Surgery of fullerenes

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Recent attempts at the synthesis of endohedral fullerenes by organic reactions, so-called “molecular surgery” methods, are surveyed. The creation of an opening on the surface of fullerene cages allowed insertion of He, H2, H2O, or CO within the cages. An effective route to “suture” an opening was established to realize new endohedral fullerene, H2@C60. Further development of this operation as well as the properties and reactions of H2@C60 are summarized. Also the application of the encapsulated H2 molecule as an NMR probe for the study of aromaticity of ionic fullerenes is described.

1. Introduction

Endohedral fullerenes, which encapsulate atom(s) or small molecule(s) within the spherical carbon cage, have attracted much attention from viewpoints of not only fundamental curiosity but also applied studies because of their unique electronic structures resulting from the interaction between the fullerene π-system and incorporated species.1-3 The fullerenes encapsulating metal atom(s), so-called endohedral metallofullerenes, were first produced by laser-vaporization of graphite-metal oxides composite materials at the early stage of fullerene research, and later by the use of arc discharge technique.1-3 In these metallofullerenes, electron transfer from the incorporated metal atom(s) to the fullerene cage gives rise to characteristic behaviors which are totally different from those of empty fullerenes. Just for example as a representative case, La@C82 (denotes that the La is encapsulated within the C82 cage) should be better described as La3+@C82− because of the electron transfer to endow La@C82 with paramagnetic character. The redox behaviors of La@C82, studied by cyclic voltammetry, revealed that abilities both as an electron acceptor and as a donor are higher than that of empty C82.4 Furthermore, La@C82 can be regarded as a “super atom”, since one-electron oxidation or reduction was shown to take place on the C82 cage while keeping the positive charge of the La atom constant.5 It should also be noted that chemical and physical properties of fullerenes encapsulating a trimetallic nitride cluster such as Sc3N@C80 have been investigated in detail since its discovery by Dorn.6 In particular, Sc3N@C80 was shown to be even applicable as a part of an electron-donor/acceptor system.7 Recently, effort was made to develop highly efficient MRI (magnetic resonance imaging) contrast agents by the use of

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water-soluble fullerene derivatives encapsulating a paramagnetic Gd(III) ion. Besides the high performance as the contrast agent, the Gd(III)-incorporated fullerenes are of great significance from the viewpoint of safety because toxic Gd(III) ion is entirely isolated by the encapsulation in the fullerene cage.

Furthermore, small band gaps of endohedral metallofullerenes make them potential candidates for practical use, for example, in the field of molecular electronics such as thin film organic field effect transistor. Actually, endohedral metallofullerenes such as La\textsubscript{2}@C\textsubscript{80} \cite{9} or nanorods of La\textsubscript{2}@C\textsubscript{82}(Ad)\textsuperscript{10} (Ad = adamantylidene) have been shown to exhibit the FET properties.

However, development of such applications has so far been hampered by severe limitations of their availability that relied on the hard-to-control method of production. The conventional methods, i.e. arc-discharge or laser vaporization of graphite rods containing metal oxides or carbides, have almost no selectivity in controlling both the cage-size and product distribution. Generally, a pure product can be available only in mg-quantities after tedious separation procedures.

As another types of endohedral fullerenes, fullerenes encapsulating a noble-gas atom, such as helium, neon, argon, krypton, and even xenon were realized by Saunders by treating the fullerene powder under forced conditions (650 °C and 3000 atm of noble gases).\textsuperscript{11} Although the occupation level of guest is so low as 0.1%,\textsuperscript{12} incorporation of an active nucleus, \textsuperscript{3}He, within fullerene cages has proven to be useful because it provides valuable information about magnetic shielding effect inside the fullerenes caused by ring currents of the spherically delocalized \pi-electrons.\textsuperscript{13,14} The NMR signal of \textsuperscript{3}He is also useful as a tool to follow chemical reactions taking place at the exterior of the fullerene cages.\textsuperscript{15}

Recent investigation revealed that the reactivity of Xe\textsubscript{2}@C\textsubscript{60} in [4+2] cycloaddition with 9,10-dimethylanthracene is substantially different from that of He\textsubscript{2}@C\textsubscript{60} because \textsuperscript{3}He electron cloud of the C\textsubscript{60} cage is pushed outward by the inside xenon atom.\textsuperscript{16} However, a low occupation level of the noble-gas atom within fullerenes (about 0.1%) and difficulty in isolation or enrichment in practical scale\textsuperscript{17} have apparently hampered the exploration of possibility of the endohedral fullerenes as functional materials. Hence, it is desirable to explore an entirely different approach that would lead to the macroscopic production of endohedral fullerenes.

