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Low-Temperature Synthesis, Thermodynamic Properties, and Electrical Conduction Properties of Lanthanum Phosphates

by

Naoyuki Hatada

Department of Materials Science and Engineering
Kyoto University
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Chapter 1
General Introduction

Recently significant research efforts have been made to develop fuel cells operating at intermediate temperatures (200–750 °C) [1-3]. This temperature range may allow for the use of non-precious metal catalysts and eliminate the need for external reforming of fuels. It may also allow for the use of less expensive components and facilitate miniaturization. Several types of phosphate based proton conductors such as ammonium polyphosphates, pyrophosphates, and cesium phosphates have been proposed for use as electrolytes in intermediate-temperature fuel cells [3]. They exhibit proton conductivities greater than $10^{-2}$ S cm$^{-1}$ at their operating temperatures. Nevertheless, there are several problems left to be solved, such as thermal decomposition at relatively low temperatures (~300 °C), mechanical instability, and water solubility.

In contrast, rare earth phosphates are attracting much attention due to their proton conduction and high-temperature chemical stability. However, further enhancement of the conductivity is still required for technological applications, while their proton conduction mechanisms are yet to be clarified. Also, knowledge regarding their optimum synthesis conditions and stabilities is insufficient. In this context, this work has focused on a comprehensive characterization of lanthanum phosphates with the objective of providing the fundamental information for developing new lanthanum-phosphate-based electrolytes for practical use. Systematic investigations were carried out on their optimum synthesis conditions, thermodynamic properties, and proton conduction properties. In this chapter, a brief overview of related work is provided.
1.1 Proton Conduction in Rare Earth Phosphates

Several rare earth phosphates doped with acceptors such as Sr have been shown to have proton conduction in the intermediate temperature range since the first report by Norby and Christiansen [4] in 1995. Norby and Christiansen showed that Ca- or Sr-doped LaPO$_4$ exhibits proton conductivities on the order of $10^{-4}$ S cm$^{-1}$ at ~700 °C in wet air while those of undoped LaPO$_4$ are ~$10^{-5}$ S cm$^{-1}$. Later, proton conduction was also found in many rare earth phosphates such as $RE$PO$_4$ ($RE$=Y, La, Ce, Pr, Nd, Sm) [5-8], La$_7$P$_3$O$_{18}$ [9], LaP$_3$O$_9$ [10], LaP$_5$O$_{14}$ [11], and Nd$_3$PO$_7$ [12] doped with acceptors. A number of studies have been carried out to date on various aspects of the proton conduction behavior in these phosphates, especially $RE$PO$_4$ and LaP$_3$O$_9$. In these studies, LaPO$_4$ was found to be the most preferable host lattice among $RE$PO$_4$ in terms of the proton conductivity [5-7]. Similarly, Sr was found to be the best dopant among alkaline earth metals for LaPO$_4$ [13] and LaP$_3$O$_9$ [14]. Figure 1.1 presents the highest conductivities of each Sr-doped lanthanum phosphate reported so far. The conductivities increase in the order of La$_7$P$_3$O$_{18}$ < LaPO$_4$ < LaP$_5$O$_{14}$ < LaP$_3$O$_9$, reaching $2 \times 10^{-4}$ S cm$^{-1}$ at 600 °C, though further improvement to ~$10^{-2}$ S cm$^{-1}$ will be required for practical applications.
Chapter 1  General Introduction

1.1.1 Proton Conduction Mechanism

The defect structures in Sr-doped LaPO$_4$ were investigated by Amezawa et al. using conductivity, MAS-NMR spectroscopy and FT-Raman spectroscopy measurements [18]. Based on the results, they proposed the following proton incorporation mechanism into Sr-doped LaPO$_4$.

When La is partially substituted with Sr, pyrophosphate ions written as (P$_2$O$_7$)$_{2p}$O$_4$ are formed by condensation of orthophosphate ions as shown in Eq. (1.1).

$\frac{1}{2}$Sr$_2$P$_2$O$_7$ $\rightarrow$ Sr'$_{La}$ + $\frac{1}{2}$ (P$_2$O$_7$)$_{2p}$O$_4$  \hspace{1cm} (1.1)

Then protons can be incorporated into LaPO$_4$ from water vapor, forming hydrogen phosphate groups, (HPO$_4$)$_{PO_4}$ as shown in Eq. (1.2).

$\frac{1}{2}$ (P$_2$O$_7$)$_{2p}$O$_4$ + $\frac{1}{2}$H$_2$O(g) $\rightarrow$ (HPO$_4$)$_{PO_4}$  \hspace{1cm} (1.2)
Based on this mechanism, the proton concentration in LaPO₄ is determined by the equilibrium between the intrinsic positive effects and water vapor in the ambient atmosphere (Eq. (1.3)). As shown in Figure 1.2, it was observed that the electric conductivity of Sr-doped LaPO₄, as well as that of Sr-doped La₇P₃O₁₈, increased with increasing water vapor pressure. This dependency agrees well with the proposed proton incorporation mechanism.

\[
K = \frac{[(\text{HPO}_4)_\text{PO}_4^\bullet]}{[\text{P}_2\text{O}_7^{\bullet\bullet\bullet}]^{1/2} p (\text{H}_2\text{O})^{1/2}}
\]  

(1.3)

Figure 1.2 Water vapor pressure dependence of the conductivities of 1% Sr-doped La₇P₃O₁₈ [9], LaPO₄ [7], LaP₃O₉ [10].

On the other hand, Sr-doped LaP₃O₉ does not follow the rule, for its conductivity is almost independent of water vapor pressure. In addition, it is also difficult to apply the defect structure model described above directly to Sr-doped LaP₃O₉ from the viewpoint of crystal structure. Figure 1.3 shows the crystal structures of undoped LaPO₄ (monazite type) and LaP₃O₉. In the LaPO₄ lattice, the constituent PO₄ tetrahedra are isolated, and they are supposed to form dimers (P₂O₇) only when La is partially substituted with divalent dopants. It is reasonable to assume that La₇P₃O₁₈ also has
isolated PO$_4$ tetrahedra considering its La/P ratio, though the crystal structure of La$_7$P$_3$O$_{18}$ has not been reported so far. In contrast, the PO$_4$ tetrahedra in undoped LaP$_3$O$_9$ are already linked in chains along the c-axis in nature. LaP$_5$O$_{14}$ was also reported to have cross-linking PO$_4$ chains [19]. Therefore, the defect structures proposed for LaPO$_4$ cannot be applied to LaP$_3$O$_9$ and LaP$_5$O$_{14}$ without appropriate modification.

![Figure 1.3 Structural arrangements of (a) undoped LaPO$_4$ (monazite type) [20] and (b) undoped LaP$_3$O$_9$ [21]. PO$_4$ groups in the crystal structures are shown as tetrahedra.](image)

Figure 1.3 Structural arrangements of (a) undoped LaPO$_4$ (monazite type) [20] and (b) undoped LaP$_3$O$_9$ [21]. PO$_4$ groups in the crystal structures are shown as tetrahedra.

![Figure 1.4 Schematic illustration of the proton transport paths in polycrystalline solid electrolytes.](image)

Figure 1.4 Schematic illustration of the proton transport paths in polycrystalline solid electrolytes.

Besides the defect structures, the proton conduction paths in undoped LaPO$_4$ have been studied
by several authors. Figure 1.4 schematically represents the two types of proton conduction paths in solid electrolytes. Jørgensen et al. observed that a part of phosphate excess during synthesis remained in the surfaces of LaPO$_4$ samples after sintering [22]. They suggested that the phosphorus-rich grain boundaries may enhance the proton conductivity. Harley et al. also reported the formation of phosphorus-rich amorphous grain boundary film by sintering of LaPO$_4$, because of residual hydrogen-phosphate-type compounds remaining from synthesis via precipitation in aqueous solutions [23]. The grain boundary conductivity was calculated to be as high as $\sim2.5\times10^{-3}$ S cm$^{-1}$ at 500 °C, from the analysis of the experimental conductivity data of the samples using the ‘bricklayer’ model. To demonstrate the effect of the presence of high conductivity grain boundary films further, they also investigated the conduction properties of a mixture of LaPO$_4$ – 15 mol% LaP$_3$O$_9$. The sample was heat treated at 1300 °C (higher than the melting temperature of LaP$_3$O$_9$) to form amorphous grain boundary film. The conductivity of the sample was measured before and after the heat treatment, and the increase in conductivity was observed with the appearance of the amorphous phases. Their work suggests that the grain boundaries provide the dominant proton conduction paths in LaPO$_4$. In contrast, theoretical calculations on the proton-transfer mechanism in LaPO$_4$ suggested there are proton transport paths in the LaPO$_4$ lattice with energy barrier of $\sim0.8$ eV, close to experimentally measured activation energies [24]. Additionally, the above mentioned experimental work also suggests that amorphous LaP$_3$O$_9$ has higher conductivity than that of crystalline LaP$_3$O$_9$. However, the investigation of the conduction properties of LaP$_3$O$_9$ glass revealed the opposite behavior [25]. The conductivity of the Sr-doped LaP$_3$O$_9$ glass significantly increased with the crystallization, while that of the undoped LaP$_3$O$_9$ glass remained unchanged.

Taking into account these inconsistencies, it can be said that the proton conduction mechanisms in lanthanum phosphates, especially condensed ones (LaP$_3$O$_9$ and LaP$_5$O$_{14}$) have not been understood at all. Clarifying them will be highly desired for improving the proton conductivities in lanthanum phosphates. For example, if the dominant proton transport paths are determined, it will be
possible to improve the proton conductivities of lanthanum phosphates by synthesizing samples with desired micro structures (fine- or coarse-grained polycrystalline samples or single crystals).

1.1.2 Solubility of Sr in LaPO$_4$ and LaP$_3$O$_9$

As described above, the proton conduction in lanthanum phosphates is generated by doping divalent cations such as Sr. Therefore, it will be important to make the most of the doping effect for improving the proton conductivity. However, the solubilities of Sr in lanthanum phosphates have been reported to be relatively low. Amezawa et al. derived the solubilities of Sr in LaPO$_4$ and LaP$_3$O$_9$ by X-ray diffraction analysis and conductivity measurements in combination [15,16]. The reported solubilities of Sr, represented by the molar ratio Sr/(La+Sr) are ~2 mol% in LaPO$_4$ and below 5 mol% in LaP$_3$O$_9$. As shown in Figure 1.5, the conductivities of Sr-doped LaPO$_4$ and LaP$_3$O$_9$ increase with increasing the Sr concentration within the concentration range below the solubilities. However, with higher Sr mixing ratios, secondary phases, Sr$_2$P$_2$O$_7$ and Sr(PO$_3$)$_2$, appeared and the conductivities cease to increase or begin to deteriorate. Thus, the low solubilities of Sr are an obstacle to enhancing the conductivities of lanthanum phosphates.
While the above mentioned study on LaPO$_4$ was conducted with samples prepared by solid state reaction and sintered at 1200 °C [15], there are several reports of different Sr doping levels with different synthesis temperatures (Table 1.1). Tyholdt et al. reported the solubility of Sr in LaPO$_4$ to be ~1% using samples sintered at 1450 °C [26]. Gallini et al. prepared 5% Sr-doped LaPO$_4$ by combustion synthesis followed by calcination at 800 °C [27]. Schatzmann et al. synthesized Sr-doped LaPO$_4$ by adding La(NO$_3$)$_3$·6H$_2$O and Sr(NO$_3$)$_2$ to heated phosphoric acid at 150 °C, followed by precipitation of LaPO$_4$ within 30 minutes [28]. With this method, at least 4% Sr-doped LaPO$_4$ was obtained. These results suggest that the solubility of Sr would be higher at lower temperatures. Therefore, to obtain highly Sr-doped lanthanum phosphates, the development and optimization of synthesis methods avoiding high temperature heat treatments might be beneficial.
Table 1.1 Sr-doping levels in LaPO\textsubscript{4} attained at various synthesis temperatures.

<table>
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<tr>
<th>Synthesis method</th>
<th>Max. temp. / °C</th>
<th>Sr-doping level (mol%)</th>
<th>Ref.</th>
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<td>Solid state reaction</td>
<td>1450</td>
<td>~1</td>
<td>[26]</td>
</tr>
<tr>
<td>Solid state reaction</td>
<td>1200</td>
<td>~2</td>
<td>[15]</td>
</tr>
<tr>
<td>Combustion synthesis</td>
<td>800</td>
<td>5</td>
<td>[27]</td>
</tr>
<tr>
<td>Precipitation in hot phosphoric acid</td>
<td>150</td>
<td>≥4</td>
<td>[28]</td>
</tr>
</tbody>
</table>

1.2 Phase equilibria in the La\textsubscript{2}O\textsubscript{3}–P\textsubscript{2}O\textsubscript{5} System and Stabilities of Lanthanum Phosphates

Various phosphates in the La\textsubscript{2}O\textsubscript{3}–P\textsubscript{2}O\textsubscript{5} system have been studied not only as proton conductors, but also as laser and luminescent materials [29-32], phosphorus diffusion sources for semiconductors [33-35], debond materials in oxide composites [36], coating materials [37], and radioactive waste forms [38,39]. The information on the phase equilibria in the La\textsubscript{2}O\textsubscript{3}–P\textsubscript{2}O\textsubscript{5} system serves as the basic knowledge needed for synthesizing various lanthanum phosphates as well as assessing their stabilities and reactivity with other materials. Two phase diagrams have so far been reported for the La\textsubscript{2}O\textsubscript{3}–P\textsubscript{2}O\textsubscript{5} system as shown in Figure 1.6 [40,41]. However, there are considerable differences in the compounds present. For example, one of the proton conductors referred to in the previous section, La\textsubscript{7}P\textsubscript{3}O\textsubscript{18}, is only present in the phase diagram by Park and Kreidler [40]. Therefore, further inspection should be necessary on the existance and stability of La\textsubscript{7}P\textsubscript{3}O\textsubscript{18}. In addition, LaP\textsubscript{3}O\textsubscript{9} and LaP\textsubscript{5}O\textsubscript{14} have been demonstrated to decompose gradually at high temperatures by the following reactions [10,40,42-44]:

\[
\text{LaP}_3\text{O}_9 (s) \rightarrow \text{LaPO}_4 (s) + 1/2 \text{P}_4\text{O}_{10} (g); \quad (1.4)
\]

\[
\text{LaP}_5\text{O}_{14} (s) \rightarrow \text{LaP}_3\text{O}_9 (s) + 1/2 \text{P}_4\text{O}_{10} (g). \quad (1.5)
\]

The decomposition temperatures have been reported to be ~800 °C [10] and ~700 °C [42], respectively. Because of these decompositions, it is difficult to obtain dense LaP\textsubscript{3}O\textsubscript{9} and LaP\textsubscript{5}O\textsubscript{14} electrolytes by conventional sintering method. In fact, dense LaP\textsubscript{3}O\textsubscript{9} samples with relative density...
higher than 0.9 have only been obtained by spark plasma sintering [10]. Moreover, the instability of LaP$_3$O$_9$ and LaP$_5$O$_{14}$ may limit the operating temperatures for the electrolytes, since the decomposition of the electrolytes during fuel cell operation may lead to the degradation of cell performance as well as corrosion of other components. However, very little quantitative information is available about the decomposition reactions. The only available experimental information, the vapor pressure of P$_4$O$_{10}$ ($p_{P4O10}$) over LaP$_3$O$_9$, was given by Ashuiko et al. [44] as follows:

$$\log (p_{P4O10} / \text{mmHg}) = 8.5641 - 13613 \times (T / K)^{-1}$$

$$\hspace{1cm} (1219 \leq T / K \leq 1404).$$

(1.6)

For accurate assessment of the stabilities of LaP$_3$O$_9$ and LaP$_5$O$_{14}$, further examination and refinement of their thermodynamic properties will be required.
Figure 1.6 Phase diagrams of the La$_2$O$_3$–P$_2$O$_5$ system presented by (a) Park and Kreidler [40] and (b) Kropiwnicka and Znamierowska [41]. Solid-solid phase transitions are not shown for simplification.

1.3 Synthesis Methods of Lanthanum Phosphates

As referred to in section 1.1, selection and development of appropriate synthesis methods will be a key to understanding the proton conduction mechanisms and improving the proton conductivity in lanthanum phosphates. In this work, most of lanthanum phosphate samples were prepared by
precipitation in condensed phosphoric acid solutions. Here, a brief description of the synthesis methods of rare earth phosphates is provided.

Rare earth phosphates have been synthesized via a variety of methods such as precipitation in aqueous solutions [45-47], precipitation in condensed phosphoric acid solutions [48,49], hydrothermal synthesis [50-52], solid state reaction [53,54], flux growth [55-57], and chemical transportation [58]. The solid state reaction or the precipitation in aqueous solutions followed by calcination have been widely used to synthesize lanthanum phosphate samples for investigating their conduction properties. By these methods, it is relatively easy to obtain polycrystalline lanthanum phosphates. It is also possible to obtain Sr-doped samples by adding Sr-containing compounds such as SrCO$_3$ to starting materials. However, it is difficult to achieve desired grain sizes, especially larger size by these methods. In addition, high temperature heat treatments are inevitable which may lower the solubility of Sr.

On the other hand, large single crystals of rare earth phosphates have been obtained by precipitation in condensed phosphoric acid solutions, hydrothermal synthesis, flux growth, and chemical transportation. Among them, the precipitation in condensed phosphoric acid solutions can be performed with relatively simple apparatus and at lower temperatures. It has been extensively investigated for synthesizing single crystals of rare earth ultraphosphates ($\text{REP}_5\text{O}_{14}$) for the purpose of using them as laser materials [48,49,59-63]. The common procedure of this method can be stated as follows: A mixture of rare earth oxide and phosphoric acid is heated to several hundred degrees Celsius (e.g. 300 °C) to dissolve the oxide into phosphoric acid completely; Then the solution is heated up to higher temperatures (e.g. 500 °C) to precipitate $\text{REP}_5\text{O}_{14}$. With this method, single crystals of NdP$_5$O$_{14}$ with size more than 3 cm were obtained [49]. Through related works, it was found that LaPO$_4$ and LaP$_3$O$_9$ precipitate at temperatures below ~300 °C [64-66]. According to Chudinova, however, it is difficult to obtain LaP$_3$O$_9$ because the temperature range of precipitation is only 50 °C [67]. Byrappa also stated that the process of formation of ortho- and polyphosphates
(REPO$_4$ and REP$_3$O$_9$) is very slow at temperatures lower than 260 °C and it is rather difficult to carry out the synthesis [68]. Nevertheless, there may be room for further improvement since the literature focusing on the precipitation behavior of LaPO$_4$ and LaP$_3$O$_9$ in condensed phosphoric acid solutions is still scarce.

1.4 Outline of This Work

In the above context, the synthesis conditions, thermodynamic properties and electrical conduction properties of lanthanum phosphates were studied throughout this work.

In Chapter 2, the precipitation conditions of LaPO$_4$, LaP$_3$O$_9$ and LaP$_5$O$_{14}$ in phosphoric acid solutions were investigated in both ambient and humidified atmospheres. The morphology of precipitates was analyzed in relation to the precipitation condition.

