_2$ signal intensity as a function of 157 nm photoirradiation time in the ASW photolysis

1. $O_2(X^3 \Sigma_g^-, v=0)$

The solid black line in Fig. 5(a) shows the time evolution of the $O_2(X^3 \Sigma_g^-, v=0)$ signal intensity at 286.127 nm recorded at $t=1.5$ $\mu$s as a function of 157 nm irradiation time on ASW. For comparison purpose, Fig. 5(a) includes the time evolution of photodesorbed O($^3P_2$) signal intensity that was recorded at the peak of TOF, i.e., $t=4.5$ $\mu$s. The behaviors of $O_2(X^3 \Sigma_g^-, v=0)$ and O($^3P_2$) signals are in accordance to each other. Previously reported time evolution of the OH signal due to the H$_2$O$_2$ photoprodut on the 157 nm photolysis ASW surface was also included in Fig. 5(a), which reflects the concentration of photogenerated H$_2$O$_2$ on the ice surface.

2. $O_2(a^1 \Delta_g, v=0)$

Figures 5(b) and 5(c) show time evolution curves of $O_2(a^1 \Delta_g, v=0)$, which were recorded at (b) $t=2.0$ $\mu$s for $O_2(T_{\text{trans}}=2500$ K) at a high rotational level REMPI wavelength of 329.965 nm, and (c) $t=10.0$ $\mu$s for $O_2(T_{\text{trans}}=250$ K) at a low rotational level REMPI wavelength of 331.215 nm. Figures 5(b) and 5(c) also show the time evolution of the secondary photoprodut H$_2$O$_2$. The curve for $O_2(T_{\text{trans}}=2500$ K) matches to that for H$_2$O$_2$. In addition, $O_2(T_{\text{trans}}=2500$ K) signals were observed promptly after 157 nm irradiation of H$_2$O$_2$/H$_2$O mixed ices, suggesting that the source of $O_2(T_{\text{trans}}=250$ K) is photochemically produced H$_2$O$_2$ by 157 nm irradiation on ASW.

On the other hand, the $O_2(T_{\text{trans}}=250$ K) component appeared promptly upon 157 nm irradiation of ASW, as shown in Fig. 5(c). These results suggest that the source of the.

FIG. 3. 2+$u$ REMPI spectrum of $O_2(a^1 \Delta_g, v=0)$ via the $d^3 \Pi_g(u'=1) \rightarrow a^1 \Delta_g(v''=0)$ transition following 157 nm photolysis of ASW. (a) $t=2.0$ $\mu$s and (b) $t=10.0$ $\mu$s. The arrows indicate the lines used when measuring the time-of-flight spectra.

FIG. 4. Time-of-flight spectra of $O_2(a^1 \Delta_g, v=0)$ following 157 nm photolysis of ASW at (a) 329.965 nm and (b) 331.215 nm. The solid curves are fits to the data derived with $T_{\text{trans}}=2500$ and 250 K.
Oxygen molecules following photolysis of ice


C. O₂ products following 193 nm photolysis of SO₂/H₂O mixed ice

193 nm irradiation of neat ASW produced no measurable REMPI signals of O₂(X ³Σg⁻, v=0) and O₂(a ¹Δg, v=0). However, as SO₂ was deposited on ASW, m/z=32 signal was observed at 285.038 nm and its intensity increased, as shown in Fig. 6, i.e., O₂ was photodesorbed via reaction (7) following 193 nm photolysis of adsorbed SO₂ on ASW. Figure 1(b) shows a REMPI spectrum of O₂ following the 193 nm photoradiation at t=1.5 μs that corresponds to the peak TOF of O₂ in Fig. 7. The TOF spectrum in Fig. 7 is characterized by T_{trans}=3000±1000 K and 500±200 K. These results are summarized in Table II. O₂(a ¹Δg, v=0) failed to be detected at 329–332 nm.

V. DISCUSSION

The 157 nm photolysis involves two primary processes; reactions (11) and (12),

H₂O(ads) + hν(157 nm) → H + OH(v=0 and 1), (11)

H₂O(ads) + hν(157 nm) → H₂ + O(¹D). (12)

In the case of 157 nm photolysis (E_{phot}=762 kJ mol⁻¹), the available energies for reaction (11) is E_{avail}(11) = 220 kJ mol⁻¹ for OH(v=0), and E_{avail}(12) = 38 kJ mol⁻¹. O(¹D) atoms via reaction (12) were successfully detected following 157 nm photolysis of ASW, but would play only a minor role in the present study, considering the small quantum yields (≤0.01 for 145–185 nm) in the gas phase, and high reactivity of O(¹D) with parent H₂O molecules by collisions with ASW to produce OH or H₂O₂.