### 2. Open-cage fullerenes

Under such circumstances it is quite appealing to produce the endohedral fullerenes by using pristine fullerenes as starting materials. In this connection, Rubin proposed a concept to realize endohedral fullerenes with the use of organic reactions, that is, "molecular surgery".\textsuperscript{18} This approach consists of a series of steps, which are (1) "incision" of the fullerene cage to form an opening on the surface, (2) insertion of some small atom(s) or molecule(s) through the opening, and (3) "suture" of the opening to reproduce the fullerene cage while retaining the guest species.

The development of this approach has been a challenging task for organic chemists for the past decade. Wudl reported, for the first time, two-step reactions to create an 11-membered-ring opening on the C\textsubscript{60} cage to give compound 1.\textsuperscript{19} Unfortunately, however, the opening was found to be not large enough even for the smallest atom, helium, to pass through.\textsuperscript{20} Rubin thereafter synthesized bisfulleroid derivative 2 with an eight-membered-ring opening and transformed it into Co(III) complex 3.\textsuperscript{21} These works are unequivocally a stepping stone to the subsequent progress in the technique of incising the fullerene cage. Herein, we summarize recent achievements in the molecular surgical technique with the hope of encouraging the progress in research on the rational chemical synthesis as well as the application of endohedral fullerenes.

#### 2.1 1,2,3,4,5,6-Hexasubstitution of fullerene C\textsubscript{60}

Rubin proposed a strategy\textsuperscript{18} toward "incising" a fullerene cage through saturation of three adjacent C=C double bonds within one hexagonal ring. The subsequent [2+2+2] cycloreversion of the planarized cyclohexane ring was expected to give a fullerene derivative with such a large opening as a 15-membered ring (Fig. 1). Thus, a thermal reaction of C\textsubscript{60} with butadiene derivative 4 having two azido groups was conducted.\textsuperscript{22} This reaction was intended to proceed via initial addition of 4 (involving both a [4+2] cycloaddition and two 1,3-dipolar additions) to the three C=C double bonds in the same hexagon of C\textsubscript{60} to form compound 5 with a saturated cyclohexane ring. However, the obtained product was instead bislactam 6 having a 14-membered-ring opening produced through a series of reactions involving rearrangements, extrusion of two N\textsubscript{2} molecules, and addition of an O\textsubscript{2} molecule followed by dehydrogenation.\textsuperscript{22}

The synthesis of bislactam 6 is of considerable significance because the size of the opening was proved to be large enough for the first time for insertion of the smallest gaseous species such as a \textsuperscript{3}He atom or a H\textsubscript{2} molecule.\textsuperscript{23} By treating a crystalline powder of 6 with 475 atm of \textsuperscript{3}He gas at 288–305 °C for 7.5 hours, a \textsuperscript{3}He atom was incorporated at an occupation level of 1.5%. A H\textsubscript{2} molecule was also introduced at an occupation level of 5% under the conditions of 100 atm of H\textsubscript{2} gas at 400 °C for 48 hours. A highly shielded \textsuperscript{1}H NMR signal for the encapsulated H\textsubscript{2} was observed at δ = −5.43 ppm. These results verified the feasibility of the strategy of molecular surgical approach.

Recently, Rubin succeeded in the synthesis of the desired 1,2,3,4,5,6-hexaadduct of C\textsubscript{60}, 7, by the reaction of bis-isobenzofuran adduct 8 with Pb(OAc)\textsubscript{4} in a substantially low concentration (1.0 × 10\textsuperscript{−7} M).\textsuperscript{24} Unfortunately, however, the retro [2+2+2] ring-opening reaction at the strained cyclohexane ring of 7 to afford open-cage fullerene 9 did not take place due to the high activation barrier for this transformation and the thermal instability of product 9 as indicated by theoretical calculations (B3LYP/6-31G*//PM3 level). The origin of the high endothermicity of this ring-opening reaction is attributable to the formation of two strained C=C double bonds in oxabenzonorbornadiene moieties of 9.

#### 2.2 Bowl-shaped fullerenes
Bisfulleroid derivative 2,21 synthesized by Rubin, possesses bridgehead C=C double bonds at the rim of the eight-membered-ring opening. Komatsu and Iwamatsu independently found that one of these double bonds in the derivatives of 2, i.e., 1025 and 1126, can be readily cleaved via addition of photochemically generated singlet oxygen to give diketone derivatives 1225 and 13,27 respectively, with a 12-membered-ring opening.28

The formation of 12 led to the successful achievement of “molecular surgery” as will be shown later, while derivative 13 was found to be a good precursor of various bowl-shaped fullerenes.29

