

# Rapid Benzylation of Cellulose in Tetra-*n*-butylphosphonium Hydroxide Aqueous Solution at Room Temperature

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**KEYWORDS:** Cellulose, Etherification, Benzylation, Tetra-*n*-butylphosphonium hydroxide, Rapid reaction

**ABSTRACT:** We found a very fast benzylation of cellulose in a tetraalkyl-onium hydroxide solvent system, 47% tetra-*n*-butylphosphonium hydroxide aqueous solution. Benzyl cellulose (BC) with a degree of substitution (DS) above 2.5 was obtained within 10 min at ambient temperature (20–25 °C) using ~9 molar equivalents of benzyl bromide to the anhydroglucose

unit. The highly efficient benzylation proceeded in the transiently stabilized emulsion state of the reaction system, and the product precipitated out as a solid sediment in the liquid medium. At 20 °C, the benzyl DS was 1.92 after 2 min, and reached 2.40 and 2.53 after 5 and 10 min, respectively. The reaction temperature had a negative correlation with the maximum DS in the explored range of 10–50 °C. No significant degradation of the cellulosic main chain was observed for highly substituted BCs (DS > 2.5). The solubility of BC in common organic solvents was also examined for products with various DSs.

## INTRODUCTION

Cellulose is the most abundant natural polymer on earth and is widely available from plant sources. It can even be an attractive candidate to replace petroleum-based organic materials. An intrinsic problem of cellulose is the intractability in materialization: cellulose cannot be melt-processed in itself and has limited industrially viable solvents. The reactivity of three hydroxyl groups in each glucose residue certainly permits a wide range of modifications of cellulose for different functionalities.<sup>1–4</sup> However, to derivatize the molecule, the hydroxyl groups must be activated and easily accessible by the reactant; hence, proper swelling or dissolution of pristine cellulose is generally required, and the reaction often needs a significant excess of solvents or intense pretreatments, such as ball milling, to increase the amount of the amorphous phase. Thus, the total production cost of the cellulose derivatives can be relatively high despite the availability of the primary resource.<sup>1–4</sup>

Recently, Abe *et al.* reported that tetraalkyl-onium hydroxide aqueous (aq.) solutions dissolved cellulose at ambient temperature.<sup>5,6</sup> Tetra-*n*-butylphosphonium hydroxide

([P<sub>4,4,4,4</sub>]OH) and tetra-*n*-butylammonium hydroxide aq. solutions dissolved over 15 wt % cellulose within a few minutes at 25 °C, without pretreatment. In the present paper, we would like to demonstrate that a [P<sub>4,4,4,4</sub>]OH aq. solution is employable as an excellent medium for facile etherification of cellulose.

Cellulose ethers, such as alkyl cellulose, hydroxyalkyl cellulose, and carboxyalkyl cellulose, are used as auxiliaries in food, cosmetics, pharmaceuticals, adhesives, etc.<sup>7–14</sup> (e.g., as a thickener or stabilizer for suitable rheological properties in fluid states). Among many possible etherifications, we chose benzylation as a model reaction of cellulose etherification in a [P<sub>4,4,4,4</sub>]OH aq. solution; rapid benzyl substitution is a comparatively difficult and challenging reaction in various derivatives of cellulose.<sup>15</sup> Practically, benzyl cellulose (BC) has some potential applications that have been claimed in technical patents: BC can be used for hemodialysis in medical treatments<sup>16</sup> and for commodities involving coatings, tubes, adhesives, inks, and binder compositions.<sup>17</sup>

Conventionally, cellulose etherification is done with a halogenated reagent under heterogeneous conditions using a concentrated NaOH aq. solution that swells cellulose fibers and simultaneously activates the hydroxyl groups of the polymer molecule. Gomberg *et al.* conducted benzylation of cotton and filter paper cellulose with benzyl chloride in a NaOH aq. solution at 100 °C and obtained BC products with benzyl degrees of substitution (DSs) of 0.43 and 1.41, respectively.<sup>18</sup> The molar ratio of the in-fed benzyl chloride to the anhydroglucose unit (AGU) of cellulose was 8.

