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Synthesis and photooxidation of oligodeoxynucleotides containing 5-dimethylaminocytosine as an efficient hole-trapping site in the positive-charge transfer through DNA duplex

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We have designed and synthesized DNA duplexes containing 5-dimethylaminocytosine (DMA C) to investigate the effects of C(5)-substituted cytosome bases on the transfer and trapping of positive charge (hole) in DNA duplexes. Fluorescence quenching experiments revealed that a DMA C base is more readily one-electron oxidized into a radical cation intermediate as compared with other natural nucleobases. Upon photoradiation of the duplexes containing DMA C, the photosensitizer injected hole migrated through the DNA bases and was trapped efficiently at the DMA C sites, where an enhanced oxidative strand cleavage occurred by hot piperidine treatment. The DMA C radical cation formed by hole transfer may undergo specific hydration and subsequent addition of molecular oxygen, thereby leading to its decomposition followed by a predominant strand cleavage at the DMA C site. This remarkable property suggests that the modified cytosome DMA C can function as an efficient hole-trapping site in the positive-charge transfer in DNA duplex.

Introduction

Photosensitized oxidation of DNA has been studied extensively in relation to the positive-charge (hole) transfer through DNA. 1-3 In general, photosensitized one-electron oxidation of DNA can form a primary base radical cation, the positive charge of which migrates through the π-stack of DNA base pairs to produce oxidative cleavage of a DNA strand at a specific base with the lowest oxidation potential such as consecutive guanine (G) sites. 3 It has been demonstrated that a hole generated in DNA can migrate over a 20-30 nm range under certain conditions. 1,4 Hole transfer through DNA has attracted particular interest in view of potential application of its properties to gene analysis of mutation and molecular wires in nanoscale electronic devices 5-6. A wide variety of modified nucleobases have been developed for investigating and promoting the hole transfer reaction in DNA duplexes. 7-15 Recent studies on hole transfer in the photosensitizer-tethered DNA duplexes showed that introduction of modified purine bases into DNA can dramatically promote the photosensitizer injected hole transfer efficiency. 7-8 These strategies provide useful information for better mechanistic understanding as well as application to DNA-based nanomaterials of the hole-transfer. On the other hand, modified purine bases such as 7-deazaguanine, 9 8-methoxyguanine, 10 and 2-amino-7-deazaadenine 11 can thermally trap the hole that was injected and migrated through the DNA duplex. These modified nucleobases have also been used as effective tools for investigating long-range hole transfer through DNA duplex. Other examples are N-cyclopropyl-modified purine bases that trap kinetically the hole migrating along the DNA duplex. 12-13 In contrast to various types of modified purine bases reported so far, there have been limited examples of chemically modified pyrimidine bases possessing a hole trapping ability, e.g., N(4)-cyclopropyl-modified cytosine that kinetically traps holes migrating along the DNA duplex 14 and trimethoxyphenyl-substituted uracil that thermally traps holes induced by ionizing radiation. 15 In this context, development of an efficient hole-trapping pyrimidine base is an intriguing area of research for further understanding of a mechanism by which hole migrates through a duplex of DNA containing diverse sequences.

In this study, we have developed DNA duplexes containing 5-dimethylaminocytosine (DMA C) as a novel electron-donating nucleobase and characterized the hole-transfer reactions in DNA possessing modified base sites of DMA C. Fluorescence quenching experiments suggested that incorporation of the dimethylaminocytosine group into the C(5) position of cytosine significantly decreases the oxidation potential of DMA C because of a strong electron-donating effect of a dimethylamino group. Gel electrophoresis analysis revealed that a photoinjected hole migrates through the DNA bases to be trapped efficiently at the DMA C site, where enhanced oxidative strand cleavage was induced. Thus, DMA C can function as an effective hole-trapping site, thereby probing the hole transfer reaction in DNA.

Results and discussion

Electron transfer oxidation property of DMA C

We evaluated the one-electron oxidation property of DMA C using a Stern-Volmer analysis of the fluorescence quenching by monomeric DMA C (5-dimethylamo-2’-deoxyctydine, d DMA C),
which was prepared from 5-bromo-2’-deoxyuridine with dimethyamine (see ESI).\textsuperscript{1} 9,10-Dicyanoanthracene (DCA) was employed as a photooxidizing fluorophore for this analysis. It has sufficient reduction potential (1.95 V versus NHE in the singlet excited state) to oxidize all of the naturally occurring nucleobases.\textsuperscript{16a} Similar analysis has been widely used for investigation of the one-electron oxidation process of DNA bases.\textsuperscript{16b}

