

**Binding Energy of Gas Molecule with Two Pyrazine Molecules as
Organic Linker in Metal-Organic Framework: Its Theoretical
Evaluation and Understanding of Determining Factors**

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Abstract

We explored the interactions of gas molecules such as H_2 , CH_4 , C_2H_4 , C_2H_6 , CO_2 , and CS_2 sandwiched by two pyrazine (Pz) molecules which were employed as a model of organic linker in the Hofmann-type metal-organic framework (MOF). The MP2.5/aug-cc-pVTZ method was employed here, because this method presents almost the same binding energy as that calculated by the CCSD(T)/aug-cc-pVDZ with MP2.5-evaluated basis set extension effects to aug-cc-pVTZ basis set. The binding energy of the gas molecule increases in the order $\text{H}_2 < \text{CH}_4 < \text{CO}_2 < \text{C}_2\text{H}_4 \approx \text{C}_2\text{H}_6 < \text{CS}_2$. The energy decomposition analysis of the interaction energy indicates that the electrostatic term presents the largest contribution to the interaction energy at the Hartree-Fock (HF) level. However, the dispersion interaction provides dominant contribution to the total binding energy at correlated level. We newly found a linear correlation between the z-component of polarizability of gas molecules and dispersion energy, where the z-axis was taken to be perpendicular to two Pz rings. These results are useful for understanding and predicting the binding energy of the gas molecule with the organic linkers of MOF.

Keywords Metal-organic framework, Binding Energy, Dispersion Interaction, Molecular Polarizability

1 Introduction

Porous coordination polymers (PCPs) also known as metal-organic frameworks (MOFs) have gained considerable attentions since last decade because of their promising possibilities for storage and selective separation of gas molecules [1-17] as well as catalyses [18]. As the name suggests, MOFs are the hybrid compounds wherein metals are connected through organic linkers. Thus, MOF with desired ability and structure can be synthesized by changing organic linkers and/or metal centers. To find efficient MOFs for capture of green house gases, many studies have been carried out [19-24]. In such works, a lot of efforts have been made to confirm and estimate porosity required to selective absorption/adsorption of gases into MOFs and also the conditions under which such absorption/adsorption of gas occurs [25-29].

However, theoretical studies of the interaction of gases such as H_2 , CO_2 , and CH_4 with MOFs have been limited so far. In a pioneering study of the interaction of H_2 with MOF-5, Sagara et al. [30, 31] calculated the binding energy of 6-7 kJ/mol at the corner site (metal site) and 4-5 kJ/mol at the linker site with MP2 theory. Density functional theory (DFT) calculations by Lee et al. [32] suggest that the interaction of H_2 with model system (benzene) is significantly different from that with actual MOF-5 framework. Also, the interaction of CO_2 with simple alcohols, ketones, esters, and amines have been investigated [33-38]. These studies suggest that the formation of a hydrogen bond is a driving force for stabilization of these complexes. More realistic system was explored with ONIOM(MP2/6-31G(d,p):HF/6-31G(d,p)) method by Pianwanit and co-workers [39]. They concluded that in MOF-5, both CO_2 and CH_4 occupy the perpendicular position to the ZnO_4 corner with the binding energies of 9.27 and 3.64 kcal/mol,

respectively. Recently, many MOFs with nitrogen containing organic linkers are reported in an attempt to form stronger interaction between the carbon of CO₂ and the nitrogen of the linker [29, 40]. For instance, Vogiatzis et al. [40] investigated the interaction between CO₂ and the N-containing organic heterocycles with CCSD(T), where the binding energy at complete basis set limit is approximately evaluated by the MP2-F12 method with 6-311++G(d,p) and aug-cc-pVTZ basis sets. They reported that the DFT-D method [41, 42] with PBS functional provides binding energies similar to the values calculated by the CCSD(T) method with MP2-F12-calculated basis set extension effects and recommended the use of the basis sets augmented with diffuse functions (e.g. aug-cc-pVTZ) [40]. One of the present authors also theoretically investigated the interactions of CO₂ and CS₂ with pyrazine and found that the binding energy of CO₂ is somewhat less than that of CS₂ [29]. In the above mentioned theoretical studies, efforts were made to examine the methodology, the model system, and the value of binding energy. Though such knowledge is indispensable in the theoretical study of MOF, it is also important to find determining factors for the interaction between gas molecule and MOF.

In the present work, we theoretically investigated the interaction of gas molecules such as H₂, CO₂, C₂H₄, C₂H₆, and CS₂ with two pyrazine (Pz) molecules which are organic linker of the recently reported Hoffmann type MOF {Fe(Pz)[Pt(CN)₄]}_n [29]. This Hoffmann type MOF is of considerable interest because the spin state conversion occurs by adsorption of some of gas molecules [29]; for instance, CS₂ induces the conversion from high spin to low spin, but CO₂ does not. Thus, the interaction energy between the gas molecule and this MOF is important to control molecular property of this MOF. Though the interaction of CO₂ with the N atom of Pz was theoretically investigated well

recently [40], gas molecule approaches not the N atom of Pz but the six-member ring of Pz in this MOF; see Scheme 1 [29]. This is because the metal atom coordinates with the N atoms of Pz. We employed this MOF here because the position of gas molecule was clearly shown by recent X-ray analysis; see also Scheme 1 [29]. First, we examined the suitable level of theory as well as basis set because the approach of gas molecule to the six-member ring of Pz has not been theoretically investigated yet. Then, we theoretically evaluated binding energies of such gas molecules as H₂, CO₂, C₂H₄, C₂H₆, and CS₂ with two pyrazine (Pz) molecules. One interesting issue here is to make comparison between the binding energy of gas molecule sandwiched by two Pz molecules and that of gas molecule with one Pz. Also, it is interesting to investigate whether the relation between binding energy and molecular property exist or not. We wish to provide theoretical answers to above mentioned issues.

2 Computational Details

Computations were carried out at MP2 to MP4(SDQ) and CCSD(T) levels of theory in combination with such basis sets as CBSB4 [43], aug-cc-pVDZ [44], and aug-cc-pVTZ [44]. The recently proposed MP2.5 method [45], which corresponds to the arithmetic mean of MP2 and MP3-calculated values, was also employed here. The basis set extension effects from aug-cc-pVDZ to aug-cc-pVTZ was evaluated at MP2 and MP2.5 levels. Molecular properties such as quadrupole moment and polarizability were calculated by the DFT method with B3LYP functional, employing aug-cc-pVTZ basis set. Counter-poise correction (CPC) [46] was made to consider the basis set superposition error (BSSE) in the binding energy. The Kitaura-Morokuma energy decomposition

analysis (EDA) [47] was carried out with the reduced variational space (RVS) scheme [48, 49]. The binding energy and molecular properties were calculated with the Gaussian 03 [50] program and the EDA was performed with GAMESS program package [51].

