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Kyoto University
Nuclear spin selection rule in the photochemical reaction of CH₃ in solid parahydrogen

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Photolysis of a methyl radical CH₃ in solid parahydrogen produces a methane molecule CH₄ via the reaction between an intermediate singlet methylene CH₂ and a parahydrogen molecule H₂. Conservation of nuclear spin during the reaction has been investigated by the intensity distribution of the rotation-vibration spectrum of methane produced by the reaction. It was found that the population of each nuclear spin state of methane just after the reaction was different from that of the statistical ratio, which indicates that a nuclear spin selection rule does exist in the reaction. However, the observed population was significantly different from the theoretically predicted ratio. The discrepancy between the experiment and the theory may indicate a breakdown of the nuclear spin conservation during the reaction, if the reaction mechanism in solid parahydrogen is the same as in the gas phase. © 2002 American Institute of Physics. [DOI: 10.1063/1.1480003]

I. INTRODUCTION

Because of the weak nuclear magnetic interactions, the total nuclear spin angular momentum I is well conserved in most physical processes such as radiative transitions and collisions. One well-known case is the par (p-H₂) and ortho (o-H₂) modifications of hydrogen molecules. The p-H₂ (I = 0) and o-H₂ (I = 1) can be treated as different molecules because the conversion between them is almost forbidden in the absence of a paramagnetic catalyst.

It has been observed that the memory of nuclear spin is conserved even in chemical processes, where particle rearrangements occur. Under the assumption of the rigorous conservation, Quack has derived selection rules on the nuclear spin modifications during chemical reactions. The selection rules play an important role in various fields, especially in interstellar chemistry. Observation of the abundance of each nuclear spin state of a molecule provides a key to understand the formation of the molecule in the interstellar space.

Despite the importance of the selection rules in various fields, few experimental studies have been reported so far on the quantitative analysis of the nuclear spin modification during chemical reactions. Nuclear spin polarization after a p-H₂ addition reaction has been observed using nuclear magnetic resonance (NMR) spectroscopy. Selective generation of o-H₂ by the photodissociation of ortho-H₂CO (Refs. 13, 14) is one evidence of the conservation. Recently, a quantitative experiment on the nuclear spin selection rule in the reactions of H₂ + H₂ → H₃ + H has been performed by Uy, Cordonnier, and Oka. They reported that the ortho--para ratio of H₃ after the reaction in a plasma can be explained with the theoretically predicted ratio under the assumption of the rigorous conservation.

We have investigated the nuclear spin modification in consecutive reactions of

\[ \text{CH}_3 + h \nu \rightarrow \text{CH}_2 + \text{H} \]  
\[ \text{CH}_2 + p \cdot \text{H}_2 \rightarrow \text{CH}_4 \]

in solid parahydrogen at liquid He temperature. Solid hydrogen is an ideal matrix for the quantitative study of chemical reactions because of (1) weak intermolecular interaction between hydrogen molecules, (2) no “cage effect” for photolysis in the solid, (3) fully quantized rotational motion of molecules embedded in the solid, and (4) narrow spectral linewidth of optical transitions. By virtue of the free rotation of the molecules involved in the above reactions, rotation-vibration transitions of the molecules gave us quantitative information on the nuclear spin modification during the reactions. The observed nuclear spin population after the reactions is compared with the theory to discuss the conservation of the nuclear spin during the reactions.

II. EXPERIMENTS

The methyl radical CH₃ was produced by in situ UV photolysis of methyl iodide CH₃I embedded in solid parahydrogen. Method of the sample preparation was the same as described in previous papers. Briefly, parahydrogen gas containing <0.01% orthohydrogen was mixed with the methyl iodide (~0.001%) at room temperature and the mixed gas was introduced in a copper optical cell kept at 8.0 K to grow a crystal. A 20 W low pressure mercury lamp with a cutoff filter Toshiba UV-25 was used for the selective photolysis of CH₃I at 253.7 nm to produce CH₃. A half-day irradiation of the 253.7 nm photons led to the photodissociation of most of the iodides into the radical. The methyl radical CH₃ thus produced is completely stable in solid parahydrogen.
After the production of the methyl radical, the radical was excited to its $B(1^3A_1^e)$ state \(^{24}\) by 193 nm UV pulses from an ArF excimer laser (6 mJ/pulse, 30 Hz). A 20 min irradiation of the UV pulses was enough to destroy most of the radical. Along with the destruction of the radical, the formation of methane CH\(_4\) was observed.