Upon electronic excitation of DCA at 390 nm in deoxygenated buffer solution, the intensity of the fluorescence emission decreased with increasing concentrations of D\textsuperscript{MAMC}. The fluorescence quenching rate constant (k\textsubscript{q}) of the singlet excited state, 1\textsuperscript{DCA}\textsuperscript{*}, was determined by the Stern–Volmer plots of the fluorescence spectral data (see Fig. S1)\textsuperscript{17} as summarized in Table 1. D\textsuperscript{MAMC} showed the highest k\textsubscript{q} value of 2’-deoxyribonucleosides evaluated: the k\textsubscript{q} values of 2’-deoxycytidine (dC), thymidine (dT), 2’-deoxyadenosine (dA), and 2’-deoxyguanosine (dG) were 2.9, 3.2, 4.6, and 5.6 \times 10\textsuperscript{10} M\textsuperscript{-1} s\textsuperscript{-1}, respectively. These results strongly suggest that D\textsuperscript{MAMC} can function as an electron-donating site, where a hole (radical cation) migrating through DNA is localized.\textsuperscript{18}

Table 1 Fluorescence quenching rate constants (k\textsubscript{q}) of DCA with 2’-deoxyribonucleosides\textsuperscript{a}

<table>
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<th>2’-Deoxyribonucleosides</th>
<th>k\textsubscript{q} (× 10\textsuperscript{10} M\textsuperscript{-1} s\textsuperscript{-1})</th>
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<td>dC</td>
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\textsuperscript{a} The excitation wavelength for DCA (25 μM) was 390 nm. Relative intensity of the emission band at 487 nm was measured in the presence of 25 μM DCA, in deoxygenated solution of 20 mM phosphate buffer (pH 7.0).\textsuperscript{19} The dynamic quenching rate constant k\textsubscript{q} was determined by the Stern-Volmer equation: k\textsubscript{q}/I = 1 + k\textsubscript{q}q[I/2’-deoxyribonucleoside], where k\textsubscript{q}/I is the ratio of the emission intensities in the absence and presence of 2’-deoxyribonucleoside and I is the lifetime of the singlet excited state of DCA (15.1 ns)\textsuperscript{20} in the absence of quencher.

Synthesis and thermal stability of \textsuperscript{DAMC}-containing ODNs duplex

Scheme 1 shows the synthesis of oligodeoxynucleotides (ODNs) containing \textsuperscript{DAMC}. 5-Bromo-2’-deoxyuridine protected by a 4,4’-dimethoxytrityl group (1) was converted to 2 by treatment with 50% dimethyamine. The 3’-hydroxyl group of 2 was then protected to afford 5-methylinodino-2’-deoxyuridine derivative (3). The 4-oxo group of the resulting 3 was transformed to an amino group to give the \textsuperscript{DAMC} derivative (4) in two steps. N-Acetyl-protected \textsuperscript{DAMC} (6) was prepared from 4 by standard acetyl protection of the exocyclic amine and subsequent desilylation of the hydroxy group, then being converted into \textsuperscript{DAMC} phosphoramidite derivative (7) that was incorporated into DNA using a DNA synthesizer and conventional phosphoramidite chemistry. Our choice of the photosensitizer was 9,10-anthraquinone (AQ), which is well known to induce hole injection to the adjacent DNA bases and thereafter positive-charge transfer through the DNA duplex.\textsuperscript{18} ODNs with AQ sensitizer were similarly prepared from a prescribed phosphoramidite derivative as in previously reported methods.\textsuperscript{18}

The ODNs used in this study are summarized in Fig. 1a.

Fig. 1 (a) Sequences and structures of ODNs used in this study. \textsuperscript{DAMC}, \textsuperscript{A\textsubscript{aq}}, \textsuperscript{A\textsubscript{aq}}, \textsuperscript{C\textsubscript{aq}}, and I denote 5-dimethylaminocytosine, 5-aminocytosine, 5-methylcytosine, 5-bromocytosine, and hypoxanthine, respectively. (b) Normalized UV melting curve of the duplex (1 μM strand concentration) measured at 260 nm in 10 mM sodium cacodylate buffer (pH 7.0) containing 100 mM NaCl: \textsuperscript{AQ}-ODN1(G)/ODN1(\textsuperscript{DAMC}) (solid line), \textsuperscript{AQ}-ODN1(G)/ODN1(\textsuperscript{DAMC}) (dotted line), and \textsuperscript{AQ}-ODN1(\textsuperscript{DAMC}) (dashed line). (c) CD spectra (5 μM strand concentration) of AQ-ODN1(G)/ODN1(\textsuperscript{DAMC}) (solid line) and AQ-ODN1(G)/ODN1(\textsuperscript{DAMC}) (dashed line) observed at 25 °C in 10 mM sodium cacodylate buffer (pH 7.0) containing 100 mM NaCl. The thermal stability of \textsuperscript{DAMC}-containing duplexes was evaluated by monitoring the melting temperatures (T\textsubscript{m}) in 10 mM sodium cacodylate buffer (pH 7.0) containing 100 mM NaCl (Fig. 1b). Introduction of the AQ moiety slightly enhanced the thermal stability of the duplex because of the π-stacking effect between the hydrophobic planar ring of AQ and the flanking bases. The T\textsubscript{m} of AQ-ODN1(G)/ODN1(\textsuperscript{DAMC}) duplex (T\textsubscript{m} = 57.0 °C) was only slightly different from that of a reference duplex AQ-ODN1(G)/ODN1(C) (T\textsubscript{m} = 56.9 °C). The circular dichroism (CD) spectra for each duplex showed a positive peak at 274 nm and a negative peak at 249 nm (Fig. 1c), indicating that the global structure of the AQ-ODN1(G)/ODN1(\textsuperscript{DAMC}) duplex was retained as a characteristic B-form. These results suggest that the dimethylamino group of \textsuperscript{DAMC} oriented toward the major groove...
does not significantly perturb the duplex structure.

