

Cobalt-catalyzed carboxylation of propargyl acetates with carbon dioxide

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The cobalt-catalyzed carboxylation of propargyl acetates with CO₂ (1 atm) is described. The reaction proceeds at room temperature with Mn powder as a reducing reagent. Various propargyl acetates are converted to the corresponding carboxylic acids in good to high yields.

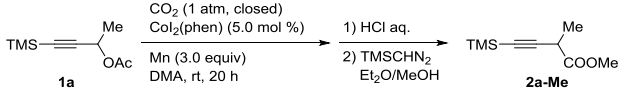
Carbon dioxide (CO₂) is an environmentally friendly raw material and its utilization as a sustainable carbon source is one of the most important challenges in homogeneous transition-metal catalysis.¹ In particular, C–C bond forming reactions using CO₂ are the most promising.^{1a} Reactions using highly reactive Grignard and organolithium reagents with CO₂ are fundamental for the C–C bond formation, but chemoselectivity with these reagents is very poor. In contrast, less reactive arylboronic esters^{2a–c} and arylzinc compounds^{2d,e} were found to react with CO₂ with good chemoselectivity in the presence of a transition-metal catalyst. However, all these organometallic compounds are mainly synthesized from the corresponding aryl halides. Thus, the direct carboxylation of aryl halides is more straightforward and efficient. We recently reported the nickel-catalyzed direct carboxylation of aryl and vinyl chlorides with CO₂ (1 atm) at room temperature,^{3a,b} while carboxylations of more reactive aryl bromides^{3c} and benzyl chlorides^{3d} as well as inert C–O bonds^{3e} were also reported.⁴

Reactions of allylic and propargylic compounds with CO₂ afford a variety of unsaturated carboxylic acids. To date, these transformations consume a stoichiometric amount of the corresponding Grignard, organolithium, or other metal reagents.⁵ These reactions also have problems in regioselectivity and chemoselectivity.

Therefore, the development of a new selective methodology is highly desirable. Allyl and propargyl esters (typically, acetates or carbonates) are well known as efficient *electrophiles* in transition-metal-catalyzed C–C bond forming reactions.⁶ In order to utilize these electrophiles with CO₂, umpolung⁷ reactivity of these esters is crucial. Actually, the reactions of *allyl* esters with CO₂ were carried out under electrochemical conditions in the presence of Pd or Ni catalysts;⁸ however, the yields and regioselectivities were low. Furthermore, there is no precedent for the carboxylation of *propargyl* esters with CO₂.⁹ Herein, we report the Co-catalyzed carboxylation of propargyl acetates with CO₂ utilizing Mn powder as a reducing agent. Various propargyl acetates were converted to the corresponding carboxylic acids under 1 atm CO₂ at room temperature.

The carboxylation of propargyl acetate **1a** was carried out under CO₂ (1 atm) at room temperature in the presence of CoI₂(phen)¹⁰ (phen

Table 1 Reaction optimization^a

		
Entry	Catalyst System: Change from the Standard Conditions	Yield of 2a-Me (%) ^b
1	none	83 (82) ^c
2	Without CoI ₂ (phen)	0
3	CoI ₂ in place of CoI ₂ (phen)	0
4	Without Mn powder	0
5	Mn Powder (0.60 mmol, 1.2 equiv)	74
6	CoBr ₂ (phen) in place of CoI ₂ (phen)	80
7	CoI ₂ (bpy) in place of CoI ₂ (phen)	76
8	CoI ₂ (PPh ₃) ₂ in place of CoI ₂ (phen)	23
9	CoI ₂ (dppe) in place of CoI ₂ (phen)	0
10	Zn in place of Mn	41
11	Mg in place of Mn	57
12 ^d	NiCl ₂ (PPh ₃) ₂ in place of CoI ₂ (phen)	7

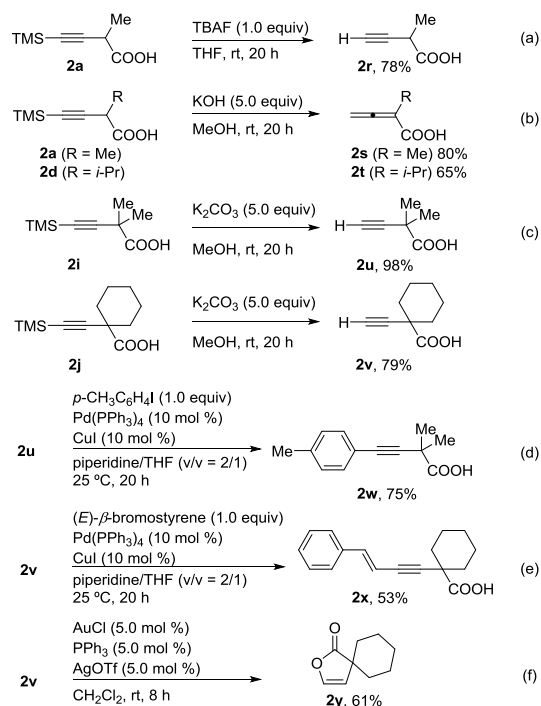
^a Reaction conditions; **1a** (0.50 mmol), CoI₂(phen) (0.025 mmol, 5.0 mol %), Mn powder (1.5 mmol, 3.0 equiv), CO₂ (1 atm), in DMA (0.50 mL), at room temperature for 20 h. ^b Determined by GC analysis. ^c Isolated yield of **2a**. ^d With Et₄Ni (0.05 mmol, 10 mol %).