In Chapter 3, the phase equilibria in the La$_2$O$_3$–P$_2$O$_5$ system were re-examined to resolve the differences in the reported phase diagrams.

In Chapter 4, thermodynamic measurements were carried out on LaP$_3$O$_9$ and LaP$_5$O$_{14}$ by transpiration method. The thermal stabilities of LaP$_3$O$_9$ and LaP$_5$O$_{14}$ were discussed using the obtained results.

In Chapter 5, synthesis of Sr-doped LaPO$_4$ and LaP$_3$O$_9$ was carried out in phosphoric acid solutions. The optimum conditions for obtaining highly Sr-doped LaPO$_4$ and LaP$_3$O$_9$ were revealed. It was confirmed that H$_2$O is certainly incorporated in the bulk by Sr doping.

In Chapter 6, direct synthesis of dense, polycrystalline Sr-doped LaPO$_4$ and LaP$_3$O$_9$ electrolytes was performed in phosphoric acid solutions. Single crystals of Sr-doped LaP$_3$O$_9$ were also synthesized. The electrical conduction properties of the precipitates were investigated by impedance spectroscopy.

In Chapter 7, a brief summary of this study is provided.
Low-Temperature Synthesis, Thermodynamic Properties, and Electrical Conduction Properties of Lanthanum Phosphates

References


Chapter 2

Synthesis of Lanthanum Phosphates in Phosphoric Acid Solutions

2.1 Introduction

As described in Chapter 1, a new synthesis method of LaPO$_4$ and LaP$_3$O$_9$ which facilitates the control of grain size is required. Also, the synthesis and densification of the compounds should be completely done at low temperatures to avoid lowering of Sr doping level and thermal decomposition. The most suited for this purpose might be the synthesis via precipitation in condensed phosphoric acid solutions. This method has been widely used for preparing single crystals of rare earth ultraphosphates, REP$_5$O$_{14}$, as laser materials. The common procedure of this method can be described as follows: A mixture of rare earth oxide and phosphoric acid is heated to several hundred degrees Celsius (e.g. 300 °C) to dissolve the oxide into phosphoric acid completely; Then the solution is heated up to higher temperatures (e.g. 500 °C) to precipitate REP$_5$O$_{14}$. Through related works, it was found that LaPO$_4$ and LaP$_3$O$_9$ precipitate at temperatures below ~300 °C [1-3]. According to Chudinova et al., the temperature ranges of precipitation are below 160 °C for LaPO$_4$ and 180–240 °C for LaP$_3$O$_9$ [1]. They also reported that increase in gas humidity from 0.013 to 0.2 atm extends the temperature range of precipitation of LaP$_3$O$_9$ up to ~300 °C [3].

However, the literature focusing on the synthesis of LaPO$_4$ and LaP$_3$O$_9$ in condensed phosphoric acid solutions is still scarce. There has been no information on the morphologies of these precipitates. It is also of interest to know the crystallographic modification of the LaPO$_4$ precipitate. There are two polymorphs known for LaPO$_4$, namely, monazite and rhabdophane. The monazite-type
LaPO$_4$ has been investigated as a proton conductor [4]. It has been typically synthesized through high-temperature heat treatments, but the nanocrystals of monazite-type LaPO$_4$ have also been synthesized at low temperatures (such as 120 °C and 150 °C) by hydrothermal method [5,6]. Meanwhile, when synthesis was carried out in aqueous solutions at temperatures below 100 °C, the rhabdophane phase (LaPO$_{4-x}$H$_2$O) was obtained [7]. It was also reported that the rhabdophane phase was obtained by heating a mixture of La$_2$O$_3$ and H$_3$PO$_4$ (P/La = 1) at 150 °C [8].

In this chapter, the precipitation conditions of LaPO$_4$, LaP$_3$O$_9$ and LaP$_5$O$_{14}$ were investigated under both ambient and humidified atmospheres. The crystal modification of the obtained LaPO$_4$ was identified by powder X-ray diffraction analysis. The morphologies of precipitates were also analyzed in relation to the precipitation condition.

### 2.2 Principles of the Synthesis Method

When phosphoric acid is heated, it gradually loses H$_2$O and condensation of phosphate anions takes place. The condensation can be expressed by the general formula:

$$n\ H_3PO_4\ (l) \rightarrow H(HPO_3)_{n}OH\ (l) + (n-1)\ H_2O\ (g).\ \ \ \ (2.1)$$

The rate of H$_2$O loss increases with increasing temperature or decreasing the water vapor pressure. Therefore, when the water vapor pressure is constant as in ambient atmosphere, the rate of H$_2$O loss or the P$_2$O$_5$ concentration in the liquid phase is determined only by temperature. This interdependence between the P$_2$O$_5$ concentration and temperature is shown in the binary phase diagram of the H$_2$O–P$_2$O$_5$ system (Figure 2.1).
Lanthanum phosphates can be precipitated in phosphoric acid containing lanthanum oxide. Chudinova et al. investigated the phase equilibria in the La$_2$O$_3$–P$_2$O$_5$–H$_2$O system, especially those between the phosphoric acid solution and lanthanum phosphates in ambient atmosphere [1]. Just like the H$_2$O–P$_2$O$_5$ system, the P$_2$O$_5$ concentration in the solution in an open system is determined only by temperature. Therefore, the P$_2$O$_5$ concentration cannot be varied arbitrary at any constant temperature. Chudinova et al. determined the solubility of La$_2$O$_3$ in such phosphoric acid solution at each temperature. They also determined the solid phase in equilibrium with the solution. The results were presented in a Gibbs triangle in a similar manner to a conventional isothermal section of a ternary phase diagram (Figure 2.2). This diagram is actually not an isothermal section but a diagram at a constant water vapor pressure.

Figure 2.1 Phase diagram of the H$_2$O–P$_2$O$_5$ system adopted from Refs. [9,10].
Figure 2.2 Phase diagram of the La$_2$O$_3$–P$_2$O$_5$–H$_2$O system at 100–500 °C by Chudinova et al. [1] The solubility of La$_2$O$_3$ in the liquid phase at each temperature can be obtained using the temperature axis below, as shown in the figure.

The relationship between conventional isothermal diagrams and the Chudinova’s phase diagram (Figure 2.2) can be explained as follows. Figure 2.3 is a schematic illustration of the isothermal sections of the La$_2$O$_3$–P$_2$O$_5$–H$_2$O phase diagram at 150 °C and 200 °C. The broad curves in the figure represent the compositions of phosphoric acid solutions in equilibrium with water vapor in the ambient atmosphere. As the solubility of La$_2$O$_3$ in the phosphoric acid solution is on the order of only a few weight percent, the P$_2$O$_5$ concentration in the solution is assumed almost equal to that given for the H$_2$O–P$_2$O$_5$ binary system (dashed line in Figure 2.1). “Point A” in the figure represents the composition of such phosphoric acid solution in equilibrium with a solid phase (one of lanthanum phosphates). This point represents the solubility of La$_2$O$_3$ in the solution at the given temperature. The position of Point A changes with temperature and its trajectory can be expressed as a curve in the temperature–composition space (Figure 2.4). By projecting the curve on the Gibbs triangle, a
diagram is obtained which provides the solubility of La$_2$O$_3$ in the solution in a wide temperature range but at a constant water vapor pressure (Figure 2.2). This type of diagram is suitable for the crystal growth of rare earth phosphates in ambient atmosphere and has been reported for several systems [1,11,12]. It should be noted that the isothermal sections of the La$_2$O$_3$–P$_2$O$_5$–H$_2$O phase diagram presented in Figure 2.3 is drawn on the assumption that the phase equilibria between the liquid phase and solid phases do not change drastically with temperature. It has proved to be true for a similar system, Nd$_2$O$_3$–P$_2$O$_5$–H$_2$O system [12-14].

![Phase diagram of the La$_2$O$_3$–P$_2$O$_5$–H$_2$O system at 150 and 200 °C. Point A represents the three phases equilibrium between the solution, solid phase, and gas phase ($p_{H2O} ~ 0.01$ atm), or the solubility of La$_2$O$_3$ in the solution.](image)

Figure 2.3 Schematic illustration of isothermal sections of the phase diagram of the La$_2$O$_3$–P$_2$O$_5$–H$_2$O system at 150 and 200 °C. Point A represents the three phases equilibrium between the solution, solid phase, and gas phase ($p_{H2O} ~ 0.01$ atm), or the solubility of La$_2$O$_3$ in the solution.
Based on Figure 2.2, the solubility of La$_2$O$_3$ in the phosphoric acid solution is found to be highest at ~190 °C. At lower and higher temperatures, the solubility of La$_2$O$_3$ decreases and several lanthanum phosphates can be precipitated: LaPO$_4$ at lower temperatures, and LaP$_3$O$_9$ and LaP$_5$O$_{14}$ at higher temperatures. Therefore, the dissolution temperature of La$_2$O$_3$ in phosphoric acid before the precipitation process was set at 190 °C in this work.

As mentioned at the beginning of this section, change in water vapor pressure also affects the composition of the phosphoric acid solution. As increasing water vapor pressure decreases the P$_2$O$_5$ concentration in the solution, it increases the temperature ranges of the precipitation of each phosphate from those in ambient atmosphere. This effect is so far confirmed by Balagina et al. for LaP$_3$O$_9$ [3]. According to Byrappa, the process of formation of ortho- and polyphosphates (REPO$_4$ and REP$_3$O$_9$) is very slow at temperatures lower than 260 °C and it is rather difficult to carry out the synthesis [15]. However, if sufficient humidification is applied, synthesis of these phosphates would be easier. In this work, the effect of humidification on the precipitation behaviors of lanthanum...
phosphates was also investigated.

2.3 Experimental

2.3.1 Synthesis

LaPO$_4$, LaP$_3$O$_9$, and LaP$_5$O$_{14}$ were synthesized by precipitation in homogeneous phosphoric acid solutions containing La$_2$O$_3$. La$_2$O$_3$ (99.99%, Nacalai Tesque) and H$_3$PO$_4$ (85%, Nacalai Tesque) were mixed at a P/La molar ratio of 15 in a perfluoroalkoxy (PFA) or glassy carbon (GC) crucible and held at 190 °C in a furnace for several days to obtain a transparent solution. Then temperature and water vapor pressure were changed and held for a predetermined period of time to precipitate the lanthanum phosphates. The temperature was set between 120 °C and 350 °C, and the water vapor pressure was set to approximately 0.01 atm (ambient atmosphere) or 1 atm (humidified atmosphere). Synthesis parameters are summarized in Table 2.1.

Table 2.1 Dissolution and precipitation conditions for the synthesis of LaPO$_4$, LaP$_3$O$_9$, and LaP$_5$O$_{14}$.

<table>
<thead>
<tr>
<th>Abbr.</th>
<th>Dissolution</th>
<th>Precipitation</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>Temp. / °C</td>
<td>Temp. / °C</td>
</tr>
<tr>
<td></td>
<td>$p_{H_2O}$ / atm</td>
<td>$p_{H_2O}$ / atm</td>
</tr>
<tr>
<td>A300</td>
<td>300</td>
<td>~0.01</td>
</tr>
<tr>
<td>A230</td>
<td>230</td>
<td>~0.01</td>
</tr>
<tr>
<td>A140</td>
<td>140</td>
<td>~0.01</td>
</tr>
<tr>
<td>H350</td>
<td>190</td>
<td>~0.01</td>
</tr>
<tr>
<td>H250</td>
<td>350</td>
<td>~1</td>
</tr>
<tr>
<td>H140</td>
<td>250</td>
<td>~1</td>
</tr>
</tbody>
</table>
For humidification, the apparatus shown in Figure 2.5 was used. The whole apparatus was placed in a furnace kept at a constant temperature. A crucible was put on a stainless steel floor plate and covered with a stainless steel box with a capacity of 2.64 L. A 4-m-long spiral tube was attached to a side wall of the box. Liquid water was fed to the box through the spiral tube so that it evaporated before reaching the box. The liquid water feed rate was 0.1 mL/min, which was equivalent to 0.19 L-H$_2$O(g)/min at 140 °C and 0.28 L-H$_2$O(g)/min at 350 °C. The apparatus was not sealed tightly.

Figure 2.5 (a) Schematic illustration of the apparatus used for humidification during the precipitation of lanthanum phosphates. The thermocouple and ribbon heater were used only for H350 condition. (b) Photo of the stainless steel box and floor plate. (c) Photo of the stainless steel box and floor plate (without the ribbon heater) placed in the furnace and connected with a metering pump on the right.
so that the inner gas was able to flow out through the narrow gap between the box and the floor plate when the inner pressure exceeded 1 atm. Therefore, the water vapour pressure inside the box was expected to be 1 atm. For H350 condition, a ribbon heater wrapped around the stainless steel box was used for heating the solution up to 350 °C while the temperature in the furnace was kept at 270 °C. The temperature inside the stainless steel box was measured by a thermocouple attached to the upper surface of the floor plate.

After the predetermined precipitation period, the crucible was cooled down to room temperature. In general, there were precipitates on the bottom of the crucible and the transparent solution above them. The precipitates were washed well with hot water (~90 °C).

2.3.2 Characterization

Phase identification was carried out via X-ray powder diffraction (XRD) analysis on PANalytical X’Pert-Pro MPD using Cu Kα radiation at room temperature. The morphology of precipitates was investigated using KEYENCE VE-7800 scanning electron microscopy (SEM).

2.4 Results and Discussion

The precipitation temperatures of lanthanum phosphates were investigated under both atmospheric and humidified conditions with the synthesis parameters shown in Table 2.1. In all of the conditions (A300, A230, A140, H350, H250, and H140), colorless precipitates were obtained at the bottom of the crucibles. Figure 2.6 shows the X-ray diffraction patterns of the precipitates. Under the atmospheric conditions, LaP$_5$O$_{14}$, LaP$_3$O$_9$, and LaPO$_4$ were obtained at 300 °C (A300), 230 °C (A230), and 140 °C (A140), respectively. This is consistent with the results reported by Chudinova et al. [1] The obtained LaPO$_4$ had the monazite-type structure, which has been reported to exhibit proton conduction. Under the humidified conditions, LaP$_3$O$_9$ was obtained at 350 °C (H350) while monazite-type LaPO$_4$ was obtained at 250 °C and 140 °C. With these results, the precipitation region...
of LaPO₄ was drawn on the temperature–pH₂O plane (Figure 2.7). As indicated in the figure, increasing water vapor pressure from 0.01 to 1 atm extends the precipitation regions of LaPO₄ and LaP₃O₉ toward higher temperatures by nearly 100 °C. At pH₂O = 1 atm, LaPO₄ can be obtained even at 250 °C, while LaP₃O₉ will be obtained at lower water vapor pressures such as 0.2 atm. Similarly, LaP₃O₉ can be obtained at 350 °C and pH₂O = 1 atm while LaP₅O₁₄ will be obtained in lower humidification levels.

![Figure 2.6 X-Ray powder diffraction patterns of lanthanum phosphate precipitates obtained under the conditions (a) A300, (b) A230, (c) A140, (d) H350, (e) H250, and (f) H140.](image-url)
Figure 2.7 Phase stability diagram of the La$_2$O$_3$–P$_2$O$_5$–H$_2$O system with a constant La/P ratio (La : P = 1 : 15). In addition to the results obtained in this research, previous results reported by Chudinova et al. [1] and Balagina et al. [3] are also shown.

Figure 2.8 shows SEM images of LaPO$_4$ and LaP$_3$O$_9$ precipitates. Under the conditions A140 and H140, LaPO$_4$ precipitates were rice-shaped particles with a size range of 500 nm–1 µm. In contrast, under H250 condition, the LaPO$_4$ precipitates were aggregates of well-shaped crystals up to 200 µm. Thus, the precipitation temperature had a marked effect on the crystal growth of LaPO$_4$. As for LaP$_3$O$_9$, crystals of ~100 µm were obtained at 230 °C (A230) while those of ~300 µm were obtained at 350 °C (A350). Therefore, the crystal growth of LaP$_3$O$_9$ was also accelerated by increasing the precipitation temperature, though the effect was not as drastic as that on LaPO$_4$. 
2.5 Conclusions

In this chapter, the precipitation conditions of LaPO₄, LaP₃O₉, and LaP₅O₁₄ in condensed phosphoric acid solutions were established under both atmospheric and humidified conditions. The obtained LaPO₄ had the monazite-type structure, which has been reported to exhibit proton conduction. Increasing water vapor pressure from 0.01 atm to 1 atm extended the precipitation regions of LaPO₄ and LaP₃O₉ toward higher temperatures by nearly 100 °C. The crystal growth of
LaPO$_4$ was significantly accelerated under such conditions (e.g. $T=250 \, ^\circ C$ and $p_{H_2O} = 1 \text{ atm}$). The crystal growth of LaP$_3$O$_9$ was also promoted moderately by the increase in precipitation temperature. These results will serve as fundamental information for further optimization of growth parameters (e.g. temperature profile, precipitation period, and initial composition) and hopefully the direct synthesis of dense lanthanum phosphate electrolytes in phosphoric acid solutions.

References


274-290.


Chapter 3
Phase Equilibria in the La$_2$O$_3$–P$_2$O$_5$ System

3.1 Introduction

Various phosphates in the La$_2$O$_3$–P$_2$O$_5$ system have been studied not only as proton conductors, but also as laser and luminescent materials [1-4], phosphorus diffusion sources for semiconductors [5-7], debond materials in oxide composites [8], coating materials [9], and radioactive waste forms [10,11]. As written in Chapter 1, two phase diagrams have so far been presented for the La$_2$O$_3$–P$_2$O$_5$ system in the literature [12,13]. However, there are considerable differences in the compounds present and they may bring some confusion. For example, La$_7$P$_3$O$_{18}$, one of the proton conductors referred to in the previous chapter, is only present in the phase diagram by Park and Kreidler [12]. Harley et al. [14] adopted the La-rich part from Ref. [13] and P-rich part from Ref. [12] for some reason to show the La$_2$O$_3$–P$_2$O$_5$ phase diagram in their paper.

As shown in Table 3.1, the phase diagrams contain four intermediate compounds in common: La$_3$PO$_7$, LaPO$_4$, LaP$_3$O$_9$, and LaP$_5$O$_{14}$. In addition, Park and Kreidler [12] reported two compounds La$_7$P$_3$O$_{18}$ and La$_2$P$_4$O$_{13}$, and Kropiwnicka and Znamierowska [13] reported La$_5$PO$_{10}$. Therefore, further inspection should be necessary on the existence and stability of these three compounds. Here, a brief description is given on these compounds.
Table 3.1 Compounds in the La$_2$O$_3$–P$_2$O$_5$ phase diagrams in the literature. They are listed in increasing order of P$_2$O$_5$ content.

