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Citation
Chemical Science (2015), 6: 3940-3951

Issue Date
2015-07

URL
http://hdl.handle.net/2433/216547

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Type
Journal Article

Textversion
publisher
Nitrogen fixation catalyzed by ferrocene-substituted dinitrogen-bridged dimolybdenum–dinitrogen complexes: unique behavior of ferrocene moiety as redox active site†

Shogo Kuriyama, Kazuya Arashiba, Kazunari Nakajima, Hiromasa Tanaka, Kazunari Yoshizawa*bc and Yoshiaki Nishibayashi*a

A series of dinitrogen-bridged dimolybdenum–dinitrogen complexes bearing metallocene-substituted PNP-pincer ligands is synthesized by the reduction of the corresponding monomeric molybdenum–trichloride complexes under 1 atm of molecular dinitrogen. Introduction of ferrocene as a redox-active moiety to the pyridine ring of the PNP-pincer ligand increases the catalytic activity for the formation of ammonia from molecular dinitrogen, up to 45 equiv. of ammonia being formed based on the catalyst (22 equiv. of ammonia based on each molybdenum atom of the catalyst). The time profile for the catalytic reaction reveals that the presence of the ferrocene unit in the catalyst increases the rate of ammonia formation. Electrochemical measurement and theoretical studies indicate that an interaction between the Fe atom of the ferrocene moiety and the Mo atom in the catalyst may play an important role to achieve a high catalytic activity.

1. Introduction

Activation and transformation of molecular dinitrogen using transition metal–dinitrogen complexes have been extensively studied to achieve novel nitrogen fixation systems under mild reaction conditions.1–3 Although a variety of stoichiometric reactivities of transition metal–dinitrogen complexes have been reported up to now, the catalytic conversion of molecular dinitrogen under ambient reaction conditions using these complexes as catalysts has so far been limited to only a few cases.1 The first successful example of the catalytic transformation of molecular dinitrogen into ammonia under ambient reaction conditions in the presence of a transition metal–dinitrogen complex was reported by Schrock and co-workers in 2003, where a molybdenum–dinitrogen complex bearing a triamidoamine ligand worked as a catalyst to produce 8 equiv. of ammonia based on the catalyst.4 In 2010, we reported another successful example of ammonia formation from molecular dinitrogen by using a dinitrogen-bridged dimolybdenum–dinitrogen complex bearing PNP-pincer ligands [Mo(N2)2(PNP)2(μ-N2)] (1a; PNP = 2,6-bis(di-tert-butylphosphinomethyl)pyridine) (Scheme 1) as a catalyst.5 More recently, Peters and co-workers have found the third successful example by using well-designed iron–dinitrogen complexes as catalysts.6

In our reaction system, a theoretical study revealed that a synergistic effect between two Mo moieties of the dinitrogen-bridged dimolybdenum–dinitrogen complex plays an important role to achieve the catalytic activity.5,6,7 An electron transfer from a Mo core to the active site of the other core via the bridging dinitrogen ligand contributes to the protonation steps of the coordinated nitrogenous ligands on the Mo atom. On the basis of the proposed catalytic pathway, we have succeeded in the development of a more effective catalyst by tuning the electronic property of the PNP-pincer ligand, where up to 52 equiv. of ammonia based on the catalyst (26 equiv. based on the Mo atom of the catalyst) were produced by using catalysts bearing an electron-donating group at the 4-position of the pyridine ring of the PNP-pincer ligand.8 Mechanistic studies suggested that the introduction of an electron-donating group to the PNP-pincer ligand enables a facile protonation on the terminal dinitrogen ligand on the Mo atom and also the suppression of the formation of a side product, molecular dihydrogen.9 However, not expectedly, the introduction of an electron-donating group such as methoxy group to the PNP-pincer ligand decreased the rate of ammonia formation probably due to the difficulty in reduction steps involved in the...
catalytic reaction. These results prompted us to develop a more efficient reaction system, where the reduction steps in the catalytic reaction can also be accelerated by suitably modifying the catalyst.

In biological nitrogen fixation, nitrogenase enzymes which are composed of two proteins named the MoFe protein and Fe protein have been known to catalyze ammonia formation from molecular dinitrogen under ambient reaction conditions (Fig. 1).

The active site of the nitrogenase, known as FeMo cofactor, receives electrons successively from P- and 4Fe4S-clusters, which are located adjacent to FeMo cofactor, to facilitate the multi-electron reduction of molecular dinitrogen into ammonia. By considering the so-far proposed reaction system of nitrogenase, we have envisaged the introduction of a ferrocene unit as a redox-active moiety to the PNP-pincer ligand of the dinitrogen-bridged dimolybdenum–dinitrogen complex, hoping to accelerate the reduction steps involved in the catalytic reaction via an intramolecular electron transfer from the Fe atom in the thiolated-bridged diiron core to the Fe atom of the ferrocene unit plays a critical role to promote the catalytic oxidation of molecular dihydrogen under ambient reaction conditions. In this article, we describe the preparation and catalytic activity of a series of dinitrogen-bridged dimolybdenum–dinitrogen complexes bearing metalloene-substituted PNP-pincer ligands 1 (Scheme 1). The dinitrogen-bridged dimolybdenum–dinitrogen complexes bearing ferrocene-substituted PNP-pincer ligands have been found to work as efficient catalysts, where up to 45 equiv. of ammonia are produced based on the catalyst (22 equiv. of ammonia based on each Mo atom of the catalyst). The unique behavior of the newly introduced ferrocene unit as a redox-active moiety in the complexes was discussed by means of experimental and theoretical methods.