Our previous study showed that O(¹P) atoms were photodesorbed immediately after 157 nm irradiation started, and are mainly formed by recombination of two OH radicals moving on the ASW surface, i.e., reaction (13),

OH + OH → H₂O + O(¹P), ∆H = -67 kJ mol⁻¹. (13)

Considering reactions (11) and (13), we will discuss three possible O₂ desorption mechanisms; exothermic and endo-
thermic reactions of OH with O(3P), i.e., reactions (14) and (15), and two O(3P) recombination reaction, i.e., reaction (16),

$$\Delta H \ (kJ \ mol^{-1})$$

$$\begin{align*}
\text{OH} + \text{O}(3P) & \rightarrow \text{O}_2(X^3\Sigma_g^-) + \text{H}, \quad -68, \quad (14) \\
\text{OH} + \text{O}(3P) & \rightarrow \text{O}_2(a^1\Delta_g) + \text{H}, \quad -26, \quad (15) \\
\text{O}(3P) + \text{O}(3P) & \rightarrow \text{O}_2(X^3\Sigma_g^+) + \text{M}, \quad -498. \quad (16)
\end{align*}$$

A. \(O_2(X^3\Sigma_g^-, v=0)\) formation mechanisms

Figure 5(a) shows that the temporal signal behaviors of \(O_2(X^3\Sigma_g^-, v=0)\) and \(O(3P)\) are in fair agreement to each other, while OH signal intensity due to the \(H_2O_2\) photoproducts on the 157 nm photoradiated ASW surface increased slowly. These results indicate that \(O_2(X^3\Sigma_g^-, v=0)\) are not produced from the accumulated \(H_2O_2\) photoproducts on the ASW surface.

On the assumptions that (a) a large amount of OH radicals are produced on the ASW surface after 157 nm laser shots on ASW, allowing successive surface reactions on the ASW surface to occur efficiently, and (b) the concentration of OH on ASW fed by primary photodissociation reaction (11) is much higher than that of \(O(3P)\) formed via secondary reaction (13), the appearance behavior of \(O_2(X^3\Sigma_g^-, v=0)\) signal intensity in Fig. 5(a) can be explained if the reaction schemes (17)–(21) proceeds on ASW,

$$\begin{align*}
\text{H}_2\text{O} + h\nu (157 \text{ nm}) & \rightarrow \text{H} + \text{OH}(v = 0 \text{ and 1}), \quad (17) \\
\text{OH} + \text{OH} & \rightarrow \text{H}_2\text{O}_2, \quad (18) \\
\text{OH} + \text{OH} & \rightarrow \text{H}_2\text{O} + \text{O}(3P), \quad (19) \\
\text{OH} + \text{O}(3P) & \rightarrow \text{O}_2(X^3\Sigma_g^-, v = 0) + \text{H}, \quad (20) \\
\text{H}_2\text{O}_2 + h\nu (157 \text{ nm}) & \rightarrow 2\text{OH}(v = 0 \text{ and 1}). \quad (21)
\end{align*}$$

Furthermore, after successive 157 nm irradiation on ASW, OH radicals are accumulated on ASW since they can exist on/in water ice at 90 K.\(^3\) These adsorbed OH on ASW can be another sources for \(O(3P)\) via reactions (19) and (22),\(^17\) and contribute to the \(O_2(X^3\Sigma_g^-, v=0)\) production, as shown in Fig. 5(a),

$$E_{\text{avail}} \ (kJ \ mol^{-1})$$

$$\begin{align*}
\text{OH}(\text{ads}) + h\nu (157 \text{ nm}) & \rightarrow \text{H} + \text{O}(3P), \quad 278. \quad (22)
\end{align*}$$

The translational temperature of \(O_2(X^3\Sigma_g^-, v=0)\) is much higher than the substrate temperature of 90 K, as shown in Table I. These results suggest that observed \(O_2(X^3\Sigma_g^-, v=0)\) originate only from the ASW surface, not from the bulk. Kimmel and their co-workers\(^35\) suggested that O₂ formation occurred at or near the ASW/vacuum interface in the low-energy electron-stimulated desorption experiments. The electron-stimulated migration of OH or OH⁻ to the vacuum interface, where they react and produce O₂, occurs via transport through the hydrogen bond network of the ASW.\(^35\)