3 Results and Discussion

3.1 Suitable Level of Theory and Basis set

We first scanned the potential energy surface (PES) at different levels of theory and basis set to find the suitable level of theory for this type of interaction. The PESs for such two systems as Pz-CO₂ and Pz-CS₂ are shown in Fig. 1, where the CO₂ and CS₂ are moved perpendicular to the six-member ring of Pz to mimic the experimental geometry [29]. Note that the negative value of the binding energy represents the energy stabilization. On the left hand side of Fig.1, the PESs at different levels are compared with the CCSD(T) values. As is clear from these PESs, the binding energy is considerably overestimated at the MP2 level, whereas it is underestimated at the MP3 level. Though the binding energy at the MP4(SDQ) level is moderately larger (more negative) than that at the MP3 level, the MP4(SDQ)-calculated PES is still considerably above the CCSD(T) values. The PES calculated at the MP2.5 level is close to that obtained at the CCSD(T) level. On the right hand side of Fig.1, the effect of basis set on the binding energy is examined. Here, the basis set extension effect was evaluated at the MP2 and MP2.5 levels. As seen in Fig. 1, when the basis set changes from aug-cc-pVDZ to aug-cc-pVTZ, the PES becomes considerably lower, suggesting that the use of aug-cc-pVTZ is necessary at least. However, the CCSD(T)/aug-cc-pVTZ calculation is computationally demanding. Thus, the CCSD(T)/aug-cc-pVDZ method is employed here

as a reference, in which the basis set extension effect from aug-cc-pVDZ to aug-cc-pVTZ is evaluated with the MP2.5 method. This method is named hereafter as CCSD(T)/aug-cc-pVDZ + MP2.5{aug-cc-pVTZ – aug-cc-pVDZ}. We found that the MP2.5/aug-cc-pVTZ-calculated PES is close to the PES calculated at the reference level, as shown in Fig. 1; though MP2.5/aug-cc-pVQZ method was recommended for evaluation of the dispersion interaction [52], the Pz-gas-Pz system is too large to perform MP2.5/aug-cc-pVQZ calculation.

In Table 1, the minimum positions of these PESs are listed with the binding energies at the minima. The minima for CO₂ and CS₂ complexes are found at 3.25 and 3.50 Å at the reference level, respectively. Other methods such as MP2.5/aug-cc-pVTZ, CCSD(T)/CBSB4 + MP2{aug-cc-pVTZ – CBSB4}, CCSD(T)/CBSB4 + MP2.5{aug-cc-pVTZ – CBSB4}, and CCSD(T)/aug-cc-pVDZ + MP2{aug-cc-pVTZ – aug-cc-pVDZ} present almost the same minimum on the PES, where the error is less than 0.03 Å. Among the best five levels of theory, the MP2.5/aug-cc-pVTZ is relatively less expensive. Further, it provides binding energy close to the reference value; see Table 1. We will use the MP2.5/aug-cc-pVTZ method for further calculations.

3.2 Binding Energies of Gas Molecules with Two Pyrazines

The geometry of the system consisting of two Pz molecules and gas molecule, which is henceforth called as Pz-gas-Pz, was optimized at the MP2/aug-cc-pVTZ level of theory, as shown in Fig.2. Here, the distance (7.2 Å) between two Pz molecules was taken to be the same as that in the Hofmann type MOF {Fe(Pz)[Pt(CN)₄]}_n. The position and the orientation of gas molecule were optimized by carrying out the manual

displacement of gas molecule along the x, y, and z-axes. A very shallow PES was observed for all these complexes. The rotation of gas molecule around the z-axis little influences the binding energy; see Supplementary Fig.S1 for details. The binding energies of these complexes at the MP2 and MP2.5 levels of theory are compared in Table 2. Significantly large differences are observed between the MP2 and MP2.5-calculated binding energies for all the systems examined here, indicating again that not the MP2 but the MP2.5 method must be applied to these systems. The calculated binding energies for these weak complexes are in the range of -1.0 to -5.5 kcal mol⁻¹ and increases in the order H₂ < CH₄ < CO₂ < C₂H₄ ≈ C₂H₆ < CS₂. It is noted that the larger binding energy of CS₂ than that of CO₂ is consistent with the previous experimental finding [29].

The binding energy of CO₂ with two Pz molecules is smaller than the twice of that with one Pz molecule: see Tables 1 and 2. On the other hand, the binding energy of CS₂ with two Pz molecules is similar to the twice of that with one Pz. This difference between CO₂ and CS₂ arises from the difference in the minimum position. In the CS₂ complex, the Pz-CS₂ distance is similar between Pz-CS₂ and Pz-CS₂-Pz systems. On the other hand, the Pz-CO₂ distance is considerably shorter in Pz-CO₂ than in Pz-CO₂-Pz. This is the reason why the binding energy of Pz-CO₂-Pz is smaller than the twice of that of Pz-CO₂. If we evaluate the binding energy of Pz-CO₂ at the same intermolecular distance as that of Pz-CO₂-Pz, the twice of the binding energy of Pz-CO₂ is almost the same as that of Pz-CO₂-Pz, as follows: The binding energy of Pz-CO₂ at 3.6 Å, which is one half of the intermolecular distance of Pz-CO₂-Pz, is 1.94 kcal mol⁻¹ at the MP2 level and 1.63 kcal mol⁻¹ at the MP2.5 level. The twice of these values are very close to the binding energy

of Pz-CO₂-Pz, as compared in Table 2. In other gas molecules, the same relation is observed; see Table 2. Interestingly, the deviation is very small. These results suggest that the binding energy of gas molecule can be approximately estimated as the sum of the interaction energies of gas with each of surrounding organic linkers. This additivity of binding energy is under investigation in various systems.

3.3 Energy Decomposition Analysis of the Binding Energy

The interactions of gas molecules with two pyrazine molecules will be discussed at the Hartree-Fock (HF) and then at the correlated levels, because the total binding energy can be divided to electrostatic (ES), exchange (EX), charge transfer (CT), polarization (PL), the dispersion interactions (DIS), and higher order coupling term. The energy decomposition analysis (EDA) [47-49] at the Hartree-Fock (HF) level provides the energy contributions of ES, EX, CT, and PL terms to the HF interaction energy. As shown in Table 3, the ES term is a major contributor to the HF interaction energy. The importance of the ES term was previously discussed in van der Waals complexes of benzene-methane and benzene-benzene [53]. This term increases in the order $H_2 < CH_4 \leq CO_2 < C_2H_6 \approx C_2H_4 < CS_2$. The EX term induces large destabilization. The stabilizing contributions of PL and CT are significantly small compared to the ES term. The combined contribution of CT and PL is in the range of 3 to 17% of the total binding energy at the MP2.5 level.

It is important to know what factors are responsible for the ES term. The ES term arises from charge-charge interaction, charge-dipole interaction, charge-quadrupole interaction, dipole-dipole interaction, and so on. All gas molecules examined here are

neutral and do not have a dipole moment. It is likely that for the ES term the next important is the quadrupole moment. Actually, a previous review proposed that the quadrupole moment would be significantly important for adsorption of gas molecule in MOF [54]. However, we could not find a linear correlation between the ES term and the quadrupole moment; see Supplementary Fig.S2. Also, we evaluated electrostatic interaction between atomic charges of gas molecule and those of two pyrazine molecules. However, thus evaluated electrostatic interaction does not show a clear relation with the ES term from EDA; see Supplementary Fig. S3 and Table S1. At this moment, we could not find determining factor for the electrostatic interaction between gas molecule and organic pillar ligands of MOF. Further study is necessary to find the determining factors for ES term.