<table>
<thead>
<tr>
<th></th>
<th></th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td>La$<em>5$PO$</em>{10}$</td>
<td>X</td>
<td></td>
</tr>
<tr>
<td>La$_3$PO$_7$</td>
<td>X</td>
<td>X</td>
</tr>
<tr>
<td>La$_7$P$<em>3$O$</em>{18}$</td>
<td>X</td>
<td></td>
</tr>
<tr>
<td>LaPO$_4$</td>
<td>X</td>
<td>X</td>
</tr>
<tr>
<td>La$_2$P$<em>4$O$</em>{13}$</td>
<td>X</td>
<td></td>
</tr>
<tr>
<td>LaP$_5$O$_9$</td>
<td>X</td>
<td>X</td>
</tr>
<tr>
<td>LaP$<em>5$O$</em>{14}$</td>
<td>X</td>
<td>X</td>
</tr>
</tbody>
</table>

3.1.1 La$_5$PO$_{10}$

La$_5$PO$_{10}$ has a composition between La$_2$O$_3$ and La$_3$PO$_7$. La$_5$PO$_{10}$ was obtained by Kropiwnicka and Znamierowska [13,15] via solid state reaction between La$_2$O$_3$ and NH$_4$H$_2$PO$_4$ at temperatures above 540 °C. They determined its congruent melting temperature to be 1870 ± 10 °C. Meanwhile, Tselebrovskaya et al. [16] presented a phase diagram of the La$_2$O$_3$–LaPO$_4$–La$_2$GeO$_5$ system at 1300 °C, in which La$_5$PO$_{10}$ was absent. El Ouenzerfi et al. [17] also observed that La$_2$O$_3$ and La$_3$PO$_7$ were in equilibrium at 1200–1400 °C, instead of forming La$_5$PO$_{10}$.

3.1.2 La$_7$P$_3$O$_{18}$

La$_7$P$_3$O$_{18}$ has a composition between La$_3$PO$_7$ and LaPO$_4$. La$_7$P$_3$O$_{18}$ has been synthesized by several researchers via solid state reactions and solidification of melts [12,18-22]. Serra et al. [20] observed an α-β phase transition at 1650 °C. According to Park and Kreidler [12], there is no evidence of a lower limit of stability. However, La$_7$P$_3$O$_{18}$ is obtainable only at temperatures above 1400 °C and it has not been synthesized at temperatures below 1200 °C [18,19,21,22]. The coexistence of La$_3$PO$_7$ and LaPO$_4$ at 1100–1200 °C was also observed by other authors [23,24]. Tselebrovskaya et al. attributed this to kinetic reasons [16,22].
3.1.3 \( \text{La}_2\text{P}_4\text{O}_{13} \)

\( \text{La}_2\text{P}_4\text{O}_{13} \) has a composition between \( \text{LaPO}_4 \) and \( \text{LaP}_3\text{O}_9 \). Park and Kreidler obtained \( \text{La}_2\text{P}_4\text{O}_{13} \) by crystallization from glass [12]. They found that it is stable at temperatures up to 755 °C. However, \( \text{La}_2\text{P}_4\text{O}_{13} \) has not been obtained by other methods even at those temperatures [25-27]. The same phenomenon was found for the neodymium analogue, \( \text{Nd}_2\text{P}_4\text{O}_{13} \), by Wong and Kreidler [27]: It was formed by crystallization of a glass, but not formed by direct solid-state reaction. They speculated that \( \text{Nd}_2\text{P}_4\text{O}_{13} \) exists in a metastable state.

3.2 Experimental

3.2.1 Sample Preparation

\( \text{La}_2\text{O}_3 \) (Nacalai Tesque, mass fraction purity 0.9999) and \( \text{H}_3\text{PO}_4 \) (Nacalai Tesque, mass fraction purity 0.85) were used to prepare samples. \( \text{La}_2\text{O}_3 \) was heated overnight at 800 °C prior to use to remove absorbed \( \text{H}_2\text{O} \) and \( \text{CO}_2 \) [28]. \( \text{LaPO}_4 \) and \( \text{LaP}_3\text{O}_9 \) were synthesized by precipitation in phosphoric acid solutions, using the procedure given in Chapter 2. The \( \text{LaPO}_4 \) precipitates were dried at 800 °C for 100 h. For preparing samples with compositions between \( \text{La}_2\text{O}_3 \) and \( \text{LaPO}_4 \) (“L9P1”, “L8P2”, “L725P275”, “L7P3”, and “L6P4”), the \( \text{La}_2\text{O}_3 \) and \( \text{LaPO}_4 \) powders were mixed in the molar ratios listed in Table 3.2. The mixtures were then ballmilled for 70–100 h in isopropyl alcohol, dried, and pressed into pellets at ~150 MPa. The sample with a composition slightly poorer in \( \text{P}_2\text{O}_5 \) than \( \text{LaP}_3\text{O}_9 \), “L25P75–δ”, was prepared from \( \text{LaP}_3\text{O}_9 \). The \( \text{LaP}_3\text{O}_9 \) precipitate was ground and pressed into a pellet at ~400 MPa. The pellet was heated at 1050 °C for 3 h, to be partially decomposed according to the following reaction [12]:

\[
\text{LaP}_3\text{O}_9 \text{ (s) } \rightarrow \text{LaPO}_4 \text{ (s) } + \frac{1}{2} \text{P}_4\text{O}_{10} \text{ (g)}. \quad (3.1)
\]
Table 3.2 Compositions of samples.

<table>
<thead>
<tr>
<th>Abbr.</th>
<th>Starting materials</th>
<th>Molar ratio</th>
<th>Remarks</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td>La : P</td>
<td></td>
</tr>
<tr>
<td>L9P1</td>
<td>La$_2$O$_3$ and LaPO$_4$</td>
<td>0.9 : 0.1</td>
<td>Between La$_2$O$_3$ and “La$<em>5$PO$</em>{10}$”</td>
</tr>
<tr>
<td>L8P2</td>
<td>La$_2$O$_3$ and LaPO$_4$</td>
<td>0.8 : 0.2</td>
<td>Between “La$<em>3$PO$</em>{10}$” and La$_3$PO$_7$</td>
</tr>
<tr>
<td>L725P275</td>
<td>La$_2$O$_3$ and LaPO$_4$</td>
<td>0.725 : 0.275</td>
<td>Between La$_3$PO$_7$ and “La$_7$P$<em>3$O$</em>{18}$”</td>
</tr>
<tr>
<td>L7P3</td>
<td>La$_2$O$_3$ and LaPO$_4$</td>
<td>0.7 : 0.3</td>
<td>Composition of “La$_7$P$<em>3$O$</em>{18}$”</td>
</tr>
<tr>
<td>L6P4</td>
<td>La$_2$O$_3$ and LaPO$_4$</td>
<td>0.6 : 0.4</td>
<td>Between “La$_7$P$<em>3$O$</em>{18}$” and LaPO$_4$</td>
</tr>
<tr>
<td>L25P75−δ</td>
<td>LaP$_3$O$_9$</td>
<td>0.25 : 0.75−δ</td>
<td>Slightly poorer in P$_2$O$_5$ than LaP$_3$O$_9$</td>
</tr>
</tbody>
</table>

3.2.2 Phase Equilibrium Experiments and Sample Characterization

Phase equilibrium experiments were carried by high-temperature heat treatments and subsequent phase identification. The sample pellets were placed in alumina crucibles, and heated at a predetermined temperature for typically 100 h in ambient atmosphere. Then the samples were cooled down in the furnace. The specific conditions for each heat treatment are provided in the following section. Phase identification was carried out via X-ray powder diffraction (XRD) analysis on PANalytical X’Pert-Pro MPD using Cu Kα radiation at room temperature. The microstructure of samples was analyzed using KEYENCE VE-7800 scanning electron microscopy (SEM).

3.3 Results and Discussion

3.3.1 La$_2$O$_3$–LaPO$_4$ Subsystem at 1000 and 1300 °C

First, the phase relationships in the La$_2$O$_3$–LaPO$_4$ subsystem were investigated at 1000 °C and 1300 °C using samples L9P1, L8P2, L725P275, and L6P4. The samples were heated at each temperature for 100 h. Figure 3.1 and Figure 3.2 show the X-ray diffraction patterns of the samples after each heat treatment. The detected phases in each sample are summarized in Table 3.3. At both temperatures and in all of the samples, reactions between LaPO$_4$ and La$_2$O$_3$ occurred and intermediate phases were formed. The observed intermediate phases were La$_3$PO$_7$ at 1000 °C,
La$_3$PO$_7$ and La$_7$P$_3$O$_{18}$ at 1300 °C. La$_3$PO$_{10}$ was not detected at either temperature.

As for La$_3$PO$_7$, Park and Kreidler observed an α-β phase transition at 935 ± 5 °C [12]. They reported the X-ray diffraction patterns of both low-temperature α-phase and high-temperature β-phase. There have also been a few slightly different X-ray diffraction data reported for La$_3$PO$_7$ [20,29]. They are very similar, and it is rather difficult to distinguish them, but the present X-ray diffraction results are best fitted by the β-phase pattern given by Park and Kreidler.

Table 3.3 Phases present in samples L9P1, L8P2, L725P275, and L6P4 before and after heat treatments at 1000 °C or 1300 °C for 100h, as analyzed by powder X-ray diffraction.

<table>
<thead>
<tr>
<th>Sample</th>
<th>Phases Before heat treatments</th>
<th>After heat treatments 1000 °C, 100 h</th>
<th>After heat treatments 1300 °C, 100 h</th>
</tr>
</thead>
<tbody>
<tr>
<td>L9P1</td>
<td>(Not investigated)</td>
<td>La$_3$O$_3^a$, La$_3$PO$_7$</td>
<td></td>
</tr>
<tr>
<td>L8P2</td>
<td>La$_2$O$_3$, LaPO$_4$</td>
<td>La$_2$O$_3^a$, La$_3$PO$_7$</td>
<td>La$_2$O$_3$, La$_3$PO$_7$</td>
</tr>
<tr>
<td>L725P275</td>
<td>La$_3$PO$_7$, LaPO$_4$</td>
<td>La$_3$PO$_7$, LaPO$_4$</td>
<td>La$_3$PO$_7$, La$_7$P$<em>3$O$</em>{18}$</td>
</tr>
<tr>
<td>L6P4</td>
<td>La$_3$PO$_7$, LaPO$_4$</td>
<td></td>
<td>La$_7$P$<em>3$O$</em>{18}$, LaPO$_4$</td>
</tr>
</tbody>
</table>

$^a$ Although La(OH)$_3$ was detected, it should have formed by hydration of La$_2$O$_3$ at room temperature after the heat treatments. Therefore, La(OH)$_3$ is expressed as “La$_2$O$_3$” in the table.
Figure 3.1 X-ray diffraction patterns of (a) L9P1 and (b)(c) L8P2 after heat treatments at (a)(b) 1300 °C or (c) 1000 °C for 100 h.
Figure 3.2 X-ray diffraction patterns of (a)(b) L725P275 and (c)(d) L6P4 after heat treatments at (a)(c) 1300 °C or (b)(d) 1000 °C for 100 h.

3.3.2 Phase Formation Process at the Composition La:P = 7:3

Since it was found in the previous subsection that La$_7$P$_3$O$_{18}$ was formed at 1300 °C while it was not formed at 1000 °C, further investigation was carried on the phase formation process at the composition La:P = 7:3. L7P3 samples (see Table 3.2) were heat treated in the temperature range...
700–1300 °C with various stepwise temperature programs shown in Figure 3.3. Phase identification and SEM observation were performed after the heat treatments.

![Figure 3.3](image)

**Figure 3.3** Temperature profiles for the heat treatments of L7P3. Occasional cooling to room temperature is not shown in the figure. Overlapping lines are slightly shifted for better visibility.

Figure 3.4 shows the X-ray diffraction patterns of L7P3 samples after various heat treatments (Temperature profile A–C in Figure 3.3). The sample held at 700 °C for 100 h (Temperature profile A) consisted of the same phases as the starting materials, La$_2$O$_3$ and LaPO$_4$, suggesting that almost no reaction occurred during the heat treatment. Weak peaks of La$_2$O$_2$CO$_3$ were also observed, but it could be because La$_2$O$_3$ reacted with CO$_2$ in the atmosphere during or after the heat treatment. On the other hand, in the sample held at 1000 °C for 100 h (Temperature profile B), La$_3$PO$_7$ and LaPO$_4$ were identified, but La$_2$O$_3$ was not observed. It indicates that the following reaction proceeded completely at 1000 °C:

$$\text{La}_2\text{O}_3 + \text{LaPO}_4 \rightarrow \text{La}_3\text{PO}_7 \text{ (700–1000 °C).}$$ (3.2)

An additional heat treatment of the sample at 1200 °C for 100 h did not have any effect on the phases present. With further heat treatment at 1250 °C for 100 h, essentially single phase La$_7$P$_3$O$_{18}$ was
formed by the following reaction:

\[ \text{La}_3\text{PO}_7 + \text{LaPO}_4 \rightarrow \text{La}_7\text{P}_3\text{O}_{18} (~1250 \degree \text{C}). \]  

(3.3)

A small amount of LaPO$_4$ still remained in the sample, but it could be due to some compositional error in preparing the initial mixture. La$_7$P$_3$O$_{18}$ was also obtained by direct heating of another L7P3 sample to 1300 °C for 30 h (Temperature profile C).

To examine the kinetic effects on the formation process of La$_7$P$_3$O$_{18}$, long-term heat treatments of a L7P3 sample were performed at temperatures lower than 1250 °C with the temperature profile D (Figure 3.3). Figure 3.5 shows the X-ray diffraction patterns of the L7P3 sample during the heat treatments. Heating the sample at 1000 °C for 1000 h and at 1100 °C for 200 h did not result in the formation of La$_7$P$_3$O$_{18}$, consistent with previous results. However, after the subsequent heating at 1200 °C for 100 h, partial formation of La$_7$P$_3$O$_{18}$ from La$_3$PO$_7$ and LaPO$_4$ was observed. Comparing with the result obtained through the temperature profile B, it could be said that the long-term heat treatments at 1000–1100 °C slightly lowered the experimental formation temperature of La$_7$P$_3$O$_{18}$ to 1200 °C.

The kinetics of the formation process was also examined by SEM observation. Figure 3.6 shows typical SEM images of L7P3 samples before and after various heat treatments. The particle size of the as-ballmilled powder, shown in Figure 3.6 (a), was several hundred nm. It was not changed by the heat treatment at 700 °C, as shown in Figure 3.6 (b). It is consistent with the X-ray diffraction result which suggests no reaction occurred at 700 °C. On the other hand, larger and smoother grains were observed in the samples heated at 1000–1300 °C, at which La$_3$PO$_7$ and La$_7$P$_3$O$_{18}$ can be formed (Figure 3.6 (c) and (d)). It indicates that diffusion of atoms is possible at such temperatures. It should be noted that the grains in the sample heated at 1300 °C for only 30 h (single phase La$_7$P$_3$O$_{18}$) were even larger than those in the sample heated at 1000–1200 °C for a total of 1300 h (La$_7$P$_3$O$_{18}$ with residual La$_3$PO$_7$ and LaPO$_4$). Therefore, such temperature difference should have a significant effect on the rate of grain growth and diffusion of atoms. Considering the results of X-ray diffraction and
SEM observation together, there is a possibility that the formation process of $\text{La}_7\text{P}_3\text{O}_{18}$ is significantly governed by kinetics at temperatures below 1300 °C even if it is thermodynamically stable. The kinetic effects can further be characterized if smaller initial particles are used.
Figure 3.4 X-ray diffraction patterns of L7P3 after various heat treatments indicated as Profiles A-C in Figure 3.3. (a) After a heat treatment at 700 °C for 100 h (Profile A); (b) After the first heat treatment at 1000 °C for 100 h (Profile B); (c) After the second heat treatment at 1200 °C for 100 h (Profile B); (d) After the third heat treatment at 1250 °C for 100 h (Profile B); (e) After a heat treatment at 1300 °C for 30 h (Profile C).
Figure 3.5 X-ray diffraction patterns of L7P3 after the heat treatments indicated as Profile D in Figure 3.3. (a) After heat treatments as 1000 °C for a total of 1000 h; (b) After additional heat treatments at 1100 °C for a total of 200h; (c) After an additional heat treatment at 1200 °C for 100 h.
Figure 3.6 SEM images of L7P3. (a) As ballmilled powder; (b) After a heat treatment at 700 °C for 100 h (Temperature profile A in Figure 3.3); (c) After successive heat treatments at 1000 °C for 1000 h, at 1100 °C for 200 h, and at 1200 °C for 100 h (Temperature profile D in Figure 3.3); (d) After a heat treatment at 1300 °C for 30 h (Temperature profile C in Figure 3.3). The phases present in the samples are shown on each image.

3.3.3 Low Temperature Stability of La$_7$P$_3$O$_{18}$

Besides the kinetic reasons discussed above, another possible reason why La$_7$P$_3$O$_{18}$ forms only at 1200 °C or higher is that La$_7$P$_3$O$_{18}$ may be thermodynamically unstable at low temperatures. To
examine this possibility, La$_7$P$_3$O$_{18}$ synthesized at 1300 °C was subjected to long-term heat treatments at 1000, 1100, and 1200 °C with stepwise temperature profiles shown in Figure 3.7.

![Figure 3.7 Temperature profiles for the heat treatments of La$_7$P$_3$O$_{18}$. Occasional cooling to room temperature is not shown in the figure. Overlapping lines are slightly shifted for better visibility.](image)

Figure 3.7 Temperature profiles for the heat treatments of La$_7$P$_3$O$_{18}$. Occasional cooling to room temperature is not shown in the figure. Overlapping lines are slightly shifted for better visibility.

Figure 3.8 shows the X-ray diffraction patterns of La$_7$P$_3$O$_{18}$ obtained after each heat treatment. As can be seen in the figure, all the patterns are essentially the same and representing that the samples consisted of La$_7$P$_3$O$_{18}$ in large part and a slight amount of La$_3$PO$_7$. No sign of the decomposition of La$_7$P$_3$O$_{18}$, the reverse reaction of Eq. (3.3), was detected. Therefore, it can be said that La$_7$P$_3$O$_{18}$, once formed at high temperatures, is practically stable at temperatures 1000–1200 °C. However, the thermodynamically stable phase(s) at these temperatures are still unknown and might be revealed by some thermodynamic measurements in the future. Figure 3.9 summarizes the results of the phase equilibrium experiments with L7P3 samples performed in the previous and this subsections.
Figure 3.8 X-ray diffraction patterns of La$_7$P$_3$O$_{18}$ during heat treatments with temperature profiles indicated in Figure 3.7. (a) After the first heat treatment at 1300 °C for 100 h; (b) After the heat treatment at 1200 °C for 100 h (Profile E); (c) After the heat treatments at 1000 °C for a total of 900 h (Profile F). (d) After the heat treatments at 1100 °C for a total of 200 h (Profile F).
3.4 LaPO₄–LaP₃O₉ Subsystem

The phase relationships between LaPO₄ and LaP₃O₉ were examined using the sample L25P75−δ. As described in 3.1.3, the existence of an intermediate compound, La₂P₄O₁₃, has been reported between LaPO₄ and LaP₃O₉, with the decomposition temperature of 755 °C. To verify whether it is formed or not, L25P75−δ was heat treated at 700 °C for 100 h.

