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DFT study of CO oxidation over Au/TiO$_2$(110): The extent of the reactive perimeter zone

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Keywords

Au catalyst, rutile TiO$_2$, O$_2$ adsorption, CO oxidation, DFT

Abstract

O$_2$ adsorption and CO oxidation on a Au/rutile TiO$_2$(110) catalyst have been examined by DFT in order to define the extent of the reactive perimeter zone. O$_2$ is found to adsorb on the five-fold coordinated Ti sites that are away from the perimeter as well as on those next to the perimeter. Both O$_2$ show similar reactivity toward a gaseous CO. CO oxidation still occurs predominantly at the perimeter because O$_2$ adsorption is strongest at the perimeter Ti site and because CO is activated on Au through back donation.
1. Introduction

Gold was traditionally regarded as the unlikeliest candidate for oxidation catalysts because of its inability to adsorb and dissociate oxygen. This view has changed drastically since gold nanoparticles (NP) supported on oxides such as TiO$_2$ were discovered to exhibit extraordinary catalytic activity for low-temperature CO oxidation [1,2] and other useful reactions [3]. Numerous experimental and theoretical efforts have been made to find the origin of catalytic activity of gold catalysts [3], especially the location of active sites. It is generally viewed that CO oxidation occurs along the perimeter of Au/oxide interface (‘perimeter hypothesis’) [4]. For the most basic case of Au/TiO$_2$, this hypothesis has recently been confirmed by experiments showing the proportionality between the rate of oxidation and the length of perimeter [5,6].

However, it is still unclear whether the reactive zone is confined to the immediate vicinity of the perimeter or includes the nearby TiO$_2$ surface. The exact determination of the active sites requires quantum-chemical approaches such as the density-functional theory (DFT) [7,8]. Earlier DFT calculations have found that, although the stoichiometric rutile TiO$_2$(110) surface is unable to adsorb O$_2$ in the absence of Au NPs [9], the five-fold coordinated Ti (Ti$^{5c}$) sites next to the perimeter are able to adsorb and activate O$_2$ [9-12]. This is because the Ti cation lowers the antibonding $\pi$ orbital of O$_2$ below the Au Fermi level, enabling electron transfer from the Au to O$_2$, and binds with the resulting O$_2$ anion [9,10]. Because the ability to lower O$_2$ $\pi^*$ is innate to TiO$_2$ [10], Ti$^{5c}$ sites away from the perimeter may also be able to adsorb and activate O$_2$.

In this study, therefore, O$_2$ adsorption and CO oxidation on a Au/rutile TiO$_2$(110) catalyst are examined by DFT in order to determine the extent of the reactive perimeter zone. The remainder of this letter is organized as follows. Section 2 describes the details of calculation. Section 3.1 examines O$_2$ adsorption and shows that O$_2$ is strongly adsorbed and activated on Ti$^{5c}$ sites both next to and away from the perimeter. Section 3.2 examines CO oxidation and shows that these O$_2$ are almost equally reactive toward a gaseous CO, but that the energy barrier is smaller at the perimeter where O$_2$ can react with CO on Au. It is concluded in Section 4 that the reactive zone is confined to the immediate vicinity of the perimeter because O$_2$ adsorption is strongest at the perimeter Ti$^{5c}$ site and because CO is activated on Au.
2. Method

Total energies and optimized geometries were calculated by a plane-wave DFT code STATE (Simulation Tool for Atom TEchnology) [13]. This code has been applied to various systems including Au(111) and Au/TiO$_2$ [14-16]. The exchange-correlation functional by Perdew, Burke, and Ernzerhof was employed [17]. The cutoff energies were 25 and 225 Ry for the Kohn-Sham orbitals and charge density, respectively. The interactions between ionic cores and valence electrons were described by ultrasoft pseudopotentials [18], with 4, 6, 10, and 11 electrons taken as valence for C, O, Ti, and Au, respectively. Spin-unpolarized calculations were performed unless otherwise noted. Transition states (TS) were searched by the force inversion method [19], and then the initial (IS) and final (FS) states were found by relaxation from positions slightly displaced from TS. Charge analysis was done by the Bader method [20] as implemented by Henkelman et al. [21]. The Bader charge is the charge inside a volume defined by the surface on which the normal derivative of the electron density vanishes. The Bader charge does not necessarily match the formal charge state, for example, the Bader charge of O in bulk rutile is calculated to be $-1.1e$. Orbitals were visualized using VESTA (Visualization for Electronic and STructural Analysis) [22].