3.4 Contribution of Dispersion Interaction to Binding Energy

The contribution of electron correlation energy to the binding energy, which is defined as the difference in the binding energy between the MP2.5 and HF levels of theory, mainly corresponds to the dispersion interaction. It should be noted that the HF interaction energy is repulsive (Table 3), indicating the important roles of the dispersion interaction in these complexes. As shown in Table 3, this contribution increases in the order $\text{H}_2 < \text{CO}_2 < \text{CH}_4 < \text{C}_2\text{H}_4 < \text{C}_2\text{H}_6 < \text{CS}_2$. It is noted that this contribution is significantly larger than the ES term, indicating that the dispersion interaction is a major contributor to the binding energies of these complexes.

We now wish to discuss possible factors determining the dispersion interaction. Because the dispersion interaction arises from the mutual excited configuration in each

moiety, it is likely that the dispersion interaction relates to the polarizability. The DFT(B3LYP)-calculated polarizabilities of these gas molecules agree well with the experimental values [55], as shown in Fig.3(A), where the correlation coefficient R^2 is 0.99; see also Table 4 for details. Though the energy contribution by correlation effect increases with increase in the mean polarizability, as shown in Fig.3(B), somewhat large deviation from the linear relation is observed ($R^2 = 0.79$). On the other hand, a very good linear correlation ($R^2 = 0.95$) is found between the z-component of polarizability and the energy contribution by correlation effect, where the z-axis is perpendicular to the six-member ring of Pz, as shown in Fig. 2 . We wish to emphasize that this interesting relation is found for the first time here.

This correlation is useful to estimate the dispersion contribution by evaluating polarizabilities of these gases. For instance, the order of binding energy can be discussed with the polarizability and the HF interaction energy. If binding energies of two gas molecules are evaluated at the correlated level, the relationship between the dispersion interaction and the polarizability can be presented. This relation provides the dispersion interaction of other gas molecule based on the polarizability. Thus, the binding energy at correlated level can be approximately estimated from the polarizability and the HF interaction energy.

4 Concluding Remarks

In this work, we theoretically investigated the interactions of gas molecules such as H_2 , CO_2 , C_2H_4 , C_2H_6 , and CS_2 with two Pz molecules which are organic linker of the recently reported Hofmann type MOF $\{Fe(Pz)[Pt(CN)_4]\}_n$. The calculations are performed at different levels of theory and basis set to find the suitable method which

provides reliable results with reasonable computational cost. We found that the MP2.5 method with aug-cc-pVTZ is a reasonable choice for this type of interaction. The binding energy of gas molecule with two Pz molecules increases in the order $\text{H}_2 < \text{CH}_4 \leq \text{CO}_2 < \text{C}_2\text{H}_4 \approx \text{C}_2\text{H}_6 < \text{CS}_2$. The additivity of the binding energy is found in these systems; in other words, the binding energy of the Pz-gas-Pz system can be approximately evaluated as the sum of each binding energy of Pz-gas which is evaluated at the same intermolecular distance as in the Pz-gas-Pz system. The energy decomposition analysis at the HF level reveals that the electrostatic interaction dominantly contributes to the interaction energy at the HF level. However, the interaction energy is repulsive at the HF level. In other words, the dispersion interaction is crucial to the total binding energies of these gas molecules. We wish to emphasize that a good linear correlation ($R^2=0.95$) is found first between the z-component of polarizability and dispersion energy, where the z-axis is perpendicular to the six-member ring of Pz. This relation is useful for discussing and predicting the binding energy of gas molecule with organic linker of MOF.

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Supplementary Information

The PES plots along x, y, and z-axes for Pz-gas-Pz four different systems are given in the Supplementary Fig.S1. A correlation between the electrostatic energy and the z-

component of quadrupole moment is given in Supplementary Fig. S2. Correlations between the ES term from EDA and the electrostatic interaction evaluated with ESP charges are provided in Supplementary Fig. S3 and Table S1.

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Figure and Table Captions

Fig.1 Comparison of potential energy surfaces of (a) Pz-CO₂ and (b) Pz-CS₂ calculated at different levels of theory and basis set

Fig.2 Optimized geometries of various complexes, (A) Pz-H₂-Pz, (B) Pz-CO₂-Pz, (C) Pz-CH₄-Pz, (D) Pz-CS₂-Pz, (E) Pz-C₂H₄-Pz, and (F) Pz-C₂H₄-Pz at the MP2/aug-cc-pVTZ level

Fig.3 Correlations between (A) experimental and theoretically calculated polarizabilities, (B) the energy contribution of correlation effect and mean polarizability, and (C) the energy contribution of correlation effect and the z-component of polarizability; see Fig.2 for x, y, and z-axes

Table 1 Minimum position (Å)^a on the potential energy surface (PES) and binding energy at the minimum position (kcal mol⁻¹)

^a The distance from the six-member ring of Pz; see Fig.1

Table 2 Binding energies (kcal mol⁻¹) of various gases with two pyrazine (Pz) molecules

^a Aug-cc-pVTZ basis sets is used

Table 3 Energy decomposition analysis (EDA) of interaction energies of various gases with two pyrazine (Pz) molecules at the Hartree-Fock level and the electron correlation effect (ΔE_{corr}) on the binding energy

^a Aug-cc-pVTZ basis sets were used

^b Contribution of mixing term to total BE_{HF} is nearly zero

^c Contribution of correlation energy evaluated at the MP2.5 level

^d Total binding energy; $BE = BE_{\text{HF}} + \Delta E_{\text{corr}}$

Table 4 Quadrupole moments and polarizabilities of various gas molecules^a

^a B3LYP/aug-cc-pVTZ

^b The electrostatic energy is obtained by energy decomposition analysis; See Table 3

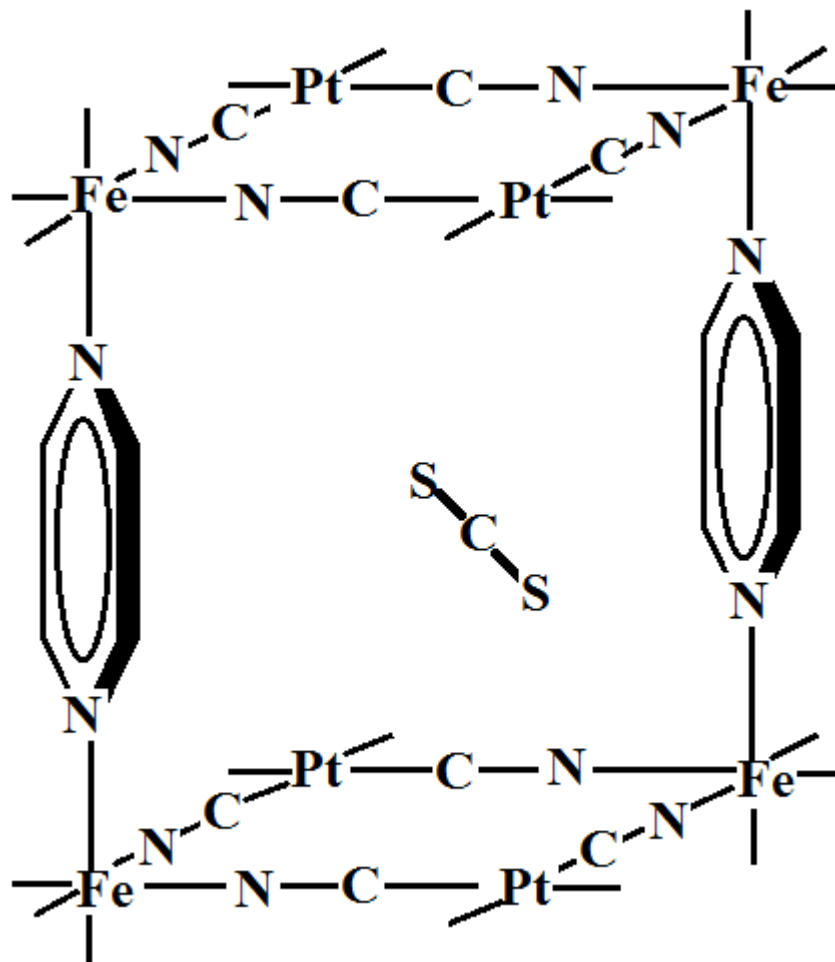
^c The difference in binding energy between the HF and MP2.5 levels

