<table>
<thead>
<tr>
<th>Title</th>
<th>High Pressure luminescence (Modern aspects of physical chemistry at high pressure: the 50th commemorative volume)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Author(s)</td>
<td>Drickamer, H. G.</td>
</tr>
<tr>
<td>Citation</td>
<td>The Review of Physical Chemistry of Japan (1980), 50: 1-18</td>
</tr>
<tr>
<td>Issue Date</td>
<td>1980</td>
</tr>
<tr>
<td>URL</td>
<td><a href="http://hdl.handle.net/2433/47101">http://hdl.handle.net/2433/47101</a></td>
</tr>
<tr>
<td>Right</td>
<td></td>
</tr>
<tr>
<td>Type</td>
<td>Departmental Bulletin Paper</td>
</tr>
<tr>
<td>Textversion</td>
<td>publisher</td>
</tr>
<tr>
<td>Textversion</td>
<td>Kyoto University</td>
</tr>
</tbody>
</table>
HIGH PRESSURE LUMINESCENCE*

By H. G. Drickamer

In this review we discuss high pressure luminescence in liquids and polymer films to 10-12 kbar pressure, as well as studies in polymer films and crystalline solids to over 100 kbar. We use these results to characterize excitations and to test theories concerning radiative and non-radiative processes.

The organic systems discussed include:
(a) several polyenes where a theory relating the effective Stokes shift to the radiative rate is tested;
(b) several dyes in liquid solution where a theory relating viscosity to luminescence efficiency is tested;
(c) polymer films involving polyvinylcarbazole, where the two excimer states are characterized.

The inorganic systems include:
(a) ZnS with appropriate dopants, where a theory of the luminescence efficiency is tested;
(b) Europium doped oxysulfides where a model describing emission from the various excited states is tested.

1. Introduction

Over the past few years high pressure luminescence studies have proved to be very effective for characterizing electronic study and electronic processes in condensed systems. A number of factors have contributed to this development including photon counters for detecting very low light levels, and very intense light sources as well as sources and detection systems permitting measurements on time scales faster than a nanosecond, and sophisticated methods for transforming and analyzing data to extract lineshapes, intensities, and lifetimes.

In this review it is my intention to present a few selected examples to illustrate applications to organic molecules, in liquid and polymeric solutions and to inorganic crystalline solids.

For localized excitations many of the features are illustrated effectively in terms of a configuration coordinate diagram (Fig. 1). The configuration coordinate can be any normal mode of motion of the system, but pressure generally couples most strongly to the totally symmetric (breathing) motion which is simply related to the volume. For molecular systems one will primarily be concerned with an intermolecular coordinate. The ground and first

(Received June 9, 1980)

*This work was supported in part by the U.S. Department of Energy under Contract DE-AC02-76-ER01198.
allowed excited electronic states are represented by potential wells (in first order, harmonic). These will, in general, have different displacements along the coordinate and possibly different force constants. Optical absorptions and emissions are vertical (Franck-Condon) processes. The horizontal lines represent vibrational sublevels of the electronic states. An excited electron may emit vibrational energy until it reaches the bottom of the (excited state) well and fluoresce. Alternatively, it may return thermally to the ground state (internal conversion, (IC)); classically, by surmounting the energy barrier, or quantum mechanically through overlap of vibrational wave functions. The relative rates of the optical and thermal processes determine, in first order, the fluorescence efficiency.

On the other hand, the electron may cross over by a similar radiationless process to an optically forbidden (e. g. a triplet) state. This intersystem crossing (ISC) can also quench fluorescence. From the triplet state, it may phosphoresce, or return thermally to the ground state. Since phosphorescence is generally slow, $10^{-6}$–$10^{-4}$ seconds, only for heavy atoms where spin-orbit coupling is large is there significant phosphorescence at room temperature.

The simplest effect of pressure is to displace the potential wells with respect to one another, vertically and/or horizontally. The first order effect of this displacement is to change the rates of the radiationless processes and thus to modify the quantum efficiency. It is also possible to perturb the shape of the potential wells. For a pure harmonic oscillator the force constant is independent of volume, but there is always some anharmonicity present. In principle, one can include anharmonic terms, but it is difficult to obtain sufficient data for accurate evaluation and anharmonic coupling terms make the normal mode analysis less valid. In any case, the single configuration coordinate model is best adapted for qualitative or simple first order quantitative treatment of data.