Complex (Fig. 2), where an intramolecular electron transfer from the Fe atom in the thiolated-bridged diiron core to the Fe atom of the ferrocene unit plays a critical role to promote the catalytic oxidation of molecular dihydrogen under ambient reaction conditions.

In this article, we describe the preparation and catalytic activity of a series of dinitrogen-bridged dimolybdenum–dinitrogen complexes bearing metalloene-substituted PNP-pincer ligands 1 (Scheme 1). The dinitrogen-bridged dimolybdenum–dinitrogen complexes bearing ferrocene-substituted PNP-pincer ligands have been found to work as efficient catalysts, where up to 45 equiv. of ammonia are produced based on the catalyst (22 equiv. of ammonia based on each Mo atom of the catalyst). The unique behavior of the newly introduced ferrocene unit as a redox-active moiety in the complexes was discussed by means of experimental and theoretical methods.
2. Results and discussion

2.1. Synthesis of molybdenum complexes bearing metalloocene-substituted PNP-pincer ligands

A series of molybdenum-substituted PNP-pincer ligands 4-R-PNP (2; 4-R-PNP = 4-substituted 2,6-bis(di-tert-butylphosphino-methyl)pyridine) was newly prepared (Scheme 1) and the detailed experimental procedure is described in the ESL.† A variety of substituents such as ferrocenyl (Fc), 2-ferrocenylethyl (EtFc), 4-ferrocenylphenyl (PhFc), and ruthenocenyl (Rc) groups were introduced to the 4-position of pyridine ring of the PNP-pincer ligand. Treatment of [MoCl3(thf)3] (thf = tetrahydrofuran) with 1 equiv. of 4-R-PNP 2 in tetrahydrofuran (THF) at 50 °C for 18 h gave the corresponding molybdenum–trichloride complexes bearing the corresponding PNP-pincer ligands [MoCl3(4-R-PNP)] (3) in high yields. The molecular structures of 3 were confirmed by X-ray crystallography (see the ESL in Figs. S3–S5†).

Reduction of 3 with 6 equiv. of Na–Hg in THF at room temperature for 12 h under an atmospheric pressure of molecular dinitrogen gave the corresponding molybdenum-substituted dinitrogen-bridged dimolybdenum–dinitrogen complexes [Mo(N2)2(4-R-PNP)]2(μ-N2) (1) in good yields. These new dinitrogen-bridged dimolybdenum–dinitrogen complexes were characterized spectroscopically. Detailed molecular structures of [Mo(N2)2(4-Fc-PNP)]2(μ-N2) (1b) and [Mo(N2)2(4-Rc-PNP)]2(μ-N2) (1e) were determined by X-ray crystallography.

ORTEP drawings of 1b (a) and 1e (b). Hydrogen atoms and solvent molecules are omitted for clarity.

The nature of metalloocene as substituent was examined by IR spectroscopy because the electronic nature of substituent at the 4-position in pyridine ring of the PNP pincer-ligand is known to affect absorptions derived from the corresponding terminal dinitrogen ligand.† Typical results are shown in Table 1. The presence of Fc and Rc groups at the 4-position in pyridine ring of the PNP-pincer ligands in 1 did not change the N=N stretching frequencies (1944 cm⁻¹ (1b) and 1944 cm⁻¹ (1e), respectively). These results indicate that the introduction of these metalloocene moieties to the pyridine ring of the PNP-pincer ligand would not affect the electronic density on the molybdenum atom in the complex 1. On the other hand, the N=N stretching frequency of [Mo(N2)2(4-EtFc-PNP)]2(μ-N2) (1c) (1939 cm⁻¹) is smaller than that of 1a and the same as that of [Mo(N2)2(4-Ph-PNP)]2(μ-N2) (1d) is larger than that of 1a, and the same as that of [Mo(N2)2(4-Ph-PNP)]2(μ-N2) (1g, 1930 cm⁻¹), indicating that the PhFc group acts as an electron-withdrawing substituent in the complex.

