High resolution NMR spectra of polyacrylonitrile and its two-unit and three-unit model compounds

Author(s): Murano, Masao

Citation: The Review of Physical Chemistry of Japan (1970), 39(2): 93-103

Issue Date: 1970-04-10

URL: http://hdl.handle.net/2433/46936

Type: Departmental Bulletin Paper

Textversion: publisher

Kyoto University
HIGH RESOLUTION NMR SPECTRA OF POLYACRYLONITRILE AND ITS TWO-UNIT AND THREE-UNIT MODEL COMPANIES

BY MASAO MURANO

The NMR spectra of PAN and PAN-ad1 were measured in NaCNS-D2O solution and nitromethane-ethylene carbonate solution.

The different spectrum was obtained in both solutions and this may be derived from the difference of the chemical shifts between the geminal protons of isotactic (meso) methylene (νmeso). But this behavior cannot be found in the model compounds. In determining the tacticity of PAN, the line width of two methylene resonance and the value of υmeso are important factors.

The relation between the NMR spectra of PAN and the model compounds is also discussed.

Introduction

The tactic structure of polyacrylonitrile (PAN) has been studied by many investigators with various methods, but it is not yet determined definitely. The high resolution NMR spectroscopy is a useful method for determining the tactic structure, so the tacticities of many high polymers have been evaluated by this method. We measured the NMR spectra of PAN and PAN-ad1 in NaCNS-D2O solution, and it was found that the tacticity in terms of diads can be determined from the methylenic proton resonance of PAN1. From the analysis of the NMR spectra of the two-unit model compounds, 2,4-dicyanopentane (2,4-DCNP), it was shown that the racemic (syndiotactic) methylene protons appear at higher field than those of meso (isotactic) methylene2. The tacticities of PAN and PAN-ad1 prepared under various polymerization conditions were determined and it was also reported that almost all PAN are rich in syndiotactic form and do not show any stereospecificity3. The tacticity is calculated, assuming that the meso methylenic protons are equivalent because the methylenic proton resonance of PAN-ad1 shows only two peaks and the methylenic proton resonance of PAN can be analyzed as the superposition of two kinds of triplets2,3).

On the other hand, Matsuzaki et al.4) measured the NMR spectrum of PAN in DMF-d7 solution and reported that PAN has a completely random configuration because PAN-ad1 shows two peaks with equal intensity.

Bargon et al.5) reported from the NMR spectrum of PAN in nitromethane-ethylene carbonate

(Received August 15, 1969)
1) R. Yamadera and M. Murano, J. Polymer Sci., B3, 821 (1965)
solution, that PAN is rich in isotactic form. In these three solutions, the isotactic methylenic protons of PAN are thought to be equivalent.

Recently, Yoshino et al. measured the NMR spectrum of PAN-α, β-d₃ in perchloric acid at 100 Mc/sec and reported the non-equivalence of isotactic methylene. As mentioned above, many different theories have been presented about the interpretation for the NMR spectrum of PAN but they are not interpreted comprehensively compared with each other, because the solvents and the samples used in these reports are different from each other.

In this report, the difference in the NMR spectra of PAN measured in various solutions using the same sample are discussed in a comprehensive interpretation about the NMR spectrum of PAN. In addition, the NMR spectra of the two-unit model (2,4-DCNP) and the three-unit model compounds, 2,4,6-tricyanoheptane (2,4,6-TCNH), are also analyzed to be compared with those of PAN.

Experimental

Polymerization

Acrylonitrile (AN) and α-deuteroacrylonitrile (AN-αd₃) were polymerized in benzene solution at 60°C with azobisisobutyronitrile as an initiator. PAN was also obtained with γ-irradiation on AN-urea canal complex at −78°C.

Preparation of model compounds

The preparation of 2,4-DCNP and 2,4,6-TCNH and the separation into stereoisomers were reported in the previous papers.

NMR spectrum

The NMR spectra of PAN and PAN-αd₃ were measured at various temperatures in 50% D₂O solution of NaCNS (NaCNS-D₂O) and in nitromethane-ethylene carbonate (1 : 1) mixed solvent.