**Hole trapping at DMA\textsubscript{C} in the positive-charge transfer through DNA duplex**

We also investigated the hole-transfer and hole-trapping properties of AQ-tethered DNA duplexes possessing a DMA\textsubscript{C} base in the middle spot of the sequence. Hole injection into the duplex was induced upon photoexcitation of AQ at the 5′-end of duplex. The duplex consisting of AQ-ODN1(G) with \textsuperscript{32}P-labeled ODN1(DMA\textsubscript{C}) in sodium cacodylate buffer solution (pH = 7.0) was photoirradiated (365 nm) at 20 °C. The photoirradiated samples were treated with hot piperidine and subjected to polyacrylamide gel electrophoresis (PAGE). The results are shown in Figs. 2a and 2b.

A predominant strand cleavage at the DMA\textsubscript{C} site was observed in the ODN1(DMA\textsubscript{C})/AQ-ODN1(G) duplex, whereas relatively weak strand cleavage occurred at the 5′-G\textsubscript{14} site of a G\textsubscript{14} doublet, which is well known as an efficient hole-trapping site \(^{(3)}\) (Fig. 2a, lanes 2–4). The cleavage yields at the DMA\textsubscript{C} site and G\textsubscript{14} site were 21% and 10%, respectively, after 20-min photoirradiation (Fig. 2b). No strand cleavage band was observed at the DMA\textsubscript{C} site in a control duplex without an AQ (Fig. 2a, lanes 6–9), indicating that strand cleavage at the DMA\textsubscript{C} due to possible direct photolysis is negligible under the present conditions. \(^{(19)}\) In the absence of piperidine treatment after photoirradiation, the strand cleavage was significantly suppressed (Fig. 2a, lane 5). These results clearly indicate that a given DNA base radical cation (hole) induced by the excited AQ migrates through the DNA duplex to be trapped efficiently at the DMA\textsubscript{C} base rather than the 5′-G site of a G-doublet. Similarly, separate experiments using the ODN2(X\textsubscript{1}/X\textsubscript{2}/X\textsubscript{3})/AQ-ODN2(Y\textsubscript{1}/Y\textsubscript{2}/Y\textsubscript{3}) duplexes showed that the cleavage efficiency of the DMA\textsubscript{C} sites in ODN2(T\textsubscript{DMA\textsubscript{C}}/T)/AQ-ODN2(A/G/A) were significantly higher...
than that of the corresponding 5'-G site in ODN2(T/G/G)/AQ-ODN2(C/C/A) (see Fig. S2) and that of the middle of the G triplet in ODN2(G/G/G)/AQ-ODN2(C/C/C) (see Fig. S3) aligned at the same distance from AQ.

It has been suggested that the hole transfer through DNA involves a short-range charge-transfer process between G or A bases, which are known to have the highest HOMO (highest occupied molecular orbital) energy levels among the naturally occurring nucleobases. In this context, competitive localization of the migrating hole at the complementary base is possibly involved in the hole trapping at DMA C. In addition, Kawai and co-workers have reported that the oxidation potential of G can be regulated by introducing a substituent to cytosine on the complementary strand. In these views, we examine the effects of complementary bases on the strand cleavage efficiency at DMA C. Hypoxanthine (I), which has a higher oxidation potential (E_{ox} = 1.50 V versus NHE) than that of G (E_{ox} = 1.29 V versus NHE), was introduced into the complementary strand (AQ-ODN1(I)). The cleavage efficiency at DMA C was compared relative to the AQ-ODN1(G) duplex. As shown in Fig. 2c, cleavage efficiencies at DMA C sites in AQ-ODN1(I)/ODN1(AQ) were virtually identical to that in the AQ-ODN1(G)/ODN1(AQ), indicating that the oxidation potentials of complementary bases have little effect on the strand cleavage at DMA C. In light of the Stern–Volmer analysis data (Table 1), this suggests that a photoinjected and migrated hole may be localized at the DMA C site rather than its complementary base, which further supports an aspect that DMA C can function as an efficient electron donor for the hole transfer through DNA duplexes.

