

Vibrational and rotational structure and excited-state dynamics of pyrene

Masaaki Baba,^{1,a)} Motohisa Saitoh,¹ Yasuyuki Kowaka,¹ Kunio Taguma,¹ Kazuto Yoshida,² Yosuke Semba,² Shunji Kasahara,² Takaya Yamanaka,³ Yasuhiro Ohshima,³ Yen-Chu Hsu,⁴ and Sheng Hsien Lin^{4,5}

¹Division of Chemistry, Graduate School of Science, Kyoto University, Kyoto 606-8502, Japan

²Molecular Photoscience Research Center, Kobe University, Nada-ku, Kobe 657-8501, Japan

³Institute for Molecular Science, Okazaki 444-8585, Japan

⁴Institute of Atomic and Molecular Science, Academic Sinica, Taipei, Taiwan 10617, Republic of China

⁵Department of Applied Chemistry, National Chiao Tung University, Hsinchu, Taiwan 300, Republic of China

(Received 15 June 2009; accepted 10 November 2009; published online 11 December 2009)

Vibrational level structure in the S_0 1A_g and S_1 $^1B_{3u}$ states of pyrene was investigated through analysis of fluorescence excitation spectra and dispersed fluorescence spectra for single vibronic level excitation in a supersonic jet and through referring to the results of *ab initio* theoretical calculation. The vibrational energies are very similar in the both states. We found broad spectral feature in the dispersed fluorescence spectrum for single vibronic level excitation with an excess energy of 730 cm^{-1} . This indicates that intramolecular vibrational redistribution efficiently occurs at small amounts of excess energy in the S_1 $^1B_{3u}$ state of pyrene. We have also observed a rotationally resolved ultrahigh-resolution spectrum of the 0_0^0 band. Rotational constants have been determined and it has been shown that the pyrene molecule is planar in both the S_0 and S_1 states, and that its geometrical structure does not change significantly upon electronic excitation. Broadening of rotational lines with the magnetic field by the Zeeman splitting of M_J levels was very small, indicating that intersystem crossing to the triplet state is minimal. The long fluorescence lifetime indicates that internal conversion to the S_0 state is also slow. We conclude that the similarity of pyrene's molecular structure and potential energy curve in its S_0 and S_1 states is the main cause of the slow radiationless transitions. © 2009 American Institute of Physics. [doi:10.1063/1.3270136]

I. INTRODUCTION

Pyrene is a prototypical molecule of the polycyclic aromatic hydrocarbons (PAHs) group, and it has been the subject of extensive spectroscopic studies in its condensed phase.^{1–7} Ever since emission of pyrene was observed from the comet P/Halley,⁸ the gaseous molecule has attracted the interest of many researchers, and several studies of the jet-cooled pyrene molecule have been reported.^{9–13} Its vibrational and rotational structures, however, have not been satisfactorily analyzed yet, and more careful measurements of jet spectroscopy and ultrahigh-resolution spectroscopy are desired to accurately determine the energy level structure of the isolated pyrene molecule. Pyrene possesses 16π electrons and is not subject to Hückel's $4n+2$ rule, which states that cyclic planar molecules in which each atom has a p orbital are aromatic if they contain $4n+2\pi$ electrons. It is of great interest to investigate pyrene's molecular structure in both the S_0 and S_1 states. In addition, it is worth noting that the nonradiative decay rate in the S_1 state is fairly slow. Fluorescence lifetime was reported to be very long ($1.4\text{ }\mu\text{s}$),⁹ and fluorescence quantum yield is expected to be high although there has been no accurate measurement for the isolated molecule. The aim of the present study is to

estimate the molecular structure of pyrene and to elucidate the cause of its slow nonradiative decay processes.

The S_1 state is expressed as one of the nearly degenerate configurations of highest occupied molecular orbital-1 (HOMO-1) to lowest unoccupied molecular orbital (LUMO) and HOMO to LUMO+1 electronic excitations. The symmetry of these two states is identical (B_{3u}) and they strongly couple with configuration interaction. The S_1 $^1B_{3u}$ state is one of the mixed states with symmetric linear combination and the energy becomes smaller in it than in that of the state of HOMO to LUMO electronic excitation (S_2 $^1B_{2u}$ state). This situation is similar to that of naphthalene and the S_1 state is designated by 1L_b .¹⁴ Relatively strong b_{1g} vibronic bands could be seen in the fluorescence excitation spectrum,^{9–13} providing evidence of strong vibronic coupling between the S_1 $^1B_{3u}$ and S_2 $^1B_{2u}$ states. The $S_1 \leftrightarrow S_0$ transition moment is small so that the radiative lifetime of the S_1 state is very long. The long fluorescence lifetime indicates that radiationless processes in the S_1 state of the isolated pyrene molecule such as predissociation, internal conversion (IC) to the ground state, and intersystem crossing (ISC) to the triplet state are all very slow.

We observed the fluorescence excitation spectrum of the S_1 $^1B_{3u} \leftarrow S_0$ 1A_g transition and dispersed fluorescence spectra for single vibronic level excitation of jet-cooled pyrene. The vibrational structures in the S_0 and S_1 states were analyzed referring to the results of *ab initio* calculation. Further,

^{a)}Author to whom correspondence should be addressed. Electronic mail: baba@kuchem.kyoto-u.ac.jp.

we observed the rotationally resolved ultrahigh-resolution spectrum of the 0_0^0 band, and rotational constants have been accurately determined for both the S_0 and S_1 states. In this article, we present these experimental results and discuss the vibrational and rotational level structure, and excited-state dynamics of the isolated pyrene molecule.

II. EXPERIMENTAL

Pyrene- h_{10} (Wako Chemical) was purified by vacuum sublimations and pyrene- d_{10} (Wako Chemical) was used without further purification. The solid sample was heated to 180 °C in a stainless steel container. The vapor was mixed with Ar gas and the mixed gas was expanded in a high-vacuum chamber through a pulsed nozzle (an automobile fuel injector) to generate a supersonic jet.

As a light source, we used a pulsed dye laser (Lambda Physik LPD 3000, $\Delta E=0.1$ cm $^{-1}$) pumped by an excimer laser (Lambda Physik LPX 105i, 308 nm, 100 mJ). Fluorescence from excited molecules in a supersonic jet was collected with a lens to a photomultiplier (Hamamatsu R927) through a glass filter (Toshiba L42) to block scattered laser light. The output was processed with a storage oscilloscope (Iwatsu-Lecroy 9362) for single photon counting. The total photon number for ten laser shots was recorded as the fluorescence intensity using the LABVIEW system. The fluorescence excitation spectrum was observed by changing the wavelength of laser light. The dispersed fluorescence spectrum was observed using a scanning monochromator (Nikon P250). The resolution was 30 cm $^{-1}$ and the accuracy of transition wavenumber was ± 6 cm $^{-1}$.