Alternatively, etherification can be achieved in homogeneous cellulose solutions, often on the laboratory scale. Isogai *et al.* carried out benzylation of cellulose in LiCl/*N,N*-dimethylacetamide as a reaction medium (85 °C) mixed with powdered NaOH as an activation

agent. The addition of a large excess of benzyl chloride (45 times in mol to AGU) produced a BC with DS = 2.8.<sup>19</sup> They also attained a complete tri-*O*-benzylation of cellulose in aprotic polar solvent systems, N<sub>2</sub>O<sub>4</sub>/*N,N*-dimethyl sulfoxide (DMSO) and SO<sub>2</sub>/diethylamine/DMSO,<sup>19,20</sup> again using a high molar excess of benzyl chloride and NaOH. A more recently discovered solvent, DMSO/tetrabutylammonium fluoride trihydrate, worked in a similar way.<sup>21,22</sup> However, the trouble with these solvent systems is the long reaction time (commonly >4 h) at elevated temperatures (>70 °C). Instead of such organic solvent systems, a NaOH/urea aq. solution was recently used as another cellulose solvent and homogeneous reaction medium for benzylation.<sup>15</sup> However, lower substituted BCs (DS = 0.29–0.54) were solely obtained by a 4 h reaction at 70 °C. The large volume fraction of water (~90%) is probably an obstacle to achieve higher DS because the water would compete with cellulose hydroxyls to react with the halide reagent.

In the present study, we applied a strong base, [P<sub>4,4,4,4</sub>]OH aq. solution, to the reaction medium with the aim of achieving faster and more efficient cellulose benzylation. This solvent dissolves cellulose with an appreciably lower content of water (<50%) compared to the NaOH/urea aq. solution and therefore is expected to be more suitable for cellulose etherification.

## EXPERIMENTAL SECTION

**Original Materials.** A 40% [P<sub>4,4,4,4</sub>]OH aq. solution was purchased from Sigma-Aldrich Co. LLC and was concentrated to 47% by evaporation at 30 °C under vacuum before use. NaOH, urea, benzyl bromide (BnBr), ethanol, pyridine (dehydrated), 4-dimethylaminopyridine (DMAP), acetic anhydride (Ac<sub>2</sub>O), chloroform, toluene, tetrahydrofuran (THF), acetone, DMSO, and chloroform-*d* (CDCl<sub>3</sub>) were purchased from Sigma-Aldrich Co. LLC; all were guaranteed

reagent-grade and used as received. Microcrystalline cellulose Avicel<sup>®</sup> PH-302, purchased from FMC Europe N. V., Belgium, was dried under vacuum for 24 h at 25 °C before use.

**Synthesis of Benzyl Cellulose.** The 47% [P<sub>4,4,4,4</sub>]OH aq. solution was tentatively adopted as the benzylating reaction medium because this was the optimum concentration to dissolve cellulose.<sup>5,6</sup> Cellulose powder (0.09 g) was dissolved at 3.0 wt % in the [P<sub>4,4,4,4</sub>]OH aq. solution (2.91 g) with gentle stirring for 30 min at 25 °C. BnBr was added dropwise to the cellulose solution, and the mixture was gently stirred under a N<sub>2</sub> atmosphere. The addition of BnBr was completed within 15 s, and the reference time of the reaction was counted from the end of the addition. In the termination of the cellulose benzylation, an excess amount of ethanol (5 mL) was poured into the reaction system to consume the unreacted BnBr. The resulting BC was purified by dissolution in pyridine (1 mL) and reprecipitation in ethanol (30 mL). Finally, the purified BC was collected by filtration with ethanol washing and vacuum-dried at 25 °C for 24 h. The in-fed amount of BnBr was varied from 3 to 15 molar equivalents to AGU. The reaction temperature and time were controlled in the range of 10–50 °C and 2–180 min, respectively. We usually prepared three BC samples for a set of variables of the BnBr:AGU ratio, reaction temperature, and time to confirm the reproducibility of the DS evaluation.

**Post-acetylation of BC.** Post-acetylation was conducted to dissolve BC in CDCl<sub>3</sub> according to previous papers<sup>23,24</sup> as follows: BC (0.05 g) and DMAP (0.025 g) were dissolved in 5 mL of pyridine under a N<sub>2</sub> atmosphere without heating. Ac<sub>2</sub>O (0.75 mL) was added into the BC solution, and the mixture was gently stirred at 60 °C for 4 h. Then, 30 mL of ethanol was added to stop the acetylation reaction. The resulting acetylated BC (a-BC) precipitated as a white powder, which was purified by dissolution–reprecipitation with pyridine and ethanol. Finally, the purified a-BC was dried under vacuum at 25 °C for 24 h.

**Fourier Transform Infrared (FT-IR) Spectroscopy.** An FT-IR spectrum of the original cellulose was measured using a Shimadzu IRPrestige-21 spectrometer. The cellulose was dried at 80 °C under vacuum for 24 h and a standard KBr pellet method was employed. For BC and a-BC products, cast thin films were prepared from their 2 wt % solutions in pyridine to collect FT-IR spectra, which were recorded using a Perkin-Elmer Spectrum RXI FT-IR spectrometer.