Figure 3.10 shows the X-ray diffraction patterns of L25P75−δ before and after the heat treatment. As can be seen in the figure, both patterns are composed of the peaks of LaPO₄ and LaP₃O₉. No other phase such as La₂P₄O₁₃ was identified. So far La₂P₄O₁₃ has been obtained only by crystallization from glass and it was reported to decompose into LaPO₄ and LaP₃O₉ at 755 °C [12]. Wong and Kreidler mentioned that La₂P₄O₁₃ and Nd₂P₄O₁₃ could not be formed by solid state reaction [27]. They speculated that Nd₂P₄O₁₃ exists in a metastable state. The present result implies that La₂P₄O₁₃ is also a metastable phase.
Figure 3.10 X-ray diffraction patterns of L25P75–δ (a) before and (b) after a heat treatment at 700 °C for 100 h.

3.4 Conclusions

In this chapter, the phase equilibria in the La₂O₃–P₂O₅ system were re-examined to verify the known phase diagrams. Two intermediate phases, La₃PO₇ and La₇P₃O₁₈, were identified in the La₂O₃–LaPO₄ subsystem. La₅PO₁₀ was not observed. Therefore, the phase diagram reported by Park and Kreidler seems more likely to be accurate in that. La₃PO₇ was formed at 1000 °C, while La₇P₃O₁₈ was only formed at 1200 °C and higher. At 1000 °C, neither the formation nor decomposition of La₇P₃O₁₈ took place within 1000 h. This behavior might be due to kinetic reasons, but it was not verified. In the LaPO₄–LaP₃O₉ subsystem, no intermediate phase such as La₂P₄O₁₃ was detected. La₂P₄O₁₃ might be a metastable phase. Based on these results, the phase diagram reported by Park and Kreidler is slightly modified and presented in Figure 3.11. The diagram shows the phases obtainable by solid state reaction, and those possibly in metastable phases are not shown.
Figure 3.11 Proposed phase diagram of the La$_2$O$_3$–P$_2$O$_5$ system.

References


473-474.


Chapter 4
Thermodynamic Properties of LaP$_3$O$_9$ and LaP$_5$O$_{14}$

4.1 Introduction

LaP$_3$O$_9$ and LaP$_5$O$_{14}$ have been demonstrated to decompose gradually at high temperatures by the following reactions [1-5]:

\[
\text{LaP}_3\text{O}_9 (s) \rightarrow \text{LaPO}_4 (s) + \frac{1}{2} \text{P}_4\text{O}_{10} (g); \quad (4.1)
\]

\[
\text{LaP}_5\text{O}_{14} (s) \rightarrow \text{LaP}_3\text{O}_9 (s) + \frac{1}{2} \text{P}_4\text{O}_{10} (g). \quad (4.2)
\]

The decomposition temperatures have been reported to be \(\sim 800 \, ^\circ\text{C} \) [1] and \(\sim 700 \, ^\circ\text{C} \) [3], respectively. The vapor produced by these reactions can be utilized for phosphorus doping of silicon wafers in the semiconductor industry [6-8]. Meanwhile, because of these decompositions, it is difficult to obtain dense LaP$_3$O$_9$ and LaP$_5$O$_{14}$ electrolytes by conventional sintering method. In fact, dense LaP$_3$O$_9$ samples with relative density higher than 0.9 have only been obtained by spark plasma sintering [1]. Moreover, the instability of LaP$_3$O$_9$ and LaP$_5$O$_{14}$ may limit the operating temperatures for the electrolytes, since the decomposition of the electrolytes during fuel cell operation may lead to the degradation of cell performance as well as corrosion of other components. However, very little quantitative information is available about the decomposition reactions. The only available experimental information, the vapor pressure of P$_4$O$_{10}$ (\(p_{\text{P}_4\text{O}_{10}}\)) over LaP$_3$O$_9$, was given by Ashuiko et al. [5] as follows:

\[
\log (p_{\text{P}_4\text{O}_{10}} / \text{mmHg}) = 8.5641 - 13613 \times (T / \text{K})^{-1} \\
(1219 \leq T / \text{K} \leq 1404). \quad (4.3)
\]

For accurate assessment of the stabilities of LaP$_3$O$_9$ and LaP$_5$O$_{14}$, further examination and refinement
of their thermodynamic properties will be necessary. In this chapter, vapor pressure measurements on LaP$_3$O$_9$ and LaP$_5$O$_{14}$ were carried out by the transpiration method [9].

4.2 Experimental

4.2.1 Sample Preparation and Phase Identification

LaP$_3$O$_9$ and LaP$_5$O$_{14}$ were synthesized by precipitation in phosphoric acid solutions [10,11]. La$_2$O$_3$ (Nacalai Tesque, mass fraction purity 0.9999) and H$_3$PO$_4$ (Nacalai Tesque, mass fraction purity 0.85) were mixed at a P/La molar ratio of 15 and kept at 463 K in the air for several days to obtain a transparent solution. Then the solution was kept at 503 or 573 K to precipitate LaP$_3$O$_9$ and LaP$_5$O$_{14}$. The precipitates were separated from the solution, washed well with hot water and dried at ~90 °C. Then they were ground with a mortar and pestle, and pressed into pellets with weight of ~300 mg and relative density of 0.50–0.75. The LaP$_3$O$_9$ and LaP$_5$O$_{14}$ pellets were heat-treated at 973 K (700 °C) and 873 K (600 °C), respectively, for 50 h in the air to eliminate residual water and phosphoric acid on the surface of samples before further experiments. Phase identification was carried out via X-ray powder diffraction (XRD) analysis on PANalytical X’Pert-Pro MPD using Cu Kα radiation.

4.2.2 Vapor Pressure Measurements

4.2.2.1 Apparatus and Procedure

The apparatus used for vapor pressure measurements is shown in Figure 4.1. The sample was placed in a platinum boat of length 110 mm and inserted into a quartz tube of diameter 25 mm. The sample was rapidly heated to the desired temperature (923–1323 K) at a rate of 5 K min$^{-1}$ under a carrier gas flow. Oxygen (O$_2$) was selected as the carrier gas to ensure that the dominant vapor species is P$_4$O$_{10}$. The details are described in the following subsection. The temperature and the
carrier gas flow rate were kept constant for a specified period of time, and then the sample was cooled down at a rate of 5 K min$^{-1}$. The carrier gas flow rate was controlled by a variable area flow meter (KOFLOC, RK1250) at the gas inlet and checked by another flow meter (Agilent Technologies, ADM 1000) at the outlet. The precision of the flow rate was ±10%, and that of the sample temperature was ±5 K.

![Diagram](image)

**Figure 4.1** (a) Schematic illustration of the apparatus used for the transpiration method. (b) Photo of the platinum boat and a sample pellet. (c) Photo of the furnace and quartz tube inserted in it, taken from the inlet side.

The vapor pressure of P$_4$O$_{10}$ over the sample, $p_{P4O10}$, was calculated using the following equation:

$$p_{P4O10} = \frac{n_{P4O10}}{n_{P4O10} + n_{O2}} P_{\text{total}},$$

(4.4)

where $n_{P4O10}$ and $n_{O2}$ are the numbers of moles of P$_4$O$_{10}$ vapor and O$_2$ flowed, respectively, and $P_{\text{total}}$
is the total pressure in the quartz tube (1 atm). The value of \( n_{P_4O_{10}} \) was deduced by

\[
n_{P_4O_{10}} = -\frac{\Delta w_{\text{sample}}}{M_{P_4O_{10}}},
\]

where \( \Delta w_{\text{sample}} \) is the weight change of the sample and \( M_{P_4O_{10}} \) is the molecular weight of \( P_4O_{10} \). The value of \( n_{O_2} \) was deduced by

\[
n_{O_2} = \frac{p_{O_2} \cdot v_{O_2} \cdot t}{R \cdot T_{\text{room}}},
\]

where \( p_{O_2} \) and \( v_{O_2} \) are the pressure (1 atm) and flow rate of \( O_2 \) at the inlet, respectively, \( t \) is the holding time, \( R \) is the gas constant and \( T_{\text{room}} \) is room temperature (298 K).

**4.2.2.2 Vapor Species**

To use the above equations, \( P_4O_{10} \) must be the dominant product of evaporation; otherwise, appropriate corrections should be made. Therefore, the partial pressures of known vapor species in the P–O system (\( P, PO, PO_2, P_2, P_2O_3, P_2O_4, P_2O_5, P_3, P_3O_6, P_4, P_4O_6, P_4O_7, P_4O_8, P_4O_9, \) and \( P_4O_{10} \)) were calculated under various conditions using thermodynamic functions given in the literature [12,13]. It should be noted that widely divergent values have been given for the standard enthalpy of formation (\( \Delta_f^\circ H^\circ \)) of \( P_4O_6 \) (g). The NIST-JANAF Thermochemical Tables [12] gives a \( \Delta_f^\circ H^\circ \) (\( P_4O_{10}, \text{ g, 298.15K} \)) value of −2214 kJ mol\(^{-1}\) while Ref. [13] gives a value of −1606 kJ mol\(^{-1}\). As pointed out by several researchers [13-16], the former value is based on a value of \( \Delta_f^\circ H^\circ \) (\( P_4O_6, \text{ l, 298.15K} \)) [17] which is not necessarily accurate. Therefore, we considered the latter value more reliable and used it for the following calculations.

Figure 4.2 (a) shows the variation of the calculated partial pressures of vapor species in the P–O system with \( O_2 \) partial pressure at 1000 K and a constant \( P_4O_{10} \) partial pressure of \( 10^{-6} \) atm. The \( P_4O_{10} \) partial pressure was fixed since it should be thermodynamically determined by LaP\(_3\)O\(_9\)/LaPO\(_4\) or LaP\(_5\)O\(_{14}\)/LaP\(_3\)O\(_9\) equilibria and independent of \( O_2 \) partial pressure. The value of \( 10^{-6} \) atm is in the
range of measured vapor pressures. From the figure, it is clear that $P_4O_{10}$ is the dominant vapor species especially under high $O_2$ partial pressures. Then the temperature dependence of the partial pressures of vapor species was examined at an $O_2$ partial pressure of 1 atm (Figure 4.2 (b)). Although partial pressures of minor vapor species increase with temperature, the partial pressure of $P_4O_{10}$ is still highest in the temperature range of 1000–1400 K. Therefore, $O_2$ was selected as the carrier gas to ensure that $P_4O_{10}$ is the dominant vapor species in the present study.

![Graphs showing vapor species pressures](image)

**Figure 4.2** Variation of the partial pressures of vapor species in the P–O system. (a) log $p_{O_2}$ dependence at temperature $T / K = 1000$ and log $(p_{P4O10} / atm) = -6$. (b) Temperature dependence at log $(p_{P4O10} / atm) = -6$ and log $(p_{O2} / atm) = 0$.

### 4.3 Results

#### 4.3.1 Equilibrium Conditions

Phase identification of LaP$_3$O$_9$ and LaP$_5$O$_{14}$ samples was carried out by X-ray diffraction
analysis before and after vapor pressure measurements. As shown in Figure 4.3, LaP$_3$O$_9$ became a mixture of LaP$_3$O$_9$ and LaPO$_4$ after vapor pressure measurements. Likewise, LaP$_5$O$_{14}$ became a mixture of LaP$_5$O$_{14}$ and LaP$_3$O$_9$. These results are consistent with the reported decomposition reactions (Eqs. (4.1) and (4.2)).

![X-ray diffraction patterns](image)

Figure 4.3 Typical X-ray diffraction patterns of (a) LaP$_3$O$_9$ and (b) LaP$_5$O$_{14}$ pellets. The upper and lower patterns in each figure were obtained before and after vapor pressure measurements, respectively. During the vapor pressure measurements, the LaP$_3$O$_9$ pellet was held at 1273 K for 10 h and the LaP$_5$O$_{14}$ pellet was held at 1073 K for 3 h. Peaks marked with ○, ▼ and × are attributed to LaPO$_4$ (PDF# 1-83-651), LaP$_3$O$_9$ (PDF# 1-84-1635) and LaP$_5$O$_{14}$ (PDF# 1-70-5341), respectively.
To establish the conditions where the carrier gas (O₂) is saturated with P₄O₁₀ vapor, variation of the apparent P₄O₁₀ vapor pressure with O₂ flow rate was examined (Figure 4.4). The plateau region, where the vapor pressure is independent of the flow rate, was found to be in the flow rate range of 20–50 sccm. In the plateau region, equilibrium vapor pressures can be obtained without any correction for diffusional transport or kinetic effects [9]. In addition, the number of sample pellets inserted simultaneously in the apparatus did not have any marked effect on the above behavior. This indicates that enough evaporation rates were achieved with only one pellet. Therefore, all further vapor pressure measurements were carried out at an O₂ flow rate of 25 sccm with one sample pellet inserted in the apparatus.

![Graph showing apparent vapor pressures of P₄O₁₀ over LaP₃O₉](image)

**Figure 4.4** Apparent vapor pressures of P₄O₁₀ over LaP₃O₉, \( p_{\text{P}_4\text{O}_{10}} \) (apparent), as a function of gas flow rate \( v_{\text{O}_2} \) at 1273 K: ○, data obtained with a pellet; ♦, data obtained with four pellets.

### 4.3.2 Temperature Dependence of Vapor Pressures

The equilibrium vapor pressures of P₄O₁₀ over La₃P₉O₁₄ and LaP₅O₁₄ were examined in the temperature range of 1123–1323K and 923–1173K, respectively. The vapor pressures obtained in each measurement are listed in Table 4.1. The logarithmic plot of the vapor pressures against
reciprocal temperature is shown in Figure 4.5. The least squares lines are given by

\[
\log \left( \frac{p_{P4O_{10}}}{\text{atm}} \right) = (5.12 \pm 1.00) - (1.22 \pm 0.12) \times 10^4 \times \left( \frac{T}{K} \right)^{-1}
\]  
(over LaP$_3$O$_9$);  \hspace{2cm} (4.7)

\[
\log \left( \frac{p_{P4O_{10}}}{\text{atm}} \right) = (4.65 \pm 1.80) - (1.00 \pm 0.19) \times 10^4 \times \left( \frac{T}{K} \right)^{-1}
\]  
(over LaP$_5$O$_{14}$),  \hspace{2cm} (4.8)

where each value is shown with two standard error range which provides 95% confidence. Our experimental data on LaP$_3$O$_9$ are in reasonable agreement with those reported by Ashuiko et al. [5].

![Figure 4.5 Temperature dependence of the vapor pressure of $P_4O_{10}$, $p_{P4O_{10}}$: ○, $p_{P4O_{10}}$ over LaP$_3$O$_9$; ●, $p_{P4O_{10}}$ over LaP$_5$O$_{14}$; ——, least-squares lines to the experimental data; - - - - - , $p_{P4O_{10}}$ over LaP$_3$O$_9$ given by Ashuiko et al. [5].](image-url)
Table 4.1 Weight losses of samples $-\Delta w_{\text{sample}}$ and the vapor pressures of $P_4O_{10}$ at gas flow rate $v_{O_2}$ and temperature $T$ with holding time $t$.

<table>
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<th>Sample</th>
<th>$v_{O_2}$/ sccm</th>
<th>$T$/ K</th>
<th>$t$/ h</th>
<th>$-\Delta w_{\text{sample}}$/ mg</th>
<th>$p_{P_4O_{10}}$ a/ atm</th>
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</tbody>
</table>

a Derived from $v_{O_2}$, $t$ and $-\Delta w_{\text{sample}}$.

4.4 Discussion

4.4.1 Thermodynamic Properties of LaP$_3$O$_9$ and LaP$_5$O$_{14}$

The thermodynamic properties of the decomposition reactions (Eqs. (4.1) and (4.2)) were determined from the results of vapor pressure measurements. The slopes and intercepts of the least-squares lines in Figure 4.5 (Eqs. (4.4) and (4.5)) yield the second-law values of standard
enthalpies and entropies of the decomposition reactions, shown in Table 4.2. The corresponding values based on the literature [5,18] are also listed in the table. The standard enthalpy of decomposition of LaP$_5$O$_{14}$ (Eq. (4.2)) was estimated from the specifications of a commercial phosphorus diffusion source made of LaP$_5$O$_{14}$ (Technegas, Inc., PhosPlus TP-250). In the brochure [18], the thicknesses of deposited glassy film using the diffusion source at various temperatures (1098–1223 K) are given. Assuming that the thickness is proportional to the vapor pressure of P$_4$O$_{10}$, one can obtain the standard enthalpy of decomposition by the second law treatment. Our present values and those from the literature are in reasonable agreement.

<table>
<thead>
<tr>
<th>Reaction</th>
<th>Source</th>
<th>Method</th>
<th>$\Delta H^o$ / kJ mol$^{-1}$</th>
<th>$\Delta S^o$ / J mol$^{-1}$ K$^{-1}$</th>
<th>$T$ / K</th>
</tr>
</thead>
<tbody>
<tr>
<td>(4.1)</td>
<td>This work</td>
<td>Second-law</td>
<td>117 ± 12</td>
<td>49 ± 10</td>
<td>1123–1323</td>
</tr>
<tr>
<td></td>
<td>Ref. [5]</td>
<td>Second-law</td>
<td>130</td>
<td>54</td>
<td>1219–1404</td>
</tr>
<tr>
<td>(4.2)</td>
<td>This work</td>
<td>Second-law</td>
<td>96 ± 18</td>
<td>45 ± 17</td>
<td>923–1173</td>
</tr>
<tr>
<td></td>
<td>Ref. [18]</td>
<td>Second-law</td>
<td>113 ± 18</td>
<td>Unavailable</td>
<td>1098–1223</td>
</tr>
</tbody>
</table>

### 4.4.2 Thermal and Chemical Stabilities of LaP$_3$O$_9$ and LaP$_5$O$_{14}$

As mentioned in the introduction, LaP$_3$O$_9$ and LaP$_5$O$_{14}$ have been considered as candidates for solid proton-conducting electrolytes in fuel cells. In general, lifetimes greater than 40000 h are required for stationary fuel cell systems [19]. To meet the requirement, electrolytes in fuel cells should have sufficient thermal and chemical stability. As for LaP$_3$O$_9$ and LaP$_5$O$_{14}$ electrolytes, their operation conditions should be limited to suppress the decomposition.

#### 4.4.2.1 Operating Temperatures for the Electrolytes in Oxygen Atmosphere

The lifetime target of stationary fuel cells is the time until 10% power degradation [19]. Since the relationship between power degradation and the decomposition ratios of the electrolytes is unknown, a decomposition ratio of 0.01 (1%) was taken as the threshold in the following discussion.
to obtain safer values. In the present study, a sign of decomposition (weight fraction loss of $5 \times 10^{-3}$; decomposition ratio of $\sim 0.01$) was detected in 50 h when the vapor pressure of $P_4O_{10}$ was about $10^{-6}$ atm (for example, see the first row in Table 4.1). Although the dimensions of electrolytes and gas flow rates in actual fuel cell systems may be different from those in this study, the above observation could be basically applicable. Therefore, for long-term ($> 40000$ h) stability, the vapor pressure of $P_4O_{10}$ should be lowered to $\sim 10^{-9}$ atm. The second-law extrapolation of the vapor pressures of $P_4O_{10}$ to $10^{-9}$ atm using Eqs. (4.4) and (4.5) gives the temperatures 864 K (591 °C) and 733 K (460 °C), respectively. Thus, the upper limits of the operating temperatures for $LaP_3O_9$ and $LaP_5O_{14}$ electrolytes are estimated to be about 600 °C and 450 °C. At these temperatures, the proton conductivity of 5% Sr-doped $LaP_3O_9$ is $\sim 2.5 \times 10^{-4}$ S cm$^{-1}$ [20] whereas that of 5% Sr-doped $LaP_5O_{14}$ is $\sim 1.3 \times 10^{-5}$ S cm$^{-1}$ [21]. Therefore, it can be said that $LaP_3O_9$ based electrolytes are more promising than $LaP_5O_{14}$ based ones.