The ultrahigh-resolution spectrum was observed using a cw single-mode ring Ti:sapphire laser (Coherent CR899-29, $\Delta E=0.0001$ cm $^{-1}$) pumped by a Nd $^{3+}$:YVO $_4$ laser (Spectra-Physics Millennia X, 532 nm, 10 W). The output was fed into an enhancement cavity (Spectra-Physics WavetrainSC, LBO) for second harmonics generation. We obtained single-mode UV laser light with a power of 20 mW. The UV laser light beam was directed across the supersonic jet at right angles, which was collimated by a skimmer (2 mm orifice diameter) and a slit (1 mm width). The distance between the nozzle and the slit was 30 cm. The residual Doppler width was estimated to be 0.0002 cm $^{-1}$ under this condition. Fluorescence was focused to a photomultiplier (Hamamatsu R565) using a pair of a spherical mirror and an ellipsoidal mirror with 90% efficiency of the solid angle. The output was processed with a gated photon counter (Stanford Research SR400). We observed a sub-Doppler fluorescence excitation spectrum by recording the change of fluorescence intensity with the laser wavelength. We simultaneously recorded the transmitted light intensity of a stabilized etalon (Burleigh CFT500) as frequency marks and a Doppler-limited absorption spectrum of the iodine in order to calibrate transition wavenumber of each rotational line. The transition wavenumber of an iodine spectral line was calibrated using the iodine atlas.¹⁵ The accuracies of relative and absolute wavenumbers were 0.0001 and 0.003 cm $^{-1}$, respectively.

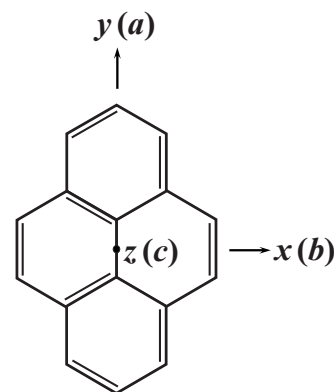


FIG. 1. Molecular structure and coordinate axes of pyrene.

III. RESULTS AND DISCUSSION

A. Vibrational structure in the S_1 $^1B_{3u}$ and S_0 1A_g states

Pyrene is considered to be a planar molecule with D_{2h} symmetry in the S_0 1A_g and S_1 $^1B_{3u}$ states, as we will show later in Sec. III B. In D_{2h} symmetry, the three coordinate axes are all equivalent because they possess one C_2 , one σ_h , and two σ_v . Here, we define the axis system as shown in Fig. 1. Pyrene has 72 normal modes and its vibrational reducible representation is reduced as

$$\Gamma = 13a_g + 12b_{1g} + 4b_{2g} + 7b_{3g} + 5a_u + 7b_{1u} + 12b_{2u} + 12b_{3u}. \quad (1)$$

The vibrational energies in the excited state can be determined by excess energies of observed bands in a fluorescence excitation spectrum. Figure 2 shows the fluorescence excitation spectra of the S_1 $^1B_{3u} \leftarrow S_0$ 1A_g transition of pyrene- h_{10} and pyrene- d_{10} in supersonic jets. We could clearly resolve and identify even weak vibronic bands al-

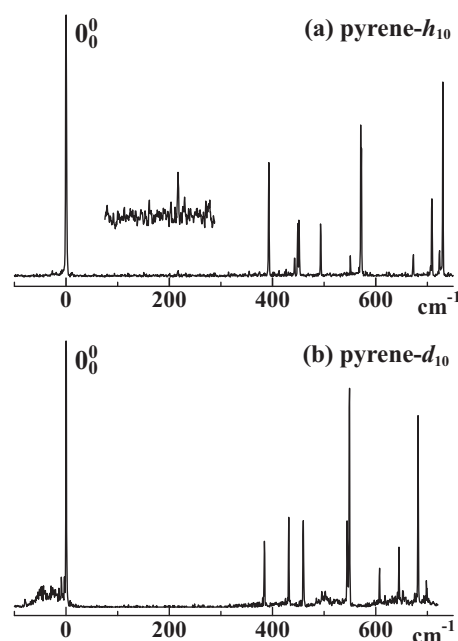


FIG. 2. Fluorescence excitation spectra of the S_1 $^1B_{3u} \leftarrow S_0$ 1A_g transition of (a) pyrene- h_{10} and (b) pyrene- d_{10} in supersonic jets.

TABLE I. Vibrational energies (cm⁻¹) and assignments of the bands observed in the fluorescence excitation spectrum of the $S_1^1B_{3u} \leftarrow S_0^1A_g$ transition of pyrene-*h*₁₀.

Band number	Vibrational energy	Intensity ^a	Band type	Assignment
1	0	m	<i>b</i>	0 ₀ ⁰ (27 204 cm ⁻¹)
2	217	vw		
3	232	vw		
4	324	vw		
5	374	vw		
6	393	s	<i>b</i>	13 ₀ ¹ (<i>a_g</i>)
7	443	w	<i>a</i>	
8	449	m	<i>a</i>	25 ₀ ¹ (<i>b_{1g}</i>)
9	452	m	<i>a</i>	
10	494	m	<i>a</i>	24 ₀ ¹ (<i>b_{1g}</i>)
11	551	w	<i>a</i>	
12	572	s	<i>b</i>	12 ₀ ¹ (<i>a_g</i>)
13	673	m	<i>a</i>	
14	709	s	<i>a</i>	
15	724	m	<i>a</i>	
16	730	s	<i>a</i>	23 ₀ ¹ (<i>b_{1g}</i>)
17	779	m	<i>b</i>	
18	780	vs	<i>b</i>	
19	782	vs	<i>b</i>	
20	783	s	<i>b</i>	
21	944	w		
22	967	w		
23	1023	w		10 ₀ ¹ (<i>a_g</i>)
24	1066	w		
25	1070	w		
26	1111	m		22 ₀ ¹ (<i>b_{1g}</i>)
27	1144	w		9 ₀ ¹ (<i>a_g</i>)
28	1155	w		
29	1162	w		
30	1246	w		20 ₀ ¹ (<i>b_{1g}</i>)
31	1250	w	<i>b</i>	8 ₀ ¹ (<i>a_g</i>)
32	1282	w		
33	1298	w		
34	1303	w		23 ₀ ¹ 12 ₀ ¹
35	1331	w	<i>b</i>	7 ₀ ¹ (<i>a_g</i>)
36	1354	w		12 ₀ ¹ 11 ₀ ¹
37	1357	w		19 ₀ ¹ (<i>b_{1g}</i>)
38	1397	m	<i>b</i>	6 ₀ ¹ (<i>a_g</i>)
39	1467	m		17 ₀ ¹ (<i>b_{1g}</i>)
40	1520	w		
41	1527	w		
42	1537	w		
43	1568	w		11 ₀ ²
44	1575	m		16 ₀ ¹ (<i>b_{1g}</i>)

^avw stands for “very weak,” w for “weak,” m for “medium,” s for “strong,” and vs for “very strong.”

though the spectral features were almost identical to those reported before.^{9–13} The excess energies of observed vibronic bands of pyrene-*h*₁₀ and pyrene-*d*₁₀ are listed in Tables I and II, respectively. In order to make reliable vibrational assignments, we observed the rotational envelopes of individual vibronic bands. The results are shown in Fig. 3. Pyrene is an asymmetric-top molecule and its rotational *a*, *b*, and *c* axes are parallel to *y* (long axis), *x* (short axis), and *z* (out of plane), respectively. In the 0₀⁰ band, two intensity maxima of *P* and *R* transitions are observed. This is the typical feature