**Nuclear Magnetic Resonance (NMR) Spectroscopy.** 400-MHz  $^1\text{H}$  NMR spectra were recorded for a-BCs in  $\text{CDCl}_3$  using a Bruker UltraShield 400 NMR spectrometer. The polymer concentration was 30 mg/mL. Tetramethylsilane was used as an internal standard. The temperature was 25 °C, and the number of scans was 16.

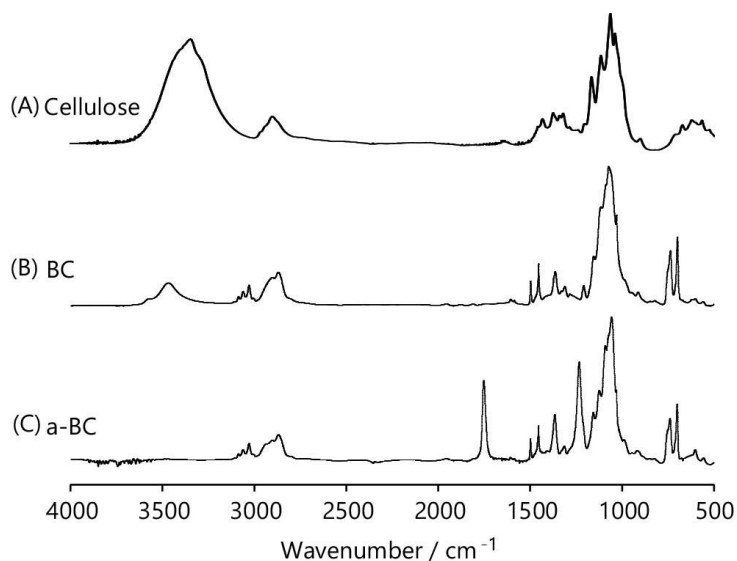
**Size Exclusion Chromatography (SEC).** SEC measurements were performed for a-BCs and the starting Avicel (nitrated according to the literature<sup>25,26</sup>) at 30 °C using an Agilent 390-MDS system (parts: 290-LC pump injector, ProStar 510 column oven, 390-MDS refractive index detector) equipped with two Agilent PLgel 5 $\mu\text{m}$  MIXED-D 300  $\times$  7.5 mm columns (part No: PL1110-6504). THF eluent was used at a flow rate of 1.0 mL/min. Each sample of cellulose derivative was dissolved in THF (5 mg/mL) at 25 °C. The solutions were filtered through a poly(tetrafluoro ethylene) syringe filter (pore size, 0.45  $\mu\text{m}$ ) before the SEC measurement.

**Solubility Test.** BC samples were vacuum-dried for 6 h at room temperature (~20 °C) before use. The dried BC powder was added to several solvents at a concentration of 10 mg/mL, and the mixtures were gently stirred for 1 h at 25 °C. The solubility was estimated by judging the amount of undissolved BC with the naked eye.

## RESULTS AND DISCUSSION

**General Aspect of Reaction and Characterization of BC Products.** When BnBr was added to the initially transparent solution of cellulose/[P<sub>4,4,4,4</sub>]OH aq., the solution became turbid. Within a few minutes, white precipitates were perceived. The floating precipitates were either fluffy or sticky, depending on the quantity of the in-fed BnBr (stickier at BnBr/AGU ratios >9). After ~10 min had elapsed in the reaction progress, the supernatant became clear. The yield of BC was 0.19 g (87%, when calculated based on AGU using the DS data of ~2.5), for example, with the following conditions: amount of starting cellulose, 0.09 g; BnBr feed, 9 molar equivalents to AGU; temperature, 25 °C; reaction time, 180 min. When the molar ratio of BnBr:AGU was changed to 6, the yield of BC (DS  $\approx$  1.7) was 0.14 g (81% in AGU).

Figure 1 illustrates FT-IR spectra measured for BC and its acetylated form (a-BC), together with a spectrum of the original cellulose. In the BC spectrum (Fig. 1B), the aromatic bands appearing at 1497 and 1454 cm<sup>-1</sup> provide evidence of successful etherification of cellulose with the benzyl substituent. The OH stretching band (3250–3750 cm<sup>-1</sup>), still observed for the BC, disappeared after the post-acetylation reaction (Fig. 1C); this confirms the full substitution of the remaining hydroxyl groups.