Thus, the opening of diketone 13 was greatly enlarged as reported by Iwamatsu. Although the detailed reaction mechanism is not clear yet, a reaction of 13 with a phenyl hydrazine proceeded at room temperature to give open-cage fullerene 14 having an sp2 methylene carbon within the rim of a 16-membered-ring opening.30 The size of the ellipsoidal opening was estimated to be 6.5 Å for a longer axis and 3.8 Å for a shorter axis (B3LYP/3-21G level). When a solution of 14 in 1,1,2,2-tetrachloroethane-d2 (TCE-d2) was heated under 133 atm of H2 gas at 100 °C for 4 hours, encapsulation of a H2 molecule took place to give H2@14 at an occupation level of 62%.31 Since the activation barrier for the escape of encapsulated H2 is only 23.7 kcal/mol reflecting the large 25 π-system to the original structure of C60 seems almost impossible by means of organic synthetic procedures.

2.3 Peroxide-mediated formation of open-cage fullerenes

On the other hand, an approach to other types of an open-cage fullerene using entirely different processes has been reported by Gan and coworkers.34 They recently reported that the reaction of fullerene C60 with tert-butyl hydroperoxide in the presence of ceric ammonium nitrate gives hexaadduct 16.35 Subsequent oxidation of 16 by irradiation with visible light in the presence of iodine resulted in the cleavage of the central five-membered-ring and afforded diketone 17 with a nine-membered-ring opening.26,27 The epoxide moiety in 17 was transformed to a vicinal diol by the use of a Lewis acid, B(C6F5)3, giving 18.

Although attempted oxidation of the vicinal diol moiety in 18 was unsuccessful, aminoketal/hemiketal derivative 19, obtained from the reaction of 18 with aniline, was found to undergo complex oxidation reactions with diazotization and the amino group of 19 were all oxidized to give products 20 with a large opening divided by an acid anhydride bridge. When 20 was reacted with hydrazine, the bridge across the opening could be cleaved to give compounds 21 and 22 having an 18-membered-ring opening. The amide group of 21 was converted to a bromoazocarbonyl moiety by the reaction with Br2 to give 23, and then the reaction of 23 with AgClO4 resulted in elimination of the azo-oxo group to afford tri-oxo open-cage fullerene 24.24

Upon 1H NMR measurement of 24 in CDCl3-CS2 (1:1), a highly shielded signal was observed at δ = −13.07 ppm. Based on the similarity of the chemical shift to that of H2O@15 (δ = −11.4 ppm) described above, the signal was assigned to an encapsulated H2O molecule, although the occupation level was less than only 5%.34

The size of the opening of 24 was further expanded by treatment with trifluoroacetic acid in the presence of ferrocene. This reaction removed two tert-butyl peroxy groups to create an aromatic hexagon, accompanied by insertion of an oxygen atom into the rim of the opening to give 25 possessing a 19-membered-ring opening. It is not surprising that the 1H NMR spectrum of 25 exhibited an intense signal of an encapsulated H2O molecule at δ = −12.8 ppm in CDCl3. The occupation level of the H2O molecule inside 25 was 57% at room temperature and it increased up to 88% upon keeping the
solution at −20 °C for several days. The encapsulated H$_2$O molecule within the fullerene cage of 24 and 25 was also unambiguously confirmed by the X-ray crystallography.

3. Molecular surgical method to endohedral fullerenes

3.1 Cage opening

In this section, we will show a series of reactions which lead us to the first successful organic synthesis of endohedral fullerene encapsulating molecular hydrogen according to the molecular surgical method.

As has been mentioned above, we synthesized the open-cage fullerene 12 by the reaction of open-cage fullerene 10 with photochemically generated singlet oxygen. Compound 10 with an 8-membered-ring opening was obtained by the reaction of C$_{60}$ with 5,6-diphenyl-1,2,4-triazine in 85% yield based on consumed C$_{60}$ (59% conversion). However, the 12-membered-ring opening in 12 was not large enough even for a molecule of H$_2$ to pass through as indicated by calculated high activation energy (51.8 kcal/mol) (B3LYP/6-31G**//B3LYP/3-21G level), and further enlargement of the opening was necessary.

For this purpose, we discovered a novel reaction to insert a sulfur atom into the rim of the opening in 12. Since the electrochemistry indicated that compound 12 with two carbonyl groups is even a better π-electron acceptor than C$_{60}$, 12 was expected to be activated by a typical electron donor such as tetrakis(dimethylamino)ethylene (TDAE). Thus, in the presence of TDAE 12 was found to react with elemental sulfur in 1,2-dichlorobenzene (ODCB) at 180 °C to give compound 26 having a 13-membered-ring opening as a single product in 77% yield. The structure of 26 was unambiguously proved by X-ray crystallography. The size of the opening of 26 was 5.6 Å for a longer axis and 3.75 Å for a shorter axis. The calculated activation energy (B3LYP/6-31G**//B3LYP/3-21G level) for the insertion of a H$_2$ molecule through the opening of 26 (30.1 kcal/mol) was lower than that calculated for bislactam 6 (41.4 kcal/mol) in support of the larger opening of 26. Thus, the encapsulation of a H$_2$ molecule to 26 was expected to be easier.