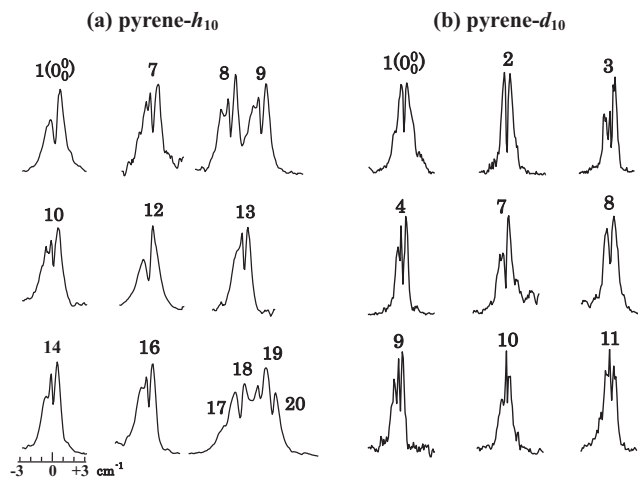
TABLE II. Vibrational energies (cm⁻¹) and assignments of the bands observed in the fluorescence excitation spectrum of the $S_1^1B_{3u} \leftarrow S_0^1A_g$ transition of pyrene-*d*₁₀.

Band number	Vibrational energy	Intensity ^a	Band type	Assignment
1	0	vs	<i>b</i>	0 ₀ ⁰ (27 283 cm ⁻¹)
2	384	m	<i>b</i>	13 ₀ ¹ (<i>a_g</i>)
3	431	m	<i>a</i>	25 ₀ ¹ (<i>b_{1g}</i>)
4	460	m	<i>a</i>	24 ₀ ¹ (<i>b_{1g}</i>)
5	495	vw		
6	502	vw		
7	544	m	<i>b</i>	
8	549	s	<i>b</i>	12 ₀ ¹ (<i>a_g</i>)
9	607	w	<i>a</i>	
10	645	w	<i>a</i>	
11	682	s	<i>a</i>	23 ₀ ¹ (<i>b_{1g}</i>)
12	698	w		

^avw stands for very weak, w for weak, m for medium, s for strong, and vs for very strong.

of a *b*-type transition in which the electronic transition moment is parallel to the *x* axis. Consequently, the S_1 state can be identified as $^1B_{3u}$. We found several *a*-type vibronic bands in which a sharp *Q* peak could be seen at the band center. The *a*-type band is identified as a transition to a *b_{1g}* vibrational level of which the transition moment is parallel to the *a*(*y*) axis. The intensity arises from vibronic interaction with the $S_2^1B_{2u}$ state.^{16–19} The *b*-type vibronic band, in contrast, is identified as a transition to an *a_g* vibrational level, the symmetry of which is the same as that of the 0₀⁰ band. Although the rotational envelope of a vibronic band observed at 780 cm⁻¹ is strange, this band shape could be reproduced as an envelope of four partially overlapped *b*-type transitions. The vibrational energies in the S_0 state can be determined by the dispersed fluorescence spectrum for single vibronic level excitation. Figure 4 shows the dispersed fluorescence spectrum that occurs when the laser wavelength is tuned to the $S_1 \leftarrow S_0$ 0₀⁰ band. The vibrational energies of observed bands are listed in Table III.

Accurate theoretical calculation is necessary for normal-

FIG. 3. Rotational envelopes of prominent bands in the fluorescence excitation spectra of the $S_1^1B_{3u} \leftarrow S_0^1A_g$ transition of (a) pyrene-*h*₁₀ and (b) pyrene-*d*₁₀.

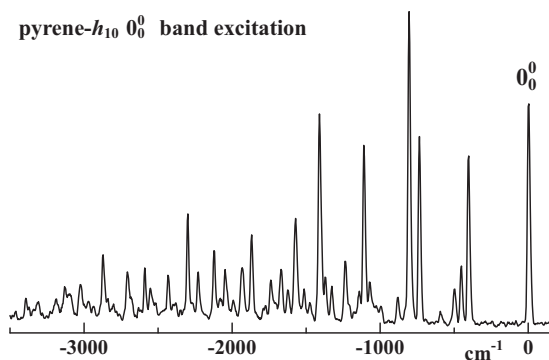


FIG. 4. Dispersed fluorescence spectrum of jet-cooled pyrene- h_{10} for excitation of the $S_1 \ ^1B_{3u} \leftarrow S_0 \ ^1A_g \ 0_0^0$ band. The resolution is 30 cm^{-1} .

mode assignments of these observed vibronic bands. For the S_0 state, we calculated the vibrational energies using the density-functional-theory (DFT) method at the B3LYP/6-311G(d,p) level with the GAUSSIAN 03 program package.²⁰

TABLE III. Vibrational energies (cm^{-1}) and assignments of the bands observed in the dispersed fluorescence spectrum of the pyrene- h_{10} for the $S_1 \ ^1B_{3u} \leftarrow S_0 \ ^1A_g \ 0_0^0$ band excitation.

Band number	Vibrational energy	Intensity ^a	Assignment
1	0	s	$0_0^0 (27\ 204 \text{ cm}^{-1})$
2	402	s	$13_1^0(a_g)$
3	451	w	$25_1^0(b_{1g})$
4	500	w	$24_1^0(b_{1g})$
5	593	vw	$12_1^0(a_g)$
6	735	s	$23_1^0(b_{1g})$
7	805	vs	$11_1^0(a_g)$
8	883	w	25_2^0
9	999	vw	24_2^0
10	1072	w	$10_1^0(a_g)$
11	1111	s	$22_1^0(b_{1g})$
12	1140	vw	$9_1^0(a_g)$
13	1212	w	$21_1^0(b_{1g})$
14	1238	w	$8_1^0(a_g)$ and $20_1^0(b_{1g})$
15	1326	w	$7_1^0(a_g)$
16	1372	w	$19_1^0(b_{1g})$
17	1410	s	$6_1^0(a_g)$
18	1476	vw	
19	1513	w	$17_1^0(b_{1g})$
20	1571	m	$16_1^0(b_{1g})$
21	1624	w	$4_1^0(a_g)$
22	1670	w	
23	1739	w	
24	1867	m	
25	1936	w	
26	2047	m	
27	2123	w	
28	2228	w	
29	2298	m	
30	2434	w	
31	2552	vw	
32	2590	w	
33	2707	w	
34	2875	w	

^avw stands for very weak, w for weak, m for medium, s for strong, and vs for very strong.