LC/ESI-TOF mass analysis in the one-electron photooxidation of DMA C to get a mechanistic insight into photooxidation and strand cleavage of DMA C

For a mechanistic understanding of the one-electron oxidation and piperidine-induced strand cleavage at DMA C, we carried out LC/ESI-TOF mass analysis in the one-electron oxidation of DMA C induced by photoexcited riboflavin, which has been reported to predominantly sensitize a Type I photooxidation. Aerobic solutions of DMA C (200 µM) and riboflavin (200 µM) in 2 mM sodium cacodylate buffer (pH 7.0) containing 10 mM NaCl were photoirradiated with 365 nm UV light, and the reaction was analyzed by LC/ESI-MS. The results are shown in Fig. 3. After 2 min irradiation, the HPLC profile exhibited two major peaks at 1.9 and 3.4 min along with the degradation of DMA C (Fig. 3a). Upon prolonged photoirradiation (ca. 10 min), DMA C was completely degraded, while the peak intensity at 3.4 min decreased with an increase in the peak at 1.9 min and was substantially disappeared. The ESI–MS analyses of these characteristic LC peaks indicates transient formation of unstable N-formylurea derivative (dNfu) (m/z 271.1223, 309.0989; calcd. for [M-H] = 301.1172; [M+H] = 301.1148) (Fig. 3b, bottom) and its intramolecularly cyclized imidazolone derivative (dImz) (m/z 271.1223, 309.0989; calcd. for [M-H] = 273.1199, [M+H] = 309.0971) (Fig. 3b, top).

Scheme 2 shows a plausible mechanism for one-electron oxidation and strand cleavage at DMA C. The primary step involves one-electron oxidation of DMA C into the corresponding radical cation (DMA C•+) intermediate by way of AQ-injected hole transfer through DNA and subsequent hole trapping at the DMA C site. By reference to the proposed reaction mechanisms for one-electron oxidation of DMA C, the hole transfer process involves the oxidation of AQ, the radical cation DMA C•+, and the transient formation of N-formylurea (dNfu) and imidazolone (dImz) derivatives. The oxidation process is followed by the opening of the 1,6-pyrimidine ring of the resulting Nfu to afford the C(5) deoxyribofuranosyl-5-hydroxy-5-methylhydantoin in the hydrolytic decomposition of 6-hydroxy-5-hydroperoxy-5,6-dihydroxyimidazoline (one of the oxidation products of dT) through the opening of the 1,6-pyrimidinone ring. In this light, the resulting C(5)-hydroperoxide intermediate of DMA C may undergo hydrolytic decomposition to give imidazoline analogue (Imz), a related degradation product to 5-hydroxy-5-methylhydantoin, as in the case of dT photooxidation. Another pathway could also be considered: N-formylurea analogue (Nfu) are likely to be generated by the disproportionation of C(5)-peroxyl radical intermediate of DMA C. The resulting Nfu may undergo hydrolysis of the formyl group and intramolecular cyclization to give Imz analogue. However, the occurrence of the bi-molecular
processes may be ruled out in the AQ-photosensitized oxidation of the DMA_C site in the restricted structure of the DNA duplex.

An attempt was also made to investigate the photooxidation of AQ-ODN1(G)/ODN1(DMA_C) duplex using ESI-TOF mass spectroscopy. The duplex (10 µM) in sodium cacodylate buffer solution (pH = 7.0) was photoirradiated (365 nm) at 20 °C for 20 min, and then subjected to ESI-TOF mass analysis. As shown in Fig. S4†, the mass spectral peak of m/z 1554.5 was observed, indicating the formation of oxidized ODN1(Imz) (calcd. for [M–4H]− = 1554.7). Irrespective of the reaction mechanisms, it is likely that a hole-trapped intermediate DMA_C•+ may be converted into an alkali-labile final product such as Imz analogue, via the hydrolytic decomposition of C(5)-hydroperoxide intermediate and/or the intramolecular cyclization of a transient species Nfu, resulting in a predominant oxidative strand cleavage at the DMA_C site. Further detailed product analysis of the hole trapping reaction of DMA_C in single-stranded DNA would be necessary for better understanding of the reactivity and the detailed reaction mechanisms.