A simple conclusion from Fig. 1 is that generally a decrease in the energy difference between the first excited state and the ground state should result in an increase in the rate of intersystem crossing and a decrease in phosphor efficiency and that an increase in the
energy difference should have the inverse effect. This phenomenon is conveniently illustrated by a study of an organic molecule (a derivative of azulene) in a rigid polymeric matrix. The system is selected because emission is observed from two excited states, one ($S_2$) distinctly more polarizable than the ground state, and the other ($S_1$) less polarizable than the ground state. As can be seen from Fig. 2a and 2b, the emission from $S_1$ shifts to lower energy by $\sim 1500 \text{cm}^{-1}$ ($\sim 0.2 \text{ eV}$) in 100 kbar and the efficiency drops by a factor of fifty; on the other hand, the emission from $S_2$ increases in energy by about 0.2 eV in 100 kbar and the efficiency increases by a factor of $\sim 40-50$.

2. Polyenes

The linear conjugated polyenes have been widely studied as prototypes for molecules of biological significance. We discuss here the effect of pressure on the radiative rate parameters for fluorescence for 1,6-diphenyl-1,3,5-hexatriene (DPH) and 1,8-diphenyl-octatetraene (DPO) in liquid solution. We mention also studies of these molecules plus retinyl acetate in polymer films.

Hudson and co-workers have shown that for many polyenes the positions of 0-0 absorption and fluorescence bands in solution are given by an expression of the form:

$$\nu = \nu_0 - \alpha L,$$

(1)

where $\nu_0$ is the energy of the transition in the gas phase, $\alpha$ is the solvent polarizability, $(n'^2-1)/(n'^2+2)$, $n$ is the solvent index of refraction, and $L$ is a constant related to the oscillator strength of the transition. The strong absorption occurs in solution 2000–3000 cm$^{-1}$ to the red.
of its gas phase position giving a value of \( L \) of about \( 10^4 \) \( \text{cm}^{-1} \) per unit change in \( \alpha \). Values of \( L \) obtained from plots of fluorescence frequency vs. \( \alpha \) are typically a factor of 5-20 smaller. Therefore as \( \alpha \) is increased, the energy gap \( \Delta E \) between the allowed \( *B_1(S_0) \) and forbidden \( *A_1(S_1) \) state rapidly decreases. Since this gap is only 1000-2000 \( \text{cm}^{-1} \) in the lower polyenes, a significant perturbation can be introduced by fairly small changes in \( \alpha \).

Andrews and Hudson have used a perturbation theory approach to derive an expression for the solvent dependence of the radiative rate of a forbidden transition in the presence of vibronic mixing. Their expression has the form:

\[
k_f = \frac{9n^3}{(n^2+2)^2} \frac{\Gamma^2}{(\Delta E)^2}.
\]

where \( \Gamma \) is constant for a given molecule and \( \Delta E \) is the energy difference between the lowest vibrational level of the emitting \( A_1 \) excited state and a value obtained from a weighted integration of the \( B_1 \) absorption spectrum over its vibronic levels. As mentioned above, \( \Delta E \) is a strong function of solvent \( \alpha \) so that Eq. (2) predicts a radiative rate which is very sensitive to solvent polarizability. Analysis of radiative rate data for DPH and trans-retinol in several solvents and over a large temperature range gave quantitative agreement with theory.

Birks and co-workers have used a similar expression:

\[
k_f = \frac{k_{f_o}V^2}{(\Delta E_{0-0})^2} = \frac{n^2k_{f_o}V^2}{(\Delta E_{0-0})^2}
\]

to describe the solvent and temperature effect on the \( S_1 \rightarrow S_0 \) radiative rate \( k_f \) in DPH and DPO. In this expression \( \Delta E_{0-0} \) is the energy difference between the 0-0 bands of absorption and emission and \( V \) is the matrix element coupling the \( S_1 \rightarrow S_0 \) and \( S_2 \rightarrow S_0 \) transitions. \( k_{f_o} \) is the \( S_2 \rightarrow S_0 \) radiative rate in cyclohexane obtained from the experimental absorption spectrum via the Birks-Dyson relation. \( k_{f_o}(=k_f/n^2) \) has been observed to be independent of solvent and temperature for many organic solutes in several solvents. Plots of \( k_f \) vs. \( (n/\Delta E_{0-0})^2 \) were found to be linear giving \( V \) values of 745 \( \text{cm}^{-1} \) for DPH and 500 \( \text{cm}^{-1} \) for DPO.