For the $S_1 \ ^1B_{3u}$ state, we calculated the vibrational energies using the restricted configuration interaction of single excitation (RCIS) method with a 6-31G(d,p) basis set. It is known that *ab initio* calculation generally yields larger vibrational energies than those obtained experimentally, and an appropriate scaling factor was proposed for various calculational methods. For instance, 0.9757 was proposed for DFT(B3LYP)/6-311G**.²¹ In order to determine the best-fit scaling factor for pyrene in the S_0 state, we performed a least-squares fit and obtained 0.9830. The scaled values are in good agreement with the observed vibrational energies. In the S_1 state, the vibrational energies are slightly different with symmetry because of interaction between the electronic states. We performed separate least-squares fits for the a_g and b_{1g} bands, and obtained 0.8923 and 0.8848, respectively. The scaled values are also in good agreement with the observed vibrational energies in the S_1 state. The resultant calculated and observed vibrational energies are listed in Table IV, and the vibronic assignments are shown in Tables I–III. The 780 cm^{-1} band is composed of four b -type bands. This band is considered to be Fermi multiplets due to the $11_1^0(a_g)$ fundamental band and the 13_2^0 and 40_2^0 overtone bands. Although these bands are supposed to be b -type, it is impossible to assign each observed band explicitly. The vibrational analysis in the S_0 and S_1 states were previously reported by Borisevich *et al.*^{11,12} based on the results of semiempirical MO/8 model calculation. We resolved the partially overlapped bands in the fluorescence excitation spectrum at the higher resolution. Accordingly, a number of vibrational energies are slightly different from theirs. Our assignments are based on the *ab initio* theoretical calculation and the analysis of rotational envelope. The obtained vibrational energies are in good agreement with the observed vibronic energies. It should be noted that the vibrational energies in the S_1 state are very similar to those in the S_0 state. It indicates that the molecular structure and the potential energy curve are almost identical in these two states.

The b_{1g} bands are remarkably strong in the fluorescence excitation spectrum, suggesting that the vibronic interaction with the $S_2 \ ^1B_{2u}$ state is strong. The $S_2 \leftarrow S_0 \ 0_0^0$ band is observed at $31\ 150 \text{ cm}^{-1}$,¹⁰ and the energy difference between the S_1 and S_2 states is only 3946 cm^{-1} . Fluorescence lifetime becomes monotonously shorter as the excess energy becomes larger, and the lifetime at 4000 cm^{-1} is reported to be $\sim 400 \text{ ns}$.⁹ As the vibronic interaction between the S_1 and S_2 states becomes stronger, the radiative decay rate becomes larger because the $S_2 \leftarrow S_0$ transition is significantly stronger. In a large molecule, intramolecular vibrational redistribution (IVR) is known to occur efficiently at high-vibrational levels of the S_1 state. We observed dispersed fluorescence spectra for excitation of prominent vibronic bands, and the results are shown in Fig. 5. Clearly, the broad emission component becomes stronger as the excess energy increases. These results indicate that the IVR process is remarkably fast at the vibrational levels over 730 cm^{-1} in the $S_1 \ ^1B_{3u}$ state of pyrene. It might be possible to estimate the IVR rate by the homogeneous width of a spectral line. However, it is ex-

TABLE IV. Calculated and observed vibrational energies (cm^{-1}) of pyrene- h_{10} .

Symmetry	Number	Vibration type	$S_0^1A_g$	$S_1^1B_{3u}$
			Calc. ^a (Obs.) ^b	Calc. ^c (Obs.) ^d
a_g	1	C–H stretch	3143	3027
	2	C–H stretch	3134	3011
	3	C–H stretch	3118	2995
	4	C–C stretch	1646 (1624)	1661
	5	C–C stretch+C–H bend	1572	1523
	6	C–C stretch	1406 (1410)	1420 (1397)
	7	Ring deform	1333 (1326)	1313 (1331)
	8	Ring deform+C–H bend	1245 (1238)	1252 (1250)
	9	C–H bend	1152 (1140)	1139 (1144)
	10	Ring deform+C–H bend	1076 (1072)	1033 (1023)
	11	Ring deform	804 (805)	780 (783)
	12	Ring breathing	587 (593)	556 (572)
	13	Ring deform	406 (402)	395 (393)
b_{1g}	14	C–H stretch	3126	2978
	15	C–H stretch	3115	2964
	16	C–C stretch	1603 (1571)	1574 (1575)
	17	C–C stretch+C–H bend	1513 (1513)	1474 (1467)
	18	C–C stretch	1415	1425
	19	C–H bend+ring deform	1381 (1372)	1397 (1357)
	20	C–H bend+ring deform	1247 (1238)	1235 (1246)
	21	C–H bend	1183 (1212)	1158
	22	C–H bend+ring deform	1112 (1111)	1098 (1111)
	23	Ring deform	739 (735)	724 (730)
	24	Ring deform	501 (500)	490 (494)
	25	Ring deform	456 (451)	446 (449)
b_{2g}	26	oop C–H bend	905	917
	27	oop C–H bend	803	808
	28	oop ring deform	530	503
	29	oop ring torsion	243	225
b_{3g}	30	oop C–H bend	978	1011
	31	oop C–H bend	963	966
	32	oop C–H bend+C–C bend	806	825
	33	oop C–H bend	759	747
	34	oop ring deform	571	507
	35	oop ring deform	500	483
	36	oop ring wagging	258	238
a_u	37	oop C–H bend	968	1006
	38	oop C–H bend	891	900
	39	oop C–H bend	675	655
	40	oop ring deform	395	378
	41	oop ring deform	149	150
b_{1u}	42	oop C–H bend	969	976
	43	oop C–H bend	844	847
	44	oop C–H bend	746	742
	45	oop ring deform	710	680
	46	oop ring deform	493	449
	47	oop ring deform	209	197
	48	Butterfly	96	90
b_{2u}	49	C–H stretch	3142	3026
	50	C–H stretch	3118	2997
	51	C–H stretch	3114	2990
	52	C–C stretch	1612	1571
	53	C–H bend+C–C stretch	1463	1440

TABLE IV. (Continued.)

Symmetry	Number	Vibration type	$S_0 \ ^1A_g$	$S_1 \ ^1B_{3u}$
			Calc. ^a (Obs.) ^b	Calc. ^c (Obs.) ^d
	54	C–H bend+ring deform	1439	1368
	55	C–H bend+ring deform	1251	1240
	56	Ring deform	1099	1059
	57	Ring deform	1001	959
	58	Ring deform	823	789
	59	Ring deform	694	671
	60	Ring deform	500	483
	61	C–H stretch	3134	3011
	62	C–H stretch	3125	3003
	63	C–H bend+C–C stretch	1621	1554
	64	C–H bend+C–C stretch	1493	1519
	65	C–H bend+C–C stretch	1438	1437
b_{3u}	66	C–H bend+C–C stretch	1325	1305
	67	C–H bend+ring deform	1215	1202
	68	C–H bend+ring deform	1189	1162
	69	C–H bend	1150	1126
	70	C–H bend+ring deform	967	928
	71	Ring deform	547	528
	72	Ring deform	354	348

^aResults of B3LYP/6-311G(d,p) are scaled by 0.9830.^bDetermined using the dispersed fluorescence spectrum with an accuracy of $\pm 6 \text{ cm}^{-1}$.^cResults of RCIS/6-31G(d,p) are scaled by 0.8848 for the b_{1g} modes and 0.8923 for other modes.^dDetermined using the fluorescence excitation spectrum with an accuracy of $\pm 1 \text{ cm}^{-1}$.

pected to be very small in the S_1 state of pyrene, and time-resolved fluorescence measurements are necessary to determine the absolute value.