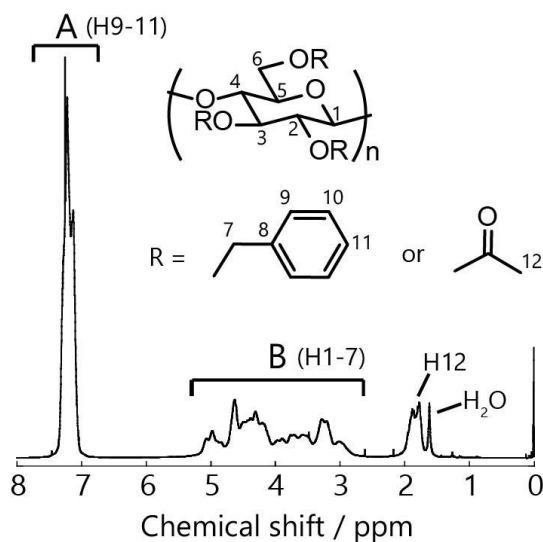
**Figure 1.** FT-IR spectra of (A) original cellulose (Avicel), (B) BC prepared from (A) in  $[P_{4,4,4,4}]\text{OH}$  aq. solution with  $\text{BnBr}$ , and (C) a-BC obtained by full acetylation of (B) with  $\text{Ac}_2\text{O}$ .

The benzyl DS of BC was determined using the  $^1\text{H}$  NMR spectrum of the corresponding a-BC sample (see Fig. 2), according to the following equation:

$$\text{Benzyl DS} = 7\mathbf{A} / (5\mathbf{B} - 2\mathbf{A}) \quad (1)$$

where **A** is the integration from 6.5 to 8.0 ppm (corresponding to the resonance intensity of the phenyl protons of the benzyl group) and **B** is the integration from 2.5 to 5.5 ppm (corresponding to the total intensity of the methylene protons of the benzyl group and the seven protons in AGU). The DS reported below is an average obtained for the evaluation of multiple similarly benzylated BCs. Integral area **A** includes a signal of  $\text{CDCl}_3$ , but this contribution was negligible to a tolerance ( $\sim 0.05$ ) of the DS quantification in the present study because BCs with high DS values ( $>1.5$ ) were mainly produced.

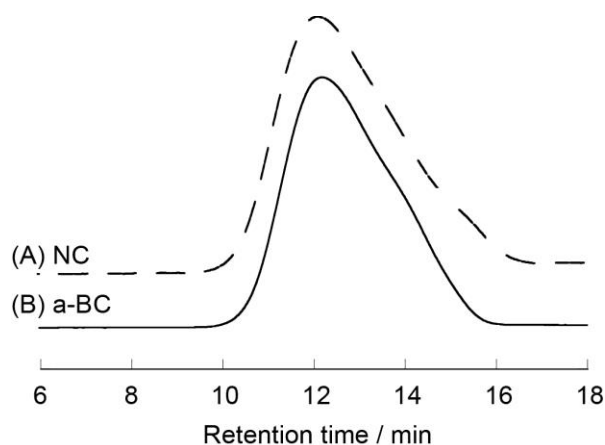




**Figure 2.**  $^1\text{H}$  NMR spectrum of an a-BC, measured in  $\text{CDCl}_3$ .

Using 9 molar equivalents of BnBr to AGU at 25 °C for 3 h, we obtained a BC product with a DS = 2.49 ( $\pm 0.06$ ). We are not aware of any reports of such an efficient benzylation at ambient temperature. Thus, the  $[\text{P}_{4,4,4,4}]\text{OH}$  aq. solution is a promising medium for etherification of cellulose. For comparison, we used a NaOH/urea aq. solution<sup>15</sup> (~90 vol % water) as a reaction medium for cellulose benzylation. Under the same conditions (BnBr:AGU = 9:1, 25 °C, and 3 h), the benzyl DS reached only 0.05, although the molarity of the alkali (NaOH) in the solution was similar to that in the  $[\text{P}_{4,4,4,4}]\text{OH}$  aq. solution. The reagent BnBr remained as a macroscopic liquid phase in the reaction system due to the immiscibility of the NaOH/urea solvent with BnBr, resulting in such a low efficiency for the benzylation. In contrast, tetraalkylphosphonium has both a hydrophilic and a hydrophobic moiety, i.e., it is of amphiphilic character. Therefore the cellulose/ $[\text{P}_{4,4,4,4}]\text{OH}$  aq. solution assumed an emulsive state immediately after the initiation of benzyl substitution by BnBr addition, whereupon further benzylation would have rapidly proceeded in the dispersed microspheres (see Fig. 6).