The peak positions obtained by computer fitting of a skewed Gaussian curve to the

![Fig. 3. Plot of energy of absorption centroid vs. solvent \( \alpha \) for DPH, DPO, and RAc in various media.](image-url)
corrected excitation spectra are plotted vs. solvent $\alpha$ in Fig. 3. The plots are fairly linear (correlation coefficient $>0.99$) in accordance with Eq. (1). Slopes are 14600 for DPH and DPO, and 29300 for RAc. These compare to slopes of 10600 for DPH and 10909 for octatetraene[10] from solvent and temperature variation studies. A slope of about 6500 for RAc can be obtained from data in Ref. (4). The reason for the very large pressure induced shifts in RAc is unknown.

The value of $C$ as determined from the center of the area under the excitation spectrum appeared to be the same for a given compound in any medium, so in Fig. 4 we plot $k^{-1/2}$ vs. $J E_{g-o}$. Figure 4 shows data for DPH and DPO in the liquid cell all to 10 kbar except for DPH in $C_6F_{14}$ (which freezes above 1.5 kbar). Over much of the experimental range, the data exhibit the linear dependence on $J E_{g-o}$ predicted by Eq. (3). For DPH in toluene the data points for which $J E_{g-o}<1000$ cm$^{-1}$ were disregarded in determining the linear fits shown. As $J E_{g-o}$ becomes small ($<1000$ cm$^{-1}$) the simple first order intensity borrowing expression begins to fail. This is to be expected as the experimental radiative rate approaches the rate calculated from the $B \leftarrow A$ absorption intensities. $\Gamma$ values calculated from the data are $1.0 \times 10^7$ cm$^{-1}$ sec$^{-1/2}$ for both DPH and DPO. This compares to a value of $0.8 \times 10^7$ cm$^{-1}$ sec$^{-1/2}$ for DPH given in Ref. (5).

The studies in polymeric media to 40 kbar also gave linear fits for plots like Fig. 4, for DPO and DPH with values of $\Gamma$ equal to $6 \times 10^7$ and $1.2 \times 10^7$ cm$^{-1}$ sec$^{-1/2}$, respectively. For retinyl acetate the model fails, possibly because there are $n^*$ states lying between the $^*B$ state to which the allowed absorption takes place and the $^*A$ state from which the emission occurs.
3. Viscosity Effects on Luminescence Efficiency

Another example of the use of high pressure concerns the viscosity dependent quenching of the fluorescence of the triphenylmethane dye crystal violet (CV) and the diphenylmethane dye auramine O (AO) in alcoholic solution.

Studies of viscosity dependent processes in fluid media are usually carried out by varying the composition or the temperature of the solvent. The problem then arises of separating the viscosity dependence from purely temperature and/or solvent effects. The use of pressure allows a significant range of viscosities to be attained in a single solvent at a single temperature. If a wider range of viscosities is desired, chemically similar solvents for which the attainable viscosities overlap can be utilized. In this way very large viscosity range can be investigated at one temperature with only a few solvents.

In this example pressures to 11 kbar were used on methanol, iso-propanol, iso-butanol, and glycerol to obtain solvent viscosities from less than $10^{-2}$ poise to more than $10^{3}$ poise at room temperature. The solute concentrations were in the range 2-4×10^{-4} M.

Forster and Hoffman (FH) have investigated the fluorescence efficiency of several triphenylmethane dyes including CV in a variety of solvents at different temperatures. FH propose a model to explain the observed dependence of $\phi$ on $\eta$ in which the dye molecule is excited to a Franck-Condon vertical state with the phenyl rings still at the ground state equilibrium angle $\theta_0$. The rings then rotate towards a new equilibrium angle $\theta_0'$ at a rate controlled by Stokes-like viscous damping. $(\theta-\theta_0')$ therefore decreases exponentially with a relaxation time proportional to $\tau$. The non-radiative viscosity dependent deactivation rate of the excited single is taken to be proportional to $(\theta-\theta_0)^2$. There is also a non-radiative viscosity independent deactivation which accounts for the limiting value of $\phi$ in highly viscous media. In solvents of very low viscosity, the model predicts a viscosity independent minimum value for $\phi$. This limiting case was not observed.

The data for CV are shown in log-log form in Fig. 5. The AO data has been displaced along the arbitrary intensity axis to show that for $\eta > 0.5$ the two dyes exhibit identical...

![Fig. 5. Log luminescence efficiency vs. log viscosity ($\eta$).]
dependence of fluorescence efficiency on solvent viscosity. The dashed line in Fig. 5 indicates
the deviation of the AO intensities at low $\eta$. The line of slope 0.7 drawn in Fig. 5
shows excellent agreement with the CV data over 3.5 orders of magnitude in viscosity. At
very low $\eta$ there is a small but definite deviation from linearity and at high $\eta$ the intensities
begin to level off. The AO data is also linear with a slope of 0.7 over a smaller range of
2.5 orders of magnitude in viscosity. It exhibits identical limiting behavior at high $\eta$, but
much larger deviations from linearity at low $\eta$. It should be noted that the deviation at low
viscosity is towards a larger viscosity dependence of the fluorescence intensity. The FH
theory predicts a leveling, i.e., a lower viscosity dependence at low viscosities. We do not at
present have an explanation for this discrepancy.