B. Rotational structure in the $S_1 \ ^1B_{3u}$ and $S_0 \ ^1A_g$ states

We observed a rotationally resolved ultrahigh-resolution fluorescence excitation spectrum of the $S_1 \ ^1B_{3u} \leftarrow S_0 \ ^1A_g \ 0_0^0$ band of jet-cooled pyrene. The whole spectrum is shown in Fig. 6. The linewidth of an observed rotational line was 0.0003 cm^{-1} , which was mainly due to the residual Doppler width of the collimated supersonic jet. The lifetime at the zero-vibrational level of the S_1 state was reported to be

$1.4 \text{ } \mu\text{s}$, and natural linewidth is expected to be 0.000004 cm^{-1} . The rotational temperature was estimated to be 12 K by the analysis of intensity distribution.

We also analyzed the rotational structure of pyrene using the A -reduced Hamiltonian method on an asymmetric-top molecule presented by Watson.²² The nonvanishing matrix elements are given by

$$\begin{aligned}
 \langle JKM | H_r^{(A)} | JKM \rangle \\
 = \frac{1}{2}(B+C)J(J+1) + \left[A - \frac{1}{2}(B+C) \right] K^2 - \Delta_J J^2(J+1)^2 \\
 - \Delta_{JK} J(J+1)K^2 - \Delta_K K^4,
 \end{aligned} \quad (2)$$

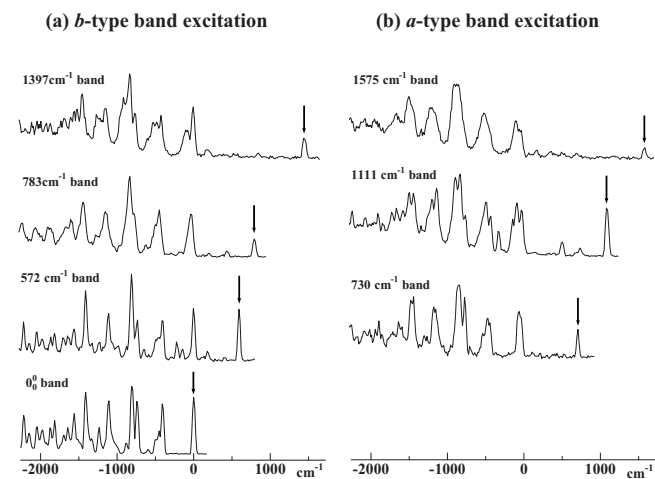


FIG. 5. Dispersed fluorescence spectra of jet-cooled pyrene- h_{10} for excitation of the $S_1 \ ^1B_{3u} \leftarrow S_0 \ ^1A_g$ prominent bands.

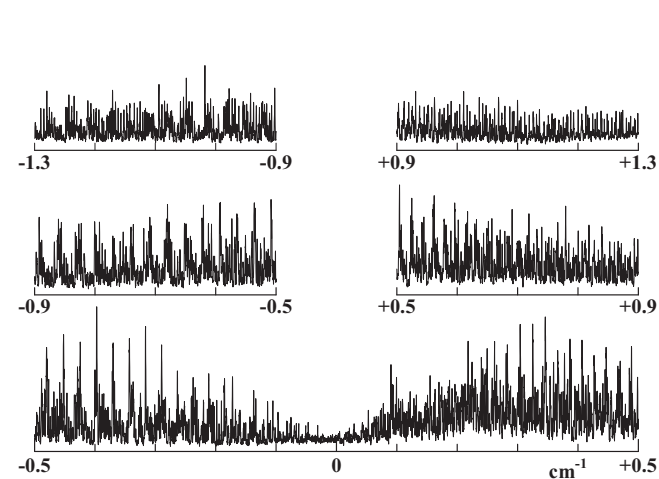


FIG. 6. Rotationally resolved fluorescence excitation spectrum of the $S_1 \ ^1B_{3u} \leftarrow S_0 \ ^1A_g \ 0_0^0$ transition of pyrene- h_{10} .

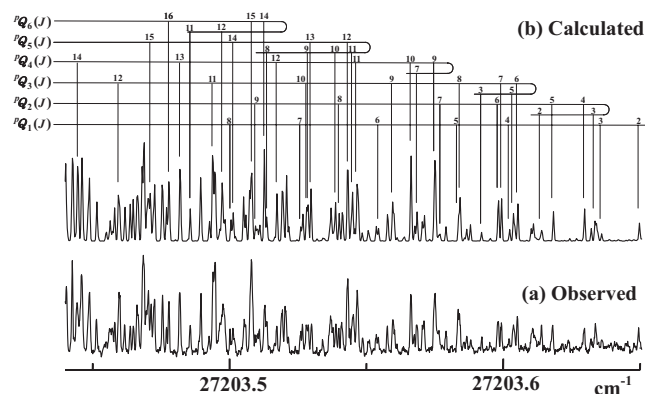


FIG. 7. (a) A part of the rotationally resolved fluorescence excitation spectrum of the $S_1 \ ^1B_{3u} \leftarrow S_0 \ ^1A_g$ 0_0^0 transition of pyrene- h_{10} is compared to (b) the calculated spectrum using the obtained constants shown in Table V.

$$\begin{aligned} \langle JK \pm 2M | H_r^{(A)} | JKM \rangle &= \left\{ \frac{1}{4}(B - C) - \delta_J J(J+1) \right. \\ &\quad \left. - \frac{1}{2} \delta_K [(K \pm 2)^2 + K^2] \right\} [J(J+1) \\ &\quad - K(K \pm 1)][J(J+1) - (K \pm 1)(K \pm 2)]^{1/2}, \end{aligned} \quad (3)$$

where $|JKM\rangle$ is the eigenfunction of a symmetric-top molecule. J is the quantum number of total angular momentum J , while K and M are quantum numbers of the projection of J along the molecule-fixed $a(y)$ axis and the space-fixed Z axis, respectively. A , B , and C are rigid-rotor rotational constants. Δ_J , Δ_{JK} , and Δ_K are symmetric-top quartic centrifugal distortion constants, and δ_J and δ_K are asymmetric-top distortion constants.

All rotational lines show b -type selection rules. The qQ lines ($\Delta K_a=0$, $\Delta J=0$) could not be seen around the band center, and the S_1 state has been confirmed to be $^1B_{3u}$. A part of the expanded spectrum in the P branch region is shown in Fig. 7. The intensity ratio with nuclear spin multiplicity is calculated to be $ee:eo:oo:oe=A_1:B_3:B_2:B_1=18:15:16:15$. We assigned 1510 rotational lines and obtained rotational constants of the S_0 and S_1 states. The results are tabulated in Table V. Although the spectrum is congested with many lines, we were able to assign most of them unambiguously. The transition energies calculated using the obtained rotational constants are in good agreement with the observed values and we could not find any energy shifts occurring through local perturbation. The calculated spectrum using the obtained rotational constants is compared to the observed one in Fig. 7.