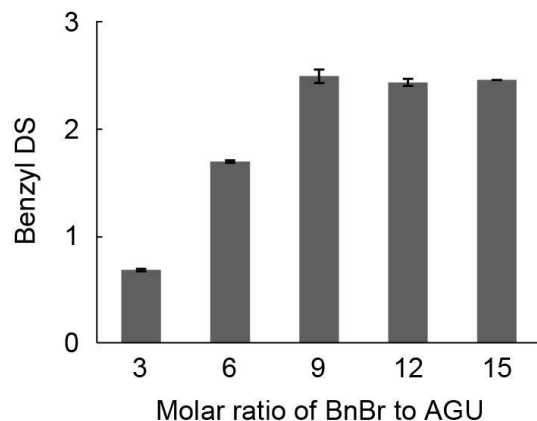
A SEC result for an a-BC (benzyl DS = 2.49) (Fig. 3B) was compared to that for a nitrated cellulose (NC) prepared from the original Avicel (Fig. 3A) to examine the potential depolymerization during the benzylation reaction. The weight-average molar mass and polydispersity of the NC were ca.  $9.6 \times 10^4$  and 3.1, respectively, when calculated based on monodisperse polystyrene. The corresponding data for the a-BC were ca.  $9.2 \times 10^4$  and 2.5. In the SEC curve of the a-BC sample, a somewhat noticeable shoulder appeared at ~14 min, which might be due to a small amount of degradation product. However, the overall elution profile was quite similar to that of the NC sample. Therefore, we can conclude that no significant depolymerization occurred in the present benzylation process. The ambient reaction temperature (25 °C for BC with DS = 2.49) might prevent degradation of the polymer chain.



**Figure 3.** SEC curve of (A) nitrated Avicel (starting cellulose) (NC) and (B) BC synthesized in  $[P_{4,4,4,4}]\text{OH}$  aq. solution with BnBr (25 °C, 3 h), then post-acetylated (a-BC).

**Effect of In-Feed Amount of BnBr.** Under fixed conditions of the reaction temperature (25 °C) and time (3 h), the DS value increased linearly with the added BnBr amount within a range of 0–9 in the molar ratio to AGU; e.g., DS = 0.68 ( $\pm 0.01$ ) and 2.49 ( $\pm 0.06$ ) for BnBr:AGU

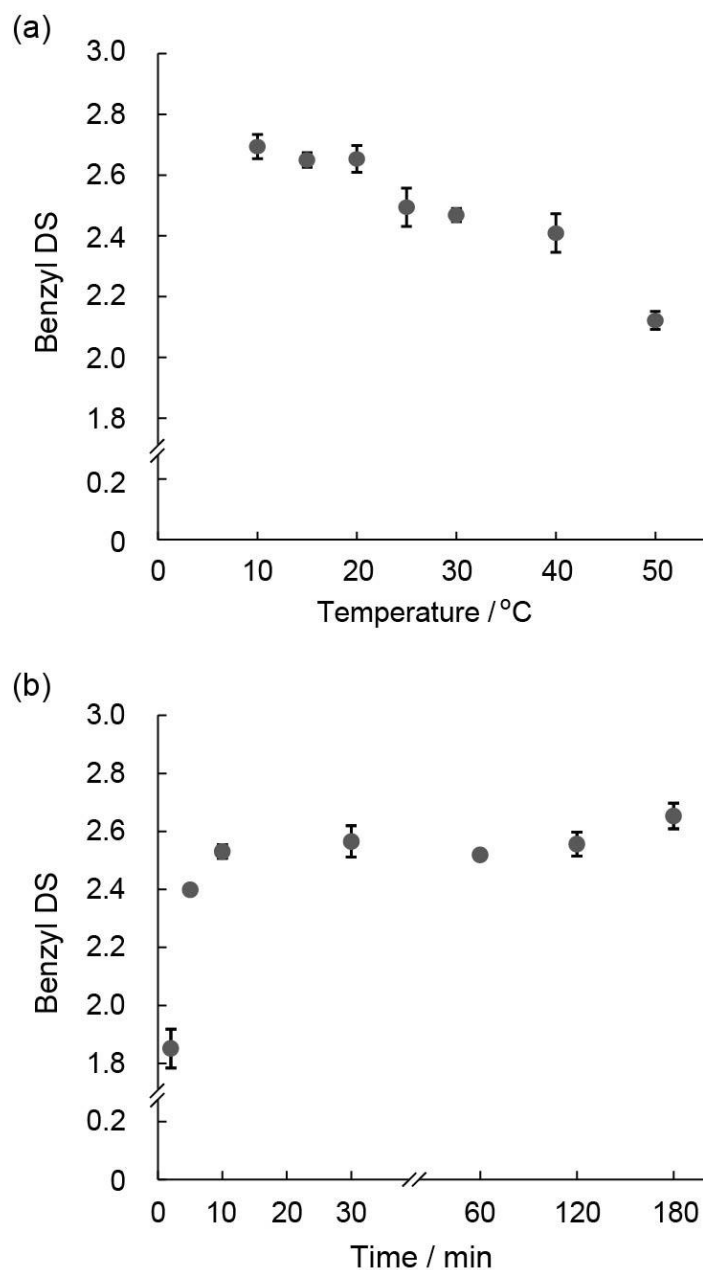
= 3 and 9, respectively (Fig. 4). However, a further increase of the molar ratio from 9 to 15 was essentially ineffective in raising the maximum DS; the reaction systems had transformed to a liquid–solid (sediment) bi-phasic one after more than 10 min without retaining the respective emulsion states (see below).