The enlargement of the opening of 12 was also found to be possible by insertion of a selenium atom in place of sulfur to the rim of the opening. In this case, stronger activation by the use of sodium alkylxanethiol as a reducing agent was needed. Thus, the reaction of 12 with elemental selenium in the presence of CH$_3$SNa in refluxing ODCB afforded open-cage fullerene 27 in 46% yield. The results of X-ray crystallography showed that the size of the opening in 27 was 5.72 Å for a longer axis and 3.88 Å for a shorter axis, which was slightly larger than that of sulfur analogue 26.

3.2 Incorporation of a H$_2$ molecule

As a guest molecule to be encapsulated in an open-cage fullerene, the smallest molecule, H$_2$, was selected. When open-cage fullerene 26 was treated with high-pressure H$_2$ gas (800 atm) at 200 °C for 8 hours in an autoclave, the incorporation of hydrogen inside the cage was realized. The incorporation of hydrogen inside the cage was realized. Thus, the reaction of

3.3 Incorporation of a He atom

Next, the insertion of the smallest noble-gas atom, $^3$He, within the fullerene cage of 26 was investigated at Yale University by collaboration with Saunders and Cross. By heating an ODCB solution of 26 under $^3$He gas (20 atm) at 80 °C for a few hours, $^3$He@26 at an occupation level of 0.1% was obtained. By monitoring the release of the $^3$He atom from $^3$He@26, the activation energy for the release was determined to be 22.8 kcal/mol, which is lower than that for the release of a H$_2$ molecule from H$_2$@26 by 11.4 kcal/mol. The release of a $^3$He atom from $^3$He@26 can take place at near room temperature with the half-life of 40.3 hours at 30 °C. Thus, it was required to develop a method to reduce the opening size of $^3$He@26 in order to prevent the encapsulated He atom from escaping.

We found that, upon sodium borohydride reduction of a carbonyl group in 26, a transannular ether-forming reaction readily takes place at room temperature to give product 28 in 86% yield. The structure of 28 was confirmed by the X-ray crystallography. Theoretical calculations (B3LYP/6-31G**//B3LYP/3-21G level) indicated that the activation...
energies for release of an encapsulated He atom from He@28 is 50.4 kcal/mol, which is more than twice as large as that from He@26, indicating that effective reduction of the opening size of 26 has occurred.

Thus, we first conducted the insertion of a He atom within the cage of 26 under the conditions of 650 atm of He gas at 90 °C for 24 hours. Then, the resulting material was immediately subjected to the sodium borohydride reduction at −20 to −25 °C to give desired product He@28 in 90% yield as a stable complex. The occupation level of the He atom was 35% based on the mass spectroscopic analysis. 44

Although the non-covalent interaction between the encapsulated He atom and the fullerene cage of 28 had been expected to be almost negligible, the NMR signal of the methine proton of He@28 showed a slight downfield shift by 0.36 Hz as compared to that of empty 28. The methine proton signal of H2@28, prepared separately, exhibited 1.9 Hz downfield shift relative to that of empty 28. These results demonstrate that the non-covalent interaction of the encapsulated H2 molecule with the fullerene cage of 28 is larger than that of the encapsulated He atom and the NMR signal of the methine proton outside the cage is a good indicator of the electronic interaction inside the cage. 44

3.4 Incorporation of two H2 molecules into an open-cage C70

Taking the thickness of π-electron cloud of fullerenes into consideration, the size of the inner cavity of C70 is estimated to be 4.6 Å along the long axis and 3.6 Å along the short axis, which is larger than that of C60 (3.6 Å in inner diameter). Therefore, it is expected to be possible to insert more than one small molecule through a newly created opening on the surface of C70. However, most studies on this line have previously been made only on C60 because of the wealth of knowledge about the chemical reactivity of C60 and also due to its higher symmetry. Thus, we challenged to prepare open-cage C70 derivative 29 46 with a 13-membered-ring opening by applying similar procedures to those used for the synthesis of the C60 analogue 26. 25 A thermal reaction of C70 with 3,6-dit-(2-pyridyl)pyridazine in refluxing 1-chloronaphthalene gave eight-membered-ring compound 30 (40%), and subsequent oxidation with singlet oxygen afforded 12-membered-ring compound 31 (49%). Then, the opening of 31 was enlarged by insertion of a sulfur atom to the rim of the opening using TDAE as a π-electron donor to give 29 (94%). The results of the X-ray crystallography of 29 showed that the opening size of 29 is almost the same as that of C60 analogue 26. 25