The obtained rotational constants in the S_0 state ($A=0.033\,914\,7$, $B=0.018\,655\,0$, and $C=0.012\,040\,6\text{ cm}^{-1}$) are similar to those obtained by *ab initio* calculation retaining D_{2h} -symmetry ($A=0.033\,767\,4$, $B=0.018\,555\,7$, and $C=0.011\,975\,1\text{ cm}^{-1}$). Inertial defect $\Delta=I_c-I_b-I_a$ is negligibly small in both the S_0 and S_1 states. Therefore, we conclude that the pyrene molecule is planar and D_{2h} -symmetrical. Pyrene is not considered to be completely aromatic because it has 16π electrons. However, the isolated pyrene molecule has been found to be highly symmetrical with three orthogonal twofold axes and symmetry planes. It is worth noting that the rotational constants in the $S_1 \ ^1B_{3u}$ state are very similar to those in the $S_0 \ ^1A_g$ state. The differ-

TABLE V. Molecular constants of the $S_0 \ ^1A_g(v''=0)$, and $S_1 \ ^1B_{3u}(v'=0)$ states of pyrene- h_{10} . Rotational constants, band origin T_0 , and standard deviation σ are in units of cm^{-1} . Moments of inertia I_a , I_b , and I_c and inertial defect Δ are in units of kg m^2 .

Pyrene- h_{10}	$S_0 \ ^1A_g(v''=0)$	$S_1 \ ^1B_{3u}(v'=0)$
A	0.033 914 7(45)	0.033 622 3(44)
B	0.018 655 0(32)	0.018 486 4(32)
C	0.012 040 6(24)	0.011 932 9(23)
$D_K(\times 10^{-8})$	-2.8(12)	-1.9(11)
$D_{JK}(\times 10^{-8})$	3.88(67)	2.99(68)
$D_J(\times 10^{-9})$	-8.2(18)	-7.0(18)
$\delta_K(\times 10^{-8})$	-6.4(10)	-5.25(96)
$\delta_J(\times 10^{-9})$	-4.24(81)	-3.53(79)
κ	-0.395 23	-0.395 70
$I_a(\times 10^{-46})$	82.539	83.256
$I_b(\times 10^{-46})$	150.05	151.42
$I_c(\times 10^{-46})$	232.49	234.58
$\Delta(\times 10^{-46})$	-0.107 30	-0.095 34
T_0	...	27 203.680(3)
σ	...	0.000 34
Fit lines	...	1510
Band type		b

ences are less than 1% for all the A , B , and C values. This indicates that structural change upon electronic excitation is very small.

C. Excited-state dynamics in the $S_1 \ ^1B_{3u}$ state

π orbitals of pyrene calculated at the B3LYP/6-31G(d,p) level are shown in Fig. 8. ϕ_{53} (HOMO) and ϕ_{54} (LUMO) are uniformly distributed over each whole π bonds although no population is present at the molecular center. The $S_2 \ ^1B_{2u}$ state is approximately expressed by the configuration of HOMO $\rightarrow$ LUMO excitation. The wave functions of ϕ_{52} (HOMO-1) and ϕ_{55} (LUMO+1) are also uniform. The $S_1 \ ^1B_{3u}$ state is expressed by the almost equally mixed configuration of ϕ_{52} (HOMO-1) $\rightarrow$ ϕ_{54} (LUMO) and

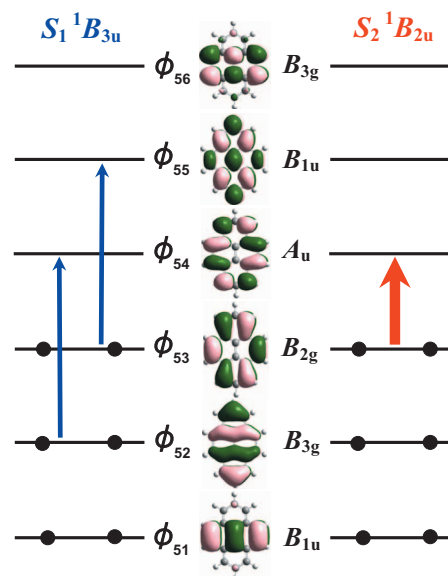


FIG. 8. π molecular orbitals of pyrene calculated at the B3LYP/6-311G(d,p) level.

$\phi_{53}(\text{HOMO}) \rightarrow \phi_{55}(\text{LUMO}+1)$ electronic excitations. The π bond order is defined by the sum over all π electrons of products of atomic orbital coefficients. Although the number of nodes in the wave function increases by two as a result of electronic excitation, the change in electronic distribution is small. The change in the bond order is further diluted among 16 electrons. Consequently, structural change resulting from electronic excitation is expected to be small.

We demonstrate that the excited-state dynamics in the S_1 state of pyrene is closely related to its geometrical structure. The fluorescence lifetime of the S_1 state of pyrene is very long, which indicates that all the radiationless transitions are very slow. Three processes that may be involved are predissociation, ISC to the triplet state, and IC to the S_0 state. The bond energies are significantly higher than the excitation energy of the $S_1 \rightarrow S_0$ transition so that predissociation does not occur at the zero-vibrational level in the S_1 state of pyrene. ISC is not expected to be efficient, according to El-Sayed's

rule that spin-orbit interaction between the $^1\pi\pi^*$ and $^3\pi\pi^*$ states is very weak for a planar molecule.²³ In order to confirm the triplet character, we observed broadening of a rotational line with the magnetic field, which is caused by the Zeeman splitting of M_J levels. The observed broadening was considerably small even at a magnetic field of 1.0 T. Accordingly, we conclude that the triplet mixing is minimal and that ISC to the triplet state is very slow.

Next we discuss the inactivity of IC in the S_1 $^1B_{3u}$ state of the isolated pyrene molecule. The IC to the S_0 state is caused by vibronic interaction. The energy difference between the zero-vibrational levels of the S_0 and S_1 states is $27\,204\text{ cm}^{-1}$, and the vibrational quantum number of the S_0 state is large in the S_1 energy region. The vibronic interaction is considered to be very weak when the equilibrium molecular structure and potential energy curves are approximately the same for both the S_0 and S_1 states.