**Figure 4.** Benzyl DS of BC synthesized in  $[P_{4,4,4,4}]OH$  aq. solution with 3 h stirring at 25 °C, showing the dependence on the molar ratio of BnBr to AGU.

**Effects of Reaction Temperature and Time.** At a constant BnBr:AGU = 9:1 in-feed, the reaction temperature had a negative correlation with the benzyl DS (Fig. 5a). The highest DS, 2.69 ( $\pm 0.04$ ), was observed at 10 °C. The DS value progressively decreased with increasing temperature. We currently have two hypotheses to explain the temperature dependence: (1) The increased mobility of the hydrophobic *n*-butyl chains of the  $[P_{4,4,4,4}]$  component at elevated temperature would reduce the surface activity to disperse the microspheres of partially benzylated cellulose. Therefore, at higher temperatures, the BC polymers would flocculate and agglomerate into a compact solid sediment at a relatively earlier stage of the benzylation

progress. (2) The selectivity of the reaction of BnBr might be reduced at higher temperatures. In the present benzylation process, the bi-products, benzyl alcohol and dibenzyl ether, are formed by a side reaction of BnBr with water (see Supporting Information, Fig. S1). As generally recognized, the hydroxyl groups in polysaccharides have a weakly acidic nature.<sup>27</sup> Therefore, BnBr preferentially reacts with the cellulose hydroxyl groups. However, the selectivity might be lowered as the temperature increases, resulting in unproductive consumption of BnBr by hydrolysis. Below 10 °C, the cellulose/[P<sub>4,4,4,4</sub>]OH aq. solution was frozen and the benzylation was virtually unsuccessful.



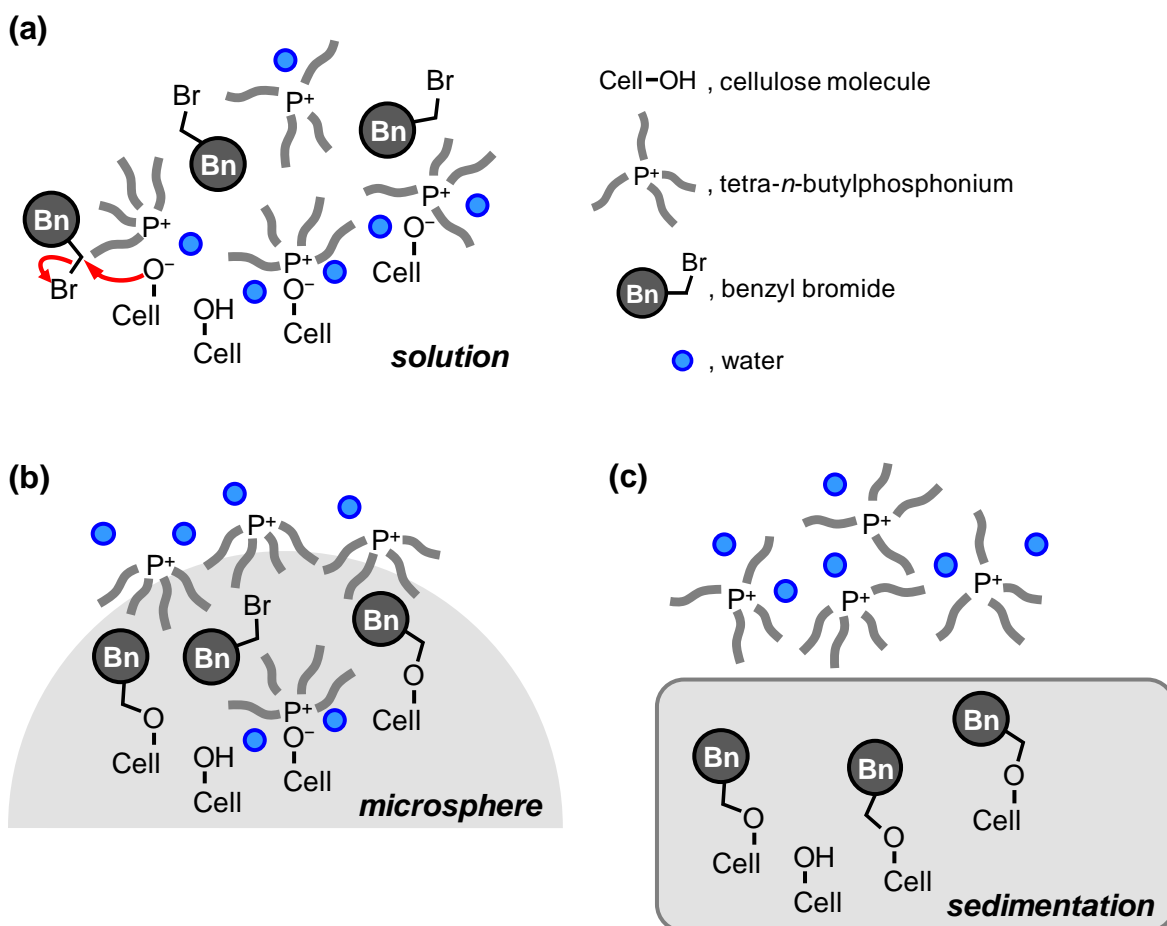
**Figure 5.** Benzyl DS of BC synthesized in  $[P_{4,4,4,4}]\text{OH}$  aq. solution with a molar ratio of  $\text{BnBr}:\text{AGU} = 9:1$  in feed: (a) the dependence on the reaction temperature with 3 h stirring, and (b) the dependence on the reaction time at 20 °C.