Infrared absorption spectra of CH\(_3\)I, CH\(_3\), and CH\(_4\) before and after the UV irradiations were observed using a FTIR spectrometer (Bruker IFS120HR) with a resolution of 0.1 cm\(^{-1}\). A KBr beamsplitter and a liquid N\(_2\) cooled MCT detector were used for the spectroscopy. All the measurements were done at 4.6 K.

### III. REACTION SCHEME

Figure 1 shows the schematic diagram of the methane formation reaction in solid \(p\)-H\(_2\). Detailed analyses of the reaction scheme in the present system have been made in previous papers. \(^{18,23}\) The UV light at 253.7 nm dissociates the methyl iodide in \(p\)-H\(_2\) crystal to produce a methyl radical and an iodine atom. The radical thus produced is stable in solid parahydrogen. Then, the UV light at 193 nm excites the radical to its electronic excited \(B\) state, in which the dissociation into a singlet methylene and a hydrogen atom occurs. The methylene thus produced quickly reacts with a nearby hydrogen molecule to produce a methane. Another possibility for the production of CH\(_4\), that is, the abstraction of a hydrogen atom from a hydrogen molecule by the radical in the excited \(B\) state could be completely ruled out by a parallel study of the photolysis of CD\(_3\)I. \(^{23}\) If the abstraction reaction took place, the formation of CD\(_3\)H should be expected in the system of CD\(_3\)I. However, the production of only CD\(_2\)H\(_2\), but no trace of CD\(_3\)H was observed as a result of the UV irradiation of CD\(_3\). Thus, it is concluded that the methane in the present system is produced only by the reaction of CH\(_3\)+H\(_2\)→CH\(_4\).

### IV. NUCLEAR SPIN POPULATION OF CH\(_4\)

Methane CH\(_4\) has three different nuclear spin modifications, that is, \(I=2(A)\), \(I=1(F)\), and \(I=0(E)\) states. Because of the symmetry requirement of the total-wavefunction of CH\(_4\), each nuclear spin state couples with particular rotational states. \(^{25}\) The \(J=0\) rotational state of CH\(_4\) couples with the \(I=2\) nuclear spin state, the \(J=1\) couples with the \(I=1\) state, the \(J=2\) couples with both the \(I=1\) and \(I=0\) states, and so on. Thus, the population of each nuclear spin state can be determined from the intensities of rotationally resolved infrared absorption spectra.

As we discussed in previous papers, \(^{26,27}\) the conversion among the three nuclear spin states of CH\(_4\) takes place in solid parahydrogen. The temporal change of the infrared absorption of the rovibrational transition revealed that the conversion from the \(I=1\) state to the \(I=2\) nuclear spin state associated with the rotational relaxation of \(J=1\rightarrow J=0\) occurred in a time scale of a few days, while the conversion from the \(I=0\) state to the \(I=1\) nuclear spin state was too fast to observe. Temporal changes of the mole fraction \(c(t)\) of the \(I=2\) nuclear spin state were well fitted with the function,

$$c(t) = \frac{c(0) - c(\infty)}{1 - e^{-kt}} + c(\infty),$$  

where \(c(0)\) and \(c(\infty)\) are the mole fraction at \(t=0\) and \(t=\infty\), respectively, and \(k\) is the conversion rate. The conversion rate depends on the temperature of the crystal. The rate at 4.6 K was found to be \(k=2.8\times10^{-3}\) min\(^{-1}\). \(^{27}\)