The IC rate is given by^{24–27}

$$W_{IC} \propto \sum_i \left| \langle \phi_{S_0} | \frac{\partial}{\partial Q_i} | \phi_{S_1} \rangle \right|^2 \left| \langle \chi_{S_0}^{v''} | \frac{\partial}{\partial Q_i} | \chi_{S_1}^{v'=0} \rangle \right|^2 \delta(E_{S_1}^{v'=0} - E_{S_0}^{v''})$$

$$\approx \frac{\sum_i \left| \langle \phi_{S_0}(r, Q_0) | \left(\frac{\partial U(r, Q)}{\partial Q_i} \right)_{Q_0} | \phi_{S_1}(r, Q_0) \rangle \right|^2}{(E_{S_1}(Q_0) - E_{S_0}(Q_0))^2} \left| \langle \chi_{S_0}^{v''}(Q_i) | \frac{\partial}{\partial Q_i} | \chi_{S_1}^{v'=0}(Q_i) \rangle \prod_{j \neq i} \langle \chi_{S_0}^{v''}(Q_j) | \chi_{S_1}^{v'=0}(Q_j) \rangle \right|^2 \delta(E_{S_1}^{v'=0} - E_{S_0}^{v''}). \quad (4)$$

Here, ϕ_{S_0} and ϕ_{S_1} represent the electronic wave functions of the S_0 and S_1 states, respectively. $\chi_{S_0}^{v''}$ is the vibrational wave function of the high-vibrational level in the S_0 state and $\chi_{S_1}^{v'=0}$ is that of the zero-vibrational level in the S_1 state. δ is the Dirac delta function and Q_i is a normal coordinate. The second and third matrix elements are nonadiabatic Franck–Condon overlap of a promoting mode and Franck–Condon overlap of other modes, respectively. In D_{2h} symmetry, the S_1 $^1B_{3u}$ state couples with the S_0 state through a b_{3u} vibration. The energy gap between the S_0 and S_1 states is large and their molecular structures are similar. The amplitude of wave function of the S_1 zero-vibrational level is large at the equilibrium nuclear positions (potential bottom), whereas that of the S_0 high-vibrational level is large near both edges of the potential energy curve. According to energy gap law,²⁴ the vibrational overlap is very small if the potential energy curves are approximately the same for the S_0 and S_1 states and the energy gap is large. We have shown that the equilibrium molecular structure is almost identical by the determined rotational constants and that the slope of the potential energy curve is also very similar by the observed vibrational energy. Therefore, the IC process in the isolated pyrene molecule is presumed to be very slow. Remarkably fast radiationless transitions such as ISC and IC with the quantum yields of 0.37 (Ref. 28) and 0.71,²⁹ respectively, were observed in condensed phases. These processes are attributed to

structural change due to intermolecular interaction with the environment. The rates of radiationless transitions are considered to be strongly dependent on the molecular structure and potential energy curves so that it is noteworthy to compare the vibrational energies of an isolated molecule with those in condensed phases. For a_g modes, the vibrational energies are approximately the same for gas and condensed phases, as shown in Table VI.^{1,5,30,31} Although the observed vibrational energies in the S_0 state are similar for gas and condensed phases, those in the S_1 state are appreciably different, supporting that the potential energy curves are affected by intermolecular interaction. Such slow nonradiative decay in the S_1 $^1B_{3u}$ state is supposed to be common for large PAHs, in which the molecular structure is not significantly changed upon electronic excitation. Predissociation does not occur because the S_1 excitation energy is less than the bond energies. The ISC to the triplet state scarcely occurs according to El-Sayed's rule. The IC rate is considered to be strongly correlated with the geometrical structure. The structural change upon $S_1 \leftarrow S_0$ excitation is small in large PAH molecules. In such cases, the overlap of vibrational wave functions is expected to be small between high-vibrational levels of the S_0 state and the zero-vibrational level of the S_1 state. Consequently, the vibronic coupling becomes very weak and the IC in the S_1 state is expected to be slow.

We assume that IC is relatively faster in the B_{2u} state of a PAH molecule with D_{2h} symmetry. For instance, the fluo-

TABLE VI. Observed a_g vibrational energies of pyrene- h_{10} for gas and condensed phases.

Mode number	Gas ^a	In biphenyl ^b	In heptane ^c	Crystal ^d
$S_0 \ ^1A_g$				
5	1571	1552	1550	1552
6	1410	1408	1405	1408
7	1326	1355		1352
8	1238	1240	1236	1243
9	1140	1144		
10	1072	1063	1065	1067
11	805	801	801	805
12	593	596	588	593
13	402	408	405	408
$S_1 \ ^1B_{3u}$				
5	1527			
6	1397	1452	1444	1450
7	1303	1339	1328	1322
8	1250	1249	1245	1237
9	1144			
10	1023	1030	1029	1030
11	783	782	779	
12	572	583	573	570
13	393	400	393	392

^aPresent work.^bReference 1.^cReference 5.^dCrystal from Ref. 30 for the $S_0 \ ^1A_g$ state and in hexane from Ref. 31 for the $S_1 \ ^1B_{3u}$ state.

rescence lifetime of anthracene- d_{10} (4 ns) is much shorter than that of anthracene- h_{10} (18 ns).^{32,33} This fact suggests that its promoting mode is C–H stretching, which includes nuclear motion along the y axis, that is, b_{2u} vibration. The IC rate in the $^1B_{3u}$ state, which couples with the $S_0 \ ^1A_g$ state through a b_{3u} vibrational mode, is presumed to be slow probably because of the small nonadiabatic Franck–Condon overlap. The observed fluorescence lifetimes of naphthalene (~ 300 ns) and pyrene (~ 1400 ns) are long in contrast to the short lifetimes of the $S_1 \ ^1B_{2u}$ states of anthracene (18 ns), tetracene (20 ns),³⁴ and perylene (8.9 ns).³⁵ It is noteworthy that the lifetime of pyrene is significantly longer than that of naphthalene in spite of the pyrene's larger molecular size and level density. The mixing with the $S_2 \ ^1B_{2u}$ state might be shortening the lifetime of the $S_1 \ ^1B_{3u}$ state. However, the energy gap between the S_1 and S_2 states is almost the same for pyrene and naphthalene. The difference in their lifetimes, therefore, cannot be attributed to vibronic coupling between the S_1 and S_2 states. Another possibility is the potential energy curve. The obtained rotational constants of the S_1 state are similar to those of the S_0 state and the change in geometrical structure (displacement) is thought to be small. We suppose that the slope of the potential energy curve (distortion) is slightly shallower in the S_1 state of naphthalene. In pyrene, the slope is nearly the same in both the S_1 and S_0 states. In this case, the nonadiabatic Franck–Condon overlap is especially small and the $S_1 \rightarrow S_0$ IC is expected to be minimal.

IVR is generally efficient in large PAHs because many vibronic levels are considerably mixed by the Coriolis inter-

action or anharmonic interaction (Fermi resonance). The threshold excess energy depends on the molecular size or the number of normal modes because the interaction strength is approximately proportional to the coupling level density. The threshold energies are 3000, 2122, and 1300 cm^{-1} for benzene,³⁶ naphthalene,³⁷ and anthracene,³³ respectively. In pyrene, IVR is remarkable in the vibronic level with an excess energy of larger than 730 cm^{-1} . Radiationless transitions are often enhanced by IVR so that the fluorescence lifetime becomes shorter at high-vibrational levels. In pyrene, however, this enhancement is not significant and the decrease in lifetime with excess energy is relatively small.