4. Excitations in Polymers

Polyvinylcarbazole is a polymer of much theoretical and practical interest and its optical
properties have been widely studied. An important feature of the emission spectrum is the
presence of two emission peaks assigned to two different excimer emissions in the regions
22.5-23.5 kK and 25-26 kK.

In this study, different films of PVCA in dilute solution with another polymer and a
film of pure PVCA were subjected to pressures up to 40 kbar, and changes in the emission
characteristics with pressure were observed and recorded. The data included peak location, the
total emission intensity as well as the relative integrated intensity of each excimer emission,
and the lifetime of such excimer as a function of pressure. The peaks shifted significantly
to lower energy with pressure. The total intensity behaves similarly for PVCA in PMMA and
in PS, as well as for the PVCA film. There is an initial small increase in intensity, followed
by a decrease at higher pressures. The only differences are in the magnitude of the increase,
and in the pressure at which the maximum occurs. The behavior of PVCA in PIB is similar
at high pressures, but there is an initial decrease, followed by a rapid increase initiating at
about 5 kbars, which is not present in the other films.

The ratio of $I_{02}/I_{01}$ was calculated and is shown in Figs. 6 and 7, for the low and
high pressure regions respectively. The semilogarithmic plots show two regions where the data
are reasonably linear. At low pressures there is a steeper slope for all four films, which
changes to a less steep one as the pressure is raised above a certain point. The transition
between regions is not sharp and occurs approximately at 2 kbar for pure PVCA, 4 kbar
for PVCA in PS, 6 bar for PVCA in PMMA, and 14 kbar for PVCA in PIB. The two
slopes indicate the existence of two mechanisms; one controlling at the “low” pressure region
and the other at the “high” pressure region. The lifetimes mirrored roughly the intensity
changes, especially at high pressure, which indicates that there was no significant effect of
pressure on radiative rates. At all but the lowest pressure the ratio of lifetimes from the
two excimer emissions was essentially independent of pressure.
It is clear from the presentation of the results that two distinct mechanisms of excimer interaction are involved. One is evident at high pressures, and the other is present only at low pressures. The pressure at which the transition from one to the other occurs is solvent dependent. An explanation of the emission behavior of PVCA can be provided, using the kinetic scheme of Fig. 8 with some appropriate simplifications and assumptions. The nomenclature is an extension to that used by Birks,\textsuperscript{15} and the mechanism is similar to that described by Johnson.\textsuperscript{16} The rate constants $k_{FM}$ and $k_{IM}$ are those for radiative and non-radiative deactivation of the excited monomer respectively, with $k_{FB_1}$, $k_{FB_2}$ and $k_{ID_1}$, $k_{ID_2}$.
High Pressure Luminescence

having the corresponding for the two excimers. \( D_1 \) and \( D_2 \) represent the excimers responsible for the 380 nm and 430 nm bands respectively. The star (*) represents an excited singlet state; unlike Johnson's nomenclature, \( D_1 \) and \( D_2 \) refer to ground state potential excimer sites and not the excited complexes, \( k_{D,p_2} \) and \( k_{D,p_1} \) are the association and dissociation rates of \( D_1 \) and \( D_2 \), with \( k_{D,p_1} \) and \( k_{D,p_2} \) representing the same quantities between the monomer and the two excimers. The mechanism, represented by the last four rate constants, involves main chain conformational changes and/or side group motion to achieve the geometrical requirements necessary for excimer formation during the lifetime of excitation. \( k_{D,D} \) and \( k_{D,D} \) correspond to single exciton migration to preformed sites along the polymer chain, which satisfy the requirements for excimer formation prior to excitation: \( k_{c,D} \) represents the rate of energy transfer of single excitons and \( f_{D_1}, f_{D_2} \) are the concentrations of preformed excimer sites. This mechanism applies to both solutions and solid films of PVCA, but some of the rate constants become insignificant in films, depending on the temperature and pressure.