^d Traceless values are presented along x, y, and z axes, respectively

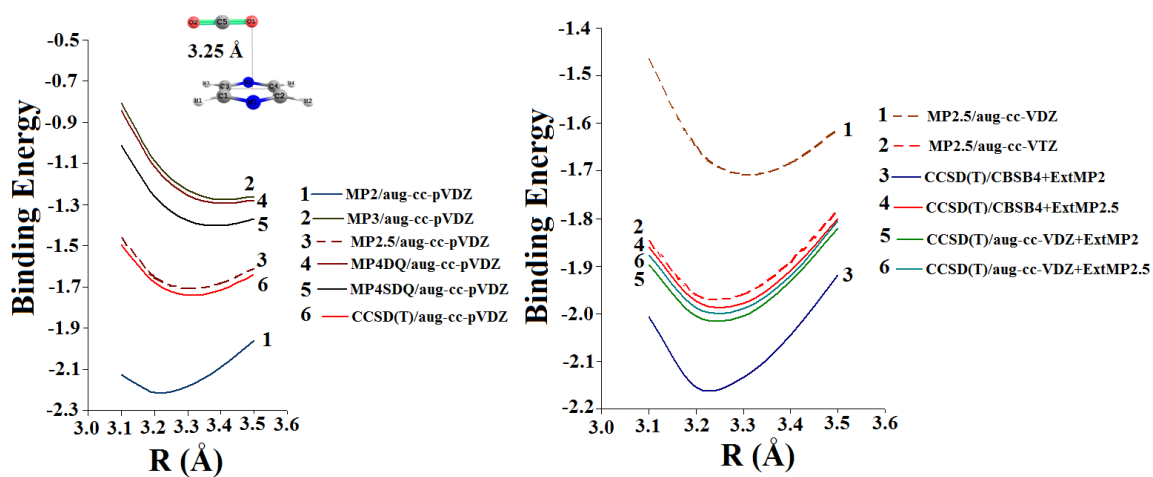
^e Values are presented along x, y, and z axes, respectively

^f Value in the parenthesis is a mean value

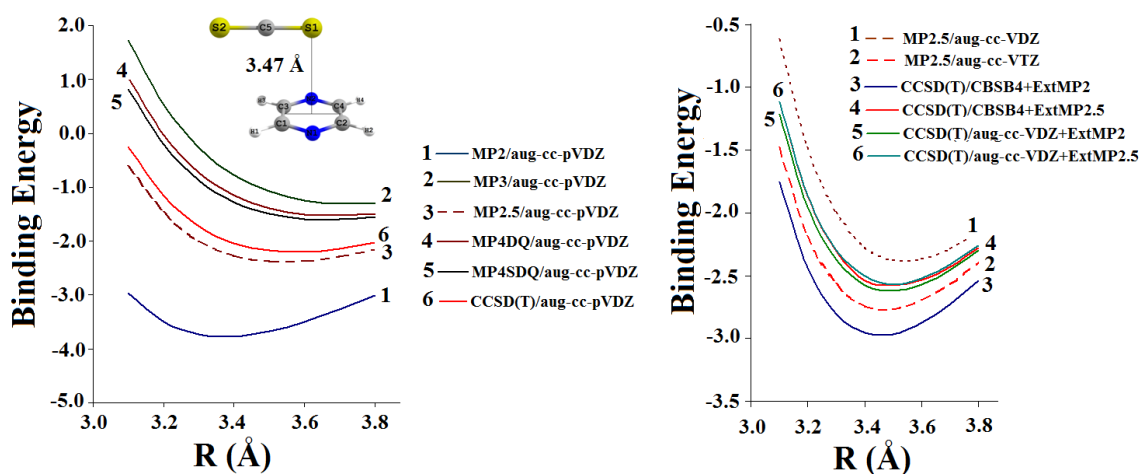
^g Experimental value



Scheme 1 The Schematic Picture of the Hoffmann type metal-organic framework; Ref. [29]



(a) Pz...CO₂ system



(b) Pz...CS₂ system

Fig.1 Comparison of potential energy surfaces of (a) Pz-CO₂ and (b) Pz-CS₂ calculated at different levels of theory and basis set