Varian A-60 and HR-60 spectrometers were used, operating at 60 Mc/sec. The NMR spectra of the model compounds were measured at various temperatures in NaCNS-D₂O, and nitromethane-ethylene carbonate. The double resonance spectra with the side band method were also measured.

Results and Discussion

NMR spectra of PAN and PAN-αd₃

The NMR spectra of PAN in NaCNS-D₂O solution and in nitromethane-ethylene carbonate solution were compared in Fig. 1. The multiplet at higher field is assigned to the methylenic protons and the multiplet at lower field to the methine protons.

The methylenic proton resonance shows mainly six peaks in both solvents and can be thought to be superposition of two kinds of triplets. But the relative intensity of them is different in each

6) T. Yoshiao, H. Kenjo and N. Hisano, J. Polymer Sci., B5, 703 (1967)
7) M. Murano and R. Yamadera, J. Polymer Sci., A-1, 6, 843 (1968)
solvent, so the tacticity in terms of diads cannot be determined by only comparing the height of their peaks. This fact can be seen more definitely in the NMR spectra of PAN-ad shown in Fig. 2. In NaCNS-D$_2$O solution, methylenic proton signal consists of two singlets with different intensity, while in nitromethane-ethylenecarbonate, two peaks are observed with almost equal intensity. Furthermore, in the case of nitromethane-ethylenecarbonate solution, the relative intensity of these two peaks depends upon temperature. The peak at higher field rapidly decreases its intensity as temperature decreases (Fig. 3), although this phenomenon was not found in NaCNS-D$_2$O solution.

NMR spectra of model compounds

The NMR spectra of model compounds of PAN have been already reported$^{9,17}$. In this report, another part of them is discussed. Fig. 4 shows the NMR spectra of the racemic form and meso form of 2,4-DCNP in NaCNS-D$_2$O solution and Fig. 5 those in nitromethane-ethylenecarbonate. In Table 1 the coupling constants and chemical shifts obtained from the analysis of these spectra are summarized. In Fig. 6 the NMR spectra of heterotactic 2,4,6-TCNH are shown, and the results obtained from the analysis of 2,4,6-TCNH are shown in Tables 2 and 3. Comparing the two kinds of
Fig. 4 NMR spectra (a) racemic 2,4-DCNP (b) meso 2,4-DCNP at 60 Mc/sec in NaCNS-D$_2$O at 25°C

Fig. 5 NMR spectra (a) racemic 2,4-DCNP (b) meso 2,4-DCNP at 60 Mc/sec in CD$_3$NO$_2$-(CH$_2$O)$_2$CO at 25°C

Fig. 6 NMR spectra of heterolactic 2,4,6-TCNH at 60 Mc/sec (a) in NaCNS-D$_2$O (b) in CD$_3$NO$_2$-(CH$_2$O)$_2$CO at 25°C
### Table 1 Chemical shifts and coupling constants of 2,4-DCNP (NaCNS-D$_2$O)

<table>
<thead>
<tr>
<th>Isomers</th>
<th>( \text{CH}_3 )</th>
<th>( \text{CH}_2 )</th>
<th>( \nu_{KL} )</th>
<th>( \text{CH} )</th>
<th>( \text{CH}_3-\text{CH} )</th>
<th>( \text{CH}_2-\text{CH} )</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td><strong>meso</strong></td>
<td>8.40</td>
<td>7.77</td>
<td>7.56</td>
<td>7.88</td>
<td>0.22</td>
<td>6.74</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td><strong>racemic</strong></td>
<td>8.40</td>
<td>7.82</td>
<td>0</td>
<td>6.72</td>
<td>7.5</td>
<td>13.5</td>
</tr>
</tbody>
</table>