### 4.4.2.2 Stabilities in the Presence of Water and Hydrogen

Proton conducting electrolytes in practical fuel cells are used in the presence of water vapor and hydrogen. Therefore, the stabilities of the $LaP_3O_9$ and $LaP_5O_{14}$ electrolytes in such conditions should also be assessed. Various vapor species can be formed by their reactions with water vapor and hydrogen:

\[
yLaP_3O_9 (s) + (x+5y−2z)H_2 (g) + (−5y+2z)H_2O (g) \rightarrow yLaPO_4 (s) + 2H_xP_yO_z (g);
\]

(4.9)

\[
yLaP_5O_{14} (s) + (x+5y−2z)H_2 (g) + (−5y+2z)H_2O (g) \rightarrow yLaP_3O_9 (s) + 2H_xP_yO_z (g).
\]

(4.10)

$H_xP_yO_z (g)$ is a general expression for the vapor species. Here, $PH$, $PH_2$, $PH_3$, $HPO$, $HPO_2$, and $HPO_3$ were considered in addition to the P–O vapor species considered in 4.2.2.2. The equilibrium partial pressures of each vapor species were calculated by these equations using our present results and
thermodynamic data in the literature. The standard enthalpies of formation and Gibbs energy functions of PH, PH$_2$, and PH$_3$ were taken from Ref. [12]. Those of HPO were taken from Refs. [13,22]. For HPO$_2$ and HPO$_3$, the standard enthalpies of formation were taken from Ref. [22] and the Gibbs energy functions were calculated in the "rigid rotator-harmonic oscillator" approximation using the molecular constants given in Ref. [22]. The calculated values of the thermodynamic functions of HPO$_2$ and HPO$_3$ are tabulated in Table 4.3. Figure 4.6 shows the calculated partial pressures over LaP$_3$O$_9$ at 864 K both in oxygen and hydrogen atmospheres as a function of water vapor pressure, $p_{H_2O}$. In oxygen atmosphere (Figure 4.6 (a)), the vapor pressure of HPO$_3$ increases with increasing $p_{H_2O}$ and exceeds the vapor pressure of P$_4$O$_{10}$. This indicates that decomposition of LaP$_3$O$_9$ will be accelerated in the presence of water vapor with evaporation of HPO$_3$. The effect of hydrogen is more drastic as can be seen in Figure 4.6 (b). In hydrogen atmosphere, the partial pressures of various vapor species in lower oxidation states such as HPO$_2$ and P$_4$ become several orders of magnitude greater than that of P$_4$O$_{10}$. Therefore, the stability of LaP$_3$O$_9$ in hydrogen atmosphere can be much poorer than that in oxygen atmosphere. As shown in Figure 4.7, the vapor pressures over LaP$_5$O$_{14}$ were also found to have the similar dependencies on surrounding atmosphere. These results suggest that the operating atmospheres for the LaP$_3$O$_9$ and LaP$_5$O$_{14}$ should be carefully chosen to prevent their decomposition. For example, in hydrogen atmospheres, decreasing $p_{H_2}$ and $p_{H_2O}$ simultaneously will lower the overall vapor pressure of phosphorus-containing vapor species as shown in Figure 4.8. Finally, as these effects were only derived from the thermodynamic calculations, future work will have to confirm these phenomena by experiment and examine kinetic aspects.
Table 4.3 Thermodynamic functions of HPO$_2$ and HPO$_3$ calculated using their molecular constants [22]. All values are in J mol$^{-1}$ K$^{-1}$ if not otherwise specified.

<table>
<thead>
<tr>
<th>T / K</th>
<th>HPO$_2$</th>
<th></th>
<th></th>
<th>HPO$_3$</th>
<th></th>
<th></th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>C$_p$</td>
<td>S</td>
<td>gef</td>
<td>C$_p$</td>
<td>S</td>
<td>gef</td>
</tr>
<tr>
<td>298.15</td>
<td>50.2</td>
<td>264.9</td>
<td>−264.9</td>
<td>62.7</td>
<td>277.8</td>
<td>−277.8</td>
</tr>
<tr>
<td>300</td>
<td>50.4</td>
<td>265.2</td>
<td>−264.9</td>
<td>62.9</td>
<td>278.2</td>
<td>−277.8</td>
</tr>
<tr>
<td>400</td>
<td>55.9</td>
<td>280.5</td>
<td>−266.9</td>
<td>71.4</td>
<td>297.5</td>
<td>−280.3</td>
</tr>
<tr>
<td>500</td>
<td>60.1</td>
<td>293.5</td>
<td>−271.0</td>
<td>77.7</td>
<td>314.1</td>
<td>−285.5</td>
</tr>
<tr>
<td>600</td>
<td>63.4</td>
<td>304.7</td>
<td>−275.7</td>
<td>82.5</td>
<td>328.7</td>
<td>−291.5</td>
</tr>
<tr>
<td>700</td>
<td>65.9</td>
<td>314.7</td>
<td>−280.6</td>
<td>86.1</td>
<td>341.7</td>
<td>−297.8</td>
</tr>
<tr>
<td>800</td>
<td>67.9</td>
<td>323.6</td>
<td>−285.4</td>
<td>89.0</td>
<td>353.4</td>
<td>−304.0</td>
</tr>
<tr>
<td>900</td>
<td>69.6</td>
<td>331.7</td>
<td>−290.1</td>
<td>91.3</td>
<td>364.1</td>
<td>−310.1</td>
</tr>
<tr>
<td>1000</td>
<td>71.0</td>
<td>339.1</td>
<td>−294.6</td>
<td>93.2</td>
<td>373.8</td>
<td>−316.0</td>
</tr>
<tr>
<td>1100</td>
<td>72.3</td>
<td>345.9</td>
<td>−299.0</td>
<td>94.9</td>
<td>382.7</td>
<td>−321.6</td>
</tr>
<tr>
<td>1200</td>
<td>73.3</td>
<td>352.3</td>
<td>−303.2</td>
<td>96.2</td>
<td>391.1</td>
<td>−327.1</td>
</tr>
<tr>
<td>1300</td>
<td>74.3</td>
<td>358.2</td>
<td>−307.2</td>
<td>97.4</td>
<td>398.8</td>
<td>−332.3</td>
</tr>
<tr>
<td>1400</td>
<td>75.1</td>
<td>363.7</td>
<td>−311.0</td>
<td>98.5</td>
<td>406.1</td>
<td>−337.3</td>
</tr>
</tbody>
</table>
Figure 4.6 Partial pressures of vapor species over LaP$_3$O$_9$ at 864 K as a function of water vapor pressure, $p_{H_2O}$. (a) Under 1 atm O$_2$; (b) under 1 atm H$_2$. 
Figure 4.7 Partial pressures of vapor species over LaP$_2$O$_{14}$ at 733 K as a function of water vapor pressure, $p_{H2O}$. (a) Under 1 atm O$_2$; (b) under 1 atm H$_2$. 

\begin{align*}
T &= 733 \text{ K} \\
p_{O2} &= 1 \text{ atm} \\
p_{H2O} &= 10^{-9} \text{ atm}
\end{align*}
4.4.3 Re-examination of the Thermodynamic Data of P₄O₁₀

Finally, validity of the thermodynamic data of P₄O₁₀ in the literature was re-examined. In this study, a value of ΔfH° (P₄O₁₀, g, 298.15 K) = −1606 kJ mol⁻¹ [13] was adopted while NIST-JANAF Thermochemical Tables [12] provide ΔfH° (P₄O₁₀, g, 298.15 K) = −2214 kJ mol⁻¹. If the value of −2214 kJ mol⁻¹ was used instead, P₄O₁₀ partial pressure over LaP₃O₉ would be as high as 10¹⁶ atm at 864 K and pH₂O/pH₂ = 1. However, such an extremely fast decomposition of LaP₃O₉ in hydrogen atmosphere has never been reported [20,23]. Therefore, the value in Ref. [13] seems more reliable, which is in agreement with several researchers [14-16]. In light of this conclusion, some work on the vapor phase equilibria in the P-O system [24] has to be reviewed again.
4.5 Conclusions

The vapor pressures of $P_4O_{10}$ produced by decomposition of LaP$_3$O$_9$ and LaP$_5$O$_{14}$ (LaP$_3$O$_9$ → LaPO$_4$ + 1/2 $P_4O_{10}$; LaP$_5$O$_{14}$ → LaP$_3$O$_9$ + 1/2 $P_4O_{10}$) were measured by the transpiration method. The equilibrium vapor pressures of $P_4O_{10}$ over LaP$_3$O$_9$ and LaP$_5$O$_{14}$ can be expressed as

$$\log \left( \frac{p_{P_4O_{10}}} {\text{atm}} \right) = (5.12 \pm 1.00) - (1.22 \pm 0.12) \times 10^4 \times \left( \frac{T} {K} \right)^{-1}$$

(over LaP$_3$O$_9$, 1123 ≤ $T$ / K ≤ 1323);

$$\log \left( \frac{p_{P_4O_{10}}} {\text{atm}} \right) = (4.65 \pm 1.80) - (1.00 \pm 0.19) \times 10^4 \times \left( \frac{T} {K} \right)^{-1}$$

(over LaP$_5$O$_{14}$, 923 ≤ $T$ / K ≤ 1173).

The standard enthalpies of the evaporation reactions are given as 117±12 kJ mol$^{-1}$ and 96±18 kJ mol$^{-1}$ in the corresponding temperature ranges. The data were evaluated with respect to the proposed use of LaP$_3$O$_9$ and LaP$_5$O$_{14}$ as solid electrolytes in fuel cells. The maximum operating temperatures for the LaP$_3$O$_9$ and LaP$_5$O$_{14}$ electrolytes were suggested to be about 600 °C and 450 °C in oxygen atmosphere. Taking into account their proton conductivities at these temperatures, it can be said that LaP$_3$O$_9$ based electrolytes are more promising than LaP$_5$O$_{14}$ based ones. In the presence of water vapor or hydrogen, however, decomposition of LaP$_3$O$_9$ and LaP$_5$O$_{14}$ would be accelerated considerably.

References


Chapter 5
Synthesis of Highly Sr-doped LaPO$_4$ and LaP$_3$O$_9$

5.1 Introduction

As mentioned in Chapter 1, the low solubilities of Sr in LaPO$_4$ and LaP$_3$O$_9$ have been reported to be relatively low. According to Amezawa et al., the solubility of Sr, represented by the molar ratio Sr/(La+Sr), is ~2 mol% in LaPO$_4$ [1] and below 5 mol% in LaP$_3$O$_9$ [2]. This is considered as an obstacle to enhancing the conductivities of lanthanum phosphates. Nevertheless, previous studies suggest that the solubility of Sr would have dependence on temperature, and higher Sr doping levels would be expected at lower temperatures. Besides temperature, water vapor pressure may also affect the solubility of Sr. When Sr is doped into a lanthanum phosphate such as LaPO$_4$, a certain amount of water would also be introduced into LaPO$_4$ as represented by the following expression:

$$(1-x)\text{LaPO}_4 + \frac{(x/2)}{}\text{Sr}_2\text{P}_2\text{O}_7 + \frac{(y/2)}{}\text{H}_2\text{O} \rightarrow \text{La}_{1-x}\text{Sr}_x\text{H}_y\text{PO}_4-(\frac{x}{2})+(\frac{y}{2}) \tag{5.1}$$

Therefore, Sr-doped LaPO$_4$ ($\text{La}_{1-x}\text{Sr}_x\text{H}_y\text{PO}_4-(\frac{x}{2})+(\frac{y}{2})$) would be stabilized with higher water vapor pressure. However, the relationship between the solubility of Sr and these parameters has not been clarified yet.

In Chapter 2, the precipitation conditions for undoped LaPO$_4$ and LaP$_3$O$_9$ in homogeneous phosphoric acid solutions were established. In this chapter, further attempts were made to synthesize Sr-doped LaPO$_4$ and LaP$_3$O$_9$ in phosphoric acid solutions. The condition dependence of the Sr doping level was analyzed to reveal the optimum condition for obtaining highly Sr-doped phosphates. The incorporation of Sr and H$_2$O in the LaPO$_4$ lattice was verified by unit cell volume change and
evaporation of H$_2$O. The incorporation of Sr in the LaP$_3$O$_9$ lattice was verified similarly. Finally, the stability of Sr-doped LaPO$_4$ at high temperatures was investigated.

5.2 Experimental

5.2.1 Synthesis

Sr-doped LaPO$_4$ and LaP$_3$O$_9$ were synthesized by precipitation in homogeneous phosphoric acid solutions containing La$_2$O$_3$ and SrCO$_3$. La$_2$O$_3$ (99.99%, Nacalai Tesque), SrCO$_3$ (99.99%, Wako Pure Chemical Industries), and H$_3$PO$_4$ (85%, Nacalai Tesque) were mixed in a perfluoroalkoxy (PFA) beaker and held at 190 °C in a furnace for several days to obtain a transparent solution. Then temperature and water vapor pressure were changed and held for a predetermined period of time to precipitate lanthanum phosphates. The temperature was set between 120 °C and 250 °C, and the water vapor pressure was set to approximately 0.01 atm (atmospheric conditions) or 1 atm (humidified conditions). Synthesis parameters are summarized in Table 5.1 and Table 5.2. The precipitation conditions are also plotted on the phase stability diagram of the La$_2$O$_3$–P$_2$O$_5$–H$_2$O system constructed in Chapter 2, Figure 5.1.
Table 5.1 Initial compositions of solutions for the synthesis of LaPO\textsubscript{4} and LaP\textsubscript{3}O\textsubscript{9}

<table>
<thead>
<tr>
<th>Abbr.</th>
<th>Starting materials</th>
<th>Initial composition</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td>Molar ratio (\text{La} : \text{Sr} : \text{P})</td>
</tr>
<tr>
<td>Sr4</td>
<td>(\text{Sr}_4\text{O}_3), (\text{SrCO}_3), (\text{H}_3\text{PO}_4)</td>
<td>0.96 : 0.04 : 15</td>
</tr>
<tr>
<td>Sr8</td>
<td></td>
<td>0.92 : 0.08 : 15</td>
</tr>
<tr>
<td>Sr12</td>
<td></td>
<td>0.88 : 0.12 : 15</td>
</tr>
<tr>
<td>Sr16</td>
<td></td>
<td>0.84 : 0.16 : 15</td>
</tr>
<tr>
<td>Sr20</td>
<td>(\text{La}_2\text{O}_3), (\text{SrCO}_3), (\text{H}_3\text{PO}_4)</td>
<td>0.80 : 0.20 : 15</td>
</tr>
<tr>
<td>Sr30</td>
<td></td>
<td>0.70 : 0.30 : 15</td>
</tr>
<tr>
<td>Sr40</td>
<td></td>
<td>0.60 : 0.40 : 15</td>
</tr>
<tr>
<td>Sr50</td>
<td></td>
<td>0.50 : 0.50 : 15</td>
</tr>
<tr>
<td>Sr60</td>
<td></td>
<td>0.40 : 0.60 : 15</td>
</tr>
<tr>
<td>Sr80</td>
<td></td>
<td>0.20 : 0.80 : 15</td>
</tr>
</tbody>
</table>

Table 5.2 Dissolution and precipitation conditions for the synthesis of LaPO\textsubscript{4} and LaP\textsubscript{3}O\textsubscript{9}.

<table>
<thead>
<tr>
<th>Abbr.</th>
<th>Dissolution</th>
<th>Precipitation</th>
<th>Product</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>Temp. / °C</td>
<td>(p_{\text{H}_2\text{O}}) / atm</td>
<td>Temp. / °C</td>
</tr>
<tr>
<td>A230</td>
<td>230</td>
<td>~0.01</td>
<td>230</td>
</tr>
<tr>
<td>A120</td>
<td>190</td>
<td>~0.01</td>
<td>120</td>
</tr>
<tr>
<td>H250</td>
<td>250</td>
<td>~1</td>
<td>250</td>
</tr>
<tr>
<td>H120</td>
<td>120</td>
<td>~1</td>
<td>120</td>
</tr>
</tbody>
</table>
Chapter 5  Synthesis of Highly Sr-doped LaPO₄ and LaP₃O₉

Figure 5.1 Phase stability diagram of the La₂O₃–P₂O₅–H₂O system with a constant La/P ratio (La : P = 1 : 15). The precipitation conditions listed in Table 5.2 are plotted in the figure.

Figure 5.2 Schematic illustration of the apparatus used for humidification during the precipitation of LaPO₄.

For humidification, the apparatus shown in Figure 5.2 was used. The whole apparatus was placed in a furnace kept at a constant temperature. A beaker was put on a stainless steel floor plate and covered with a stainless steel box with a capacity of 2.64 L. A 4-m-long spiral tube was attached to a side wall of the box. Liquid water was fed to the box through the spiral tube so that it evaporated
before reaching the box. The liquid water feed rate was 0.1 mL/min, which was equivalent to 0.18 L-H$_2$O(g)/min at 120 °C and 0.24 L-H$_2$O(g)/min at 250 °C. The apparatus was not sealed tightly, so that the inner gas was able to flow out through the narrow gap between the box and the floor plate when the inner pressure exceeded 1 atm. Therefore, the water vapor pressure inside the box was expected to be 1 atm.

After the predetermined precipitation period, the beaker was cooled down to room temperature. In general, there were precipitates on the bottom of the beaker and the transparent solution above them. The residual solution was collected for the chemical analysis. The precipitates were washed well with hot water (~90 °C).

5.2.2 Characterization

Phase identification was carried out via X-ray powder diffraction (XRD) analysis on PANalytical X’Pert-Pro MPD using Cu-Kα radiation at room temperature. The unit cell volume of samples was evaluated by the Rietveld method using X’Pert HighScore Plus (Version 2.2c) software. The morphology of precipitates was investigated using KEYENCE VE-7800 scanning electron microscope (SEM). Chemical analysis of solutions and precipitates was performed by inductively coupled plasma atomic emission spectrometry (ICP-AES) on Seiko Instruments SPS4000.

The amount of H$_2$O evaporated from LaPO$_4$ was determined by the Karl Fischer titration method. The samples were dried in advance at 600 °C for 50 h in the air ($p_{H_2O} \sim 0.01$ atm) to eliminate residual water and phosphoric acid on the surface of samples. The protons responsible for the proton conduction were still expected to remain in the samples, judging from the reported $p_{H_2O}$ dependence of the conductivity at 600 °C [3]. Then each sample was held at 800 °C under dry nitrogen flow (200 mL/min) to be dehydrated. The nitrogen which contains the released H$_2$O was transferred to the Karl Fischer moisture titrator to determine the amount of H$_2$O evaporated from the sample. For this, Kyoto Electronics Manufacturing MKC-510N (Karl-Fischer moisture titrator) and
ADP-512S (evaporator) were utilized.