In the present reaction system, the initial nuclear spin population of CH\(_4\) just after the reaction can be obtained from the value of \(c(0)\) in Eq. (3). Since the population in the \(I=0\) nuclear spin state is negligible, the values of \(c(0)\) and \((1-c(0))\) correspond to the population of the \(I=2\) state and the \(I=1\) state, respectively. Note that the population in the \(I=1\) nuclear spin state we observed should be considered as the sum of the population of the \(I=1\) and \(I=0\) nuclear spin states after the reaction, since all the \(J=2\) and \(I=0\) state had relaxed to the \(J=1\) state within our experimental time scale. \(^{27}\)

Figure 2 shows the \(v_3\) infrared absorption of CH\(_4\) produced by the 193 nm UV irradiation. Panel (a) in Fig. 2 is the spectrum observed just after the UV irradiation for 20 min, while panel (b) is the spectrum observed 1000 min after the UV irradiation. It is clearly seen that intensities in the transition from the \(J=1\) state greatly decreased after 1000 min, while the intensity of the \(R(0)\) transition increased. The initial mole fraction \(c(0)\) can be determined by the least-squares fitting of the temporal change of the intensities of each rotational branches by the use of Eq. (3).

A plot of the mole fraction \(c(t)\) of methane produced by the photolysis is shown in Fig. 3. The time when the second UV irradiation at 193 nm started was taken as the origin of time, that is, \(t=0\). From the least-squares fitting, the initial mole fraction of the \(I=2\) state was obtained to be \(c(0) = 0.19\pm0.01\) with the rate constant of \(k=3.9\times10^{-3}\) min\(^{-1}\). The rate constant was consistent with that observed in pure CH\(_4\)/p-H\(_2\) system reported previously. \(^{27}\) Thus, although the iodine atoms and the methyl radicals were present in the sample, \(^{28}\) their effect on the nuclear spin conversion process of CH\(_4\) is negligible.
In Fig. 3, the period of the 193 nm UV irradiation is drawn by a broken curve. Exactly speaking, the temporal change of the mole fraction $c(t)$ during the irradiation does not obey the function in Eq. (3), since methane molecules were produced continuously during the irradiation. However, the irradiation time of 20 min is short enough compared with the nuclear spin conversion rate of $4 \times 10^{-3}$ min$^{-1}$. If we take into account the correction, the mole fraction of the $I=2$ state at the time $t=0$ must be greater than that determined without the correction, but the difference is estimated to be 5% at most. Therefore, it is safely concluded that the population of each nuclear spin state of CH$_4$ just after the consecutive reactions (1) and (2) was $(I=2):(I=1):(I=0)=0.2:0.8:0.0$ with the uncertainty of 5%.

![Image](image_url)
TABLE II. (a) Ratio of the nuclear spin isomers of CH₃ produced by the abstraction reaction of CH₃ + H₂ → CH₄ + H. (b) Ratio of the nuclear spin isomers of CH₄ after the reaction of CH₃ + H → CH₄.

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<tr>
<td></td>
<td>(I = 3/2) : (I = 1/2)</td>
<td></td>
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<tr>
<td>0</td>
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<td>0 : 1</td>
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| 3/2 | I/2 | 5 : 3 : 0  
| 1/2 | I/2 | 0 : 3 : 1  

\[ ^1 \text{CH}_3(I = 1) + \text{H}_2(I = 0) \rightarrow \text{CH}_4(I = 1). \quad (4) \]

If the nuclear spin selection rule during the chemical reactions is rigorous, only the \( I = 1 \) nuclear spin state of CH₄ should be produced in the above reaction system, that is, the initial population of CH₄ after the reactions should be \( (I = 2):(I = 1):(I = 0) = 0:0:1:0.0 \). However, we observed the significant population in the \( I = 2 \) state of CH₄ just after the reaction. One may think that the production of the \( I = 2 \) state is due to the nuclear spin conversion in the solid after the reactions. However, we can rule out the possibility, because it needs more than 1 h to produce 20% of the \( I = 2 \) methane by the nuclear spin conversion from the \( I = 1 \) methane at 4.6 K. Thus, the \( I = 2 \) nuclear spin state must be produced during the reactions.