The insertion of H2 molecule(s) was carried out by applying 890 atm of H2 gas at 230 °C for 8 hours. The successful encapsulation of molecular hydrogen was clearly demonstrated by appearance of a new intense signal in the 1H NMR spectrum at unusually high field, δ = −16.51 ppm in ODCB-d8. This signal was assigned to the resonance of H2@29 based on the mass spectroscopic data. Noteworthy is that a small signal was also observed at δ = −15.22 ppm. Upon low-temperature NMR studies, this signal completely disappeared by cooling the solution to −60 °C and reappeared as two new signals at −80 °C, which became sharp at −100 °C with the chemical shift of δ = −12.87 ppm and −17.38 ppm. The observed dynamic behavior is interpreted by assuming the positional exchange of two H2 molecules encapsulated in open-cage 29, i.e., (H2)2@29. The line-shape analysis of the low-temperature NMR spectra gave $\Delta E = 8.0$ kcal/mol for this positional exchange. The yield of H2@29 and (H2)2@29 was 97% and 3%, respectively, as estimated from the integrated areas of the NMR signals.

In accord with the results of the X-ray crystallography, i.e., the opening of 29 being almost the same size as that of 26, the rate for the release of a H2 molecule from H2@29, monitored at temperatures above 160 °C, was almost as fast as that from the C60 analogue H2@26. 28 The activation energy was determined to be 33.8 kcal/mol.

3.5 Closure of the opening

As described above, we have succeeded in introducing a H2 molecule into the carbonaceous cage of 26 at the occupation level of 100% 26. 28 Subsequently, we developed a method to “suture” the 13-membered-ring opening of H2@26 to complete the molecular surgery operation. 47-51 Prior to this study, there had been no report for the attempt at suturing a once formed opening on the fullerene cage.

Apparently, the first step for the size-reduction of the 13-membered-ring in H2@26 should be the removal of the sulfur atom. We first conducted an oxidation of the sulfide unit of H2@26 by m-chloroperbenzoic acid in order to make the sulfur atom readily removable. The reaction proceeded at room temperature to give sulfoxide derivative H2@32 almost quantitatively. Then, irradiation of a solution of H2@32 in benzene with visible light at room temperature caused the elimination of the SO unit to give product H2@12 having a 12-membered-ring opening in 42% yield. The encapsulated H2 molecule was completely retained throughout these two steps because the two reactions were conducted at room temperature. 26 This removal of a sulfur atom made the distance between two carbonyl carbons across the opening closer from 3.89 Å for H2@26 to 3.12 Å for H2@12 as shown by calculations (B3LYP/6-31G* level). The reductive coupling of the two carbonyl groups efficiently proceeded by McMurry reaction using Ti(0) 52 at 80 °C to give product H2@10 with an eight-membered-ring opening in 88% yield. Here, it is to be noted that the MALDI-TOF mass spectrum of H2@10 already exhibited an intense peak of H2@C60 together with a smaller molecular ion peak of H2@10.

The final step to remove all the remaining organic addends on the C60 cage was performed by simply heating a powder of H2@10 (245 mg) in a vacuum-sealed tube placed in an electric furnace at 340 °C for 2 hours. Separation of the reaction mixture was carried out by flash column chromatography over silica gel eluted with carbon disulfide to afford a pure solution containing desired H2@C60 (118 mg, 67% yield), contaminated by 9% of empty C60. Subsequently, complete separation of H2@C60 from empty C60 was achieved by recycling HPLC on semipreparative CosmOSil Buckyprep columns (two directly connected columns; 250 mm length, 10 mm i.d.; mobile phase, toluene; flow rate, 4 mL/min) to give H2@C60 as a pure material after 20 recycles (total retention time, 399 minutes; the retention time for empty C60, 395...
minutes). The adsorption mechanism of the Buckyprep column is largely based on the π-π interaction with pyrrenyl groups in the stationary phase. A very weak van der Waals interaction between the encapsulated H₂ molecule and the C₆₀ π-system must have contributed to this separation.

The ¹³C NMR spectrum of pure H₂@C₆₀ exhibited a signal at δ = 142.84 ppm in ODCB-d₆, which is very slightly downfield shifted by 0.078 ppm relative to that of empty C₆₀.