Figure 5b shows the time dependence of the benzyl DS of the BC products obtained at 20 °C in a fixed ratio of BnBr:AGU = 9:1. The DS was 1.92 ( $\pm 0.04$ ) within 2 min after completing the BnBr addition, and it reached 2.40 ( $\pm 0.01$ ) and 2.53 ( $\pm 0.02$ ) at 5 and 10 min, respectively. This is a very fast reaction compared to the conventional benzylation protocols that require stirring for several hours at an elevated temperature to achieve a DS value  $>2$ .<sup>19–22</sup> In the present system, however, further prolonged reaction up to 3 h had little impact on the DS growth. Therefore it follows that the cellulose benzylation substantially finished within 10 min in the [P<sub>4,4,4,4</sub>]OH aq. solution.

Eventually, the highest DS attained in this reaction system by varying the temperature, time, and BnBr in-feed amount was 2.69 ( $\pm 0.04$ ). The saturation of the benzyl substitution may be ascribed to the essentially poor compatibility of the product with the aqueous solvent; roughly, in a mild ambient condition (20–25 °C), BCs with DS  $> 2.4$  precipitate from the emulsive reaction medium so as to form a compact sedimentation.

**Model Scheme of Cellulose Benzylation in [P<sub>4,4,4,4</sub>]OH aq. Solution.** Using illustrations depicted in Figure 6, we summarize a possible scheme of the present benzylation as follows: As soon as the BnBr reagent is added into the cellulose/[P<sub>4,4,4,4</sub>]OH aq. solution, the droplets are finely dispersed because of the good lyophilic nature of the cellulose solvent. Consecutively, the BnBr reagent readily migrates to attack cellulose hydroxyls activated as alkoxide ions by the phosphonium cation; then, the benzylation reaction starts (Fig. 6a). Within the first minute of the reaction, the lightly benzylated cellulose molecules become hydrophobic to some extent and the initial BC products huddle together to form a number of microspheres in the [P<sub>4,4,4,4</sub>]OH aq.

medium, whereupon the tetra-*n*-butylphosphonium component functions as a surfactant (Fig. 6b). The benzylating BnBr is partly trapped in and can also diffuse into the lipophilic BC micelles, and hence further benzylation proceeds with fast reaction kinetics in the confined space. However, the ascent of benzyl DS (e.g., approaching 2.5) diminishes the compatibility of BC with the tetraalkyl-onium hydroxide aq. solution, which collapses the BC spheres, followed by macroscopic separation into a solution–solid bi-phasic system (Fig. 6c). In this final stage, the cellulose benzylation no longer occurs, except for a possible local reaction in the interfacial region of the two phases, because the BnBr liquid hardly penetrates into the aggregated BC sediment.



**Figure 6.** Model scheme of the benzylation of cellulose in  $[P_{4,4,4,4}]OH$  aq. solution: (a) beginning of benzylation after BnBr addition; (b) growth of benzylation in a temporarily stabilized micelle; (c) ending of benzylation accompanied by a solution–sediment separation.