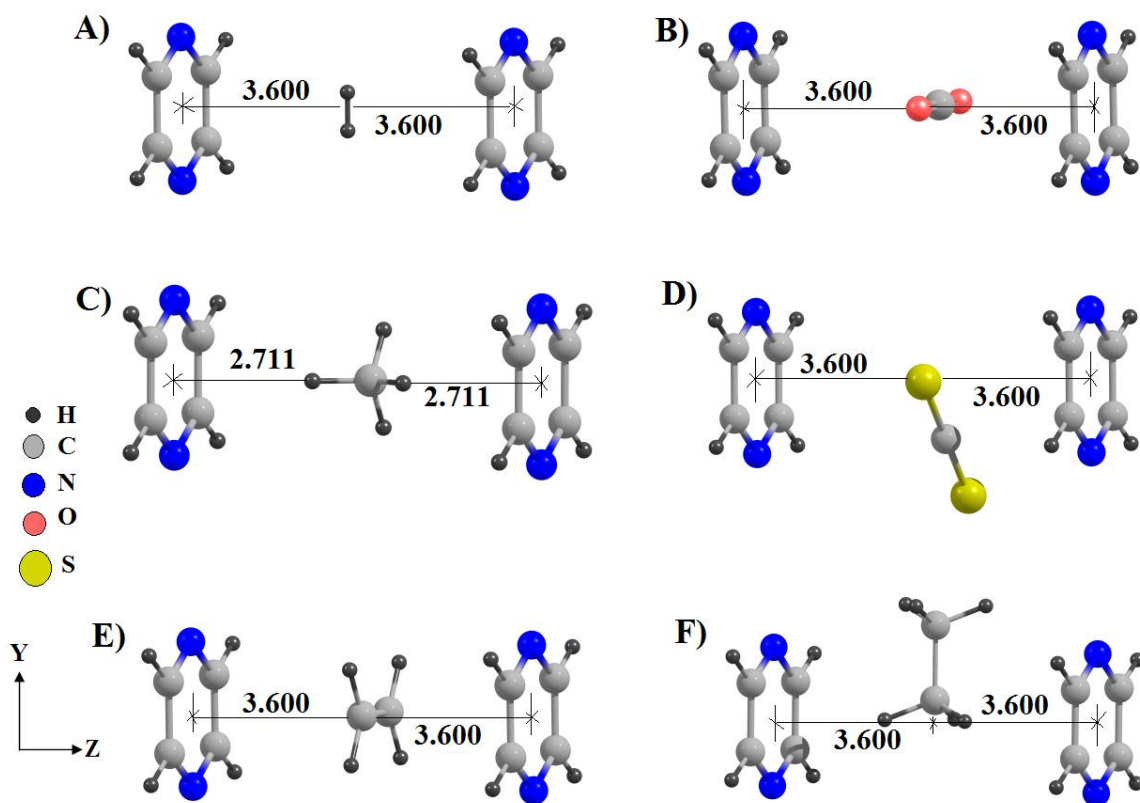


Fig.2 Optimized geometries of various complexes, (A) Pz-H₂-Pz, (B) Pz-CO₂-Pz, (C) Pz-CH₄-Pz, (D) Pz-CS₂-Pz, (E) Pz-C₂H₄-Pz, and (F) Pz-C₂H₄-Pz at the MP2/aug-cc-pVTZ level

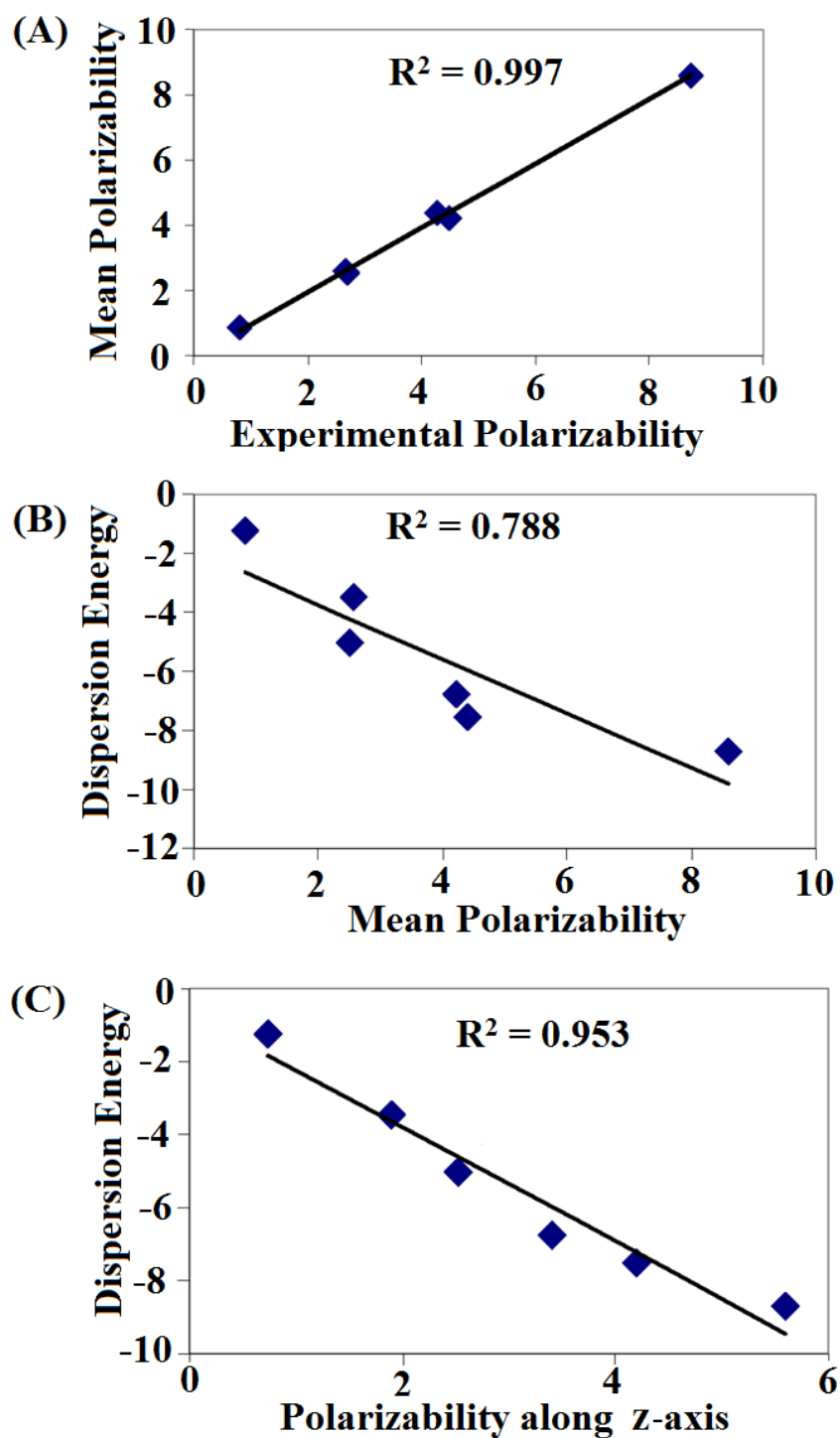


Fig.3: Correlations between (A) experimental and theoretically calculated polarizabilities, (B) the energy contribution of correlation effect and mean polarizability, and (C) the energy contribution of correlation effect and the z-component of polarizability; see Fig.2 for the x, y, and z-axes

Table 1: Minimum position (\AA)^a on the potential energy surface (PES) and binding energy at the minimum position (kcal mol^{-1})

Level of theory	Minimum position		Binding Energy	
	CO ₂	CS ₂	CO ₂	CS ₂
MP2/CBSB4	3.41	3.57	-1.28	-2.08
MP2.5/CBSB4	3.51	3.74	-1.03	-1.33
CCSD(T)/CBSB4	3.51	3.78	-1.04	-1.24
MP2/aug-cc-pVDZ	3.22	3.37	-2.21	-3.78
MP2.5/aug-cc-pVDZ	3.32	3.54	-1.70	-2.39
CCSD(T)/aug-cc-pVDZ	3.32	3.57	-1.73	-2.19
CCSD(T)/CBSB4 + MP2(aug-cc-pVDZ-CBSB4)	3.29	3.53	-1.87	-2.52
CCSD(T)/CBSB4 + MP2.5(aug-cc-pVDZ-CBSB4)	3.31	3.57	-1.72	-2.20
MP2/aug-cc-pVTZ	3.15	3.30	-2.55	-4.37
MP2.5/aug-cc-pVTZ	3.25	3.47	-1.96	-2.78
CCSD(T)/CBSB4 + MP2(aug-cc-pVTZ-CBSB4)	3.24	3.46	-2.14	-2.78
CCSD(T)/CBSB4 + MP2.5(aug-cc-pVTZ-CBSB4)	3.25	3.49	-1.98	-2.60
CCSD(T)/aug-cc-pVDZ + MP2(aug-cc-pVTZ - aug-cc-pVDZ)	3.25	3.49	-2.03	-2.63
CCSD(T)/ aug-cc-pVDZ + MP2.5(aug-cc-pVTZ - aug-cc-pVDZ)	3.25	3.50	-2.00	-2.57