2.2. Electrochemical property of molybdenum complexes bearing metalloocene-substituted PNP-pincer ligands

Cyclic voltammograms (CVs) of PNP-pincer ligands 2b–e in THF with [NBu4]BArF † (ArF † = 3,5-(CF3)2C6H3) as supporting electrolyte showed oxidation processes assigned to metalloocene moieties. Typical results are shown in Table 2. The PNP-pincer ligand bearing a ferrocene moiety, 2b, showed a reversible one-electron oxidation wave derived from the Fe(II/III) couple at E1/2 = +0.10 V vs. FeCp2/2+ (Fig. 4). The PNP-pincer ligands bearing ferrocene

Table 1 Yield of 1 and IR absorbance of terminal dinitrogen ligand in 1a

<table>
<thead>
<tr>
<th>Complex</th>
<th>R</th>
<th>Yield of 1 (%)</th>
<th>ν_N=N/cm⁻¹</th>
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<tr>
<td>1a</td>
<td>H</td>
<td>63</td>
<td>1944</td>
</tr>
<tr>
<td>1b</td>
<td>Fc</td>
<td>74</td>
<td>1944</td>
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<td>1c</td>
<td>EtFc</td>
<td>62</td>
<td>1939</td>
</tr>
<tr>
<td>1d</td>
<td>PhFc</td>
<td>32</td>
<td>1951</td>
</tr>
<tr>
<td>1e</td>
<td>Rc</td>
<td>52</td>
<td>1944</td>
</tr>
<tr>
<td>1f</td>
<td>Me</td>
<td>—</td>
<td>1939</td>
</tr>
<tr>
<td>1g</td>
<td>Ph</td>
<td>—</td>
<td>1950</td>
</tr>
</tbody>
</table>

* IR absorbance was measured in a THF solution of 1. † Based on 3. ‡ Ref. 9.
units such as Et<sup>Fe</sup> and Ph<sup>Fe</sup> groups, 2e and 2d, were oxidized at $E_{1/2} = -0.03$ V (2e) and $+0.04$ V (2d), respectively. On the other hand, the PNP-pincer ligand bearing a Rc moiety, 2e, showed one irreversible oxidation wave at $E_a = +0.79$ V.<sup>16</sup>

Electrochemical studies were also carried out on the molybdenum–trichloride complexes bearing a substituted PNP-pincer ligand 3. Typical results are shown in Table 2 and the CVs of 3 are shown in Fig. 4. Non-substituted complex [MoCl<sub>3</sub>(PNP)]<sup>3a</sup> has one reversible wave at $E_{1/2} = +0.14$ V assignable to Mo(0/1), while the CV of molybdenum–trichloride complex bearing a ferrocene-substituted PNP-pincer ligand, 3b, showed two successive reversible one-electron waves at $+0.10$ and $+0.42$ V. We have performed density functional theory (DFT) calculations to assign the observed oxidation waves. The electronic ground state of the cation of 3b ([3b]<sup>+</sup>) is a triplet state, where the Mulliken spin density is highly localized at the Mo atom (1.85). The removal of an electron from 3b changes the formal charge of the Mo atom from +3 (d<sup>3</sup>) to +4 (d<sup>2</sup>), and thus the oxidation at $E_{1/2} = +0.10$ V is assignable to Mo(0/1). On the other hand, the dication of 3b ([3b]<sup>2+</sup>) adopts a quartet state as the ground state of the cation, where the spin densities at the Mo and Fe atoms are calculated to be 1.82 and 1.27, respectively. This result suggests that the second oxidation should occur at the Fe atom and the oxidation at $E_{1/2} = +0.42$ V is assignable to Fe(0/1).

In the molybdenum–trichloride complex 3b, the $E_{1/2}$ value of Fe(0/1) of the ferrocene moiety shifts by $+0.32$ V upon coordination. This is evidence of the electronic interaction between the Mo and Fe centers in the complex 3b. In fact, similar shifts of the $E_{1/2}$ values of Fe(0/1) of the ferrocene moiety were observed in several transition metal complexes bearing 4-ferrocenylpyridine and 4′-ferroenyl-2,2′,6′,2′′-terpyridine as ligands, indicating an electronic interaction between the metal and Fe centers in these complexes.<sup>17,18</sup>

The electronic interaction between the Mo and Fe centers in 3b is also supported by the results of spectroelectrochemical measurements. A weak absorption appeared in the near-IR region at around 800–1800 nm when 3b was oxidized to [3b]<sup>+</sup> (Fig. S8 and S9†). A time-dependent BP86 calculation of [3b]<sup>+</sup> predicted three electron transitions at 3188, 2276 and 1073 nm. Fig. 5 depicts electron density difference maps (EDDMs) for these transitions, where EDDMs present isosurface plots of loss (light green) and gain (purple) of electron density for the corresponding electron transitions. Although the observed absorption of [3b]<sup>+</sup> in the near IR region was weak, the calculated EDDMs clearly show that these low-energy electron transitions were assignable to the metal-to-metal charge transfer (MMCT) from Fe(0) of the ferrocene to Mo(0). The relationship between electronic couplings and intramolecular MMCTs has already been observed in some transition-metal complexes bearing a 4-ferroenylpyridine scaffold as a ligand.<sup>18</sup> In fact, intramolecular MMCTs between the Fe atom in the ferrocene moiety and the Ru atom were observed in ruthenium complexes bearing 4-ferroenylpyridines<sup>19a,b</sup> and/or 4′-ferroenyl-2,2′,6,2′′-terpyridine as ligands.<sup>19c</sup>

The CVs of the molybdenum–trichloride complexes bearing Et<sup>Fe</sup> and Ph<sup>Fe</sup> groups, 3c and 3d, showed similar two merged reversible waves. On the basis of the result of the CV of 3b, we