### Table 2 Chemical shifts and coupling constants of heterotactic 2,4,6-TCNH (NaCNS-D$_2$O)

<table>
<thead>
<tr>
<th>Isomers</th>
<th>( \text{CH}_3 )</th>
<th>( \text{CH}_2 )</th>
<th>( \nu_{KL} )</th>
<th>( \text{CH} )</th>
<th>( \text{CH}_3-\text{CH} )</th>
<th>( \text{CH}_2-\text{CH} )</th>
</tr>
</thead>
<tbody>
<tr>
<td><strong>meso part</strong></td>
<td>8.26</td>
<td>7.51</td>
<td>7.42</td>
<td>7.01</td>
<td>0.10</td>
<td>6.50</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td>6.52</td>
<td></td>
</tr>
<tr>
<td><strong>racemic part</strong></td>
<td>7.55</td>
<td>0</td>
<td>6.51</td>
<td>7.5</td>
<td>13.5</td>
<td>7.5</td>
</tr>
</tbody>
</table>

* The difference in chemical shifts between two protons in methylene resonance

### Table 3 Chemical shifts and coupling constants of heterotactic 2,4,6-TCNH (CD$_3$NO$_2$-(CH$_2$O)$_3$CO)

<table>
<thead>
<tr>
<th>Isomers</th>
<th>( \text{CH}_3 )</th>
<th>( \text{CH}_2 )</th>
<th>( \nu_{KL} )</th>
<th>( \text{CH} )</th>
<th>( \text{CH}_3-\text{CH} )</th>
<th>( \text{CH}_2-\text{CH} )</th>
</tr>
</thead>
<tbody>
<tr>
<td><strong>meso part</strong></td>
<td>8.51</td>
<td>7.81</td>
<td>7.78</td>
<td>7.87</td>
<td>0.19</td>
<td>6.90</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td>6.92</td>
<td></td>
</tr>
<tr>
<td><strong>racemic part</strong></td>
<td>7.95</td>
<td>7.92</td>
<td>0.07</td>
<td>6.99</td>
<td>7.3</td>
<td>14.2</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td>7.5</td>
<td>7.5</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td>7.0</td>
<td>6.9</td>
</tr>
</tbody>
</table>

<table>
<thead>
<tr>
<th>Isomers</th>
<th>( \text{CH}_3 )</th>
<th>( \text{CH}_2 )</th>
<th>( \nu_{KL} )</th>
<th>( \text{CH} )</th>
<th>( \text{CH}_3-\text{CH} )</th>
<th>( \text{CH}_2-\text{CH} )</th>
</tr>
</thead>
<tbody>
<tr>
<td><strong>meso part</strong></td>
<td>8.51</td>
<td>7.81</td>
<td>7.78</td>
<td>7.97</td>
<td>0.19</td>
<td>6.90</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td>6.92</td>
<td></td>
</tr>
<tr>
<td><strong>racemic part</strong></td>
<td>7.95</td>
<td>7.92</td>
<td>0.07</td>
<td>6.99</td>
<td>7.3</td>
<td>14.2</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td>7.5</td>
<td>7.5</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td>7.0</td>
<td>6.9</td>
</tr>
</tbody>
</table>

* The difference in chemical shifts between two protons in methylene resonance
Table 3 Chemical shifts and coupling constants of 2,4,6-TCNH
(NaOCH - D$_2$O)

<table>
<thead>
<tr>
<th>Isomers</th>
<th>Chemical shifts (ppm)</th>
<th>Coupling constants (cps)</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>CH$_3$</td>
<td>CH$_2$</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>syndio</td>
<td>8.26</td>
<td>7.55</td>
</tr>
<tr>
<td>hetero</td>
<td>8.26</td>
<td>7.55</td>
</tr>
<tr>
<td>iso</td>
<td>8.26</td>
<td>7.52</td>
</tr>
</tbody>
</table>

methylene in 2,4-DCNP, the racemic form appears higher than meso form (Table 1) in both solvents. The heterotactic-2,4,6-TCNH shows the same (Table 2). From this fact, the two kinds of methylene in PAN can be assigned as follows: the higher field peak to racemic methylene and the lower to meso one. In meso 2,4-DCNP, no remarkable difference can be found between both solvents, while the racemic one shows different behavior in both solvents. In NaOCH - D$_2$O, racemic 2,4-DCNP shows a spectrum which means the complete equivalence of methylenic protons ($\nu_{KL}=0$), and it does not depend on temperature. But in nitromethane-ethylene carbonate solution, it shows a nonequivalent spectrum ($\nu_{KL} \neq 0$) and this tendency is clearer at low temperature. At high temperature it becomes equivalent (Fig. 7). The same tendency is found in DMF solution. This fact may mean that the conformation of the racemic form is fixed to some extent in these solvents at low temperature.