^a The PESs shown in Fig.1 are fitted to polynomial equation of n^{th} degree ensuring the good correlation coefficient ($R^2 = 0.9999$). This polynomial equation is then used for locating minimum on the PES

Table 2: Binding energies (kcal mol⁻¹) of various gases with two pyrazine (Pz) molecules

(A) Binding Energy ^a						
	Pz-H ₂ -Pz	Pz-CH ₄ -Pz	Pz-CO ₂ -Pz	Pz-CS ₂ -Pz	Pz-C ₂ H ₄ -Pz	Pz- C ₂ H ₆ -Pz
MP2	-1.24	-3.14	-3.92	-7.82	-5.06	-4.99
MP2.5	-1.16	-2.58	-3.30	-5.49	-4.02	-4.06
(B) Twice of Binding Energy of Pz-gas system ^a						
MP2	-1.24	-3.13	-3.88	-7.92	-5.03	-4.97
MP2.5	-1.16	-2.57	-3.26	-5.60	-3.99	-4.05

^a Aug-cc-pVTZ basis sets is used

Table 3: Energy decomposition analysis (EDA) of binding energies of various gases with two pyrazine (Pz) molecules at the Hartree-Fock level and the electron correlation effect (ΔBE_{corr}) on the binding energy

	EDA at the HF level ^{a,b}					ΔBE_{corr} ^c	BE^d
	ES	EX	PL	CT	BE_{HF}		
H₂	-0.27	0.39	-0.02	-0.02	0.08	-1.24	-1.16
CO₂	-1.33	1.83	-0.22	-0.13	0.15	-3.47	-3.30
CS₂	-2.62	6.48	-0.24	-0.45	3.17	-8.86	-5.49
CH₄	-1.29	4.10	-0.15	-0.23	2.42	-5.00	-2.58
C₂H₄	-2.07	5.42	-0.27	-0.34	2.74	-6.76	-4.02
C₂H₆	-2.04	6.20	-0.21	-0.46	3.48	-7.54	-4.06

^a Aug-cc-pVTZ basis sets were used

^b Contribution of mixing term to total BE_{HF} is nearly zero

^c Contribution of correlation energy evaluated at the MP2.5 level

^d Total binding energy; $BE = BE_{\text{HF}} + \Delta BE_{\text{corr}}$

Table 4: Quadrupole moments and polarizabilities of various gas molecules^a

Gases	Electrostatic Energy ^b (kcal mol ⁻¹)	Dispersion Energy ^c (kcal mol ⁻¹)	Quadrupole Moments ^{d,e} (Debye - Å)	Polarizability ^e (Å ³)
H₂	-0.27	-1.24	-0.2036, -0.2036, 0.4071	0.74, 1.03, 0.74 (0.84) ^f , [0.79] ^g
CO₂	-1.33	-3.47	1.4926, 1.4926, -2.9852	1.90, 3.89, 1.90 (2.56), [2.65]
CS₂	-2.62	-8.86	-1.1613, -1.1613, 2.3227	5.60, 14.58, 5.60 (8.59), [8.74]
CH₄	-1.29	-5.00	0.0000, 0.0000, 0.0000	2.52, 2.52, 2.52 (2.52), [2.68]
C₂H₄	-2.07	-6.76	-2.2040, 1.0775, 1.1264	3.86, 5.38, 3.40 (4.21), [4.48]
C₂H₆	-2.04	-7.54	0.2372, 0.2371, -0.4743	4.19, 4.79, 4.19 (4.39), [4.26]

^a B3LYP/aug-cc-pVTZ.

^b The electrostatic energy is obtained by energy decomposition analysis; See Table 3

^c The difference in binding energy between the HF and MP2.5 levels

^d Traceless values are presented along x, y, and z axes, respectively

^e Values are presented along x, y, and z axes, respectively

^f Value in the parenthesis is a mean value

^g Experimental value

Binding Energy of Gas Molecule with Two Pyrazine Molecules as Organic Linker in Metal-Organic Framework: Its Theoretical Evaluation and Understanding of Determining Factors

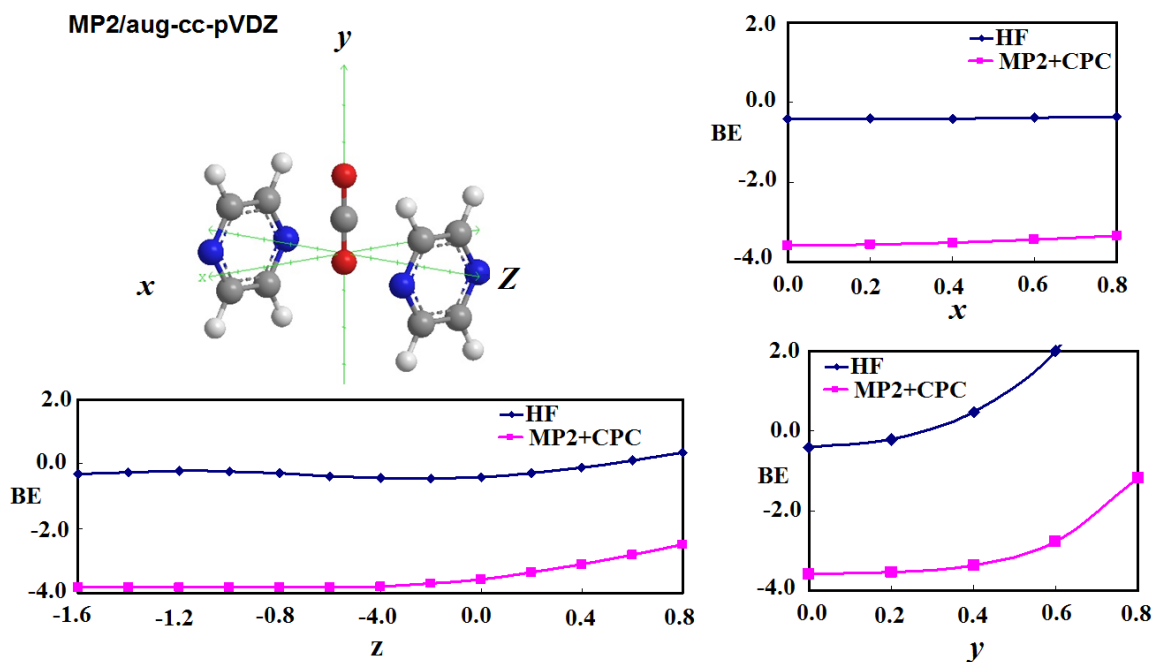
Milind M. Deshmukh and Shigeyoshi Sakaki*