C(5)-substituent effects on the hole-trapping efficiency

We examined the electronic substituent effects on the hole-trapping ability of C(5)-modified cytosines. As an electronic substituent incorporated into cytosine at C(5), we focused on four typical types of dimethylamino, amino, methyl, and bromo groups. The ODN containing 5-aminocytosine (AmC) was synthesized from the corresponding ODN bearing 5-bromocytosine (BrC) according to previously reported methods. Duplexes consisting of AQ-ODN1(G) with 32P-labeled ODN1(X) (X = DMA_C, AmC, 5-methycytosine (McC), normal C, and BrC) were photoirradiated, treated with hot piperidine, and then subjected to PAGE under the conditions described above. As shown in Fig. 4, strand cleavage at the G14 site was observed in each photoirradiated ODN, whereas the band intensities in ODN1(DMA_C) and ODN1(AmC) were slightly lower than those observed in other ODNs. Intense strand cleavages at DMA_C and AmC in ODN1(DMA_C) and ODN1(McC), respectively, were observed after 20-min irradiation; the corresponding cleavage band intensities were of a background level at the BrC in ODN1(BrC), normal C in ODN1(C), and BrC in ODN1(BrC) (Fig. 4, lanes 5, 9, 13, 17, and 21). These results clearly indicate that DMA_C as well as AmC can efficiently trap the AQ-photoinjected holes in the positive-charge transfer through DNA duplexes. Incorporation of a strongly electron-donating group such as dimethylamino and amino groups, but not the methyl group, at the C(5) position of cytosine may substantially reduce oxidation of ODN1(DMA_C) and ODN1(AmC) were slightly lower than those observed in other ODNs. Intense strand cleavages at DMA_C and AmC in ODN1(DMA_C) and ODN1(McC), respectively, were observed after 20-min irradiation; the corresponding cleavage band intensities were of a background level at the BrC in ODN1(BrC), normal C in ODN1(C), and BrC in ODN1(BrC) (Fig. 4, lanes 5, 9, 13, 17, and 21). These results clearly indicate that DMA_C as well as AmC can efficiently trap the AQ-photoinjected holes in the positive-charge transfer through DNA duplexes. Incorporation of a strongly electron-donating group such as dimethylamino and amino groups, but not the methyl group, at the C(5) position of cytosine may substantially reduce oxidation of ODN1(DMA_C) and ODN1(AmC) were slightly lower than those observed in other ODNs. Intense strand cleavages at DMA_C and AmC in ODN1(DMA_C) and ODN1(McC), respectively, were observed after 20-min irradiation; the corresponding cleavage band intensities were of a background level at the BrC in ODN1(BrC), normal C in ODN1(C), and BrC in ODN1(BrC) (Fig. 4, lanes 5, 9, 13, 17, and 21). These results clearly indicate that DMA_C as well as AmC can efficiently trap the AQ-photoinjected holes in the positive-charge transfer through DNA duplexes. Incorporation of a strongly electron-donating group such as dimethylamino and amino groups, but not the methyl group, at the C(5) position of cytosine may substantially reduce oxidation of ODN1(DMA_C) and ODN1(AmC) were slightly lower than those observed in other ODNs. Intense strand cleavages at DMA_C and AmC in ODN1(DMA_C) and ODN1(McC), respectively, were observed after 20-min irradiation; the corresponding cleavage band intensities were of a background level at the BrC in ODN1(BrC), normal C in ODN1(C), and BrC in ODN1(BrC) (Fig. 4, lanes 5, 9, 13, 17, and 21). These results clearly indicate that DMA_C as well as AmC can efficiently trap the AQ-photoinjected holes in the positive-charge transfer through DNA duplexes. Incorporation of a strongly electron-donating group such as dimethylamino and amino groups, but not the methyl group, at the C(5) position of cytosine may substantially reduce oxidation of ODN1(DMA_C) and ODN1(AmC) were slightly lower than those observed in other ODNs. Intense strand cleavages at DMA_C and AmC in ODN1(DMA_C) and ODN1(McC), respectively, were observed after 20-min irradiation; the corresponding cleavage band intensities were of a background level at the BrC in ODN1(BrC), normal C in ODN1(C), and BrC in ODN1(BrC) (Fig. 4, lanes 5, 9, 13, 17, and 21). These results clearly indicate that DMA_C as well as AmC can efficiently trap the AQ-photoinjected holes in the positive-charge transfer through DNA duplexes. Incorporation of a strongly electron-donating group such as dimethylamino and amino groups, but not the methyl group, at the C(5) position of cytosine may substantially reduce oxidation.
the DNA synthesizer were purchased from Glen Research. The reagents for experimental purposes were purchased from standard suppliers and used without further purification. The reagents for the DNA synthesizer were purchased from Glen Research.