IV. SUMMARY

We have analyzed vibronic structure in the $S_0 \ ^1A_g$ and $S_1 \ ^1B_{3u}$ states of the isolated pyrene molecule by fluorescence excitation and dispersed fluorescence spectra, as well as by *ab initio* calculation. The vibrational energies of the S_0 and S_1 states are similar and totally symmetric a_g vibronic bands are weaker than the 0_0^0 band. This suggests that the stable geometrical structures of pyrene are almost identical in both states. The b_{1g} bands were observed with appreciable intensity through vibronic interaction with the $S_2 \ ^1B_{2u}$ state. The rotational structure has been analyzed using ultrahigh-resolution fluorescence excitation spectroscopy. We accurately determined the rotational constants at the zero-vibrational levels of the $S_0 \ ^1A_g$ and $S_1 \ ^1B_{3u}$ states. It has been shown that the molecule is planar and that its geometrical structure does not significantly change upon electronic excitation. The vibronic coupling between the S_1 zero-vibrational level and S_0 high-vibrational levels is very small and the IC to the S_0 state becomes very slow. Predissociation does not occur in the S_1 state of PAHs, and ISC is also considered to be very slow. Consequently, all of the possible radiationless transitions are slow and fluorescence lifetime is accordingly very long in the S_1 state of the isolated pyrene molecule. The IVR rate is expected to be proportional to the density of coupling vibrational levels and the broad emission can be observed in the dispersed fluorescence spectrum even for the excitation of a low excess energy level at 730 cm^{-1} .

ACKNOWLEDGMENTS

This work was supported by the JSPS Asian CORE program on Frontiers of Materials, Photo, and Theoretical Molecular Sciences, and also by the Joint Studies Program (2008–2009) of the Institute for Molecular Science. The authors are grateful to Professor Hiroki Nakamura (Institute for Molecular Science, Japan) and Professor Yuh-Lin Wang (Institute of Atomic and Molecular Science, R.O.C.). The authors thank to M. Aoyama, M. Kondo, T. Yano, N. Mizutani, and M. Suzui (Equipment Development Center, Institute for Molecular Science, Japan), and K. Mitsui and N. Okada (Advanced Technology Center, National Astronomical Observatory, Japan) for creating up a mirror setup for high-efficiency fluorescence collection.

¹A. Bree and V. V. B. Vilks, *Spectrochim. Acta, Part A* **27**, 2333 (1971).²W. H. van Leeuwen, J. Langelaar, and J. D. W. van Voorst, *Chem. Phys. Lett.* **13**, 622 (1972).

- ³L. Peter and G. Vaubel, *Chem. Phys. Lett.* **21**, 158 (1973).
- ⁴F. W. Langkilde, E. W. Thulstrup, and J. Michl, *J. Chem. Phys.* **78**, 3372 (1983).
- ⁵A. Bree, A. Leyderman, P. R. Salvi, and C. Taliani, *Chem. Phys.* **110**, 211 (1986).
- ⁶C. M. Jones and S. A. Asher, *J. Chem. Phys.* **89**, 2649 (1988).
- ⁷T. Kamisuki, H. Kataoka, and S. Maeda, *Spectrochim. Acta, Part A* **54**, 277 (1998).
- ⁸J. Clairemidi, P. Bréchnignac, G. Moreels, and D. Pautet, *Planet. Space Sci.* **52**, 761 (2004).
- ⁹E. A. Mangle and M. R. Topp, *J. Phys. Chem.* **90**, 802 (1986).
- ¹⁰N. Ohta, H. Baba, and G. Marconi, *Chem. Phys. Lett.* **133**, 222 (1987).
- ¹¹N. A. Borisevich, L. B. Vodovatov, G. G. D'yachenko, V. A. Petukhov, and M. A. Semyonov, *J. Appl. Spectrosc.* **62**, 482 (1995).
- ¹²N. A. Borisevich, L. B. Vodovatov, G. G. D'yachenko, V. A. Petukhov, and M. A. Semyonov, *Opt. Spectrosc.* **78**, 213 (1995).
- ¹³G. Rouillé, S. Krasnokutski, F. Huysken, T. Henning, O. Sukhorukov, and A. Staicu, *J. Chem. Phys.* **120**, 6028 (2004).
- ¹⁴J. R. Platt, *J. Chem. Phys.* **17**, 484 (1949).
- ¹⁵S. Gerstenkorn and P. Luc, *Atlas du Spectre d'Absorption de la Molécule D'Iode* (CNRS, Paris, 1978).
- ¹⁶G. Orlandi and W. Siebrand, *J. Chem. Phys.* **58**, 4513 (1973).
- ¹⁷K. Ohno, *Chem. Phys. Lett.* **64**, 560 (1979).
- ¹⁸K. Ohno, *J. Chem. Phys.* **95**, 5524 (1991).
- ¹⁹F. Negri and G. Orlandi, *Theor. Comput. Chem.* **16**, 5129 (2005).
- ²⁰M. J. Frisch, G. W. Trucks, H. B. Schlegel *et al.*, GAUSSIAN 03, Revision B.05, Gaussian, Inc., Pittsburgh PA, 2003.
- ²¹C. W. Bauschlicher and S. R. Langhoff, *Spectrochim. Acta, Part A* **53**, 1225 (1997).
- ²²J. K. G. Watson, in *Vibrational Spectra and Structure*, edited by J. R. Durig (Dekker, New York, 1977), Vol. 6, Chap. 1.
- ²³M. A. El-Sayed, *J. Chem. Phys.* **38**, 2834 (1963).
- ²⁴W. Siebrand, *J. Chem. Phys.* **46**, 440 (1967).
- ²⁵S. H. Lin, *J. Chem. Phys.* **44**, 3759 (1966).
- ²⁶C. H. Chin, H. Y.-J. Shin, H.-W. Wang, Y.-L. Chen, C.-C. Wang, S. H. Lin, and M. Hayashi, *J. Chin. Chem. Soc. (Taipei)* **53**, 131 (2006).
- ²⁷M. Bixon and J. Jortner, *J. Chem. Phys.* **48**, 715 (1968).
- ²⁸H. Hirano and T. Azumi, *Chem. Phys. Lett.* **86**, 109 (1982).
- ²⁹N. I. Nijegorodov and W. S. Downey, *J. Phys. Chem.* **98**, 5639 (1994).
- ³⁰A. Bree, R. A. Kydd, T. N. Misra, and V. V. B. Vilks, *Spectrochim. Acta, Part A* **27**, 2315 (1971).
- ³¹P. R. Salvi, E. Castellucci, P. Foggi, and R. Quacquarelli, *Chem. Phys.* **83**, 345 (1984).
- ³²S. Okajima, B. E. Forch, and E. C. Lim, *J. Phys. Chem.* **87**, 4571 (1983).
- ³³W. R. Lambert, P. M. Felker, and A. H. Zewail, *J. Chem. Phys.* **81**, 2209 (1984).
- ³⁴A. Amirav, U. Even, and J. Jortner, *J. Chem. Phys.* **75**, 3770 (1981).
- ³⁵A. J. Kaziska, S. A. Wittmeyer, A. L. Motyka, and M. R. Topp, *Chem. Phys. Lett.* **154**, 199 (1989).
- ³⁶K. G. Spears and S. A. Rice, *J. Chem. Phys.* **55**, 5561 (1971).
- ³⁷S. M. Beck, J. B. Hopkins, D. E. Powers, and R. E. Smalley, *J. Chem. Phys.* **74**, 43 (1981).