Eq. (4) is slightly modified from Johnson's Eq. (6):

\[
\frac{I_{D_2}}{I_{D_1}} = \frac{k_{F,D_1}}{k_{F,D_2}} \frac{k_{D,D}(f_{D_2}+f_{D_1}) + k_{D,p_1}}{k_{D,p_2}(f_{D_2}+f_{D_1}) + k_{D,p_1}}.
\]  

(4)

In our case, \( f_{D_2} \) and \( f_{D_1} \) include the probability of capture of an exciton by a \( D_2 \) or \( D_1 \) site; they are therefore effective concentrations. In order to simplify Eq. (2), one can consider the following limiting cases:

(a) \( k_{D,p_2} \ll k_{D,p_1} \ll k_{D,p_1} \) (i.e. no rotations are allowed to any appreciable extent), and Eq. (2) becomes

\[
\frac{I_{D_2}}{I_{D_1}} = \frac{k_{F,D_1}}{k_{F,D_2}} \frac{f_{D_2}}{f_{D_1}}.
\]  

(5)

where \( \tau_1 = 1/k_{D,p_1} \) and \( \tau_2 = 1/k_{D,p_2} \).

If the ratio of radiative rates and the ratio of lifetimes are not varying, \( I_{D_2}/I_{D_1} \) depends only on the change of the relative probability of capture of an exciton by a \( D_2 \) or \( D_1 \) site.

(b) \( k_{D,p_2} \gg k_{D,p_1} \gg k_{D,p_1} \) (i.e. a rearrangement between \( D_1 \) and \( D_2 \) is allowed, and proceeds efficiently). In this case Eq. (2) becomes

\[
\frac{I_{D_2}}{I_{D_1}} = \frac{k_{F,D_2} k_{D,p_1}}{k_{F,D_1} k_{D,p_2}}.
\]  

(6)

If the ratio of the radiative rates does not change, \( I_{D_2}/I_{D_1} \) depends on the rotational equilibrium constant between \( D_1 \) and \( D_2 \).

Any intermediate cases, where one of \( k_{D,p_1} \) is negligible relative to the other, cannot occur in a plastic medium, because it would deplete the respective population and resulting emission from one of the sites; this follows because there is no other way of forming or depleting sites.

In order to distinguish between mechanisms, it is useful to know the glass transition...
temperatures ($T_\sigma$) of the different polymer films, and these are shown in Table 1. On a molecular scale, the glass transition temperature is the temperature above which the polymer has acquired sufficient thermal energy for rotational motion, or considerable torsional oscillation, to occur about the majority of bonds in the main chain. Below this temperature, this form of motion takes place infrequently and the majority of in-chain groups have fixed conformations. Some transitions can still occur at $T < T_\sigma$; Boyer\textsuperscript{21} has classified those and proposed that they be called $T_\text{rel}$.

If such a transition is taking place between $D_2$ and $D_1$ at atmospheric pressure, one would eventually expect it to stop at some higher pressure, which should correlate inversely with $T_\sigma$. In other words, the higher $T_\sigma$, the lower the pressure, at room temperature, that the transition between $D_2$ and $D_1$ should stop. This, in fact, is observed with our data if we assume that the $D_2 \leftrightarrow D_1$ transition is the controlling factor at low pressures. From Figs. 6 and 7 it is evident that the pressures at which the change in slope occurs have the following order:

$$P(\text{PVCA in PIB}) > P(\text{PVCA in PMMA}) > P(\text{PVCA in PS}) > P(\text{PVCA film}).$$

Also from Table 1:

$$T_\sigma(\text{PIB}) > T_\sigma(\text{PMMA}) > T_\sigma(\text{PS}) > T_\sigma(\text{PVCA}).$$

This correlation leads one to conclude that mechanism (b) dominates at low pressures, and only mechanism (a) is present at high pressures. We can now examine each pressure region separately.

(a) At the high pressure region, $\tau_2/\tau_1$ is approximately constant.\textsuperscript{14} The radiative rates ($k_{\text{rel}}$) are also constant with pressure. One can then take the logarithm of both sides of Eq. (5) to get

$$\ln \frac{I_{D_2}}{I_{D_1}} = \text{Constant} + \ln \frac{f_{D_2}}{f_{D_1}}. \quad (7)$$

The number of preformed excimer sites cannot be changing, since all pathways of forming or destroying them do not exist at high pressures. What must be changing is the relative probability of capture. At equilibrium, this can be represented by a constant $k_\text{rel}$, which is related to known thermodynamic quantities by
High Pressure Luminescence

\[ k_c \propto e^{-(\Delta G_c/RT)} \]  

and

\[ \frac{\partial \ln(k_c)}{\partial P} = -\frac{\partial \ln(I_{D_c}/I_{D})}{\partial P} = -\frac{\Delta V_c}{RT}. \]

where \( \Delta G_c \) is the change of the Gibbs free energy for the above process. The volume extracted from this treatment is a thermodynamic equilibrium quantity, and represents the difference in volume of the system between the \( D_1 \rightarrow D_1^* \) and \( D_2 \rightarrow D_2^* \) transitions. This process can be expected to extend to the low pressure region also, but its contribution is small relative to process (b) (see Fig. 7), and to a first approximation it can be subtracted out.