**Solubility of BC in Conventional Solvents.** The solubility of the BC products in organic solvents varied depending on the attained benzyl DS, as exemplified in Table 1. BCs with relatively low DSs (e.g., BC-1 and BC-2) were only soluble in DMSO and pyridine. With an increase in the DS, BCs became soluble in chloroform and THF. Somewhat interestingly, however, the BC (BC-4) showing the highest DS of 2.69 was not soluble in DMSO. Such a tendency in the solubility of BC is, in part, different from that of the previous reports: e.g., BCs of  $DS < 1.22$  were not soluble in any of the conventional solvents,<sup>21</sup> and a DS value  $\geq 1.7$  was necessary to solubilize BC prepared in a classical heterogeneous way.<sup>28</sup> It may be taken by comparison that, in the present preparation, the solubility of BC was relatively high even at a considerably lower degree of benzylation. The difference in the degree of polymerization between the starting celluloses could be partly responsible; the previous works used cotton linter as the cellulose material, while we used a microcrystalline Avicel. The difference in the distribution patterns of the benzyl ether substituents may also influence the solubility. Further characterization of BC and other cellulose ethers synthesized in the  $[P_{4,4,4,4}]OH$  aq. medium, including specification of the substitution distribution, is under way. Additionally, a preliminary thermal analysis indicated that a highly benzylated BC ( $DS = 2.5$ ) exhibited a transition associated with an anisotropic phase above 130 °C besides the glass transition ( $T_g = \sim 90$  °C), while a BC with  $DS = 1.7$  only showed a  $T_g$  signal at  $\sim 120$  °C in a temperature range below



200 °C. In parallel with the structural characterization, detail studies on the thermal and optical properties are also now in progress.

**Table 1.** Solubility of BC samples having various benzyl DSs in conventional organic solvents.

No.	Preparation condition			Benzyl DS	Solubility <sup>b</sup> (at 10 mg/mL)					
	Feed <sup>a</sup>	Temp. (°C)	Time (min)		DMSO	Pyridine	Chloroform	THF	Acetone	Toluene
BC-1	3	25	180	0.68	S	S	I	I	I	I
BC-2	6	25	180	1.70	S	S	I	P	I	I
BC-3	9	25	180	2.49	S	S	S	S	P	P
BC-4	9	10	180	2.69	I	S	S	S	I	P

<sup>a</sup> Molar ratio of in-fed BnBr to AGU for cellulose benzylation in [P<sub>4,4,4,4</sub>]OH aq. solution. <sup>b</sup>

Notations: S, soluble; P, partially soluble; I, insoluble.

## CONCLUSION

We found a new route for cellulose benzylation without the usual activation of the natural polymer with NaOH; that is, BCs with DSs  $\geq 2.5$  were readily obtained with BnBr in a [P<sub>4,4,4,4</sub>]OH aq. solution as the reaction medium, within 10 min at room temperature (20–25 °C) via an emulsion state. The ionic solution is expected to be applicable as a good solvent for many other etherifications of cellulose.

## **ASSOCIATED CONTENT**

### **Supporting Information**

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acssuschemeng.7b00492.

<sup>13</sup>C NMR spectral data of bi-products formed in the present benzylation reaction (PDF)

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### **Author Contributions**

The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

### **Notes**

The authors declare no competing financial interest.

## **ACKNOWLEDGEMENTS**

This work was financed by a Grant-in-Aid for Scientific Research (A) (No. 26252025 to YN) from the Japan Society for the Promotion of Science. M. A. acknowledges the financial support by a Grant-in-Aid for JSPS research fellows (No. 15 J02611) and for Young Scientist Research (B) (No. 16K18731) from the Japan Society for the Promotion of Science. The authors thank Dr. I. Otsuka for the help with the SEC measurements.

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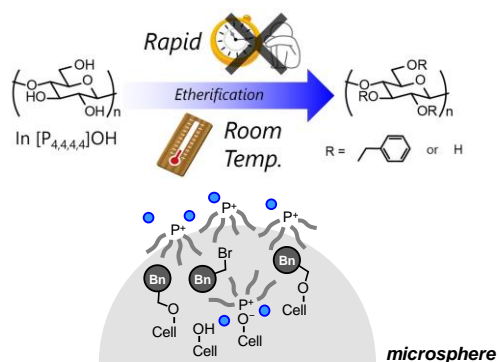
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## Graphical Abstract



Benzyl cellulose with a degree of substitution above 2.5 was obtained with benzyl bromide in a tetra-*n*-butylphosphonium hydroxide aqueous solution, within 10 min at room temperature.