4 The ¹H NMR signal of the encapsulated H₂ molecule appeared at δ = −1.44 ppm in ODCB-d₆, which is 5.98 ppm upfield shifted from dissolved free H₂ molecule. This value is comparable to the 6.36 ppm upfield shift of a ³He NMR signal of ³He@C₆₀, suggesting that this nearly 6 ppm upfield shift is a universal value corresponding to the magnetic field at the central position in the C₆₀ core. The IR spectrum of H₂@C₆₀ was almost the same as that of empty C₆₀, exhibiting four absorption bands at 1429.2, 1182.3, 576.7, and 526.5 cm⁻¹ (to be compared with 1429.2, 1182.3, 575.7, and 526.5 cm⁻¹ for empty C₆₀). Only the band at 576.7 cm⁻¹ of H₂@C₆₀, corresponding to an out-of-plane vibration mode, is higher in energy than that of C₆₀ by 1.0 cm⁻¹. This might be interpreted as evidence of a very slight repulsive interaction between the C₆₀ cage and the inner H₂ molecule. The UV-vis spectrum of H₂@C₆₀ was almost the same as that of C₆₀.

The cyclic voltammetry and differential pulse voltammetry of H₂@C₆₀ at room temperature exhibited three reversible reduction waves at E₁/₂ = −1.13, −1.54, and −1.99 V vs Fc/Fc⁺ in ODCB and one irreversible oxidation peak at Epa = +1.62 V in TCE, which were virtually the same as those of empty C₆₀. However, when more negative potential was applied using the solution in toluene-acetonitrile (5.4 : 1) under vacuum at −10 °C, the fourth, fifth, and sixth reduction waves became observable, which were more cathodic than the reduction of empty C₆₀ by 0.04, 0.07, and 0.15 V, suggesting that H₂@C₆₀ becomes more difficult to be reduced as it acquires more than three electrons.

H₂@C₆₀ is thermally stable. Upon heating H₂@C₆₀ at 500 °C for 10 minutes under vacuum, no decomposition or no release of encapsulated H₂ molecule was observed at all.

### 4. Organic derivatization of H₂@C₆₀

In order to examine the effect of encapsulated H₂ molecule on the reactivity of the outer C₆₀ cage, the solid-state mechanochemoactivation dimerization of H₂@C₆₀ (occupation level of 91%) was conducted under the same conditions reported previously. The dumbbell-shaped dimer, (H₂@C₆₀)₂, was obtained in 30% isolated yield similarly to the reaction of empty C₆₀. Apparently the inside hydrogen does not affect the reactivity of the C₆₀ cage. The NMR signal for the inside H₂ molecule was observed as a singlet at δ = −4.04 ppm, which is 8.58 ppm upfield shifted from that of free H₂ similarly to the case for ³He@C₆₀ (8.81 ppm upfield shift from the signal of free ³He). Three additional fullerene derivatives H₂@C₆₆, H₂@C₆₄, and H₂@C₆₂ were also synthesized and their NMR signals for the encapsulated H₂ molecule in ODCB-d₆ appeared at δ = −3.27, −4.30, and −4.64 ppm, respectively.

Since the values of chemical shift change sensitively according to the difference in structures of the organic addends, the encapsulated H₂ molecule within C₆₀ cage can also be used as a good probe to investigate the chemical reactions at the exterior of the cage, just as the ³He atom inside C₆₀ (occupation level of 0.1%) has been used for this purpose.

Furthermore, several organic and organometallic derivatives of H₂@C₆₀ were prepared at the University of Tokyo by the group of Nakamura. The NMR signals for the encapsulated H₂ molecule appeared at δ = −10.39 ppm for compound H₂@C₃₆ in CDCl₃-CH₂Cl₂, δ = −9.79 ppm for potassium cyclopentadendione H₂@C₇₆ in THF-d₈, δ = −10.44 ppm for bucky ferrocene H₂@C₇₈ in CDCl₃-CH₂Cl₂, and δ = −10.77 ppm for tetraminofullerene epoxide H₂@C₇₉a in CDCl₃. Although the ¹H NMR chemical shifts for the encapsulated H₂ molecule of amphiphilic derivative H₂@C₉₀b were measured in a variety of solvents, such as THF-d₈, DMSO-d₆-toluene-d₈ (1:1), DMSO-d₆, and D₂O-DMSO-d₆ (1:1), no specific solvent effect on the chemical shift was observed.