Fig. 1 shows the model of a Au/TiO$_2$(110) catalyst used in the calculations. The TiO$_2$ surface was represented by a periodic slab consisting of four rutile (110) trilayers. The top three trilayers were allowed to move while the bottom one was frozen during geometrical optimization. The periodic images of the slab were separated by 1.39 nm of vacuum. The stoichiometric surface was examined because oxygen vacancies would be readily repaired under oxygen-rich conditions typical with CO oxidation. The perimeter of a Au/TiO$_2$ catalyst was modeled by placing a Au rod on the rutile surface with the orientation of Au(111)[110]/TiO$_2$(110)[001]. This is the orientation found by TEM experiments, and the calculated distance between the Au NP and the Ti layer (0.34 nm) is consistent with the measured one (0.33 nm) [23]. The rod consists of three Au(111) layers, the top layer being two-atom wide and the bottom two layers three-atom wide. The Au atoms at the rim of the rod are in register with the bridging O. This model is the same as that used by Green et al. [12], but Au atoms were now allowed to move in three dimensions. Ti$^{6c}$ sites are distinguished by the order of proximity to the perimeter, such
as first, second, and third nearest sites (Fig. 1). Large 4×3 and 6×3 cells, sampled with a 1×2 k-point mesh, were used, where the first and second dimensions refer to TiO$_2$[110] and [001].

The stability of an adsorption geometry is expressed in terms of adsorption energy $E_{\text{ad}}$, defined as the total energy of the adsorption geometry minus the total energy of Au/TiO$_2$(110) minus the total energy of gaseous O$_2$ and/or CO. Here, a gaseous molecule was calculated using a 24×24×24 Bohr$^3$ cell (O$_2$ calculated with spin polarization). According to the above definition, the lower the adsorption energy, the more stable the adsorbate. The energy barrier of reaction was calculated as the total energy of the transition state minus the total energy of the initial state. For a reaction with a gaseous CO, the total energy of the initial state was calculated as the sum of the total energies of the O$_2$-adsorbed geometry and a gaseous CO.

Because periodic boundary conditions are used, interactions with the periodic images of the Au rod might result in the overestimation of O$_2$ adsorption energies. To examine this, we have performed several test calculations for O$_2$ adsorbed side-on to the nearest Ti$^{5c}$ site (see Section 3.1), varying the width of the cell. The adsorption energy, bond length, and Bader charge on O$_2$ are reported in Table 1. As can be seen here, the results are practically the same for 3×3 and wider cells, and the error is not very large even at 2×3. In addition, dependence on k-point mesh and slab thickness (examined using the 2×3 cell) is also small (Table 1).

3. Results and discussion

3.1 O$_2$ adsorption

We begin with confirming that O$_2$ cannot be adsorbed on the stoichiometric TiO$_2$ surface in the absence of the rod, using the 4×3 cell. Fig. 2a shows the equilibrium position for O$_2$ placed above the Ti$^{5c}$ site. Negligible adsorption energy (−0.04 eV) is obtained for this O$_2$. The O$_2$ bond length is practically unchanged from that (125.4 pm) in the gas phase. O$_2$ remains in the triplet state, with no charge transferred from the surface. These results are reasonable because, although the Ti cation may be able to lower O$_2$ π*, the transfer from the valence band of TiO$_2$ can only result in overall destabilization. Thus, without the Au rod, the stoichiometric surface is practically
unable to adsorb and activate $\text{O}_2$.

Next, we examine $\text{O}_2$ adsorption at the perimeter of the Au/TiO$_2$ interface, using the 4×3 cell. First, weak adsorption energy (−0.22 eV) is obtained for $\text{O}_2$ bridging two Au sites at the perimeter (Fig. 2b). Small negative charge and residual spin on $\text{O}_2$ indicate insufficient activation of $\text{O}_2$. In contrast, much stronger adsorption energy (−1.03 eV) is obtained for $\text{O}_2$ bridging Ti$^{5c}$ and Au at the perimeter (Fig. 2c), reproducing an earlier result (−1.01 eV) [12]. The longer O-O bond and negative charge on $\text{O}_2$ indicate that this $\text{O}_2$ is activated to a superoxide state [9]. Most of the charge (−0.51e) is on the lower O, drawn to the Ti cation. Even stronger adsorption energy (−1.74 eV) is obtained for $\text{O}_2$ adsorbed side-on to the Ti$^{5c}$ site next to the perimeter (Fig. 2d). Both O of the $\text{O}_2$ are now bonded to the Ti cation, enhancing the adsorption energy. Even longer O-O bond and more negative charge indicate that this $\text{O}_2$ is activated to a peroxide state [9]. (Negligible spin polarization was found for the superoxide and peroxide states.) In the following, we focus on the side-on $\text{O}_2$ because this is the more stable and further activated species.