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### Table 2 Electrochemical data of PNP-pincer ligands 2 and molybdenum–trichloride complexes 3 bearing PNP-pincer ligand<sup>a</sup>

<table>
<thead>
<tr>
<th>Compd.</th>
<th>R</th>
<th>Mo(0/1)/V</th>
<th>Fe(0/1)/V</th>
<th>$\Delta E(Fe)/V$</th>
</tr>
</thead>
<tbody>
<tr>
<td>2b</td>
<td>Fc</td>
<td>—</td>
<td>+0.10</td>
<td>—</td>
</tr>
<tr>
<td>2c</td>
<td>Et&lt;sup&gt;Fe&lt;/sup&gt;</td>
<td>—</td>
<td>—0.03</td>
<td>—</td>
</tr>
<tr>
<td>2d</td>
<td>Ph&lt;sup&gt;Fe&lt;/sup&gt;</td>
<td>—</td>
<td>+0.04</td>
<td>—</td>
</tr>
<tr>
<td>2e</td>
<td>Rc</td>
<td>—</td>
<td>+0.79&lt;sup&gt;c&lt;/sup&gt;</td>
<td>—</td>
</tr>
<tr>
<td>3a</td>
<td>H</td>
<td>+0.14</td>
<td>—</td>
<td>—</td>
</tr>
<tr>
<td>3b</td>
<td>Fc</td>
<td>+0.10</td>
<td>+0.42</td>
<td>+0.32</td>
</tr>
<tr>
<td>3c</td>
<td>Et&lt;sup&gt;Fe&lt;/sup&gt;</td>
<td>+0.05</td>
<td>+0.19</td>
<td>+0.22</td>
</tr>
<tr>
<td>3d</td>
<td>Ph&lt;sup&gt;Fe&lt;/sup&gt;</td>
<td>+0.06&lt;sup&gt;e&lt;/sup&gt;</td>
<td>+0.13&lt;sup&gt;e&lt;/sup&gt;</td>
<td>+0.09</td>
</tr>
<tr>
<td>3e</td>
<td>Rc</td>
<td>+0.09&lt;sup&gt;e&lt;/sup&gt;</td>
<td>—</td>
<td>—</td>
</tr>
</tbody>
</table>

<sup>a</sup> Measured by cyclic voltammetry in THF at 0.1 V s<sup>−1</sup> with [NBu<sub>4</sub>]<sup>+</sup>BARF<sub>4</sub> as supporting electrolyte. Potentials are indicated as $E_{1/2}$ values vs. FeCp<sub>2</sub><sup>0/1</sup>.<sup>b</sup> $\Delta E(Fe)$ is the difference of the Fe<sup>0/1</sup> potential between 2 and 3.<sup>c</sup> $E_{1/2}$ value of ruthenocene oxidation.<sup>d</sup> Measured by differential pulse voltammetry (DPV).<sup>e</sup> Oxidation of ruthenocene was not observed in the electrochemical window.

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**Fig. 4** Cyclic voltammograms of 2b and 3a–e.
These results indicate that the electronic communication is weakened by the introduction of ethylene or phenylene group into the linkage between the ferrocene and the pyridine of the PNP-pincer ligand. Similar phenomena have already been observed in a rhodium complex bearing 4-(4-ferrocenylphenyl) pyridine as a ligand and diruthenium complexes bearing a bridging dipyridine ligand. In the latter case, the peak to peak separation in the CVs of the dinuclear ruthenium complexes decreased with the increase of the metal separation of the two Ru centers in the complex.

The CV of the molybdenum–trichloride complex bearing a Re-substituted PNP-pincer ligand, 3e, showed only one reversible wave assignable to Mo(m/v) in the electrochemical window, indicating that the Re group in 3e did not act as a redox active moiety, in contrast to the ferrocene units in 3b, 3c and 3d (vide supra).

2.3. Evaluation of electronic interaction between Fe and Mo atoms in ammonia formation

The electrochemical and theoretical studies revealed the presence of electronic interaction between the Mo and Fe atoms in the series of molybdenum–trichloride complexes bearing a ferrocene-substituted PNP-pincer ligand. Such interactions could enhance the catalytic ability of the dinitrogen-bridged dimolybdenum complexes via an electron transfer from the Fe atom in the ferrocene moiety to the Mo atom of the complexes bearing nitrogenous ligands that are transformed into ammonia. Among these molybdenum complexes proposed as reactive intermediates in the catalytic cycle in our previous study,7 we have found that a dinitrogen-bridged dimolybdenum–hydrazidium (−NNH₂) complex met our hypothesis. In the reaction pathway starting from 1a, dinitrogen-bridged dimolybdenum–hydrazide(2−) complex [Mo(PNP)(OTf)(NNH₂)−N≡N−Mo(PNP)(N₃)₂] (1a) is protonated to yield the corresponding dinitrogen-bridged dimolybdenum–hydrazidium complex [Mo(PNP)(OTf)(NNH₂)−N≡N−Mo(PNP)(N₃)₂]^+ (IIa), which is then reduced with cobaltocene to afford ammonia and the dinitrogen-bridged dimolybdenum–nitrile complex [Mo(PNP)(OTf)(N≡N−Mo(PNP)(N₃)₂)] (IIIa) (Scheme 2). The protonation of 1a is slightly exothermic (ΔE = −1.6 kcal mol⁻¹) with very low activation energy of 1.9 kcal mol⁻¹, while the reduction of IIa followed by the N–N bond cleavage is extremely exothermic (ΔE = −156.1 kcal mol⁻¹).7 We have now carried out DFT calculations to examine the impact of the introduction of ferrocene to the pyridine ring of the PNP-pincer ligand. Fig. 6(a) presents an optimized structure of the ferrocene-substituted dinitrogen-bridged dimolybdenum–hydrazidium complex [Mo(4-Fc-PNP)(OTf)(NNH₂)−N≡N−Mo(4-Fc-PNP)(N₃)₂]^+ (IIb).