Arasa \textit{et al.} performed molecular dynamics simulations of the UV photodissociation of ASW at 10–90 K. They reported that OH desorbs into the vacuum only when they were generated from the top 3 ML of ASW at 90 K and its probability is <0.05 per adsorbed UV photon, whereas most of OH are trapped on/in ASW.\(^5\) Andersen \textit{et al.}\(^57\) predicted that OH radicals in bulk ASW formed following photodissociation of ASW move only less than 0.5 nm, while OH radicals formed in the top 3 ML of ASW were able to move up to more than 6 nm until they were trapped. This large mobility of OH on the top ASW makes secondary reactions of OH with adsorbed species possible. These models support our proposed reaction schemes.

B. \(O_2(a^1\Delta_g, v=0)\) formation mechanisms

The time evolution curve for \(O_2(a^1\Delta_g, T_{\text{trans}}=250 \text{ K})\) recorded at \(t=10.0 \ \mu\text{s}\) appears promptly, as shown in Fig. 5(c). Furthermore, Fig. 5(d) shows that \(O_2(a^1\Delta_g, v=0), O(3P_2)\), and \(O_2(X^3\Sigma_g^-, v=0)\) were photodesorbed immediately after the 157 nm irradiation started. Since vibrationally excited \(OH(v=1)\) are produced from primary photodissociation of \(H_2O\),\(^15\) \(O_2(a^1\Delta_g, T_{\text{trans}}=250 \text{ K})\) can be produced via reaction of vibrationally excited \(OH(v=1)\) and \(O(3P)\), as Lunt \textit{et al.} observed in the gas phase, i.e., reaction (23).\(^23\)

$$\Delta H \ (kJ \ mol^{-1})$$

$$\begin{align*}
\text{OH}(v = 1) + \text{O}(3P) & \rightarrow \text{O}_2(a^1\Delta_g, v = 0) + \text{H}, \quad -17. \quad (23)
\end{align*}$$

On the other hand, Fig. 5(b) shows that \(O_2(a^1\Delta_g, T_{\text{trans}}=2500 \text{ K})\) was not observed immediately after 157 nm photoradiation, and the appearance behavior is in accordance with the concentration of photogenerated \(H_2O_2\) on the ice surface. Furthermore, the \(O_2(a^1\Delta_g, T_{\text{trans}}=2500 \text{ K})\) was observed promptly after 157 nm irradiation of the \(H_2O_2/H_2O\) mixed ices. These results suggest that \(H_2O_2\) photoproduction on ASW is a source for the \(O_2(a^1\Delta_g, T_{\text{trans}}=2500 \text{ K})\) formation. Hama \textit{et al.}\(^15\) previously reported that \(OH(v=0 \text{ and 1}, T_{\text{trans}}=7500 \text{ K})\) are produced from the secondary photodissociation of \(H_2O_2\) that is photolytically accumulated on the ASW surface by 157 nm irradiation, i.e., reaction (24),

$$\begin{align*}
\text{H}_2\text{O}_2(\text{ads}) + h\nu (157 \text{ nm}) & \rightarrow 2\text{OH}(v = 0 \text{ and 1}, T_{\text{trans}}=7500 \text{ K}), \quad (24)
\end{align*}$$

where \(E_{\text{avail}}(24)=493 \text{ kJ mol}^{-1}\). These \(\text{OH}(v=0 \text{ and 1}, T_{\text{trans}}=7500 \text{ K})\) can proceed in endothermic reactions on ASW surface.\(^16\) Thus, \(O_2(a^1\Delta_g, T_{\text{trans}}=2500 \text{ K})\) can be produced from the hot photoproducots such as \(OH\) from \(H_2O_2\) via reaction (25),

$$\begin{align*}
\text{OH}(v = 0 \text{ and 1}, T_{\text{trans}}=7500 \text{ K}) + \text{O}(3P) & \rightarrow \text{O}_2(a^1\Delta_g, T_{\text{trans}}=2500 \text{ K}) + \text{H}. \quad (25)
\end{align*}$$