(b) In the low pressure region \( k_{FP}/k_{FD} \) is also approximately constant, and taking the logarithm of both sides of Eq. (4) gives

\[ \ln \frac{I_{D_1}}{I_{D}} \approx \text{Constant} + \ln \frac{k_{FP}}{k_{FD}}. \]  

A change in the relative intensities is then only dependent on the change of the equilibrium distribution of the two sites. Similarly

\[ \frac{k_{FP}}{k_{FD}} \propto e^{-(\Delta G_{FP}/RT)}. \]  

and

\[ \frac{\partial \ln(k_{FP}/k_{FD})}{\partial P} = \frac{\partial \ln(I_{D}/I_{D_1})}{\partial P} = -\frac{\Delta V_{12}}{RT}. \]  

\( \Delta V_{12} \) represents the change in volume of the system for \( D_1^* \rightarrow D_2^* \).

One can extract these volumes from the slopes of the different lines in Figs. 6 and 7. The approximate molar volume \( (V_{\text{mol}}) \) can also be calculated from the density \((1.509 \text{ g/cm}^3)\) and the molecular weight \((193.2 \text{ g/mole})\) of the monomer: \( V_{\text{mol}} = 182.4 \text{ cm}^3/\text{mole} \). \( \Delta V \)'s and \( \Delta V/\text{V}_{\text{mol}} \) of PVCA in all polymer films are given in Table 2.

The first observation one can make is that there is a net reduction of the system volume for both processes. \( \Delta V_{12} \), which represents the process of \( D_1^* \rightarrow D_2^* \), is at its largest only 3% of the total molar volume, which indicates that a very small rearrangement is necessary to achieve this. This confirms that the initial assumption of some movement in the polymer films is

**Table 2. \( \Delta V \)'s of PVCA in all host polymers**

<table>
<thead>
<tr>
<th>Polymer</th>
<th>( \Delta V_c ) (cm(^3)/mole)</th>
<th>( \Delta V_c/V_{\text{mol}} )</th>
<th>( \Delta V_{12} ) (cm(^3)/mole)</th>
<th>( \Delta V_{12}/V_{\text{mol}} )</th>
</tr>
</thead>
<tbody>
<tr>
<td>PVCA film</td>
<td>-2.0</td>
<td>0.011</td>
<td>-5.1</td>
<td>0.028</td>
</tr>
<tr>
<td>PVCA in PS</td>
<td>-1.6</td>
<td>0.009</td>
<td>-5.0</td>
<td>0.027</td>
</tr>
<tr>
<td>PVCA in PMMA</td>
<td>-1.1</td>
<td>0.006</td>
<td>-3.9</td>
<td>0.021</td>
</tr>
<tr>
<td>PVCA in PIB</td>
<td>-0.9</td>
<td>0.005</td>
<td>-2.1</td>
<td>0.013</td>
</tr>
</tbody>
</table>
possible at low pressures. A second observation is that the change in volume for the two processes is solvent dependent. The order of decreasing $T_x$ is followed by the decreasing $JV$'s, which implies that the less rigid solvents can rearrange with less volume change.

The observations of the total intensity variation with pressure also confirm the assumptions made about the nature of the controlling mechanisms. Emission of the $D_1$ excimer is efficient ($\tau_2/\tau_1$), and any shift of the equilibrium distribution towards $D_2$ can be expected to increase the total intensity. The extent and pressure range where the increase occurs, follows the same order as that of the existence of mechanism (b). The largest increase in intensity is observed for PVCA in PIB for which equilibrium between $D_1$ and $D_2$ is affected by pressure over the largest range. The smallest increase in intensity is observed for the PVCA film, for which mechanism (b) is in effect over the shortest range of pressure. The eventual decrease in intensity can be attributed to an increase of the non-radiative rate ($k_{np}$) with pressure. This is due to the fact that radiationless transitions to the ground state increase as the separation of the two states decreases. From the above argument one would expect the order of quenching by internal conversion to be:

$$Q(\text{PVCA film}) > Q(\text{PVCA in PS}) > Q(\text{PVCA in PMMA}) > Q(\text{PVCA in PIB}),$$

and this is experimentally observed.