Complementary ODNs were purchased from Invitrogen unless otherwise noted. Monomeric DMA-C (dDMA-C) was prepared from 5-bromo-2'-deoxyuridine as detailed in ES1. The NMR spectra were obtained on a JEOL JMN-AL300 (300 MHz), a JEOL JMN-EX-400 (400 MHz), or a JEOL JMN-ECX-400 (400 MHz) spectrometer. J values are given in Hz. The FAB mass spectra were recorded on a JEOL JMS-SX102A spectrometer, using a nitrobenzyl alcohol or glycerol matrix. Matrix-assisted laser desorption ionization time-of-flight (MALDI–TOF) mass spectrometry of ODNs was performed on a JEOL LMS-ELITE MALDI–TOF mass spectrometer with 2,3,4,5-tetradeuterobenzoyloxyacetonitrile as the matrix. ESI–TOF mass spectra were obtained on a Bruker Daltonics microTOF focus–KE.

5'-[(4,4′-Dimethoxytrityl)-5-methylaminol-2'-deoxyuridine (2).

5'-(Dimethoxytrityl)-5-bromo-2'-deoxyuridine (1) was prepared as described elsewhere. A solution of 1 (3.66 g, 6.00 mmol) in 10 mL of 50% aqueous dimethylamine was sealed in a 10 mL vial and heated for 60 h at 60 °C. The residue was evaporated under reduced pressure. The crude product was purified by column chromatography (SiO2, 100% ethyl acetate) to give 2 (1.79 g, 52%) as a white foam. 1H NMR (CDCl3, 400 MHz) δ 8.34 (s, 1H), 7.34–7.09 (9H), 6.76–6.75 (4H), 6.30 (t, 1H, J = 6.6 Hz), 4.24 (t, 1H, J = 2.9 Hz), 3.97 (d, 1H, J = 3.4 Hz), 3.71 (s, 6H), 3.36 (dd, 1H, J = 4.4, 10.4 Hz), 3.28 (dd, 1H, J = 3.2, 10 Hz), 2.34 (s, 6H), 2.29–2.14 (2H); 13C NMR (CDCl3, 400 MHz) δ 158.8, 149.0, 144.4, 139.5, 131.0, 130.2, 129.2, 128.9, 128.2, 127.9, 127.1, 113.3, 113.2, 85.5, 84.6, 72.6, 63.6, 55.3, 42.2, 38.8; FABMS m/z 574 [(M+H)+]; HRMS calcd. for C32H32N2O7Si 574.2553, found 574.2557.

3'-O-(tert-Butyldimethylsilyl)-5'-O-(4,4′-dimethoxytrityl)-5-methylaminol-2'-deoxyuridine (3).

A solution of 2 (1.70 g, 3.00 mmol), tert-butyldimethylsilylchloride (900 mg, 5.97 mmol), and imidazole (820 mg, 12.0 mmol) in DMF (7 mL) was stirred for 6 h at room temperature. The reaction mixture was quenched by water and extracted with ethyl acetate. The organic layer was washed with brine, dried over Na2SO4, filtered, and concentrated. The crude product was purified by column chromatography (SiO2, 100% ethyl acetate) to give 3 (1.95 mg, 94%) as a white solid: 1H NMR (CDCl3, 400 MHz) δ 8.45 (s, 1H), 7.33–7.13 (9H), 6.74–6.71 (4H), 6.26 (dd, 1H, J = 5.9, 6.0 Hz), 4.34 (3H, J = J = 3.2 Hz), 3.88 (dd, 1H, J = 3.2, 6.8 Hz), 3.69 (s, 6H), 3.32 (dd, 1H, J = 3.4, 10.2 Hz), 3.12 (dd, 1H, J = 3.8, 10.6 Hz), 2.35 (s, 6H), 2.20–2.15 (1H), 2.05–1.98 (1H), 0.77 (s, 9H), –0.04 (6H); 13C NMR (CDCl3, 400 MHz) δ 158.6, 149.0, 144.4, 135.5, 130.0, 129.1, 128.1, 127.9, 127.0, 113.1, 81.6, 72.2, 62.9, 55.2, 42.3, 40.7, 25.6, 17.9, –4.7, –4.9; FABMS m/z 688 [(M+H)+]; HRMS calcd. for C38H40N4O8Si 688.3418, found 688.3412.

3'-O-(tert-Butyldimethylsilyl)-5'-O-(4,4′-dimethoxytrityl)-5-methylaminol-2'-deoxyuridine (4).