Lunt \textit{et al.}\(^23\) reported 0.025 for the yield of \(O_2(a^1\Delta_g\text{) formation via reaction of } \text{OH}(v=1) \text{ with } O(3P) \text{ in the gas phase}, \) and the majority of the products are \(O_2(X^3\Sigma_g^-)\). Lique \textit{et al.}\(^15\) pointed out a possiblility of the nonadiabatic transition for the \(OH+O\) reaction. The potential energy surface of \(OH+O\) reaction corresponding to the electronically excited 2A¹ state correlates only with the \(O_2(a^1\Delta_g) + \text{H} \text{ product channel through the electronically excited state of the}$$
O2(2A′) complex, but nonadiabatic electronic relaxation of this complex to the lowest 2A'' state can provide O2(X 3Σg−) +H product channel.21 Our previous observations of vibrational distribution of OH(v=1/v=0) produced via reactions (17) and (24) were both 0.2 ± 0.1.15 These previously reported results suggest that the relative product ratio of O2(a 1Δg,v=0) would be small in the present study.

C. Other possible O2 formation mechanisms

To consider a contribution of highly exothermic recombination reaction of two O(1P), we measured the translational and internal energy distributions of O2 following 193 nm photoexcitation of SO2 adsorbed on ASW since O(1P) were efficiently produced from the 193 nm photolysis of SO2.49 As shown in Figs. 1(b), 6, and 7, 193 nm photolysis of SO2 on ASW yielded O2 products via subsequent reactions (26) and (27) and they were desorbed into the vacuum,

SO2(ads) + hν(193 nm) → SO + O(1P), (26)

\[ \Delta H \text{ (kJ mol}^{-1} \text{)} \]

O(1P) + O(1P) + M → O2(X 3Σg−) + M, \[ -498. \] (27)
The available energy for reaction (26) at 193 nm (E_{\text{photon}} = 620 \text{ kJ mol}^{-1}) was calculated using dilute solution phase thermodynamic data for the adsorbed SO2(ads).20 The REMPI spectrum of Fig. 1(b) may be simulated by a sum of the sequence bands of the C-3Πa(v′=v+2)→X 3Σg(v,v') transitions mainly for v'=2 and 3 [Fig. 1(g)]. In addition, O2(a 1Δg,v=0) was failed to be detected at 329–332 nm. These results indicate that highly exothermic reaction (27) yields mainly vibrationally excited O2(X 3Σg−,v=2 and 3), and plays only a minor role in the formation of O2(X 3Σg−,v=0) and O2(a 1Δg,v=0) from neat ASW photolysis at 157 nm.

Hydroperoxyl radical, HO2, is known to be photodissociated to OH and O(1D or 3P) at excitation wavelengths of 228–417 nm in the gas phase.50–52 Our previous studies showed no contribution of HO2 photoproducts on ASW to production of OH and O(1D and 3P) following 157 nm photolysis of ASW, suggesting that the amount of HO2 photoproducts on the ice surface was negligible at 90 K. Therefore, the O2 formation following reactions of HO2 + H or HO2 + OH are considered to be unlikely sources for the observed O2 ejection from the ASW surface in the present experiments.3,35,54 Theoretical and experimental studies reported that the probabilities for the photodissociation processes of HO2 to H + O2 fragments were estimated to be small, since these photoprocesses are prohibited by potential barriers.52,55

VI. CONCLUSION

We have measured time-of-flight and rotational spectra of the photodesorbed O2(X 3Σg−,v=0) and O2(a 1Δg,v=0) following pulsed 157 nm irradiation of ASW at 90 K, and the translational and internal energy distributions were obtained. Measured translational energies of O2(X 3Σg−,v=0) and O2(a 1Δg,v=0) suggested that the observed O2(X 3Σg−,v=0) and O2(a 1Δg,v=0) fragments originate only from the ASW surface, not from the ice bulk. The exothermic barrierless reaction of OH and O(1P) is proposed as the most plausible source for O2(X 3Σg−,v=0) formation. The reaction of translationally and internally excited OH with O(1P) plays a key role in the O2(a 1Δg,v=0) formation. The contribution of recombination reaction of two O(1P) to O2 formation was found to be negligible. O(1P) would react with OH much more frequently than they encounter another O(1P) to produce O2 since concentration of primary product OH on ASW is much higher than that of O(1P) formed via secondary reaction on ASW.

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