For electronic transitions of IIb calculated by the TD-BP86 method (Fig. 6(b)), a low-energy transition at 1174 nm is assignable to the electron transfer from the Fe atom of the ferrocene moiety to the Mo atom. Fig. 6(c) and (d) describe the EDDM of the transition at 1174 nm and molecular orbitals responsible for the transition (HOMO−2 and LUMO). This can be assigned to a MMCT transition from the Fe atom to the Mo atom. It should also be noted that the LUMO of IIb has

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**Fig. 5** Electron density difference maps (EDDMs) of the electron transitions at 3188, 2276, 1073 nm of [3b]⁺ calculated with a time-dependent BP86 method. Isosurface plots of loss and gain of electron density are presented in light green and purple, respectively.

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assigned the former waves (E₁/₂⁺: +0.05 and +0.06 V) to Mo(m/v) and the latter (E₁/₂⁻: +0.19 and +0.13 V) to Fe(m/v) in the CVs of 3e and 3d (Table 2). In both cases, oxidation of the Mo atom at first may occur in the anodic scan and then oxidation of the Fe atom of the ferrocene unit in the PNP-pincer ligand may proceed sequentially. The E₁/₂⁺ values of Fe(m/v) of the ferrocene units shift by +0.22 and +0.09 V, respectively, upon coordination.
2.4. Catalytic formation of ammonia from molecular dinitrogen

2.4.1. Catalytic reactions by using metalocene-substituted dinitrogen-bridged dimolybdenum–dinitrogen complexes. We carried out the catalytic reduction of molecular dinitrogen into ammonia using the complex 1 as a catalyst, according to the following procedure of the previous method. To a mixture of 1 and 2,6-lutidinium trifluoromethanesulfonate (288 equiv. to 1) as a proton source in toluene was added a solution of cobaltocene (216 equiv. to 1; CoCp₂) as a reductant in toluene via a syringe pump at room temperature over a period of 1 h, and the resulting mixture was stirred at room temperature for another 19 h under 1 atm of dinitrogen. After the reaction, the formation of ammonia together with molecular dihydrogen was observed, the amounts of ammonia and molecular dihydrogen being determined by indophenol method and GC, respectively. Typical results are shown in Table 3. The yields of ammonia and molecular dihydrogen were estimated on the basis of cobaltocene. No formation of other products such as hydrazine was observed in all cases.

The reaction by use of 1b as a catalyst gave 37 equiv. of ammonia and 36 equiv. of molecular dihydrogen, respectively, based on the catalyst (Table 3, run 1). The amount of produced ammonia was higher than that obtained by using 1a as a catalyst under the same reaction conditions (Table 3, run 2). For comparison, we carried out the catalytic reaction by using 1a as a catalyst in the presence of 2 equiv. of ferrocene to 1a, where 21 equiv. of ammonia were produced based on the catalyst (Table 3, run 3). These results indicated that the ferrocene moiety directly connected to the pyridine ring of the PNP-pincer ligand plays an important role to promote the catalytic reaction more efficiently. As shown in the previous section, the electronic density of the Mo atom of 1b is almost the same with that of 1a (Table 1). The presence of ferrocene at the PNP-pincer ligand in 1 did not affect the nature of the coordinated terminal dinitrogen ligand in 1. Thus, the ferrocene moiety in 1b may not act as an electron-donating group but rather as a redox active group for assisting to promote the reduction steps of the coordinated nitrogenous ligands on the Mo atom more smoothly.

Next, we carried out the catalytic ammonia formation by using other ferrocene-substituted dinitrogen-bridged dimolybdenum–dinitrogen complexes as catalysts. In the reactions using 1c and 1d as catalysts, lesser amounts of ammonia compared to that by use of 1b was formed (30 equiv. and 10 equiv. of ammonia based on the catalysts, respectively) (Table 3, runs 4 and 5). The catalytic activity of 1 decreased with the
decrease of the electrochemical interaction between the Mo atom and the Fe atom of the ferrocene unit in the PNP-pincer ligands (Table 4), and actually a good relationship between $\Delta E(\text{Fe})$ [the difference of the Fe($\mu$/m) potential ($E_{1/2}$) between 2 and 3] and $\Delta NH_3$ [the difference of the catalytic activity of 1] was observed as shown in Fig. 7. This result also shows that the complex having a larger $\Delta E(\text{Fe})$ value has a higher catalytic activity for ammonia formation.