5. Zincblende Emission

Crystals have the zincblende-like structure, with appropriate dopants, are utilized commercially in television screens, light emitting diodes and a variety of other applications.

We present here a study of ZnS as a prototype although similar results are obtained for related systems. Figure 9 presents a schematic outline of the principle electronic states and excitations. The basic excitation is from the top of the valence band to the bottom of the conduction band. Two types of dopants are added (in concentrations $\sim 0.01$ atom %). A Cu$^+$ or Ag$^+$ ion acts as a hole trap. It is a deep trap, 1-1.5 eV above the valence band, and thus is strongly localized. There is a variety of evidence that its position vis-a-vis the top of the

![Fig. 9. Schematic diagram of band structure and impurity levels—ZnS doped with donors and acceptors.](image_url)
The Review of Physical Chemistry of Japan Vol. 50 (1980)

High Pressure Luminescence

Valence band is quite independent of pressure. A second compensating ion (e.g., Cl\textsuperscript{-}, Al\textsuperscript{3+}) acts as an electron trap with a trap depth $E_a$ of $\sim 0.1-0.2$ eV. This trap depth is sensitive to conditions, and the delocalization of the electron trap wave function is strongly dependent on $E_a$. The luminescence efficiency depends directly on the overlap and thus on the delocalization of $E_a$. The relevant equation for the intensity\textsuperscript{26,27} can be written:

$$I(r) \propto \int r^2 W(r) G(r) \, dr,$$

where $W(r)$, the radiative recombination rate, is given by

$$W(r) = \text{const} \times (r/a^*)^{3N-9} \exp(-2r/Na^*).$$

Here

$$N = \frac{E^{1/2}}{E_a}, \quad E^* = \frac{e^2}{2a^*\varepsilon}, \quad a^* = \frac{h^2}{m^*\varepsilon^2}.$$

$m^*$ is the effective electronic mass and $\varepsilon$ is the dielectric constant. The distribution function $G(r)$ used corresponded to a random distribution of impurities.

The major pressure dependence of the intensity arises through $N$ and thus through $E_a$. We use here two independent ways of measuring $E_a$ as a function of pressure. First, we measure the shift of the absorption edge and of the emission peak with pressure and take the difference of these as the change in $E_a$ with pressure. It is assumed that the hole trap energy $E_h$ is independent of pressure. There are several pieces of evidence\textsuperscript{23} that this is a reasonable assumption, but establishing a small number from the difference between two large numbers always involves an element of risk. We later check the measurement of $E_a$ by a second method.\textsuperscript{28} Figure 10 shows the shift of the absorption edge and the luminescence emission

**Fig. 10.** Shift of absorption edge and of luminescence emission peak with pressure—doped ZnS.

**Fig. 11.** Comparison of relative values calculated and measured luminescence efficiency—doped ZnS.
with pressure. It is clear that $E_D$ increases significantly as pressure increases. In Fig. 11 we compare the relative efficiency as calculated from Eq. (13) with the measured values. The agreement is remarkably good. (In the lower part of the figure are shown two lifetimes which are essentially independent of pressure. The complex decay was approximated as a double exponential. The independence of pressure is consistent with our assumption that $E_a$ is pressure independent).

In spite of the satisfactory comparison of theory with experiment, it is desirable to check the pressure effect on $E_D$ by a second method. If the sample of ZnS is irradiated at say 77 K, no emission is observed. Heating the sample after irradiation generates an emission which maximizes at some temperature $T_m$. Of the heating is performed so that the temperature-time curve is linear, it is straightforward to extract $E_D$ from the heating rate and $T_m$. In Fig. 12 values of the change in $E_D$ with pressure from thermoluminescence data are compared with the value from the steady state measurements discussed earlier. The discrepancy is nowhere greater than 0.01 eV which is surely within the error of the experiments and of the theory. It is encouraging that rather different heating rates generated the value for $E_D$.

These results illustrate rather well the ability of pressure to test a model of a material of both theoretical and practical interest.

6. Eu$^{3+}$ in Oxysulfides

A second type of useful solid state phosphor involves an oxysulfide doped with Eu$^{3+}$. Here we discuss primarily La$_2$O$_2$S: Eu, with a few remarks about Y$_2$O$_2$S: Eu. These materials are used in fluorescent lighting and as prototypes for lasers.