A solution of 3 (1.94 g, 2.82 mmol), triethylamine (1.2 mL, 8.62 mmol), 4-dimethylaminopyridine (1.04 g, 8.46 mmol), and 2,6-trisopropylbenzensulfonyl chloride (2.56 g, 8.46 mmol) in anhydrous acetonitrile (24 mL) was stirred at room temperature for 16 h. The mixture was cooled in an ice bath. Aqueous ammonia (28%, 35 mL) was added, and the mixture was stirred for 3 h at ambient temperature. The reaction mixture was concentrated under reduced pressure and was taken up in ethyl acetate. The organic layer was washed with brine, dried over Na2SO4, filtered, and concentrated. The crude product was purified by column chromatography (SiO2, 4% methanol–chloroform) to give 4 (1.87 g, 97%) as a white solid: 1H NMR (CDCl3, 400 MHz) δ 7.55 (s, 1H), 7.54–7.23 (9H), 6.92–6.89 (4H), 6.46 (4H), 6.46 (dd, 1H, J = 6.6 Hz), 7.9 Hz), 4.05 (dd, 1H, J = 3.0, 7.0 Hz), 3.87 (s, 6H), 3.57 (dd, 1H, J = 3.2, 10.8 Hz), 3.26 (dd, 1H, J = 3.2, 10.8 Hz), 2.53–2.49 (1H), 2.38 (6H), 2.17–2.10 (1H), 0.89 (s, 9H), 0.03 (6H); 13C NMR (CDCl3, 400 MHz) δ 163.4, 158.6, 144.4, 135.7, 135.5, 130.1, 129.8, 128.2, 127.9, 127.0, 113.2, 113.1, 86.3, 85.9, 72.0, 62.7, 55.2, 44.2, 42.0, 25.7, 17.9, –4.7, –5.0; FABMS m/z 687 [(M+H)+]; HRMS calcd. for C32H32N2O7Si 687.3578, found 687.3580.

3'-O-(tert-Butyldimethylsilyl)-5'-O-(4,4′-dimethoxytrityl)-N-acetyl-5-methylaminol-2'-deoxyuridine (5).

A solution of 4 (1.87 g, 2.73 mmol), acetic anhydride (645 μL,
5′-O-(4,4′-Dimethoxytrityl)-N-acetyl-5-methylamino-2′-deoxycytidine (6).

Tert-butyl ammonium fluoride (1.0 M, 3.5 mL, 3.5 mmol) was added to a solution of 5 (1.28 g, 1.76 mmol) in dry THF (24 mL) and the mixture was stirred at room temperature for 1 h. The reaction mixture was diluted with water and extracted with ethyl acetate. The organic layer was washed with brine, dried over Na₂SO₄, and concentrated. The crude product was purified by column chromatography (SiO₂; 50% ethyl acetate–hexane) to give 6 (1.29 g, 65%) as a yellow solid. ¹H NMR (CDCl₃, 400 MHz) δ 7.82 (s, 1H), 7.43–7.24 (9H), 6.84–6.87 (4H), 6.26 (t, 1H, J = 6.6 Hz), 4.40 (dd, 1H, J = 3.4, 7.0 Hz), 4.02 (dd, 1H, J = 3.6, 6.8 Hz), 3.79 (s, 6H), 3.54 (dd, 1H, J = 3.2, 11.6 Hz), 3.22 (dd, 1H, J = 3.4, 10.6 Hz), 2.71 (s, 3H), 2.58–2.53 (1H), 2.29 (s, 6H), 2.12–2.05 (1H), 0.82 (s, 9H), –0.03 (6H); ¹³C NMR (CDCl₃, 400 MHz) δ 158.7, 144.3, 135.5, 135.3, 130.1, 130.0, 129.1, 128.2, 127.9, 127.1, 113.2, 113.1, 86.7, 86.4, 71.7, 62.4, 44.8, 42.1, 25.7, 17.9, 4.7, –5.0; FABMS m/z 729 [(M+H)+]; HRMS calcd. for C₁₃H₇₅N₁₇O₂₇Si 729.3684, found 729.3701.

Fluorescence quenching.

Quenching experiments of the fluorescence of the photosensitizer were carried out on a Shimadzu RF-5300PC spectrophotometer. The excitation wavelength for DCA (25 µM) was 390 nm and the monitoring wavelength was that corresponding to the respective emission band at 487 nm. The dynamic quenching rate constant kₚ was determined by the Stern–Volmer equation: I₀/I = 1 + kₚq[l/2-deoxyribonucleoside] where I₀/I is the ratio of the emission intensities in the absence and presence of 2′-deoxyribonucleoside and τ₀ is the lifetime of the singlet excited state of DCA (15.1 ns) in the absence of quencher.

Photooxidative cleavage reaction and PAGE analysis.