The significant population in the \( I = 2 \) nuclear spin state just after the reaction needs to be explained. One possible explanation for this is that the nuclear spin selection rule is partially violated in the present reaction system. Another explanation is that the reaction mechanism in solid parahydrogen is different from that in the gas phase. In the following, we shall discuss the latter possibility in detail.

If the reaction between a methylene and a hydrogen molecule is not the insertion reaction, but an abstraction reaction, the production of the \( I = 2 \) nuclear spin state can be expected as follows:

\[ \text{CH}_2(I = 1) + \text{H}_2(I = 0) \rightarrow \text{CH}_3(I = 3/2, 1/2) + \text{H}(I = 1/2) \]

\[ \text{→ CH}_4(I = 2, 1, 0). \quad (5) \]

Table II summarizes the nuclear spin selection rule involved in the above stepwise reaction. It is seen that the \( I = 2 \) CH₄ is produced by the reaction between the \( I = 3/2 \) CH₃ radical and \( I = 1/2 \) H atom. The stepwise reaction of the \( I = 1 \) CH₂ and \( I = 0 \) H₂ results in the nuclear spin population of CH₃ as \( (I = 2):(I = 1):(I = 0) = 0.42:0.50:0.08 \), or \( (I = 2):(I = 1):(I = 0) = 0.42:0.58:0.00 \) after the fast relaxation from the \( I = 0 \) state to the \( I = 1 \) state. Thus, a combination of the insertion reaction [reactions (1) and (2)] and the abstraction reaction [reaction (5)] might explain the observed nuclear spin population of CH₄, even if the nuclear spin selection rule is rigorous.

It is known that, in the gas phase, the reaction between a singlet methylene \(^1\text{CH}_2\) and a hydrogen molecule \( \text{H}_2 \) is the insertion reaction, while the reaction between a triplet methylene \(^3\text{CH}_2\) and a hydrogen molecule \( \text{H}_2 \) is the abstraction reaction.\(^{32,33}\) Thus, if part of the singlet methylenes \(^1\text{CH}_2\) relax to the triplet state before the reaction with \( \text{H}_2 \), the \( I = 2 \) CH₄ is produced by the stepwise reaction in reaction (5). However, since the relaxation rate from the singlet \(^1\text{CH}_2\) to the triplet \(^3\text{CH}_2\) is two orders of magnitude slower than the rate of the insertion reaction,\(^{31,34}\) it seems unlikely that the triplet methylene was formed in solid parahydrogen. Moreover, it is important to note that no infrared absorption corresponding to \(^3\text{CH}_2\) (Ref. 35) was observed at any stage of the present system.

Nevertheless, because of the proximity between CH₃ and \( \text{H}_2 \) in the solid, the possibility of the stepwise reaction in reaction (5) cannot be completely ruled out. The experimental result is definitive and it is clear that the nuclear spin selection rule is violated in the present system, if the reaction mechanism is the same as in the gas phase. Further studies are necessary for more quantitative understanding of the nuclear spin selection rule during chemical reactions.

VI. CONCLUSION

The population of the \( I = 2 \) nuclear spin state of CH₄ just after the consecutive reactions of \( \text{CH}_3 + h\nu \rightarrow \text{CH}_2 + \text{H}, \text{CH}_2 + \text{H}_2 \rightarrow \text{CH}_4 \) was found to be 0.2. The population is smaller than that of the high temperature limit of 0.3, which indicates that there exists a selection rule of nuclear spin during the reactions. However, the observed population is significantly different from the theoretically predicted population of 0.0. The discrepancy between the theory and the experiment may be due to a breakdown of the nuclear spin conservation during the chemical reactions. Studies on deuterated system are presently underway in order to understand the reason for the unexpected population of the \( I = 2 \) nuclear spin state of methane after the reactions.

ACKNOWLEDGMENT

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1 G. Herzberg, Molecular Spectra and Molecular Structure (Krier, Malabar, 1991), Vols. I–III.

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