The excitation lies in the ultraviolet at about 4-4.2 eV, and involves a partial charge transfer from sulfur to europium. The excitation is transferred to excited $^5D$ level of Eu$^{3+}$
which are split by spin-orbit coupling into the series $^5D_0$, $^5D_1$, and $^5D_2$ in order of increasing energy. Since this levels be from 1.9 to 2.9 eV above the ground state one would expect to observe emission from all four levels. In the lanthanum salt only $^5D_0$ and $^5D_1$ emit at room temperature, with a weak $^5D_2$ emission at very low temperature. Struck and Fonger have proposed a solution for this anomaly based on the observed temperature coefficient of emission intensity: their proposal is illustrated in the configuration coordinate diagram of Fig. 13. The charge transfer (CT) state does indeed feed all the $^1D$ levels of Eu$^{2+}$, but the radiative lifetime of the $^5D$ levels is long and the energy barrier to backfeeding into the CT state very low so the latter process governs. Backfeeding from $^1D_2$ also occurs with a slightly smaller efficiency, so that emission is observed at one atmosphere only from $^5D_1$ and $^5D_0$.

If we consider the CC diagram of Struck and Fonger to represent a slice along the totally symmetric coordinate, then pressure experiments should provide a test of this model: since the CT state lies at a larger value of the configuration coordinate, it should shift energy with pressure and thus affect the backfeeding from $^1D_2$ and $^1D_3$ to the CT state. Such high pressure experiments have been performed. As can be seen from Fig. 14, the CT state does indeed shift to higher energy with pressure by over 0.25 eV (2000 cm$^{-1}$) in 100 kbar. The effect on the emission from the $^5D$ levels is summarized in Fig. 15. The emission

![Fig. 13. Schematic configuration coordinate diagram. La$_2$O$_2$S: Eu.](image1)

![Fig. 14. Shift of CT (charge transfer) state with pressure—La$_2$O$_2$S: Eu.](image2)
from $^4D_0$ and $^4D_1$ decreases as pressure increases, with a small but definite leveling at intermediate pressures. The $^4D_2$ emission increases rapidly for $\sim 40$ kbar levels, and then drops beyond $\sim 90$ kbar. The $^4D_3$ emission appears with sufficient intensity for quantitative measurement at $70-75$ kbar and increases rapidly in intensity at higher pressures. As far as we could estimate, the total intensity of emission from all $^5D$ levels is essentially independent of pressure. The effect of pressure on the lifetimes is exhibited in Fig. 16. The $^5D_0$ and $^5D_1$ lifetimes are pressure independent since only the feeding is affected. The $^5D_2$ lifetime increases when the intensity increases and then levels. The $^5D_3$ lifetime increases rapidly with pressure as would be expected.

From the shift of the CT peak one can calculate the increase in the energy barrier for backfeeding with pressure. If we add this number to the atmospheric pressure estimate of Struck and Fonger we can predict the increase in intensity of emission from $^5D_2$ and $^5D_3$ with pressure using a simple classical model. In Fig. 17 we compare prediction and measurement; the agreement is quite satisfactory.

The yttrium salt exhibits qualitatively similar behavior to the lanthanum, except that the CT peak initially lies at about $1000$ cm$^{-1}$ (0.125 eV) higher energy for the $\text{Y}_2\text{O}_3\text{S}:\text{Eu}$. The situation at one atmosphere corresponds roughly to that for $\text{La}_2\text{O}_3\text{S}:\text{Eu}$ at $\sim 50$ kbar. Indeed there is significant $^5D_1$ emission at room temperature. Its intensity decreases with increasing pressure, at first slowly and then more rapidly. The $^5D_2$ emission appears at $\sim 25$ kbar in contact to $\text{La}_2\text{O}_3\text{S}:\text{Eu}$ where it appears near $75$ kbar. It increases rapidly in intensity with pressure. In Fig. 18 exhibit the comparison between calculation and experiment obtained by the same technique discussed for Fig. 17 above. Agreement is, again, excellent.
High Pressure Luminescence

Fig. 17. Calculated and measured emission intensity vs. pressure La$_2$O$_3$: Eu ($^4$D$_0$ and $^4$D$_1$ levels of Eu).

Fig. 18. Calculated and measured emission intensities vs. pressure Y$_2$O$_3$: Eu ($^5$D$_3$ level of Eu).

It should be mentioned that all of the above discussion applies to dilute (~0.1% Eu) materials. Above about 0.5-1.0% Eu significant concentration quenching effects occur. The reader is referred to the original paper for a discussion of how concentration quenching affects the pressure data.

While the results presented here constitute only a small fraction of the variety of studies we have made, they illustrate both the breadth and power of high pressure as a tool for understanding luminescence in particular and electronic structure in general.

References

19) Private communication with Dr. J. Pochan.

School of Chemical Sciences
and
Materials Research Laboratory
University of Illinois at Urbana-Champaign
Urbana, Illinois 61801
U. S. A.