Fukui Institute of Fundamental Chemistry, Kyoto University, Nishihiraki-cho, Sakyo-ku,
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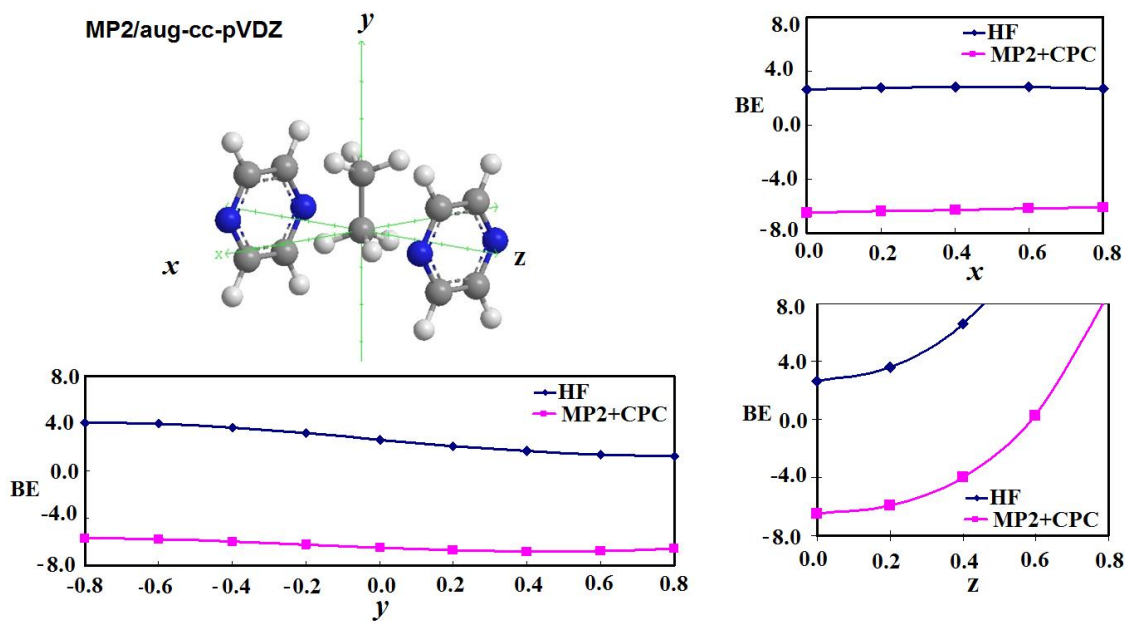
Supplementary Information

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Supplementary Information:

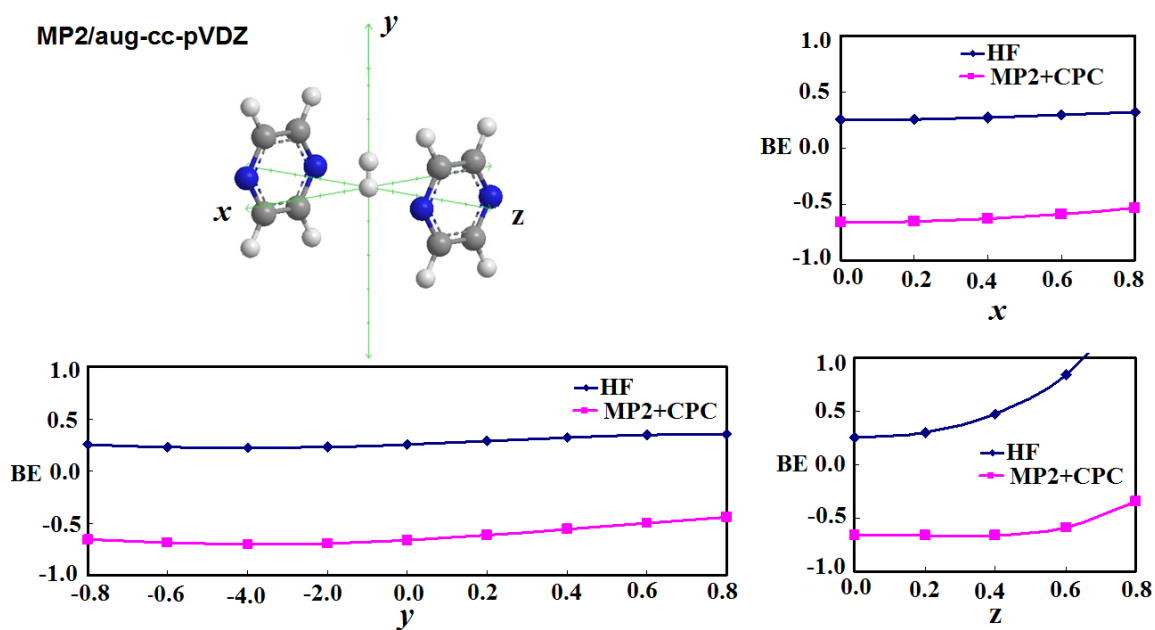


(a) PZ-CO₂-Pz

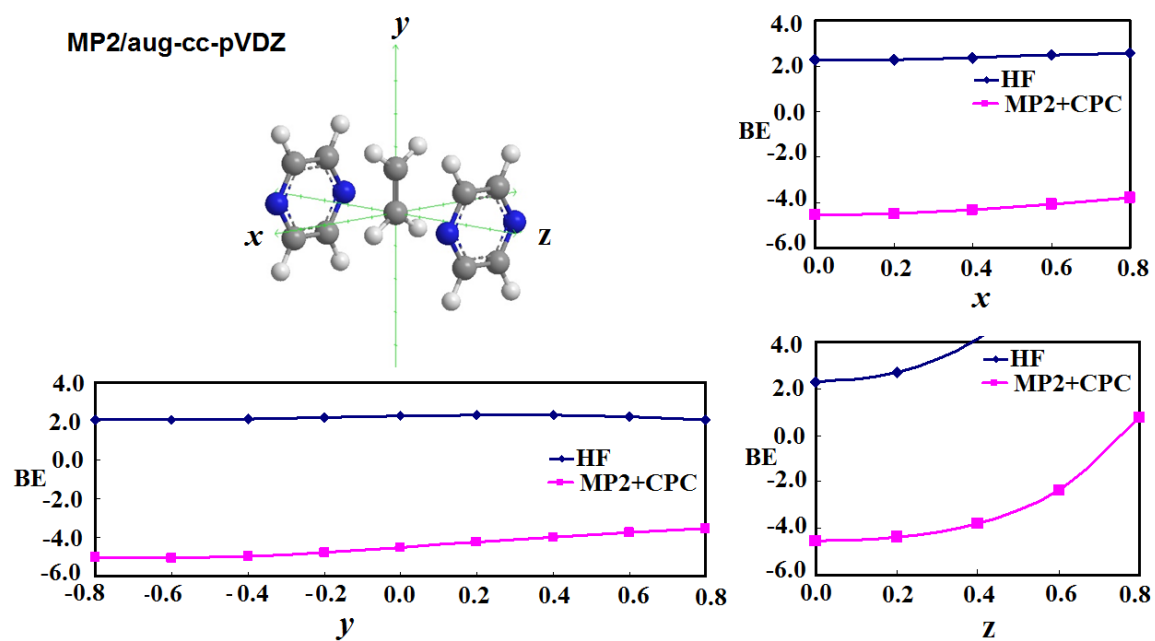


(b) PZ-C₂H₆-Pz

Supporting Information Figure S1: Effect of displacement of gas molecules along x-, y- and z-directions on the binding energy.