### 5.2.3 Calculation of Precipitation Ratios

For better understanding of the precipitation behavior, the concentration values of La and Sr in the residual solutions were converted into “precipitation ratios”. Here, the precipitation ratio of element M, $R_M$, is defined by the following formula:

$$R_M = \frac{n_{M, \text{prec}}}{n_{M, \text{init}}}, \tag{5.2}$$

where $n_{M, \text{prec}}$ and $n_{M, \text{init}}$ are the quantities of M in the precipitate and initial solution, respectively.

For example, $R_{\text{La}}$ will be 1 (=100 %) when all La in the initial solution precipitates as LaPO$_4$ or LaP$_3$O$_9$. For calculating $R_{\text{La}}$ and $R_{\text{Sr}}$ values, the following method was used. First, variables $C_{i, \text{La}}$, $C_{i, \text{Sr}}$, $C_{f, \text{La}}$, and $C_{f, \text{Sr}}$ are defined to represent the molar ratios of La, Sr, and P in the initial and residual solutions as shown in Eqs. (5.9) and (5.10):

$$\text{La} : \text{Sr} : \text{P} = C_{i, \text{La}} : C_{i, \text{Sr}} : 1 \text{ (In the initial solution)}, \tag{5.3}$$

$$\text{La} : \text{Sr} : \text{P} = C_{f, \text{La}} : C_{f, \text{Sr}} : 1 \text{ (In the residual solution after precipitation)}. \tag{5.4}$$

The values of $C_{i, \text{La}}$ and $C_{i, \text{Sr}}$ are readily obtained from the mixing ratio of the starting materials and those of $C_{f, \text{La}}$ and $C_{f, \text{Sr}}$ can be derived from the ICP analysis on the residual solution. Assuming that the molar ratio of P/(La+Sr) in the precipitates is $k$ ($k=1$ for Sr-doped LaPO$_4$ and $k=3$ for Sr-doped LaP$_3$O$_9$), $C_{i, \text{La}}$, $C_{i, \text{Sr}}$, $C_{f, \text{La}}$, and $C_{f, \text{Sr}}$ can be correlated by Eq. (5.11) by considering the molar ratio in the residual solution:

$$C_{f, \text{La}} : C_{f, \text{Sr}} : 1 = C_{i, \text{La}} (1 - R_{\text{La}}) : C_{i, \text{Sr}} (1 - R_{\text{Sr}}) : 1 - k \left( C_{i, \text{La}} R_{\text{La}} + C_{i, \text{Sr}} R_{\text{Sr}} \right). \tag{5.5}$$

By solving this equation, the precipitation ratios ($R_{\text{La}}$ and $R_{\text{Sr}}$) are given as follows:

$$R_{\text{La}} = - \frac{(C_{f, \text{La}} C_{i, \text{Sr}} - C_{f, \text{Sr}} C_{i, \text{La}}) k + C_{i, \text{La}} - C_{f, \text{La}}}{(C_{f, \text{Sr}} + C_{f, \text{La}}) C_{i, \text{La}} k - C_{i, \text{La}}}, \tag{5.6}$$
\[ R_{Sr} = \frac{(C_{f,La} C_{i, Sr} - C_{f, Sr} C_{i, La}) k - C_{i, Sr} + C_{f, Sr}}{(C_{f, Sr} + C_{f, La}) C_{i, Sr} k - C_{i, Sr}} \]  

(5.7)

5.3 Results and Discussion

5.3.1 Sr-doped LaPO\(_4\)

5.3.1.1 Precipitation Behavior in Low Sr Concentration Solutions

First, attempts to synthesize Sr-doped LaPO\(_4\) were carried out using solutions with the compositions Sr4 and Sr8. Experiments were performed in the precipitation region of LaPO\(_4\) (see Figure 5.1) under both atmospheric (A120) and humidified (H250, H120) conditions. The precipitation period was varied to observe changes in the composition of precipitates and solutions with time. In each condition, single phase LaPO\(_4\) was precipitated, as confirmed by X-ray diffraction.
Figure 5.3 Change in composition of solutions with time under the conditions (a) A120, (b) H250, and (c) H120. Inset in (c) is a logarithmic plot.
Figure 5.3 shows the change in the composition of the solutions with time measured by ICP. The concentrations of La and Sr decreased with time. In particular, under H120 condition, the concentrations decreased rapidly by 2-3 orders. This indicates that the solubility of LaPO$_4$ in phosphoric acid solutions is especially low under H120 condition, or the driving force for the precipitation is strong under that condition. This result corresponds well to the fact that H120 condition is quite deep in the precipitation region of LaPO$_4$ in Figure 5.1. Under A120 condition, the changes in the concentrations of La and Sr fluctuated. This is likely due to the fluctuation of water vapor pressure in the atmosphere.

For better understanding of the precipitation behavior, these concentration values were converted into precipitation ratios defined by Eq. (5.8). Figure 5.4 shows the changes in $R_{La}$ and $R_{Sr}$. $R_{Sr}$ was greater than $R_{La}$ in the early stages of the precipitation regardless of the initial composition of the solution. This suggests the formation of Sr-richer LaPO$_4$ than the initial solutions.

![Figure 5.4 Change in the precipitation ratios of La and Sr observed along with the progression of the precipitation of LaPO$_4$.](image-url)
Figure 5.5 shows the Sr doping levels in the LaPO$_4$ precipitates analyzed by ICP. Here, the Sr doping level, $x_{\text{Sr, prec}}$, represents the following molar fraction:

$$x_{\text{Sr, prec}} \equiv \frac{n_{\text{Sr, prec}}}{n_{\text{La, prec}} + n_{\text{Sr, prec}}} \times 100 \, \%.$$  

(5.8)

The Sr doping level showed a strong dependence on $R_{\text{La}}$. In good agreement with the precipitation ratio changes described above (Figure 5.4), the Sr doping levels in the precipitates were higher in the low $R_{\text{La}}$ region, reaching 20 mol% at maximum. This value is one order of magnitude higher than the reported solubilities of Sr in LaPO$_4$ [1,4]. The precipitation temperature and $p_{\text{H}_2\text{O}}$ did not seem to have a significant effect on this trend, although they affect the attainable precipitation ratio limited by thermodynamics, and the Sr doping level accordingly.

### 5.3.1.2 Precipitation Behavior in High Sr Concentration Solutions

Because of the rapid decrease in the Sr doping level with progression of the precipitation, it would be difficult to obtain uniformly Sr-doped LaPO$_4$ in the above solutions. Since the Sr doping
levels shown in Figure 5.5 are average values, the actual Sr doping levels in the final precipitates ($R_{La} = \sim 100\%$) are thought to be close to 0 %. The decrease in Sr doping level seems inevitable, as Sr in the solutions is nearly “exhausted” in early stages of the precipitation. A way to avoid the exhaustion of Sr in the solutions may be to increase the Sr concentration in the solutions. Therefore, the precipitation behavior of LaPO$_4$ in solutions Sr12, Sr16, Sr20, Sr40, Sr60, and Sr80 was investigated at H250 condition.

Single phase LaPO$_4$ was precipitated in all the solutions except Sr80, as confirmed by X-ray diffraction. In the Sr80 solution, SrH$_3$P$_2$O$_7$ precipitated instead of LaPO$_4$. Figure 5.6 represents the Sr doping levels in precipitates. As shown in Figure 5.6 (a), the Sr doping level in LaPO$_4$ increased with increasing the initial Sr concentration in the solutions. The dependence was not linear, however, and the increase in the doping level became moderate with higher initial Sr concentrations (Sr20, Sr40 and Sr60). The Sr doping level tended to decrease with increasing the precipitation ratio of La as observed previously. Nevertheless, the variation range relative to the doping level became smaller with higher initial Sr concentrations. As shown in Figure 5.6 (b), the Sr doping level relative to the initial Sr concentration in the solutions, ($x_{Sr, prec} / x_{Sr, init}$), decreased with increasing the initial Sr concentration in the solutions. With Sr20, Sr40, and Sr60 compositions, ($x_{Sr, prec} / x_{Sr, init}$) was not larger than 1. Thus, the exhaustion of Sr cannot happen with those compositions. These results indicate that uniformly and highly Sr-doped LaPO$_4$ with doping levels of \sim 20\% can be obtained in the solutions with Sr concentrations higher than that of Sr20.
5.3.1.3 Effectiveness of Sr Doping

To examine whether Sr was doped correctly and H$_2$O was incorporated in the bulk as desired, the following two methods were used: (1) The amount of H$_2$O evaporated from samples at 800 °C in dry N$_2$ was measured by the Karl Fischer titration method. (2) The unit cell volume of samples was evaluated before and after the H$_2$O evaporation measurement using powder X-ray diffraction patterns. Prior to these measurements, samples were dried at 600 °C in the air for 50 h.
Figure 5.7 (a) H$_2$O evaporation from Sr-doped LaPO$_4$ under N$_2$ at 800 °C as measured by the Karl Fischer titration (KFT) method. The samples were held in the air at 600 °C for 50 h beforehand. (b) Unit cell volume of Sr-doped LaPO$_4$ as determined by Rietveld analysis using powder X-ray diffraction patterns. The open circles (○) represent the cell volume of samples before the titration (just after held in the air at 600 °C for 50 h). The filled squares (■) represent the cell volume of samples after the titration.

Figure 5.7 (a) represents the H$_2$O evaporation from LaPO$_4$ (expressed in percentage of the initial weight of each sample) with different Sr doping levels (0, 4, 8 and 18 mol%). Although the plots are not linear, H$_2$O evaporation monotonically increased with increasing the Sr doping level. This should indicate that H$_2$O was incorporated in LaPO$_4$ with Sr on the synthesis. Figure 5.7 (b) shows the unit cell volume of samples before and after the H$_2$O evaporation measurement. In both cases, LaPO$_4$ with higher Sr doping level had larger unit cell volume. This would indicate that Sr substituted for La in the monazite structure, since the ionic radius of Sr$^{2+}$ in ninefold coordination (1.31 Å) is larger than that of La$^{3+}$ (1.21 Å) [5]. After the H$_2$O evaporation measurement, the unit cell volume of Sr-doped LaPO$_4$ significantly decreased while that of undoped LaPO$_4$ remained unchanged. This would reflect the evaporation of H$_2$O from the bulk Sr-doped LaPO$_4$. In Figure 5.7 (a) the small H$_2$O evaporation was observed even from undoped LaPO$_4$, but it could be due to some residual water on
the surface or grain boundaries of the sample.

These results suggest that Sr successfully substituted for La and H₂O was certainly incorporated in the bulk as desired.

### 5.3.1.4 Thermal Stability of Highly Sr-doped LaPO₄

To evaluate the stability of highly Sr-doped LaPO₄, 13.4 % Sr-doped LaPO₄ and 20.1 % Sr-doped LaPO₄ synthesized in phosphoric acid solutions were heat treated at higher temperatures. The samples were first held at 600 °C for 80 h in the atmosphere, then the 13.4% doped sample was additionally held at 1200 °C for 5 h. Figure 5.8 shows the X-ray diffraction patterns obtained after each heat treatment. While no secondary phase was detected after the first heat treatment at 600 °C, the presence of secondary phases, Sr₂P₂O₇ and Sr₃La(PO₄)₃, was clearly detected after the second heat treatment at 1200 °C. Amezawa et al. reported Sr₂P₂O₇ as the secondary phase at 1200 °C and Sr₃La(PO₄)₃ at above 1350 °C [3]. Tyholdt et al. also reported both as the secondary phases at 1450 °C [4]. The fact that both Sr₂P₂O₇ and Sr₃La(PO₄)₃ were detected in the present study is somewhat incompatible with the previous study by Amezawa, but it could be due to some kinetic reason. The present results indicate that the solubility of Sr in LaPO₄ would be sufficiently high (at least 20.1 %) at 600 °C, which is in the operating temperature range for the electrolyte. However, the solubility of Sr is considerably lower than 13.4 % at 1200 °C, which is a typical sintering temperature for this material. Therefore, a new synthesis process of LaPO₄ electrolytes based on precipitation in homogeneous phosphoric acid solutions, which enables synthesizing dense LaPO₄ electrolytes without sintering, should have a considerable advantage. The reason why the solubility of Sr is higher at lower temperatures has not been clarified, but it is possible to speculate that Sr in the LaPO₄ lattice is destabilized with evaporation of water at high temperatures. In other words, the reverse reaction of Eq. (5.1) can proceed at higher temperatures.
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5.3.2 Sr-doped LaP₃O₉

The precipitation behavior of LaP₃O₉ in phosphoric acid solutions was also investigated at A230 condition. Single phase LaP₃O₉ was obtained in the solutions Sr4, Sr8, Sr12, Sr16, Sr20, Sr30, Sr40, and Sr50 whereas Sr(PO₃)₂ formed in the solutions Sr60 and Sr80. Figure 5.9 represents the Sr doping levels in precipitates. As shown in Figure 5.9 (a), the Sr doping level in LaP₃O₉ increased with increasing the initial Sr concentration in the solutions, reaching 14 mol% at maximum. This value is approximately three times higher than the previously attained value [2]. Similar to the case
of LaPO$_4$, the Sr doping level tended to decrease with increasing the precipitation ratio of La. However, as shown in Figure 5.6 (b), the Sr doping level relative to the initial Sr concentration in the solutions, ($x_{\text{Sr, prec}} / x_{\text{Sr, init}}$), was always less than 1. Thus, in contrast to the case of LaPO$_4$, the exhaustion of Sr is not thought to happen. The reason for the decrease in the Sr doping level is currently unknown, but the problem might be less serious than that with LaPO$_4$. In the same time, as is the case for LaPO$_4$, uniformly and highly Sr-doped LaP$_3$O$_9$ can be easily obtained in the solutions with higher Sr concentrations.

![Graph](image)

**Figure 5.9** (a) Dependence of the Sr doping level in LaP$_3$O$_9$ on the precipitation ratio of La. (b) Sr doping level in LaP$_3$O$_9$ relative to the initial Sr concentration in the solutions.

In order to see whether Sr was successfully incorporated in the LaP$_3$O$_9$ lattice, unit cell volume changes were determined by Rietveld analysis using powder X-ray diffraction data. As shown in Figure 5.10, LaP$_3$O$_9$ with higher Sr doping level had larger unit cell volume. This would indicate that Sr substituted for La in the LaP$_3$O$_9$ lattice, since the ionic radius of Sr$^{2+}$ in eightfold coordination
(1.26 Å) is larger than that of La$^{3+}$ (1.16 Å) [5].

![Figure 5.10 Unit cell volume of Sr-doped LaP$_3$O$_9$ as determined by Rietveld analysis using powder X-ray diffraction patterns.](image)

5.4 Conclusions

In this chapter, attempts were made to synthesize Sr-doped LaPO$_4$ and LaP$_3$O$_9$ in phosphoric acid solutions. The condition dependence of the Sr doping level was analyzed to reveal the optimum condition for obtaining highly Sr-doped phosphates.

The Sr doping level in LaPO$_4$ and LaP$_3$O$_9$ depended on the precipitation ratio of La, $R_{La}$. The Sr doping level gradually decreased with the progression of the precipitation, which was especially noticeable for low Sr concentration solutions. In the case of LaPO$_4$, Sr in the solutions was thought to be exhausted in early stages of the precipitation. However, with high Sr concentrations in the solutions, such as $\left(\frac{Sr}{La+Sr}\right) \geq 20$ mol%, the change in the Sr doping level was relatively small. In such solutions, almost uniformly Sr doped samples are expected to be obtained. The maximum Sr doping levels attained for LaPO$_4$ and LaP$_3$O$_9$ were ~23 % and ~14 %, respectively. These values are
several times higher than previously reported solubilities of Sr in LaPO₄ and LaP₃O₉.

The incorporation of Sr and H₂O in the LaPO₄ lattice was verified by unit cell volume change and evaporation of H₂O. Sr partially substituted for La in monazite-type LaPO₄ when synthesized in phosphoric acid solutions, as well as by solid state reaction as reported in the literature. On synthesis, H₂O was certainly incorporated in the bulk by Sr doping. The incorporation of Sr in the LaP₃O₉ lattice was also verified by unit cell volume change.

The stability of Sr-doped LaPO₄ at high temperatures was also investigated. The solubility of Sr in LaPO₄ was found to be sufficiently high (at least 20.1 %) even at 600 °C, which is in the operating temperature range for the electrolyte. However, it becomes considerably lower than 13.4 % at 1200 °C. The reason why the solubility of Sr is higher at lower temperatures has not been clarified, but it is possible to speculate that Sr in the LaPO₄ lattice is destabilized with evaporation of water at high temperatures.

References

Chapter 6

Electrical Conduction Properties of Sr-doped LaPO$_4$ and LaP$_3$O$_9$

6.1 Introduction

As mentioned in Chapter 1, the proton conductivity enhancement in lanthanum phosphates is desirable, while the proton conduction mechanisms are still unclear. In previous chapters, the precipitation conditions have been established for synthesizing highly Sr-doped LaPO$_4$ and LaP$_3$O$_9$ with various grain sizes. These techniques should broaden the way to investigate the proton conduction mechanisms, as well as to improve the electrolyte performances. Nevertheless, if they are heat treated at typical sintering temperatures to obtain dense electrolytes, they will decompose or Sr doping level will decrease (see previous chapters). Therefore, the direct synthesis of dense electrolytes at low temperatures is necessary for investigating their electrical conduction properties.

In this chapter, attempts were made to obtain dense, polycrystalline plates of Sr-doped LaPO$_4$ and LaP$_3$O$_9$ in phosphoric acid solutions by gradually providing the driving force for the precipitation. The synthesis of large single crystals of Sr-doped LaP$_3$O$_9$ was also tried with a similar strategy. Then the electrical conduction properties of the precipitates were investigated by impedance spectroscopy.

6.2 Experimental

6.2.1 Synthesis

Sr-doped LaPO$_4$ and LaP$_3$O$_9$ were synthesized in phosphoric acid solutions based on the method
given in the previous chapter. La$_2$O$_3$ (99.99%, Nacalai Tesque), SrCO$_3$ (99.99%, Wako Pure Chemical Industries), and H$_3$PO$_4$ (85%, Nacalai Tesque) were mixed in glassy carbon crucibles and heated at 190 °C for several days to dissolve all solid compounds in phosphoric acid. Then the solutions were gradually cooled or heated in ambient ($p_{\text{H}_2\text{O}} = \sim 0.01$ atm) or humidified ($p_{\text{H}_2\text{O}} = \sim 1$ atm) atmospheres to precipitate lanthanum phosphates. The synthesis parameters are given in Table 6.1. These parameters were chosen based on the solubility of lanthanum oxide in phosphoric acid solutions at each temperature. The following requirements were considered: (1) No precipitate should be formed at the initial temperature in the precipitation stage; (2) The desired precipitate should be formed at the final temperature. As a result, a moderate driving force for the precipitation should begin to be provided at a certain point during the cooling or heating process. After the temperature had reached the final temperature, the solutions were held isothermally for several days. For humidification, the apparatus used in Chapter 2 was used.