### 5. Utilization of the encapsulated H₂ molecule as an NMR probe

Saunders reported that the ³He NMR signal of hexaion ³He@C₆₀ appeared at dramatically high field (δ = −48.7 ppm relative to the signal of free ³He), reflecting the strong shielding effect of C₆₀. This was apparently due to the high aromaticity of C₆₀. Theoretical as well as experimental studies indicated that all of the hexagons and pentagons of C₆₀ showed diamagnetic ring currents. Among the other possible anionic states of C₆₀, dianion C₆₀⁻ was particularly important in synthetic chemistry for introduction of two functional groups on the C₆₀ cage. Although the “2(N + 1)² rule”, describing the spherical aromaticity of C₆₀ symmetrical fullerenes, proposed by Hirsh, predicts that the 62-π-electron system should not have high aromaticity, little had been known about the aromaticity of C₆₀⁻ prior to our study.

Thus, we carried out the generation of H₂@C₆₀ by using CH₃SNa as a reductant in CD₃CN under vacuum. The ¹H NMR signal of the encapsulated H₂ molecule of H₂@C₆₀ was observed at surprisingly low field such as δ = −26.36 ppm. This is downfield shifted by 27.8 ppm relative to that of neutral H₂@C₆₀ (δ = −1.44 ppm in ODCB-d₆). This result demonstrates that the overall aromaticity within the cage of C₆₀ decreases drastically upon two-electron reduction. The NICS (nucleus independent chemical shifts) calculations (B3LYP/6-31G* level) for all the hexagons and pentagons of C₆₀ showed diamagnetic ring currents. Because there exist more hexagons than pentagons in C₆₀, the antiaromatic character of hexagons overwhelms the aromatic character of pentagons, resulting in the strong deshielding effect inside the cage. This is the first time that the reversal of aromaticity/antiaromaticity for hexagons and pentagons of fullerenes was observed.

The similar reversal was observed even for the dianion of highly functionalized C₆₀ such as open-ring fullerene H₂@C₆₆. The NMR signal of H₂@C₆₆ was observed at δ = −8.10 ppm in CD₃CN, which is downfield shifted by 15.4 ppm relative to that of neutral H₂@C₆₀ (δ = −7.25 ppm in ODCB-d₆).
NICS calculations (B3LYP/6-31G* level) showed that the aromatic and antiaromatic characters of hexagons and pentagons are mostly reversed in the same way as those for H2@C60,[26] in spite of the highly ruptured π-system in H2@C60.

While the aromaticity of the spherically conjugated system of C60 is proved to be drastically changed when C60 acquired extra electrons,[30,36] the research on the magnetic properties of fullerene derivatives are still quite limited. Thus, we prepared dichloromethyl-C60 cation[33] and (1-octynyl)-C60 anion[34] encapsulating a H2 molecule (H2@C60 and H2@C60)[35] as these compounds appeared to be well suited for the study of the difference in aromaticity between the monofunctionalized C60 cation and anion. The cation H2@C60 was generated in CF3SO3H as a stable species and the NMR signal of the encapsulated H2 molecule was observed at δ = −2.89 ppm. The signal was downfield shifted by 1.73 ppm from a H2 signal of the corresponding neutral compound H2@C60[30,36] (δ = −4.62 ppm in CS2-CDCl3 (1:1)), indicating that the aromaticity of 40 H2@C60 was slightly decreased as compared to that of 42. On the other hand, an NMR signal for the encapsulated H2 molecule of anion H2@C60[36] which was also generated as a stable species in THF-d8, appeared at δ = −0.60 ppm. This resonance was shifted to lower field than that for a H2 molecule of neutral counterpart H2@C60[43] (δ = −4.75 ppm in CS2-CDCl3 (1:1)), again indicating the decrease in aromaticity. The relatively small difference in the chemical shifts of the encapsulated H2 molecule between H2@C60[40] and H2@C60 (absolute ∆δ value, 2.29 ppm) demonstrates that the aromaticity of the fullerenes are affected to a comparative degree in these cationic and anionic systems.

6. Physical properties of the H2 molecule inside C60

The encapsulated H2 molecule in H2@C60 is isolated from the outside environment by the surrounding fullerene cage because the opening is so small that only a He atom or a H2 molecule can go through. Actually, the nuclear spin-lattice relaxation time (T1) of the encapsulated H2 of H2@C60 upon the 1H NMR measurements were not affected by the presence of molecular oxygen as a paramagnetic species in the solution.[35] The T1 values of the encapsulated H2 molecule and one of the pyridyl proton of H2@C60 in ODCB-d4 are 0.2 s and 3.9 s under vacuum and 0.2 and 0.9 s in an oxygen-saturated solution, respectively. In H2@C60, synthesized by complete restoration of the opening of H2@C60, the encapsulated H2 is completely isolated from the outside. As judged from the difference in chemical shift of 13C NMR (∆δ = 0.078 ppm; vide supra), the interaction of the encapsulated H2 and the outer C60 cage in H2@C60 appears to exist but should be very weak.