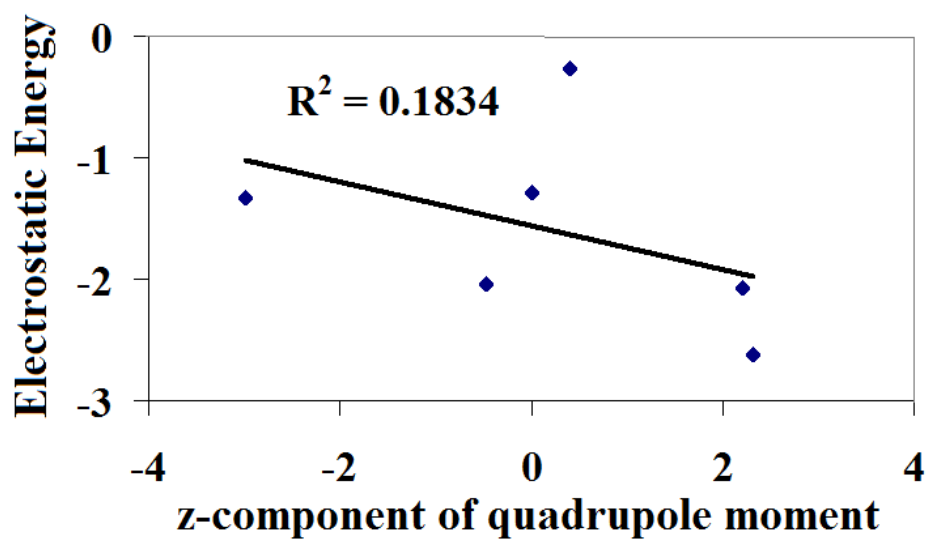


(c) PZ-H₂-Pz



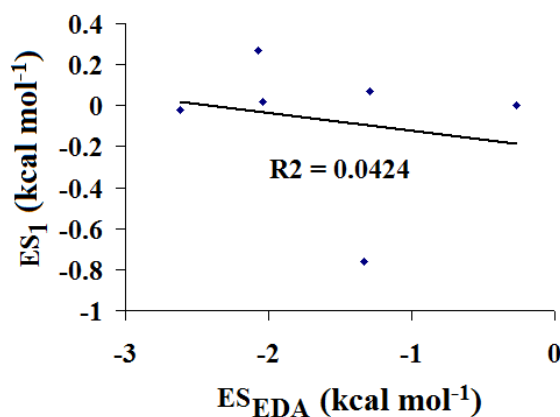
(d) PZ-C₂H₄-Pz

Supporting Information Figure 1: Continued....

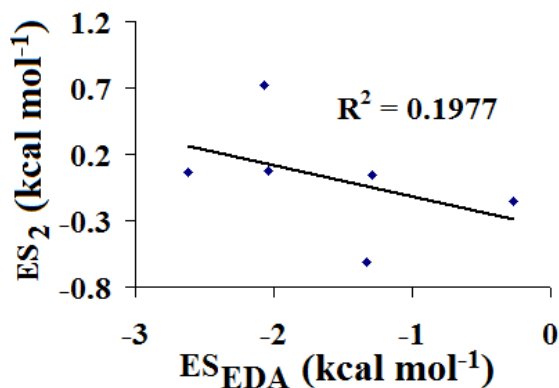


Supporting Information Figure S2: Correlation plots between electrostatic contribution at the HF level and the z-component of quadrupole moment.

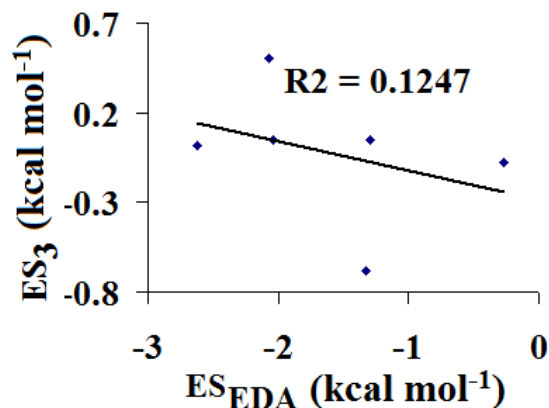
(A) ES term of EDA (ES_{EDA}) vs. electrostatic interaction (ES_1) between atomic charges^a of gas molecule and electrostatic potential of two pyrazine molecules



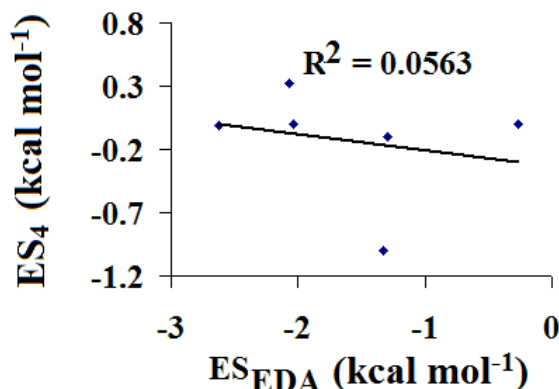
(B) ES term of EDA (ES_{EDA}) vs. electrostatic interaction (ES_2) between atomic charges^a of two pyrazine molecules and electrostatic potential of gas molecule



(C) ES term of EDA (ES_{EDA}) vs. average of the electrostatic interactions ($ES_3 = \{ES_1 + ES_2\}/2$)



(D) ES term of EDA (ES_{EDA}) vs. electrostatic interaction (ES_4)^b evaluated as charge-charge interactions



Supporting Information Figure S3: Correlation plots between ES term by EDA (ES_{EDA}) and electrostatic interactions such as charge-charge and charge-electrostatic potential.

^a Electrostatic potential fitted (ESPFIT) charges are employed.

$$^b ES_4 = \sum_{i \in \text{Gas}} \sum_{j \in \text{Pz-Pz}} \frac{q_i * q_j}{|r_i - r_j|}$$

Supporting Information Table S1: Electrostatic interaction between gas molecule and the two pyrazine molecules calculated with charge-charge ($q_i \cdot q_j / r_{ij}$)^a and charge-electrostatic potential ($q \cdot V$).

Complex	$q_{\text{gas}} \cdot V_{(\text{Pz-Pz})}$ ^b	$q_{(\text{Pz-Pz})} \cdot V_{\text{gas}}$ ^c	Average ($q \cdot V$) ^d	$\sum_{i \in \text{gas}} \sum_{j \in \text{Pz-Pz}} \frac{q_i \cdot q_j}{ r_i - r_j }$	ES Term from EDA
Pz-CO ₂ -Pz	-0.76	-0.61	-0.68	-1.001	-1.33
Pz-CS ₂ -Pz	-0.02	0.06	0.02	-0.016	-2.62
Pz-H ₂ -Pz	0.00	-0.15	-0.08	0.00	-0.27
Pz-CH ₄ -Pz	0.07	0.04	0.05	-0.100	-1.29
Pz-C ₂ H ₄ -Pz	0.27	0.72	0.5	0.319	-2.07
Pz-C ₂ H ₆ -Pz	0.02	0.08	0.05	0.003	-2.04

^a Electrostatic potential fitted (ESPFIT) charges are employed.

^b q_{gas} represents atomic charges of gas molecules. $V_{(\text{Pz-Pz})}$ represents the electrostatic potential due to two Pz molecules calculated at the atomic centers of gas molecules.

^c $q_{(\text{Pz-Pz})}$ represents atomic charges of two pyrazine molecules. V_{gas} represents the electrostatic potential due to gas molecule calculated at the atomic centers of two pyrazine molecules.

^d The average value of second and third column.