**Table 6.1 Synthesis parameters for Sr-doped LaPO$_4$ and LaP$_3$O$_9$ electrolytes.** $p_{\text{H}_2\text{O}}$ is water vapor pressure, $T_{\text{initial}}$ and $T_{\text{final}}$ denote temperatures at which slow cooling/heating started and ended, respectively, $\beta$ is the heating rate, and $t_{\text{isothermal}}$ is the duration of the isothermal heating.

<table>
<thead>
<tr>
<th>Product</th>
<th>Initial composition</th>
<th>Precipitation parameters</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>La : Sr : P</td>
<td>$p_{\text{H}<em>2\text{O}}$ / atm, $T</em>{\text{initial}}$ / °C, $T_{\text{final}}$ / °C, $\beta$ / °C h$^{-1}$, $t_{\text{isothermal}}$ / day</td>
</tr>
<tr>
<td>(a) Sr-doped LaPO$_4$ polycrystalline plates</td>
<td>0.48 : 0.12 : 15</td>
<td>$\sim 1$, 250, 200, $-1$, 1</td>
</tr>
<tr>
<td>(b) Sr-doped LaP$_3$O$_9$ polycrystalline plates</td>
<td>0.80 : 0.20 : 15</td>
<td>$\sim 0.01$, 190, 250, 1, 8</td>
</tr>
<tr>
<td>(c) Sr-doped LaP$_3$O$_9$ single crystals</td>
<td>0.40 : 0.10 : 15</td>
<td>$\sim 1$, 300, 360, 0.5, 4</td>
</tr>
</tbody>
</table>

**6.2.2 Characterization**

Phase identification was carried out via X-ray diffraction (XRD) analysis on PANalytical X’Pert-Pro MPD using Cu-K$_\alpha$ radiation at room temperature. The morphology and composition of
precipitates were investigated using KEYENCE VE-7800 scanning electron microscope (SEM) equipped with energy dispersive X-ray spectrometer (EDS, EDAX Genesis-XM2). Chemical analysis was also performed by inductively coupled plasma atomic emission spectrometry (ICP-AES) on Seiko Instruments SPS4000. It has been found that, in general, the difference between the values for the Sr doping level of the same sample derived from EDS and ICP analyses are within 3%.

6.2.3 AC Impedance Spectroscopy

For measuring the electrical conductivity of the polycrystalline plates, silver paste (Fujikura Kasei, D-550) was painted on each surface of the precipitates. A silver wire and a silver mesh were attached to each surface using the silver paste. Then the sample was fixed to the sample holder shown in Figure 6.1 (a). For measuring the electrical conductivity of the single crystals, platinum wires were attached to parallel surfaces using the silver paste. Then the sample was fixed to the sample holder shown in Figure 6.1 (b).
Figure 6.1 Schematic illustrations and typical photographs of the sample holders for conductivity measurements. (a) For polycrystalline plates; (b) For single crystals.

The conductivity of the samples was measured by impedance spectroscopy. A.C. impedance data were collected over the frequency range of 10 Hz–10 MHz using a frequency response analyzer (Solartron SI 1260) at an applied voltage of typically 100 mV. Humidified hydrogen, argon or their mixture was flowed over the samples at typically 20 sccm. The fitting of the impedance data was performed using the commercial software package, ZView Version 2.9c (Scribner Associates).

6.3 Results and Discussion

6.3.1 Direct Synthesis of Electrolytes in Phosphoric Acid Solutions

6.3.1.1 Sr-doped LaPO₄ Polycrystalline Plates

Figure 6.2 shows a typical photograph and SEM image of Sr-doped LaPO₄ precipitated on the bottom surface of the crucible with the synthesis parameters given in Table 6.1 (a). The precipitate
formed a plate which can be handled with tweezers or by hand. The thickness of the plate was about 150 \( \mu \text{m} \). As can be seen in the cross-sectional image of the plate (Figure 6.2 (b)), the plate mainly consisted of densely packed, elongated grains of ~100 \( \mu \text{m} \) arranged in the direction perpendicular to the surfaces. In addition, a closer look at the lower part of the plate, which was in contact with the crucible, reveals the existence of a thin layer consisted of relatively small grains. The overall geometry of the cross-section resembles the general microstructure of ingots, with the lower part and upper part of the plate corresponding to the chill zone and columnar zone in an ingot, respectively. This suggests that heterogeneous nucleation took place on the bottom surface of the crucible and then large columnar grains were formed upward as a result of competitive growth.

Figure 6.3 shows the X-ray diffraction patterns obtained from each surface of the Sr-doped LaPO\(_4\) plate. The relative intensities of the peaks were significantly different from those expected from randomly oriented powders, indicating the existence of preferred orientations of the grains. While the (111) planes were oriented parallel to the surfaces in the lower part, the (100) planes were oriented parallel to the surfaces in the upper part. Therefore, the perpendicular direction to the (100) planes is considered to be the preferred growth direction for LaPO\(_4\) crystals.

By EDS analysis, the Sr doping level was found to be 13% at the lower surface and 18% at the upper surface. These results show that, dense, highly-oriented and coarse-grained polycrystalline plates of highly Sr-doped LaPO\(_4\) can be directly synthesized in phosphoric acid solutions, which cannot be obtained by conventional solid state reaction.
6.3.1.2 Sr-doped LaP$_3$O$_9$ Polycrystalline Plates

Figure 6.4 shows a typical photograph and SEM image of Sr-doped LaP$_3$O$_9$ precipitated on the bottom surface of the crucible with the synthesis parameters given in Table 6.1 (b). Similarly to the
In the case of Sr-doped LaPO₄, the Sr-doped LaP₃O₉ precipitate formed a plate with a thickness of about 250 µm and it had the chill zone and columnar zone. (As the plate was precipitated by heating, “chill zone” might not be appropriate.) Figure 6.5 shows the X-ray diffraction patterns obtained from each surface of the Sr-doped LaP₃O₉ plate. While the (110) planes tended to be oriented parallel to the surfaces in the lower part, the (001) planes were strongly oriented parallel to the surfaces in the upper part. Therefore, the perpendicular direction to the (001) planes is considered to be the preferred growth direction for LaP₃O₉ crystals. As LaP₃O₉ has an orthorhombic crystal structure, the preferred growth direction is identical to the c-axis. By EDS analysis, the Sr doping level was found to be 11% at the lower surface and 5% at the upper surface. In summary, as is the case for Sr-doped LaPO₄, dense, highly-oriented and coarse-grained polycrystalline plates of highly Sr-doped LaP₃O₉ can also be directly synthesized in phosphoric acid solutions.

Figure 6.4 Typical (a) photograph and (b) SEM image of the polycrystalline plate of Sr-doped LaP₃O₉.
Figure 6.5 X-ray diffraction patterns of a polycrystalline plate of Sr-doped LaP$_3$O$_9$: (a) upper surface of the plate; (b) lower surface of the plate. The numbers in parentheses represent the Miller indices of peaks.

### 6.3.1.3 Sr-doped LaP$_3$O$_9$ Single Crystals

Figure 6.6 shows a typical photograph and SEM image of Sr-doped LaP$_3$O$_9$ precipitated at the bottom of the crucible with the synthesis parameters given in Table 6.1 (c). Transparent, column crystals with well-developed morphology were obtained. The crystals were nearly 2 mm in length, sufficiently large for conductivity measurements. This result confirms that the “high-humidity and high-temperature” condition established in Chapter 2 facilitates the growth of LaP$_3$O$_9$ crystals. By X-ray diffraction, it was found that the c-axis of the LaP$_3$O$_9$ lattice was always directed parallel to the longer direction of the crystals (Figure 6.7). This is consistent with the preferred growth direction of LaP$_3$O$_9$ suggested in the previous subsection (c-axis). As LaP$_3$O$_9$ has an orthorhombic crystal structure, the ab plane is perpendicular to the c-axis, but it was difficult to distinguish crystallographic directions in the ab plane by crystal morphology. By ICP analysis, the Sr doping level was found to be 5.1%, sufficient for generating protonic charge carriers.
6.3.2 Electrical Conduction Properties

6.3.2.1 Sr-doped LaPO$_4$ Polycrystalline Plates

Figure 6.8 shows the impedance data of a Sr-doped LaPO$_4$ plate collected at 500 °C under various atmospheres with different hydrogen partial pressures $p_{H_2}$ and water vapor pressures $p_{H_2O}$. Two impedance arcs were observed in the impedance spectra shown in Figure 6.8 (a). The
impedance spectra were analyzed using an equivalent circuit model \((R_1Q_1)(R_2Q_2)\) where \(R\) is a resistor and \(Q\) is a constant phase element. The resistances corresponding to high- and low-frequency arcs, \(R_1\) and \(R_2\), showed different dependences on hydrogen and water partial pressures. Figure 6.8 (b) shows the deviation in \(R_1\) and \(R_2\) on changing the atmosphere. While \(R_1\) was independent of hydrogen partial pressure, \(R_2\) increased with decreasing hydrogen partial pressure (atmospheres (1)–(3) in the figure).

In the present atmospheres, the forward and reverse reactions of the following equation are mainly expected to take place at both electrodes:

\[
2H^+ + 2e = H_2 (g)
\]  

(6.1)

where \(H^+\) represents a proton in the electrolyte. Generally, at low overpotentials, the charge transfer resistance \(R_{ct}\) is inversely proportional to the exchange current density \(i_0\) as follows:

\[
R_{ct} = \frac{RT}{nFi_0}
\]  

(6.2)

where \(R\), \(T\), \(n\), and \(F\) are the gas constant, absolute temperature, number of electrons in the reaction, and Faraday’s constant, respectively. The exchange current density depends on the concentrations of reactants. For the reaction of Eq. (6.1), the dependence is given as

\[
i_0 \propto c_{H^+}^{1-\alpha} \cdot p_{H_2}^{\alpha} \quad (0 < \alpha < 1)
\]  

(6.3)

where \(c_{H^+}\), \(p_{H_2}\), and \(\alpha\) are the concentration of protons in the electrolyte, hydrogen partial pressure, and charge transfer coefficient, respectively. Combining Eqs. (6.2) and (6.3) yields

\[
R_{ct} \propto \frac{1}{c_{H^+}^{1-\alpha} \cdot p_{H_2}^{\alpha}}
\]  

(6.4)

Therefore, if \(c_{H^+}\) is constant, changing hydrogen partial pressure from \(p_{H_2,i}\) to \(p_{H_2,f}\) should result in the charge transfer resistance change from \(R_{ct,i}\) to \(R_{ct,f}\) according to the following relationship:

\[
\frac{R_{ct,f}}{R_{ct,i}} = \left( \frac{p_{H_2,i}}{p_{H_2,f}} \right)^\alpha.
\]  

(6.5)
In the present measurements, the value of $R_2$ increased from 21.7 kΩ to 30.1 kΩ with decreasing hydrogen partial pressure from 0.9 atm to 0.45 atm. This behavior fits well to Eq. (6.5) with an $\alpha$ value of ~0.47. Therefore, $R_2$ would surely correspond to the charge transfer resistance $R_{ct}$. Accordingly, $R_1$ would be attributed to the electrolyte. As mentioned above, $R_1$ did not depend on hydrogen partial pressure. Kitamura et al. reported that the conductivity of Sr-doped LaPO$_4$ was almost independent of hydrogen partial pressure [1]. Hence it is reasonable to assume $R_1$ as the resistance of the Sr-doped LaPO$_4$ plate. For further analysis, the conductivity of the Sr-doped LaPO$_4$ plate, $\sigma$, was calculated from $R_1$.

When $p_{H2}$ and $p_{H2O}$ were varied simultaneously, from atmosphere (3) to (4) in Figure 6.8, both $R_1$ and $R_2$ changed. As it was found that $R_1$ did not depend on $p_{H2}$, the decrease in $R_1$ can be attributed to the increase in $p_{H2O}$. It suggests that the charge carrier concentration in the electrolyte increased with increasing $p_{H2O}$, and it is consistent with the idea that the protons are introduced into Sr-doped LaPO$_4$ through the equilibrium with water vapor [2] (See Chapter 1). As shown in Figure 6.9, the conductivity of the Sr-doped LaPO$_4$ plate exhibited a similar dependence on water vapor pressure to those reported for 2% or 3% Sr-doped LaPO$_4$ [3]. The slope of the log $\sigma$ – log $p_{H2O}$ plot was 0.09, much smaller than 1/2. As discussed by Amezawa et al. [3], it implies that the sample was mostly hydrated in the present atmospheres.
Figure 6.8 (a) Impedance spectra of a polycrystalline plate of Sr-doped LaPO$_4$ at 500 °C under various atmospheres ((1)-(4)). Inset is an enlarged view of the high-frequency arc. (b) Time dependence of the resistance of Sr-doped LaPO$_4$ at 500 °C under various atmospheres. $R_1$ and $R_2$ represent the resistances corresponding to the high- and low-frequency arcs in the impedance spectra, respectively.
Assuming a linear relationship between the conductivity derived from $R_1$ and the proton concentration in the electrolyte, a proton concentration increase of 1.11 times is expected with increasing $p_{H_2O}$ from 0.1 atm to 0.3 atm (Figure 6.8 and Figure 6.9). It is also possible to estimate the proton concentration change by the charge transfer resistance. From Eq. (6.4), the following relationship is derived for the variation in charge transfer resistance caused by the concentration changes:

$$\frac{R_{ct,f}}{R_{ct,i}} = \frac{(c_{H^{+},i})^{1-\alpha}}{(c_{H^{+},f})^{1-\alpha}} \left(\frac{p_{H_2,i}}{p_{H_2,f}}\right)^{\alpha}$$  \hspace{1cm} (6.6)

where the suffixes “i” and “f” represent “before” and “after” the concentration changes. Using $R_{ct,f} / R_{ct,i} = 22.4/21.7$, $p_{H_2,i} / p_{H_2,f} = 0.9/0.7$, and the $\alpha$ value of 0.47 derived above, the proton concentration change is given as

$$\frac{c_{H^{+},f}}{c_{H^{+},i}} = 1.18.$$

This is in a fair agreement with the value of 1.11 derived from the conductivity of the electrolyte.
indicating the current interpretation is valid. Thus, the proton conduction was confirmed in the Sr-doped LaPO$_4$ plate which was directly synthesized in a phosphoric acid solution.

Figure 6.10 shows the temperature dependence of the conductivities of Sr-doped LaPO$_4$ plates. Besides the above used plate (“Plate A”), two additional plates (“Plate B” and “Plate C”) were used for the conductivity measurements. All the plates were obtained in the same solution. The conductivities of the three samples were of the same order of magnitude as those reported in the literature [3], and showed similar temperature dependences. The activation energies for charge transport, derived from the slopes of the plots, were 0.80 ± 0.05 eV at 400–500 °C while Amezawa et al. reported a value of 0.86 eV for 2% Sr-doped LaPO$_4$ at 500–600 °C [3].

![Figure 6.10](image_url)

**Figure 6.10** Temperature dependence of the conductivity of Sr-doped LaPO$_4$ plates. “Plate A”, “Plate B”, and “Plate C” are different samples obtained from the same crucible.

However, as obvious from the figure, the three samples exhibited different absolute values for the conductivity although they should be equal. The conductivities of each sample shown in Figure
6.10 were essentially reproducible through repeated heating/cooling cycles, but there was even a case where the conductivity changed through heating/cooling cycles (Not shown on the figure). The reason for the conductivity differences might be, for example, inadequate contact between the electrolyte and electrodes, or breakage of the electrolyte due to its insufficient mechanical strength, but the exact reason is not clear at present. Although the atmospheres were slightly different in which each measurement was done, it would not result in the severalfold differences of the conductivities. For further studies of the effect of Sr doping level, grain size, and crystallographic orientation on the conduction properties, it is required to establish the experimental conditions where reproducible results can be obtained.

6.3.2.2 Sr-doped LaP₃O₉ Polycrystalline Plates

Figure 6.11 shows the impedance data of a Sr-doped LaP₃O₉ plate collected at 500 °C with different hydrogen partial pressures. Two partly overlapping arcs were identified in the impedance spectra shown in Figure 6.11 (a). Fitting to the impedance spectra was carried out using an equivalent circuit model \((R_1)(R_2Q_2)\) at low frequencies (100–1000 Hz) where the low frequency arc was dominant. The high frequency intercept of the fit arc with the real axis, \(R_1\), was regarded as the resistance corresponding to the high frequency arc. The resistances corresponding to the high- and low-frequency arcs, \(R_1\) and \(R_2\), showed analogous dependences on hydrogen partial pressure to those observed for Sr-doped LaPO₄ plates: While \(R_1\) remained constant, \(R_2\) increased with decreasing hydrogen partial pressure. The variation in \(R_2\) fits well to Eq. (6.5) with an \(\alpha\) value of ~0.39. For these reasons, \(R_1\) and \(R_2\) can be assigned to the electrolyte resistance and charge transfer resistance, respectively. However, because of the overlap of two impedance arcs, it was difficult to obtain accurate \(R_1\) values at lower temperatures and hence the temperature dependence of the conductivity was not analyzed. The conductivity of the plate at 500 °C calculated from \(R_1\) was \(1.4 \times 10^{-5}\) S cm⁻¹, while Amezawa et al. reported a value of \(8.2 \times 10^{-5}\) S cm⁻¹ for “5% Sr-doped LaP₃O₉” [4]. For
clarifying the reason for this difference as well as the effects of Sr doping level, grain size, and crystallographic orientation on the conduction properties, further experiments with sufficient reproducibility need to be carried out in future work.
Figure 6.11 (a) Impedance spectra of a polycrystalline plate of Sr-doped LaP$_3$O$_9$ at 500 °C under various atmospheres ((1)-(3)). Solid lines are the fitted curves to the low-frequency arc using an equivalent circuit model ($R_1$)($R_2Q_2$) where $R$ is a resistor and $Q$ is a constant phase element. $R_1$ and $R_2$ represent the resistances corresponding to the high- and low-frequency arcs. (b) Time dependence of the resistance of Sr-doped LaP$_3$O$_9$ at 500 °C under various atmospheres.
6.3.2.3 Sr-doped LaP$_3$O$_9$ Single Crystals

Figure 6.12 shows the impedance spectra of a Sr-doped LaP$_3$O$_9$ single crystal measured along the c-axis at 500 °C with different hydrogen partial pressures. Two impedance arcs were identified as was the case for Sr-doped LaPO$_4$ and LaP$_3$O$_9$ plates. The resistances corresponding to the high- and low-frequency arcs, $R_1$ and $R_2$, showed analogous dependences on hydrogen partial pressure to those observed for Sr-doped LaPO$_4$ and LaP$_3$O$_9$ plates: While $R_1$ remained constant, $R_2$ increased with decreasing hydrogen partial pressure. The variation in $R_2$ fits well to Eq. (6.5) with an $\alpha$ value of $\sim$0.49. For these reasons, $R_1$ and $R_2$ can be assigned to the electrolyte resistance and charge transfer resistance, respectively. The impedance spectra measured in the ab plane also exhibited two arcs and the resistance corresponding to the high-frequency arc was taken as the electrolyte resistance.