To investigate the nature of such interaction, the T1 values of H2 molecule encapsulated in C60 cage as well as those of free H2 molecule were measured in Columbia University by the group of Turro for the first time.[66] The T1 values of free H2 at 300 K were found to depend significantly on the organic solvent, for example, from 1.44 s (benzene) to 0.84 s (CCl4). A somewhat larger variation of T1 values was observed for H2@C60: from 0.118 s (benzene) to 0.046 s (CCl4), which are 12–18 times smaller than those for free H2. However, the value of T1 for both H2 and H2@C60 does not significantly change between the solutions in benzene-d6 and benzene-d4.

Therefore, the dominating interactions determining H2 and H2@C60 nuclear relaxation are concluded to be intramolecular. On the other hand, the T1 value for both H2 and H2@C60 was found to be temperature dependent with the maximum value observed at ~ 240 K. This kind of temperature dependence of T1 is consistent with two different relaxation mechanisms dominantly operating at different temperature ranges, that is, below and above 240 K, for both H2 and H2@C60. Qualitatively, the dipole-dipole interaction accounts for the observed increase in T1 with temperature below 240 K, whereas the spin-rotation interaction accounts for the observed decrease in T1 with temperature above 240 K. These facts and consideration derived therefrom imply that the H2 in both environments rotates through large angles between collisions with the solvent shell or with the walls of the C60 cage.

Although the encapsulated H2 molecule in C60 is completely isolated from the outside, it can communicate with the outside world.[67] First, no differences in the triplet life times were observed for C60, H2@C60, and D2@C60 upon irradiation of laser pulse. Thus, the interaction of encapsulated H2 and D2 with the paramagnetic walls of the fullerene is too weak to be determined by triplet lifetime measurements. However, clear differences in reactivity were observed for the quenching of singlet oxygen O2 by C60, H2@C60, and D2@C60. The absolute quenching rate constants kq of O2 by H2@C60, D2@C60, and C60 were determined using a time-resolved method in CS2 to give the values of kq(H2@C60) = 1.5 × 107 M–1 s–1, kq(D2@C60) = 0.49 × 106 M–1 s–1, and kq(C60) = 0.38 × 105 M–1 s–1, respectively. The results demonstrate that both H2@C60 and D2@C60 are better quenchers than empty C60. Importantly, the O2 can sense the difference between encapsulated H2 and D2. The rate constants for quenching of O2 by free H2 and D2 in CC14 were also measured to afford the values of kq(H2) = 0.81 × 103 M–1 s–1, kq(D2) = 0.024 × 104 M–1 s–1, which are significantly smaller than the values by H2@C60 and D2@C60. This is a unique example of an encapsulated guest having a significantly larger rate constant for quenching than the free guest. Since O2 might form an exciplex with the outer surface of fullerene, it is speculated that this unique behavior can be attributed to a significant life time to provide an opportunity for O2 and the encapsulated H2 to interact for a considerable period of time.

Interaction of the encapsulated H2 molecule with another species outside the fullerene cage is also seen for the interaction with nitroxide radicals. In the presence of paramagnetic nitroxide radicals, bimolecular contribution to the spin-lattice relaxation rate, 1/T1, for the protons of H2 and H2@C60 dissolved in toluene-d8 was investigated.[68] The measured relaxation rates depended on the concentration of the nitroxide, [S], according to the relationship: 1/T1 = 1/T1,0 + R1[S], where T1,0 is the relaxation time in the absence of paramagnetic relaxant and R1 (M–1 s–1) is the second-order relaxation coefficient, or relaxibility. It was found that the
relaxation effect of the paramagnets is enhanced 5-fold in H2@C60 compared to free H2 under the same conditions.

Concluding remarks

In this article we have outlined recent progress toward synthesizing endohedral fullerenes not by a physical method but by the stepwise transformations of fullerenes by organic reactions. This molecular surgical operation was shown to be effective for realization of a new endohedral fullerene, H2@C60. This method can be applied to the preparation of endohedral fullerenes encapsulating gas atoms or molecules with sizes comparable or smaller than a H2 molecule. While the preparation of bowl-shaped compounds by the wide incision of a C60 cage provided access to the encapsulation of molecules larger than H2, i.e., H2O and CO, the operation to suture the opening must be a highly difficult task. Of course the most important and ultimate goal of the present method would be to develop a route to the encapsulation of metal atoms affording endohedral metallofullerenes. However, as far as the present method is used, the insertion of metal ions such as Li+ and Na+ through the opening is difficult because of their strong coordination with two carbonyl oxygen atoms at the opening. Further development of the technique for the chemical transformations of the fullerene cage will be crucial for this project to be accomplished.

Notes and references