The reaction using the ruthenocene-substituted dinitrogen-bridged dimolybdenum–dinitrogen complex 1e as a catalyst under the same reaction conditions gave only 25 equiv. of ammonia based on the catalyst (Table 3, run 6). The catalytic activity of 1e was almost the same with that of 1a and lower than that of 1b. This result indicates that the ruthenocene moiety in 3e did not act as a redox active group in the present reaction system.

2.4.2. Nature of ferrocene-substituted dinitrogen-bridged dimolybdenum–dinitrogen complexes. We previously reported the nature of substituents at PNP-pincer ligands in dinitrogen-bridged dimolybdenum–dinitrogen complexes affects the catalytic activity for ammonia formation.3 Especially, the dinitrogen-bridged dimolybdenum–dinitrogen complex bearing 4-methoxy-substituted PNP-pincer ligands [Mo(N2)2(4-MeO-PNP)(N2)]2+ (1h) worked as the most effective catalyst toward the catalytic ammonia formation under ambient reaction conditions.3 However, the formation of ammonia by using 1h as a catalyst proceeded sluggishly. For comparison with the catalytic activity of 1b toward ammonia formation, we previously carried out the catalytic reaction by using 1h as a catalyst under the present reaction conditions, where 34 equiv. of ammonia were formed based on the catalyst (Table 3, run 7).3 In addition, when larger amounts of CoCp2 (288 equiv. to 1) and [LutH]OTf

Table 3 Molybdenum-catalyzed reduction of molecular dinitrogen into ammonia under ambient reaction conditionsa

<table>
<thead>
<tr>
<th>Run</th>
<th>Cat.</th>
<th>R</th>
<th>$NH_3$ (equiv.)</th>
<th>$NH_3$ (%)</th>
<th>$H_2$ (equiv.)</th>
<th>$H_2$ (%)</th>
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<tbody>
<tr>
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<td>Fe</td>
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<td>51</td>
<td>36h</td>
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<td>H</td>
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</tr>
<tr>
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<td>H</td>
<td>21f</td>
<td>29</td>
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</tr>
<tr>
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<td>1h</td>
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</tbody>
</table>

a To a mixture of the catalyst (1: 0.010 mmol) and [LutH]OTf (288 equiv. to the catalyst) as proton source in toluene (1.0 mL) was added a solution of CoCp2 (216 equiv. to catalyst) as a reductant in toluene (4.0 mL) at room temperature over a period of 1 h, followed by stirring the resulting mixture at room temperature for another 19 h under an atmospheric pressure of dinitrogen. b Mol equiv. relative to the catalyst. Averages of multiple runs are shown. c Yield based on CoCp2. d Ref. 9. e FeCp2 (2 equiv. to 1a) was added. f Variation of ±2 equiv. between experiments. g Variation of ±3 equiv. between experiments. h Variation of ±1 equiv. between experiments.
were used as reductant and proton source under the same reaction conditions, 45 equiv. (1b as a catalyst) and 43 equiv. (1h as a catalyst) of ammonia were produced based on the catalysts, respectively (Scheme 3). This result indicated that the catalytic activity of 1b is almost the same as that of 1h, which was so far known to work as the most effective catalyst toward the formation of ammonia under ambient reaction conditions.

To obtain more information on the catalytic behavior of 1b, we monitored the time profile of catalytic reactions by using 1b as a catalyst. Typical results are shown in Fig. 8 together with the time profiles of catalytic reactions by using 1a and 1h as catalysts. The catalytic conversion of molecular dinitrogen into ammonia in the presence of a catalytic amount of 1b proceeded more rapidly than that with 1a and 1h, and the turnover frequency (TOF) for ammonia, which was determined by mols of ammonia produced in initial 1 h per catalyst, by 1b as a catalyst was 23 h⁻¹/C₀. Typical results are shown in Table 5. On the other hand, the catalytic formation of ammonia in the presence of catalytic amounts of 1a and 1h proceeded more slowly than that of 1b under the same reaction conditions. In fact, the TOFs for ammonia formation using 1a and 1h as catalysts, 17 h⁻¹ and 7 h⁻¹, respectively, were lower than that using 1b. These results indicate that the ferrocene-substituted dinitrogen-bridged dimolybdenum–dinitrogen complex 1b has the highest performance from the viewpoint of both the catalytic ability and the rate for ammonia formation.

Previously we disclosed that the formation of ammonia and molecular dihydrogen is complementary in the present reaction system. To obtain more information on the catalytic activity of 1b toward the formation of molecular dihydrogen, we monitored it in the catalytic reaction using 1b as a catalyst. Typical

Table 4 Difference of catalytic activity of dinitrogen-bridged dimolybdenum–dinitrogen complexes and difference of \( E_{1/2} \) value of Fe(II/III) couple upon coordination.