We now examine how the $\text{O}_2$ adsorption energy changes with distance from the perimeter, using the 6×3 cell. Figs. 2e-2g show $\text{O}_2$ adsorbed side-on to the Ti$^{5c}$ site that is first, second, and third nearest to the perimeter, respectively. The adsorption is strongest at the first site (−1.70 eV), but the second (−1.43 eV) and third (−1.39 eV) sites are not much higher in energy. Moreover, similar $\text{O}_2$ bond lengths and Bader charges are found for the three sites. Thus, Ti$^{5c}$ sites away from the perimeter also have the ability to adsorb and activate $\text{O}_2$, and this ability appears to weaken very slowly beyond the second nearest site. (There should still be a limit to $\text{O}_2$ adsorption on remote sites because the probability of electron transfer will decrease with distance.)

To explain the above results, we have calculated projected density of states (PDOS) for the Au/TiO$_2$ surface before $\text{O}_2$ adsorption and the geometries of Figs. 2e-2g. Before $\text{O}_2$ adsorption, the PDOS of the three Ti$^{5c}$ sites are nearly indistinguishable (Fig. 3a). The Fermi level is located slightly below the empty 3d states of Ti$^{5c}$, indicating that the HOMO of the Au rod is close to the conduction band of TiO$_2$. This could also imply a small amount of electron transfer to TiO$_2$, but the amount would not be so large as to
cause significant model-size dependence (see Table 1) or artificial dipole interaction across the vacuum layer (dipole correction [24] for Au/TiO$_2$ in a 2×1 cell shows that its stability is overestimated by a mere 4 meV).

Fig. 3b shows PDOS for O$_2$ adsorption on the nearest Ti$^{5c}$ site. The O 2p levels of O$_2$ appear below the Fermi level, indicating electron transfer from the Au rod to O$_2$. The peak around −1.7 eV is from O 2p$_{z}$ and is identified as out-of-plane O$_2$ π*. The peak around −0.4 eV is from O 2p$_{x}$ and 2p$_{y}$ and is identified as in-plane O$_2$ π*. These states are strongly mixed with the 3d orbitals of Ti$^{5c}$, confirming that the Ti cation stabilizes O$_2$ π* [10]. Moreover, the lower peak also coincides with that of the three-fold coordinated O (O$^{3c}$) immediately below O$_2$, indicating that this O$_2$ π* is stabilized so much as to come close to the valence band of TiO$_2$. This is reasonable considering that they are both Ti-stabilized O 2p states. Because Au HOMO is close to the TiO$_2$ conduction band while at least one of O$_2$ π* is brought close to the TiO$_2$ valence band, electrons are transferred from Au to O$_2$ through the TiO$_2$ conduction band. O$_2$ thus becomes anionic and is bound strongly to the Ti cation. Au states, on the other hand, do not appear to mix strongly with O$_2$ π*, in fact, Au PDOS is mostly unchanged from that before O$_2$ adsorption (Fig. 3a).

PDOS for O$_2$ adsorption on the second and third nearest Ti$^{5c}$ sites (Figs. 3c and 3d, respectively) are similar to that for the nearest site (Fig. 3b). Again, strong interaction with the Ti cation lowers the O 2p levels of O$_2$ below the Fermi level, causing electron transfer from the Au rod to O$_2$. The out-of-plane O$_2$ π* also mixes with O$^{3c}$ p$_{z}$. The only difference is a small shift of O$_2$ π* peaks to lower energy when O$_2$ is next to the perimeter (Fig. 3b). This is not caused by orbital interaction with Au because Au PDOS is almost identical in Figs. 3b-3d. Moreover, there appears to be little difference between the PDOS of the three Ti$^{5c}$ sites before O$_2$ adsorption (Fig. 3a). Hence the shift is attributed to the influence of the positive charge left on the Au rod after electron transfer to O$_2$.

In summary, we have confirmed strong adsorption of O$_2$ to the Ti$^{5c}$ site of TiO$_2$ in the presence of the Au rod. The adsorption is strong also at Ti$^{5c}$ sites away from the perimeter. This is because the Ti cation is chiefly responsible for stabilizing O$_2$ π*. The
Au rod plays an essential but passive role of electron reservoir, raising the Fermi level so that O$_2$ π* is filled. Still, the adsorption is strongest at the perimeter Ti$^{5c}$ site because the positive charge left on the rod provides additional stabilization to the O$_2$ anion.