This phenomenon can be found in the racemic part of heterotactic 2,4,6-TCNH.

The chemical shifts of hydrogen attached to 4th position in 2,4,6-TCNH, which can be thought to represent the configuration of triads in PAN, are shown in Table 3. The methine protons show the isotactic-CH, heterotactic-CH and syndiotactic-CH resonance, respectively, from the higher field. The difference in chemical shifts between geminal protons of isotactic methylene ($\nu_{KL}$) takes smaller value in isotactic 2,4,6-TCNH than in meso 2,4-DCNP. This fact may be important to analyze

![Fig. 7 Methylenic proton resonance of racemic 2,4-DCNP at 60 Mc/s/sec in CD$_3$NO$_2$-(CH$_2$)$_2$CO](a) at 25°C (b) at 120°C
the NMR spectrum of PAN. This tendency is found between the two-unit and three-unit model compounds of polyvinylchloride.

**NMR line width**

As shown in Fig. 2, the NMR spectrum of PAN-ads, in nitromethane-ethylene carbonate has two peaks and the line width of the *racemic* methylene resonance at higher field is larger than that of the *meso* methylene resonance at lower field. This difference in line width becomes more definite at low temperature. The same difference between two kinds of methylene has been found in polyvinylchloride and polymethylmethacrylate, and it has been reported that there is a difference in the mobility of solution between the isotactic segment and the syndiotactic segment and the peak of the syndiotactic part shows broader than that of the isotactic part. This phenomenon is not so remarkable in NaCNS-D$_2$O solution. As mentioned above, the model compound shows the same behavior: the methylene protons of racemic 2,4-DCNP and the racemic part in 2,4,6-TCNH show the non-equivalent resonance in nitromethane-ethylene carbonate solution, while they show the equivalent resonance in NaCNS-D$_2$O solution. From these facts, the difference in line width between the isotactic methylene and the syndiotactic one may be attributed to the conformation of PAN in solution.

**Tactility of PAN**

The tacticity in terms of diads is determined from the ratio of the area of the *meso* methylene resonance and the *racemic* methylene resonance. The line width is an important factor in comparing the area. Changing the value of the line width, the spectrum of PAN-ads was constructed by superposing the syndiotactic methylene upon the isotactic one in order to get a good agreement with the observed spectrum. As shown in Fig. 8, the tacticity is greatly changed by slight change in line width. Fig. 8(b) shows the best agreement with the observed spectrum. The same superposition was tried for PAN and in Fig. 9(b) the best agreement was obtained.

In the previous reports, the tacticity was determined assuming that the isotactic methylene is equivalent in NaCNS-D$_2$O solution. In this report, the tacticity is obtained with the assumption.

---

that the isotactic methylene is non-equivalent. The observed spectrum shows six peaks in the methylene resonance. This spectrum can be obtained by superposing the six main peaks of isotactic methylene upon the triplet of syndiotactic methylene as well as by superposing two kinds of triplet. (Fig. 10) The methylene resonance of PAN-ad₄ is a superposition of a quartet upon a singlet if the isotactic methylene is non-equivalent. In this case, one of the main two peaks of isotactic methylene has almost the same chemical shifts as that of syndiotactic methylene, so the spectrum mainly consists of two peaks. (Fig. 11)

The difference between the two spectra assuming equivalence and non-equivalence of isotactic methylene is small peaks appeared both side of the main peaks. In the observed spectra of PAN and PAN-ad₄ in NaCNS-D₄O solution, four and two small peaks can be found, and these peaks become larger in 100 Mc/sec (Fig. 12). νₓₓ of isotactic methylene can be calculated from these small peaks to be νₓₓ = 9 at 60 Mc/sec and 15 cps at 100 Mc/sec. This value is almost the same as that obtained from meso 2,4-DCNP, but pretty larger than that of isotactic 2,4,6-TCNH. But the detection of these peaks is very difficult and there may remain some problem in determining whether these small peaks represent the properties of meso or racemate form.
peaks are assigned to the isotactic methylene or to the impurities such as heterolinkages.