ODNs were 5′-³²P-labeled by phosphorylation with 4 µL of [γ-³²P]ATP (PerkinElmer) and 4 µL of T4 polynucleotide kinase (Nippon Gene). The reaction mixtures were purified using QIAquick Nucleotide Removal Kit (QIAGEN) to remove excess unincorporated nucleotide. Complementary ODNs (1 µM) were annealed in 10 mM sodium cacodylate buffer (pH 7.0) containing 100 mM NaCl by heating to 90 °C, followed by slow cooling to room temperature. The samples were irradiated at 365 nm UV light with a transilluminator (FTI-20L, Funakoshi, Tokyo) at 20 °C on exposure to air. After irradiation, the DNA samples were precipitated by adding 10 µL of herring sperm DNA (1 mg/mL), 10 µL of 3 M sodium acetate and 800 µL of ethanol. The precipitated DNA was resolved in 50 µL of 10% piperidine (v/v), heated at 90 °C for 20 min, and concentrated. The radioactivity of samples was assayed using an Aloka 1000 liquid scintillation counter and the dried DNA pellets were resuspended in loading buffer (a solution of 80% formamide (v/v), 1 mM EDTA, 0.1% xylene cyanol, and 0.1% bromophenol blue). All reactions, along with Maxam–Gilbert G+A sequencing reactions, were heat denatured at 90 °C for 3 min and quickly chilled on ice. The samples (3–10 × 10⁷ cpm) were loaded onto 20% polyacrylamide/7 M urea sequencing gels and electrophoresed at 1900 V for 60–90 min, transferred to a cassette, and stored at –80 °C with Fuji X-ray film (RX-U). The gels were analyzed using autoradiography with ATTO CS Analyzer (version 3.0). Individual yields were calculated relative to each total band intensity. The average of three independent measurements for each sample is indicated.
LC/ESI–MS analysis of one-electron photooxidation of 5-dimethylamino-2-deoxyxysteinylsensitized riboflavin. A solution of 5-dimethylamino-2-deoxyxysteine (dODMA) (200 μM) and riboflavin (200 μM) in 2 mM sodium cacodylate buffer (pH 7.0) containing 20 mM NaCl (total volume 100 μL) was exposed to 365 nm UV light at 0 °C and subjected to LC/ESI–MS analysis. LC/ESI–MS (negative mode) was performed with a Thermo Exactive equipped with Thermo Accela LC system. The column eluents were monitored by UV absorbance at 260 nm. The solvent mixture of triethylammonium acetate buffer solution (0.1 M, pH 7.0) containing 5 vol% acetonitrile was delivered as the mobile phase.

Melting temperature (Tm) of hybridized ODNs. 1 μM of appropriate ODNs was dissolved in 10 mM sodium cacodylate buffer (pH 7.0) containing 100 mM NaCl. UV melting curves were recorded on a JASCO-V630 spectrophotometer equipped with a multi-cell block and a Peltier temperature controller. Melting curves were obtained by monitoring the UV absorbance at 260 nm with the temperature being increased from 4 °C to 80 °C at a rate of 1 °C min⁻¹.

CD spectrum of hybridized ODNs. 5 μM duplexes were dissolved in 10 mM sodium cacodylate buffer (pH 7.0) containing 100 mM NaCl. CD spectra of the solutions were recorded at 20 °C on a JASCO J-700 spectrophotometer, using a UV cell with 0.1 cm path length.

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Notes and references

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† Electronic Supplementary Information (ESI) available; [synthesis, Stem–Volmer plots of fluorescence quenching, PAGE analysis, and ESI-TOF mass analysis]. See DOI: 10.1039/b000000x
15. Monomeric dODMA has the absorption band at wavelengths shorter than 350 nm. The photoirradiation at 365 nm can selectively excite AQ sensitizer.
20. Although we have not yet accomplished full assignment in the absence of comprehensive structural information on the photooxidation products by means of NMR spectroscopy, dImz may exist as the 5R* and 5S* forms, as in the case of 2-deoxyribofuranosyl-5-hydroxy-5-methylhydantoin, a related degradation product to dImz. 25
21. We could not detect other photooxidation products under the present LC conditions. At the present stage, however, another possibility cannot be ruled out that other minor degradation products could be formed.
24. In addition to predominant hydration at C(6) of dODMA-C*, the minor pathway could be considered: the 76dODMA-C* may undergo competitive deprotonation of the exocyclic amine group. 25 In this context, other piperidine-stable oxidation products may be produced along with the formation of piperidine-labile Imz.
The Imz site would be piperidine-labile in view of the fact that similar imidazolone analogs have been shown to be alkali-sensitive, resulting in strand cleavage at these sites. See K. Kino, I. Saito and H. Sugiyama, *J. Am. Chem. Soc.*, 1998, **120**, 7373–7374.


The ionization potentials of DMA-C and Am-C were estimated to be 8.58 and 8.64 eV, respectively, using the previously reported methods. This indicates that the oxidation potential of DMA-C is not much different than that of Am-C. The ionization potentials of DMA-C and Am-C were estimated by ab initio calculations at the B3LYP/6-31G* level. Ab initio calculations were performed for DMA-C and Am-C derivatives in which the sugar units were replaced by methyl groups.
5-Dimethylaminocytosine (DMAc) can function as an efficient hole-trapping site in the anthraquinone (AQ) photosensitizer-injected positive-charge transfer though DNA duplex.