3.2 CO oxidation

Because O$_2$ molecules are activated to a similar extent whether they are on the first or second nearest Ti$^{5c}$ sites, we expect them to react similarly to a gaseous CO. On the other hand, O$_2$ on the nearest site can also react with CO adsorbed on the rim of the Au rod (Au-CO). We thus examine three cases of CO oxidation: (i) the nearest O$_2$ reacting with Au-CO (Figs. 4a-4c), (ii) the nearest O$_2$ reacting with a gaseous CO (Figs. 4d-4f), and (iii) the second nearest O$_2$ reacting with a gaseous CO (Figs. 4g-4i). The results presented in this section were obtained using the 4×3 cell.

For the reaction of Au-CO with the nearest O$_2$, we find the initial and transition states shown in Figs. 4a and 4b, respectively. A search for the final state resulted in CO$_2$ desorption as shown in Fig. 4c, so CO$_2$ desorbs spontaneously after reaction. The energy barrier is calculated to be 0.22 eV, consistent with the low-temperature catalytic activity of Au/TiO$_2$. In the initial state (Fig. 4a), CO is adsorbed head-on to the perimeter Au site. The Au-O$^{br}$ bond there is broken as a result. The plot of a Kohn-Sham orbital directly below the Fermi level shows the hybridization between Au states and an antibonding π orbital of CO, indicating back donation from the Au rod to CO (Fig. 5). On the other hand, the state of O$_2$ activation is unaffected by CO adsorption: The O$_2$ bond length and the charge on O$_2$ are similar to those before CO adsorption (Table 2). In proceeding from the initial to transition state, the distance between O$_2$ and CO is greatly decreased, whereas neither the O$_2$ bond length nor the charge on O$_2$ is changed much (Table 2). That is, the transition state is reached without further activation of O$_2$. The energy barrier is thus mostly due to the exchange repulsion between O$_2$ and CO, and not due to the cost of breaking the O$_2$ bond.

For the reaction of a gaseous CO with the nearest O$_2$, we find the transition state shown in Fig. 4e. Relaxation from this transition state resulted in desorption of CO and CO$_2$ as shown in Figs. 4d and 4f, respectively, so a gaseous CO reacts directly with the O$_2$ anion to produce a gaseous CO$_2$. The calculated energy barrier (0.56 eV) is small, but
not as small as that (0.22 eV) for the reaction with Au-CO. Because the gaseous CO reacts directly with O₂, the reaction depends heavily on O₂ activation, requiring more activation energy. In fact, the O-O bond is stretched further and O₂ is made more negative at the transition state (Table 2).

The reaction of a gaseous CO with the second nearest O₂ is very much similar to that with the nearest O₂. The calculated transition state is shown in Fig. 4h. Relaxation from this transition state resulted in desorption of CO and CO₂ as shown in Figs. 4g and 4i, respectively, so a gaseous CO reacts directly with the O₂ anion to produce a gaseous CO₂. The calculated energy barrier (0.57 eV) is almost equal to that for the nearest O₂ (0.56 eV). The stretching of the O-O bond and the more negative charge on O₂ at the transition state indicate that the reaction requires further activation of O₂ (Table 2).

4. Conclusions
We have performed DFT calculations for O₂ adsorption and CO oxidation on the rutile TiO₂(110) surface supporting a Au rod. First, we have confirmed that O₂ is strongly adsorbed and activated on the Ti⁵c site next to the perimeter (Eₐd ~ −1.7 eV). The large energy generated by adsorption will assist desorption of contaminants from the oxide surface, releasing additional Ti⁵c sites for O₂ adsorption. Then, we have found strong adsorption energies (Eₐd ~ −1.4 eV) also for Ti⁵c sites away from the perimeter. This is because the Ti cation is mostly responsible for O₂ adsorption, by stabilizing O₂ π*, enabling electron transfer from the Au rod to O₂, and binding the resulting O₂ anion. The role of the Au rod is primarily to donate electrons to O₂ although positive charge left on it provides some additional stabilization to the O₂ anion.