<table>
<thead>
<tr>
<th>Run</th>
<th>Complex bearing ferrocene moiety</th>
<th>R</th>
<th>( \text{NH}_3^b ) (equiv.)</th>
<th>( \Delta \text{NH}_3^b ) (equiv.)</th>
<th>( \Delta E_{1/2} \slash V )</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>1b</td>
<td>Fc</td>
<td>37</td>
<td></td>
<td></td>
</tr>
<tr>
<td>2</td>
<td>1c</td>
<td>EtFc</td>
<td>30</td>
<td></td>
<td></td>
</tr>
<tr>
<td>3</td>
<td>1d</td>
<td>PhFc</td>
<td>10</td>
<td></td>
<td></td>
</tr>
</tbody>
</table>

\( ^a \) Ref. 9. \(^b \) The amount of ammonia based on the catalyst. \(^c \) Difference of \( E_{1/2} \) value of the Fe(II/III) couple between the corresponding PNP-pincer ligand 2 and the molybdenum–trichloride complex 3. See Table 2. \(^d \) Difference between the amount of ammonia catalyzed by 1b and that by 1a. \(^e \) Difference between the amount of ammonia catalyzed by 1c and that by 1f. \(^f \) Difference between the amount of ammonia catalyzed by 1d and that by 1g.
results are shown in Fig. 9. The formation of molecular dihydrogen proceeded simultaneously with ammonia formation, where the TOF for molecular dihydrogen using $1b$ as a catalyst is 17 h$^{-1}$ (Table 5). As shown in the previous paper, the TOF for molecular dihydrogen using $1a$ as a catalyst is 23 h$^{-1}$. Thus, the TOF of $1b$ for molecular dihydrogen is lower than that of $1a$. These results indicate that $1b$ can promote the formation of ammonia and, on the other hand, suppress the formation of molecular dihydrogen.

Based on the unique behavior and the DFT calculation result of $1b$, we consider that the presence of the ferrocene moiety directly connected to the pyridine ring of the PNP-pincer ligand in the catalyst increases the rate of the reduction steps of the coordinated nitrogenous ligand on the Mo atom. In this reaction system, the electronic interaction between the Fe atom of the ferrocene moiety and the Mo atom of the complex plays an important role to promote the catalytic formation of ammonia more smoothly, where the ferrocene moiety works as an intramolecular reductant, clarified by the reversible one-electron redox behavior, toward the high-oxidation state of the Mo complexes such as hydrazidium complexes. The observed effect of direct introduction of a ferrocene moiety to the pyridine ring

Table 5 TOF and TON values for the formation of ammonia and molecular dihydrogen catalyzed by $1$$^a$

<table>
<thead>
<tr>
<th>Complex</th>
<th>R</th>
<th>TOF (NH$_3$)/h$^{-1}$</th>
<th>TOF (H$_2$)/h$^{-1}$</th>
<th>TON (NH$_3$)</th>
<th>TON (H$_2$)</th>
</tr>
</thead>
<tbody>
<tr>
<td>$1b$</td>
<td>Fe</td>
<td>23$^b$</td>
<td>17$^c$</td>
<td>37$^c$</td>
<td>36$^b$</td>
</tr>
<tr>
<td>$1a$</td>
<td>H</td>
<td>17</td>
<td>23</td>
<td>23</td>
<td>46</td>
</tr>
<tr>
<td>$1h$</td>
<td>MeO</td>
<td>7</td>
<td>5</td>
<td>34</td>
<td>36</td>
</tr>
</tbody>
</table>

$^a$ TOF and TON are determined based on the catalyst. Average of multiple runs is shown. See Table 3 for reaction conditions. Values of $1a$ and $1h$ are from ref. 9. $^b$ Variation of ±1 equiv. between experiments. $^c$ Variation of ±2 equiv. between experiments.

![Fig. 8 Time profiles of ammonia formation with $1a$ (red), $1b$ (black) and $1h$ (green) as catalysts.](image)

![Fig. 9 Time profiles of the formation of ammonia (solid line) and molecular dihydrogen (dashed line) with $1b$ as a catalyst.](image)

![Fig. 10 The role of substituent to the PNP-pincer ligand toward catalytic ammonia formation from molecular dinitrogen.](image)
of the PNP-pincer ligand is in sharp contrast to the effect observed by introduction of an electron-donating group such as methoxy group to the pyridine ring of the PNP-pincer ligand. In the latter case, the electron-donating group accelerates the protonation step of the coordinated nitrogenous ligands on the Mo atom of the complexes in the catalytic reaction (Fig. 10), while a ferrocene moiety increases the rate of the reduction step.\textsuperscript{a} At present, we can not exclude another possibility of the role of the ferrocene moiety in the catalyst and further work is necessary to clarify the exact role of the ferrocene moiety in the catalyst. However, the result described in the present paper provides the first successful example of the substantial improvement of the catalytic reactivity by the presence of the ferrocene moiety as a redox active site in the catalytic reduction of molecular dinitrogen into ammonia under ambient reaction conditions.\textsuperscript{24}

The electron transfer from the Fe atom of the ferrocene moiety to the active site of the Mo atom in the complex plays an important role to promote catalytic ammonia formation effectively. As presented in the introduction part, a similar electron transfer may occur at the active site of nitrogenase enzyme in biological nitrogen fixation (Fig. 1). We believe that the mechanistic insight obtained in the present paper may provide valuable information to understand the mechanism of biological nitrogen fixation by nitrogenase.

