

ABSTRACTS

all the samples were analysed chemically. The results were as follows.

(1) By X-ray investigations, it has been shown that only the pure spinel structure is preserved in the composition 1:1 to 1:2 ($\text{ZnO}:\text{Fe}_2\text{O}_3$), and free Fe_2O_3 is found to be more than 1:3. On the contrary, in all samples which contain the excess ZnO more than 1:1, both the spinel and ZnO structured are observed. From these results, it seems that the zine-ferrite makes a solid solution with Fe_2O_3 readily to some extent, but with ZnO very slightly at high temperature.

(2) Different from the samples containing the excess Fe_2O_3 over the stoichiometric amount, the samples with higher ZnO contents have no ferromagnetism.

(3) Pure Fe_2O_3 is not decomposed to Fe_3O_4 when heated at 1250°C . But when it is mixed with ZnO, a part of Fe_2O_3 is converted to Fe_3O_4 . This tendency is marked for the samples with higher Fe_2O_3 contents. In the samples containing excess ZnO, the formation of Fe_3O_4 from Fe_2O_3 is negligible by heating at 1250°C .

(4) Fe_3O_4 thus formed, is almost reoxidized to Fe_2O_3 while cooling down in the air.

(5) From our systematic studies, we ought to consider that the ferro-magnetism of zine-ferrite with higher Fe_2O_3 contents is not due to the formation of magnetite, but preferable to dissolved Fe_2O_3 in the 1:1 sample at high temperature.

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On the Recovery of Copper from Leached Solution of Pyrite Cinder

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When copper leached from pyrite cinder is recovered by scrap cementation method, recovery of copper, state of cemented copper and the scrap loss are affected by the methods of agitation and conditions of solution such as Fe^{2+} conc., Cu^{2+} conc., pH value etc.. In this work, these influences have been studied by statistical methods of analysis. Agitation was performed by (1) pneumatic agitation (method A) and (2) air-tight mechanical agitation (method B). Design of experiment was the analysis of variance by Latin square whose factors were Fe^{2+} conc. (1~10 g/l), Cu^{2+} conc. (0.5~2 g/l) and pH value (0.5~2.5) of the solution. Results are summarized as follows:

(1) Mean value of the recovery of copper in method B is higher than in method A, and difference of mean scrap loss in both methods is not significant.

(2) Significant factors on the recovery of copper in A are Fe^{2+} conc. and Cu^{2+} conc., on the other hand, in B, none of above factors is significant.

(3) On the scrap loss, Fe^{2+} conc. is significant and Cu^{2+} conc. seems to be significant in B, but these factors are not so in A.

(4) Recovery of copper in B is higher than A, when scrap is repeatedly used for cementation.

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