Although the small peaks are also found on both sides of the main peaks in nitromethanee-
ethylene carbonate solution (Figs. 1(b) and 2(b)), if one takes the assumption that the isotactic me-
thylene is equivalent in this solution and non-equivalent in NaCNS-D$_2$O solution, the resultant tactic-
city is in good agreement in both solutions. But this assumption cannot be confirmed from the analy-
ysis of the model compounds. (Table 1)

As shown in Fig. 13, the methine proton resonance decoupled from the methylenic protons con-
sists of three peaks and these peaks may represent the configuration in terms of triads. From the

---

**Fig. 11** The superposition of two kinds of methylene assuming (a) the equivalence of isotactic methylene ($b=0.025$ ppm, $s=75:75$) and (b) non-equivalence of isotactic methylene ($b=0.025$ ppm, $s=45:55$) to get a good agreement with the observed methylene resonance of PAN-ads in NaCNS-D$_2$O

**Fig. 12** NMR spectrum of PAN-ads at 100 Mc/sec in NaCNS-
D$_2$O at 120°C

**Fig. 13** The methine proton resonance of PAN at 60 Mc/sec decoupled from the methylenic protons in CD$_3$NO$_2$-(CH$_2$O)$_2$CO at 150°C
results of 2,4,6-TCNH (Table 3), these three peaks may be assigned to syndiotactic-CH, hetero-
tactic-CH and isotactic-CH from the lower field. This spectrum shows that PAN is syndiotactic rich
polymer.

Polyvinylchloride is one of the polymers studied by the NMR technique to determine the
tactility, and there have been reported many opinions about the spectrum of the isotactic methylene.
The methylene resonance decoupled from the methine protons\(^\text{11}\) and the analysis of the three-unit
model compounds\(^\text{12}\) show that the isotactic methylene is to be equivalent, but recently the 100 Mc/sec
spectrum has been shown to be non-equivalent by Jobensen et al.\(^\text{12}\) who first reported that the iso-
tactic methylene is equivalent. Thus, there remains some problem about this point. In the case of
PAN, some detailed studies will be necessary to solve this problem.

**NMR spectra of PAN obtained from other conditions**

The study of the NMR spectra of PAN which may have different tacticity will give us a new
information about the analysis of the NMR spectra. For this object, PAN is obtained with γ-irra-
diation on AN-urea canal complexes. Fig. 14 shows the NMR spectra of this polymer.

![Fig. 14](image)

In NaCNS-D\(_2\)O solution (Fig. 14(a)), the spectrum consists of two kinds of triplets with equal
intensity. Assuming the equivalence of isotactic methylene, one can get 50/50 for the value of the
ratio syndio/isotactic. While taking the non-equivalence of isotactic methylene, the two kinds of trip-
lets are assigned to the isotactic methylene, and so this polymer is almost isotactic, which is in good
agreement with the result reported by Yoshino et al.\(^\text{13}\)

In nitromethane-ethylene carbonate solution (Fig. 14(b)), the different spectrum was obtained. If
one assumes that the isotactic methylene is equivalent in this solution, this spectrum shows that this
PAN is isotactic rich.

\(\text{11) F. A. Bovey, E. W. Anderson, D. C. Douglass and J. M. Mason, } \textit{J. Chem. Phys.}, \textbf{39}, 1199 (1963); S. Satoh, } \textit{J. Polymer Sci.}, \textbf{A, 2}, 5221 (1964)\)

If this polymer is isotactic rich, some different behaviors from the ordinary PAN may be expected in the infrared spectrum, x-ray diffraction pattern and density. And these studies may be important to determine the tactic structure of PAN.

Acknowledgment

The author wishes to express his sincere thanks to Dr. R. Yamadera of this Institute and Prof. H. Hatano of Kyoto University for valuable discussion and encouragement of this research.

Kataha Research Institute
TOYOBO Company Ltd.
Hokkaido
Otsu, Japan