O₂ adsorbed on the second nearest Ti⁵c site is almost as active as that on the perimeter Ti⁵c site, in fact, we have found almost identical energy barriers (0.57 and 0.56 eV, respectively) for the reaction with a gaseous CO. At room temperature and below, however, CO oxidation still occurs predominantly at the perimeter for two reasons. First, the adsorption is strongest at the perimeter because of the influence of the positive charge of the Au rod on the O₂ anion. Even the small energy difference (~0.3 eV) translates to a Boltzmann factor of 10⁴ and more at low temperatures. The probability of electron transfer should also decrease with distance from the perimeter. Second, the
barrier is much smaller (0.22 eV) for the reaction between O\textsubscript{2} on the perimeter Ti\textsuperscript{5c} site and CO on the Au rod because CO is activated by back donation from Au. At higher temperatures, on the other hand, O\textsubscript{2} appears to react on the surface of a Au NP [6]. Under such conditions, the Au NP is likely to donate electrons directly to O\textsubscript{2} on its surface, instead of transferring them to O\textsubscript{2} on Ti\textsuperscript{5c} sites. Thus, at least with the conventional TiO\textsubscript{2}-supported Au NP catalysts, the catalytic activity of remote Ti\textsuperscript{5c} sites makes comparatively little contributions to the rate of CO oxidation. The remote sites can still be useful when the perimeter sites are poisoned with contaminants, for example.

In conclusion, Ti\textsuperscript{5c} sites away from the perimeter of Au/TiO\textsubscript{2} also have the ability to adsorb and activate O\textsubscript{2}. CO oxidation still occurs predominantly along the perimeter because O\textsubscript{2} adsorption is strongest at the perimeter Ti\textsuperscript{5c} site and because CO is activated on Au. Thus, the reactive zone is confined to the immediate vicinity of the perimeter.

Acknowledgments
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References
Tables

Table 1. Variation of O$_2$ adsorption energy $E_{ad}$, O$_2$ bond length $l$, and Bader charge $q$ on O$_2$ with cell width, k-point mesh, and the number of trilayers $L$.

<table>
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<th>$l$ (pm)</th>
<th>$q$ (e)</th>
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Table 2. Variation of interatomic distance $d$ (pm) and Bader charge $q$ (e) during CO oxidation over Au/TiO$_2$, calculated using the 4×3 cell. O$_2$(i): O$_2$ on the nearest Ti$^{5c}$ site. O$_2$(ii): O$_2$ on the second nearest Ti$^{5c}$ site. Au-CO: CO on the Au rod. CO(g): CO from the gas phase.

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<th>$d$(O-O)</th>
<th>$d$(C-O)</th>
<th>$d$(O$_2$-CO)</th>
<th>$d$(Au-CO)</th>
<th>$q$(O$_2$)</th>
<th>$q$(CO)</th>
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Fig. 1. Perspective (a) and plan (b) views of the Au/TiO$_2$(110) model. Ti$^{5c}$, Ti$^{6c}$, O$^{br}$, and O$^{3c}$ denote five-fold coordinated Ti, six-fold coordinated Ti, two-fold coordinated bridging O, and three-fold coordinated in-plane O, respectively.
Fig. 2. O$_2$ on stoichiometric TiO$_2$(110) (a), O$_2$ at the perimeter of Au/TiO$_2$(110) (b-d), and O$_2$ on Ti$^{5c}$ sites (e-g). (a)-(d) are calculated in 4×3 and (e)-(g) in 6×3 cells. Adsorption energies, O$_2$ bond lengths, and Bader charges on O$_2$ (for spin-polarized systems, the difference of up and down Bader charges in parentheses) are displayed in the figure. O$_2$ in (b) is parallel to the rod axis.
Fig. 3. Projected density of states (PDOS): Before O$_2$ adsorption (a), O$_2$ on the first (b), second (c), and third (d) nearest Ti$^{5c}$ sites. Au$^{pm}$ denotes the Au atom at the perimeter of the rod. O$^{ad}$ denotes the O of O$_2$ that is nearer to the perimeter. O$^{3c}$ denotes the three-fold coordinated in-plane O nearest to the perimeter in (a) and the one directly below O$^{ad}$ in (b)-(d). The Fermi level is at 0 eV. Refer to Fig. 1 for definition of axes. Calculated using the 6×3 cell.
Fig. 4. CO oxidation: $\text{O}_2$ on the nearest Ti$^{5c}$ site reacting with Au-CO (a-c), $\text{O}_2$ on the nearest Ti$^{5c}$ site reacting with a gaseous CO (d-f), and $\text{O}_2$ on the second nearest Ti$^{5c}$ site reacting with a gaseous CO (g-i). Adsorption energies are displayed for IS and TS, and those at the desorption limit for snapshots (marked with an asterisk). Viewed along the rod. Calculated using the 4×3 cell.
Fig. 5. Highest occupied orbital (at the sampling k point) of the initial state of the reaction between $\text{O}_2$ and Au-CO (isosurface plot of the real part; plan view).