3. Conclusion

We have prepared and characterized a series of dinitrogen-bridged dimolybdenum–dinitrogen complexes bearing metalloocene-substituted PNP-pincer ligands spectroscopically and theoretically. The complex bearing a ferrocene group on the pyridine ring of the PNP-pincer ligand 1b has been found to work as the most efficient catalyst toward the catalytic formation of ammonia from molecular dinitrogen, where up to 45 equiv. of ammonia were produced based on the catalyst (22 equiv. of ammonia based on each Mo atom of the catalyst). The catalytic activity of 1b is almost the same with that of 1h, which was previously reported to work as the most effective catalyst toward ammonia formation under ambient reaction conditions.\textsuperscript{a} Mechanistic study indicated that the presence of ferrocene as a redox active moiety in the PNP-pincer ligand increases the rates of the reduction step involved in the catalytic process of ammonia formation via an intramolecular electron transfer from the Fe atom of the ferrocene to the active site of the Mo atom in the complex. The present result is another useful innovation to the previous finding that the introduction of an electron-donating group to the pyridine ring of the PNP-pincer ligand accelerated the protonation step involved in the catalytic reaction.\textsuperscript{9} We believe that the result described in this article provides useful information to design a more effective catalyst toward the catalytic formation of ammonia from molecular dinitrogen under ambient reaction conditions. Further work is currently in progress to develop a more effective reaction system.\textsuperscript{25}

Acknowledgements

The authors are grateful to M. Takahashi and Prof. H. Houjou in The University of Tokyo for spectroelectrochemical analysis. Y.N. thanks Grants-in-Aid for Scientific Research (nos 26288044, 26620075 and 26105708) from the Japan Society for the Promotion of Science (JSPS) and the Ministry of Education, Culture, Sports, Science and Technology of Japan (MEXT). K.Y. thanks Grants-in-Aid for Scientific Research (nos 22240528 and 24109014) from the Japan Society for the Promotion of Science (JSPS) and the Ministry of Education, Culture, Sports, Science and Technology of Japan (MEXT) and the MEXT Projects of “Integreted Research on Chemical Synthesis” and “Elements Strategy Initiative to Form Core Research Center”. H.T. thanks a Grant-in-Aid for Research Activity Start-up (nos 26888008) from the Japan Society for the Promotion of Science (JSPS). S.K. is a recipient of the JSPS Predoctoral Fellowships for Young Scientists. We also thank the Research Hub for Advanced Nano Characterization at The University of Tokyo for X-ray analysis.

Notes and references


Independently, Batista and co-workers have recently reported a theoretical study on our reaction pathway, and they also concluded bimetallic complexes as reactive intermediates: Y.-H. Tian, A. W. Pierpont and E. R. Batista, Angew. Chem., Int. Ed., 2014, 53, 4512.

8 Independently, Batista and co-workers have recently reported a theoretical study on our reaction pathway, and they also concluded bimetallic complexes as reactive intermediates: Y.-H. Tian, A. W. Pierpont and E. R. Batista, Inorg. Chem., 2014, 53, 4177.


21 Separately, we confirmed the direct conversion of molecular dinitrogen into ammonia by using $^{15}\text{N}_2$ gas in place of normal $^{14}\text{N}_2$ gas (see the ESI† for details).

22 Separately, we also investigated catalytic reactions by using 1b as a catalyst with another reductant such as CrCp$_2^*$ (Cp$^*$ = $^5\text{C}_5\text{Me}_5$) having a higher reducing ability and proton sources such as [2-PicH]OTf (2-Pic = 2-picoline) and [ColH]OTf (Col = 2,4,6-collidine) under the same reaction conditions. In all cases, however, the yield of ammonia was lower compared with the result using CoCp$_2$ and [LutH]OTf as a reductant and a proton source, respectively. See the ESI† for details.

23 By comparison of the IR absorbance of the terminal dinitrogen ligand in 1a with that in 1d (vide supra), the PhFc-moiety in 1d is considered to act as an electron-withdrawing group, suppressing the rate of catalytic formation. This may be the reason why a lower amount of ammonia was produced under the same reaction conditions.

24 (a) Previously, George and Tisdale found a substantial improvement of the stoichiometric formation of ammonia by the protonation of a molybdenum–dinitrogen complex bearing a ferrocenyldiphosphine as an auxiliary ligand.\(^{24b}\)

More recently, we found that a molybdenum–dinitrogen complex bearing two ferrocenyldiphosphines as auxiliary ligands worked as an effective catalyst toward the formation of silylamine from molecular dinitrogen under ambient reaction conditions.\(^{24b}\) Electronic interaction between the Fe atom of the ferrocenyldiphosphine and the Mo atom was not observed in both cases; (b) T. A. George and R. C. Tisdale, *J. Am. Chem. Soc.*, 1985, **107**, 5157.