![Figure 6.12 (a) Impedance spectra of a Sr-doped LaP$_3$O$_9$ single crystal measured along the c-axis at 500 °C.](image)

<table>
<thead>
<tr>
<th>Atmosphere</th>
<th>$\rho_{H_2}$ / atm</th>
<th>$\rho_{H_2O}$ / atm</th>
</tr>
</thead>
<tbody>
<tr>
<td>(1)</td>
<td>0.95</td>
<td>0.05</td>
</tr>
<tr>
<td>(2)</td>
<td>0.05</td>
<td>0.05*</td>
</tr>
</tbody>
</table>

*Balanced with Ar.

Figure 6.13 shows the temperature dependence of the conductivities of Sr-doped LaP$_3$O$_9$ single crystals measured along the c-axis or in the ab plane, using two crystals for each direction. The specific crystallographic directions in the ab plane might be different because they were not identified. All the samples were synthesized in the same crucible. The measured conductivities were distributed over a range of 1 to 5 times the highest values reported by Amezawa et al. for Sr-doped LaP$_3$O$_9$. The activation energies for charge transport were calculated to be $0.78 \pm 0.04$ eV at 300–
500 °C and there was no significant dependence on the crystallographic direction. The obtained activation energies were also comparable with the value reported by Amezawa et al. for 1% Sr-doped LaP$_3$O$_9$ at 300–500 °C (0.84 eV) [5]. On the other hand, the absolute values for the conductivity tended to be higher along the c-axis than in the ab plane, though exact directions in the ab plane were not specified.

![Graph](image-url)

**Figure 6.13 Temperature dependence of the conductivity of Sr-doped LaP$_3$O$_9$ single crystals.**

These results provide a clue to determine the dominant proton conduction paths in Sr-doped LaP$_3$O$_9$. In the present measurements, the grain boundary conductivity is considered to be negligible compared with those of polycrystalline samples synthesized by solid state reaction, because the samples were composed of significantly larger crystals (even if they were not perfect single crystals). Therefore, the fact that Sr-doped LaP$_3$O$_9$ single crystals exhibited comparable or larger conductivities to those reported for polycrystalline samples suggests that, there are substantial proton conduction paths in the LaP$_3$O$_9$ lattice. The observed anisotropy of the conductivity also supports this
idea, since the crystal structure of LaP$_3$O$_9$ is highly anisotropic as shown in Chapter 1. If grain boundary conduction were more dominant than bulk conduction in polycrystalline LaP$_3$O$_9$, the increase in grain size should lead to the decrease in the conductivity. Thus, the present result would be difficult to be explained. Therefore, it can be concluded that grain growth and orientation control might be a promising strategy for improving the conductivity of Sr-doped LaP$_3$O$_9$ electrolytes.

### 6.4 Conclusions

In this chapter, dense, highly-oriented and coarse-grained polycrystalline plates of highly Sr-doped LaPO$_4$ and LaP$_3$O$_9$ were directly synthesized in phosphoric acid solutions by slow cooling or heating of the solutions. The thicknesses of the plates were 150–250 µm. Single crystals of Sr-doped LaP$_3$O$_9$ larger than 1 mm were also obtained by slow heating of the solution in the “high-humidity and high-temperature” condition.

The electrical conduction properties of the precipitates were investigated by impedance spectroscopy. For all samples, the resistance corresponding to the proton conduction in electrolytes was identified. It was demonstrated in principle that the Sr-doped LaPO$_4$ and LaP$_3$O$_9$ polycrystalline plates, which were directly synthesized in phosphoric acid solutions, also function as proton conducting electrolytes as well as those prepared by solid state reaction in previous studies.

The results on Sr-doped LaP$_3$O$_9$ single crystals suggested that there are substantial proton conduction paths in the LaP$_3$O$_9$ lattice. Therefore, it can be said that grain growth and orientation control might be a promising strategy for improving the conductivity of Sr-doped LaP$_3$O$_9$ electrolytes.

Finally, it should be noted that there is much room for further characterization of the proton conduction properties of Sr-doped LaPO$_4$ and LaP$_3$O$_9$ precipitates in terms of reproducibility, effects of doping level and crystallographic direction, and so on.
References


Chapter 7
Summary

Lanthanum phosphates are attracting much attention due to their proton conduction. However, further proton conductivity enhancement is desirable, while the proton conduction mechanisms such as the dominant proton conduction paths are still unclear. Their thermodynamic characteristics, which are necessary for evaluating their stabilities, have also been scarce.

In this work, systematic investigations have been carried out to establish the new, low-temperature synthesis method of lanthanum phosphate electrolytes in phosphoric acid solutions. The synthesis method enabled significant increase in Sr doping levels as well as vigorous crystal growth compared with conventional solid state reaction method. This method may serve for revealing the proton conduction mechanisms, and enhancing the proton conductivity. The phase equilibria and thermodynamic properties of the La$_2$O$_3$–P$_2$O$_5$ system were also characterized and their stabilities were evaluated. Finally, the electrical conduction properties of lanthanum phosphate electrolytes synthesized in phosphoric acid solutions were investigated.

Chapter 2 Synthesis of Lanthanum Phosphates in Phosphoric Acid Solutions

In this chapter, the precipitation conditions of LaPO$_4$, LaP$_3$O$_9$, and LaP$_5$O$_{14}$ in condensed phosphoric acid solutions were established under both atmospheric and humidified conditions. The obtained LaPO$_4$ had the monazite-type structure, which has been reported to exhibit proton conduction. Increasing water vapor pressure from 0.01 atm to 1 atm extended the precipitation regions of LaPO$_4$ and LaP$_3$O$_9$ toward higher temperatures by nearly 100 °C. The crystal growth of
LaPO$_4$ was significantly accelerated under such conditions (e.g. $T=250$ °C and $p_{H2O} = 1$ atm). The crystal growth of LaP$_3$O$_9$ was also promoted moderately by the increase in precipitation temperature. These results will serve as fundamental information for further optimization of growth parameters (e.g. temperature profile, precipitation period, and initial composition) and the direct synthesis of dense lanthanum phosphate electrolytes in phosphoric acid solutions.

Chapter 3 Phase Equilibria in the La$_2$O$_3$–P$_2$O$_5$ System

In this chapter, the phase equilibria in the La$_2$O$_3$–P$_2$O$_5$ system were re-examined to verify the known phase diagrams. Two intermediate phases, La$_3$PO$_7$ and La$_7$P$_3$O$_{18}$, were identified in the La$_2$O$_3$–LaPO$_4$ subsystem. La$_8$PO$_{10}$ was not observed. Therefore, the phase diagram reported by Park and Kreidler seems more likely to be accurate in that. La$_3$PO$_7$ was formed at 1000 °C, while La$_7$P$_3$O$_{18}$ was only formed at 1200 °C and higher. At 1000 °C, neither the formation nor decomposition of La$_7$P$_3$O$_{18}$ took place within 1000 h. This behavior might be due to kinetic reasons, though it was not verified. In the LaPO$_4$–LaP$_3$O$_9$ subsystem, no intermediate phase such as La$_2$P$_4$O$_{13}$ was detected. Based on these results, the most likely representation of the phase diagram is given.

Chapter 4 Thermodynamic Properties of LaP$_3$O$_9$ and LaP$_5$O$_{14}$

LaP$_3$O$_9$ and LaP$_5$O$_{14}$ have been demonstrated to decompose gradually at high temperatures. However, very little quantitative information is available about the decomposition reactions. In this chapter, for accurate assessment of the stabilities of LaP$_3$O$_9$ and LaP$_5$O$_{14}$, vapor pressure measurements on LaP$_3$O$_9$ and LaP$_5$O$_{14}$ were carried out by the transpiration method. The equilibrium vapor pressures of P$_4$O$_{10}$ over LaP$_3$O$_9$ and LaP$_5$O$_{14}$ were obtained to be

$$\log \left( \frac{p_{P_4O_{10}}} {\text{atm}} \right) = (5.12\pm1.00) - (1.22\pm0.12) \times 10^4 \times \left( \frac{T} {K} \right)^{-1}$$  
(over LaP$_3$O$_9$, 1123 ≤ $T / K$ ≤ 1323);

$$\log \left( \frac{p_{P_4O_{10}}} {\text{atm}} \right) = (4.65\pm1.80) - (1.00\pm0.19) \times 10^4 \times \left( \frac{T} {K} \right)^{-1}$$
The standard enthalpies of the evaporation reactions are given as $117\pm12$ kJ mol$^{-1}$ and $96\pm18$ kJ mol$^{-1}$ in the corresponding temperature ranges. The data were evaluated with respect to the proposed use of LaP$_3$O$_9$ and LaP$_5$O$_{14}$ as solid electrolytes in fuel cells. The maximum operating temperatures for the LaP$_3$O$_9$ and LaP$_5$O$_{14}$ electrolytes were suggested to be about 600 °C and 450 °C in oxygen atmosphere. Taking into account their proton conductivities at these temperatures, it can be said that LaP$_3$O$_9$ based electrolytes are more promising than LaP$_5$O$_{14}$ based ones. In the presence of water vapor or hydrogen, however, decomposition of LaP$_3$O$_9$ and LaP$_5$O$_{14}$ is expected to be accelerated considerably.

**Chapter 5 Synthesis of Highly Sr-doped LaPO$_4$ and LaP$_3$O$_9$**

In Chapter 2, the precipitation conditions for undoped LaPO$_4$ and LaP$_3$O$_9$ in homogeneous phosphoric acid solutions were established. In this chapter, further attempts were made to synthesize Sr-doped LaPO$_4$ and LaP$_3$O$_9$ in phosphoric acid solutions. The condition dependence of the Sr doping level was analyzed to reveal the optimum condition for obtaining highly Sr-doped phosphates.

The Sr doping level in LaPO$_4$ and LaP$_3$O$_9$ depended on the precipitation ratio of La, $R_{La}$. The Sr doping level gradually decreased with the progression of the precipitation, which was especially noticeable for low Sr concentration solutions. In the case of LaPO$_4$, Sr in the solutions was thought to be exhausted in early stages of the precipitation. However, with high Sr concentrations in the solutions, such as $\left(\frac{\text{Sr}}{\text{La}+\text{Sr}}\right) \geq 20$ mol%, the change in the Sr doping level was relatively small. In such solutions, almost uniformly Sr doped samples are expected to be obtained. The maximum Sr doping levels attained for LaPO$_4$ and LaP$_3$O$_9$ were ~23% and ~14%, respectively. These values are several times higher than previously reported solubilities of Sr in LaPO$_4$ and LaP$_3$O$_9$. 
The effectiveness of Sr doping in LaPO$_4$ was verified by observing unit cell volume changes by doping and the evaporation of H$_2$O which had been incorporated into LaPO$_4$ during the synthesis. Sr partially substituted for La in monazite-type LaPO$_4$ when synthesized in phosphoric acid solutions, as well as by solid state reaction as reported in literature. On synthesis, H$_2$O was certainly incorporated in the bulk by Sr doping.

The stability of Sr-doped LaPO$_4$ at high temperatures was also investigated. The solubility of Sr in LaPO$_4$ was found to be sufficiently high (at least 20.1 %) even at 600 °C, which is in the operating temperature range for the electrolyte. However, it becomes considerably lower than 13.4 % at 1200 °C. The reason why the solubility of Sr is higher at lower temperatures has not been clarified, but it is possible to speculate that Sr in the LaPO$_4$ lattice is destabilized with evaporation of water at high temperatures.

Chapter 6 Electrical Conduction Properties of Sr-doped LaPO$_4$ and LaP$_3$O$_9$

The proton conductivity enhancement in lanthanum phosphates is desirable, while the proton conduction mechanisms are still unclear. In this chapter, dense, plate-like precipitates of Sr-doped LaPO$_4$ and LaP$_3$O$_9$ were directly synthesized in phosphoric acid solutions by slow cooling or heating of the solutions. Single crystals of Sr-doped LaP$_3$O$_9$ larger than 1 mm were also obtained by slow heating of the solution in the “high-humidity and high-temperature” condition.

The electrical conduction properties of the precipitates were investigated by impedance spectroscopy. For all samples, the resistance corresponding to the proton conduction in electrolytes was identified. It was demonstrated in principle that the Sr-doped LaPO$_4$ and LaP$_3$O$_9$ plates which were directly synthesized in a phosphoric acid solutions also function as proton conducting electrolytes as well as those prepared by solid state reaction in previous studies.

The results on Sr-doped LaP$_3$O$_9$ single crystals suggested that there are substantial proton
conduction paths in the LaP$_3$O$_9$ lattice. Therefore, it can be said that grain growth and orientation control might be a promising strategy for improving the conductivity of Sr-doped LaP$_3$O$_9$ electrolytes.
List of Publications

Lanthanum Phosphates

Papers

- “Precipitation Behavior of Highly Sr-doped LaPO₄ in Phosphoric Acid Solutions”
  Naoyuki Hatada, Yoshitaro Nose, Akiko Kuramitsu, and Tetsuya Uda,
- “First-principles Thermodynamics of La₂O₃-P₂O₅ Pseudobinary System”
  Kazuaki Toyoura, Naoyuki Hatada, Yoshitaro Nose, Tetsuya Uda, and Isao Tanaka,
- “Phase Equilibria in the La₂O₃–P₂O₅ System”
  Naoyuki Hatada, Takashi Nagai, Yoshitaro Nose, and Tetsuya Uda,
  (in preparation).
- “Vapor Pressure Measurements on Lanthanum Polyphosphate and Ultraphosphate by Transpiration Method”
  Naoyuki Hatada, Kazuaki Toyoura, Yoshitaro Nose, and Tetsuya Uda,
  (in preparation).
- “Crystal Growth of LaP₃O₉ in phosphoric acid solutions”
  Takayuki Onishi, Naoyuki Hatada, Yoshitaro Nose, and Tetsuya Uda,
  (in preparation).
- “High proton conduction in heavily Sr-doped LaP₃O₉ single crystals”
  Naoyuki Hatada, Takayuki Onishi, Kazuaki Toyoura, Yoshitaro Nose, and Tetsuya Uda,
  (in preparation).
Oral Presentations

International Conferences

• “Synthesis of Lanthanum Polyphosphate in Condensed Phosphoric Acid Solutions”
  Naoyuki Hatada, Yoshitaro Nose, and Tetsuya Uda,

• “Synthesis of Lanthanum Orthophosphate and Polyphosphate in Condensed Phosphoric Acid Solutions”
  Naoyuki Hatada, Yoshitaro Nose, and Tetsuya Uda,

• “Synthesis of Highly Sr-doped Lanthanum Orthophosphate and Polyphosphate in Phosphoric Acid Solutions”
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  The 15th International Conference on Solid State Protonic Conductors, Santa Barbara, USA, August, 2010.

• “Synthesis of Sr-doped LaP$_3$O$_9$ Films in Phosphoric Acid Solutions and Their Proton Conduction Properties”
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• “Thermal Stability of LaP$_3$O$_9$ and LaP$_5$O$_{14}$ Proton Conductors”
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Domestic Conferences in Japan

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  Naoyuki Hatada, Takahito Nakajima, Yoshitaro Nose, and Tetsuya Uda,
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  2010 Fall Meeting of the Mining and Materials Processing Institute of Japan, Fukuoka, September, 2010.

• “Dependence of the Sr-doping level in LaPO$_4$ Precipitated in Phosphoric Acid Solutions on the Precipitation Ratio of La and Direct Synthesis of Plate-like Electrolytes”
  Naoyuki Hatada, Akiko Kuramitsu, Yoshitaro Nose, and Tetsuya Uda,

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• “Crystal Structures and Proton Conductivities in Lanthanum Phosphates”
  Kazuaki Toyoura, Naoyuki Hatada, Yoshitaro Nose, and Tetsuya Uda,
  The 78th Meeting of the Electrochemical Society of Japan, Yokohama, March, 2011 (cancelled).

• “Investigation of the Thermal Stabilities of Proton Conductors LaP$_3$O$_9$ and LaP$_5$O$_{14}$ by Transpiration Method”
  Naoyuki Hatada, Kazuaki Toyoura, Yoshitaro Nose, and Tetsuya Uda,
  2011 Fall Meeting of the Mining and Materials Processing Institute of Japan, Sakai (Osaka), September, 2011.

• “Evaluation of the Thermodynamic Properties of LaP$_3$O$_9$ and LaP$_5$O$_{14}$ by Third-Law Analysis”
  Naoyuki Hatada, Kazuaki Toyoura, Yoshitaro Nose, and Tetsuya Uda,

• “Phase Stabilities of the La$_2$O$_3$ – P$_2$O$_5$ Pseudo-Binary System Based on First Principle Thermodynamics”
  Kazuaki Toyoura, Naoyuki Hatada, Yoshitaro Nose, Tetsuya Uda, and Isao Tanaka,

• “Evaluation of the Proton Conduction Properties of Sr-doped Lanthanum Polyphosphate Using Single Crystals”
  Takayuki Onishi, Naoyuki Hatada, Akiko Kuramitsu, Yoshitaro Nose, and Tetsuya Uda
Poster Presentations

International Conferences

• “Synthesis of Lanthanum Polyphosphate Single Crystals”
  Naoyuki Hatada, Tetsuya Uda, Yoshitaro Nose, and Yasuhiro Awakura,

• “Synthesis of Highly Sr-doped Lanthanum Orthophosphate in Phosphoric Acid Solutions”
  Naoyuki Hatada, Akiko Kuramitsu, Yoshitaro Nose, and Tetsuya Uda,

• “Vapor Pressure Measurements on LaP$_3$O$_9$ by the Transpiration Technique”
  Naoyuki Hatada, Yoshitaro Nose, and Tetsuya Uda,

• “Correlation between Crystal Structures and Proton Conduction in Lanthanum Phosphates - A First Principles Study”
  Kazuaki Toyoura, Naoyuki Hatada, Yoshitaro Nose, and Tetsuya Uda,
  2011 MRS Fall Meeting & Exhibit, Boston, USA, November, 2011.
Trirubidium Hydrogen Bisulfate, $Rb_3H(SO_4)_2$

Paper

- “High Temperature Properties of $Rb_3H(SO_4)_2$ at Ambient Pressure: Absence of a Polymorphic, Superprotonic Transition”

Cesium Metaphosphate, $CsPO_3$

Paper

- “Phase Relationship of $CsH_2PO_4$-$CsPO_3$ System and Electrical Properties of $CsPO_3$”
  Yu-ki Taninouchi, Naoyuki Hatada, Tetsuya Uda, and Yasuhiro Awakura,

Oral Presentation

Domestic Conference in Japan

- “Electrical conductivity and phase stability of dehydrated liquid of $CsH_2PO_4$”
  Yu-ki Taninouchi, Naoyuki Hatada, Tetsuya Uda, and Yasuhiro Awakura,
  The 74th Meeting of the Electrochemical Society of Japan, Noda (Chiba), March, 2007.
Poster Presentation

Domestic Conference in Japan

• “Electrical Conductivity Mechanism of Cesium Metaphosphate”

Naoyuki Hatada, Yu-ki Taninouchi, Tetsuya Uda, and Yasuhiro Awakura,

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