= 1,10-phenanthroline) and Mn powder (3.0 equiv) in DMA (*N,N*-dimethylacetamide) (Table 1). The yield of the corresponding carboxylic acid (**2a**) was determined by gas chromatographic (GC) analysis after derivatization to the corresponding methyl ester (**2a-Me**). Under the standard conditions, **2a-Me** was obtained in 83% yield (entry 1). Compound **2a** was isolated from the reaction mixture in 82% yield. Without CoI₂(phen), **2a-Me** was not obtained (entry 2). In the absence of phen (i.e., CoI₂ as the catalyst), **1a** was not converted (entry 3). Without the addition of Mn powder, the carboxylation did not proceed at all (entry 4). When the amount of Mn powder was reduced to 1.2 equiv, the yield of **2a-Me** was decreased to 74% (entry 5). CoBr₂(phen) was also a good catalyst and afforded **2a-Me** in 80% yield (entry 6). Employing CoI₂(bpy) (bpy = 2,2'-bipyridine) as a catalyst, **2a-Me** was obtained in 76% yield (entry 7). Thus, phen and bpy show comparable efficiency as the ligand. In contrast, CoI₂(PPh₃)₂ and CoI₂(dppe) (dppe = 1,2-diphenylphosphinoethane) suppressed the carboxylation considerably (entries 8 and 9). Other reducing agents such as Zn powder and Mg turnings gave **2a-Me** in only moderate yields (entries 10 and 11). NiCl₂(PPh₃)₂, which was an

Table 2 Cobalt-catalyzed carboxylation of various propargyl acetates^a

$ \begin{array}{c} \text{R}^1 \text{---} \text{C} \equiv \text{C} \text{---} \text{C}(\text{R}^2)(\text{R}^3) \text{---} \text{OAc} \\ \mathbf{1} \end{array} \xrightarrow[\text{Mn (3.0 equiv), DMA, rt, 20 h}]{\text{CO}_2 \text{ (1 atm, closed), CoI}_2(\text{phen}) \text{ (5.0 mol \%)} } \xrightarrow{\text{HCl aq.}} \begin{array}{c} \text{R}^1 \text{---} \text{C} \equiv \text{C} \text{---} \text{C}(\text{R}^2)(\text{R}^3) \text{---} \text{COOH} \\ \mathbf{2} \end{array} $			
Entry	Substrate 1	Product 2	Yield (%) ^b
1			80
2			75
3			79
4			65
5			85
6			73
7			87
8 ^c			71
9 ^c			80
10 ^c			80
11 ^c			46
12			40
13			88
14			55
15			57
16			26

^a Reaction conditions; propargyl acetate (**1**, 0.50 mmol), CoI₂(phen) (0.025 mmol, 5.0 mol %), Mn powder (1.5 mmol, 3.0 equiv), CO₂ (1 atm), in DMA (0.50 mL), at room temperature for 20 h. ^b Isolated yield. ^c CoI₂(bpy) (0.025 mmol, 5.0 mol %) was used as a catalyst.

**Scheme 1** Derivatization of carboxylated products.

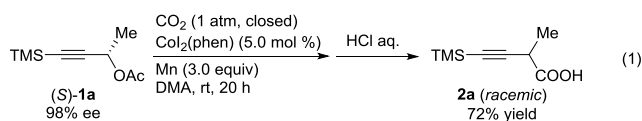
efficient catalyst for the carboxylation of aryl chloride,^{3a} did not show good catalytic activity (entry 12). Other nickel catalysts such as NiBr₂(bpy) and NiI₂(phen) were not efficient. With regard to the choice of solvent, 1,3-dimethyl-2-imidazolidinone (DMI) and DMF were also suitable, while reactions in THF and toluene afforded **2a-Me** in 26% and 0% yields, respectively (Table S1).¹¹

The carboxylation of various propargyl acetates was carried out in the presence of CoI₂(phen) or CoI₂(bpy) as a catalyst (Table 2). The carboxylation reaction of acetates of secondary alcohols (**1b–h**) bearing TMS in the R¹ position proceeded smoothly and afforded the corresponding carboxylic acids (**2b–h**) in high isolated yields (entries 1–7). It is noteworthy that ester (**1e**), chloro (**1f**), terminal alkene (**1g**), and furan (**1h**) functionalities were compatible in the reaction (entries 4–7). When the carboxylation of acetates derived from tertiary alcohols was examined with CoI₂(phen), conversion of starting material remained low. In that case, CoI₂(bpy) was found to be a good catalyst and provided the carboxylated products (**2i–l**) in good to high yields (entries 8–11). Amide functionality (**1l**) was also tolerated in the reaction (entry 11). The acetate of primary propargylic alcohol (**1m**) also provided the corresponding carboxylic acid (**2m**) in moderate yield (entry 12). A substituent on the alkyne carbon of **1** (R¹, Table 2) affects the carboxylation considerably. As the substituent R¹ became less bulky, the yields of the carboxylated products (**2**) decreased; **1n** (R¹ = TBS, TBS = *tert*-butyldimethylsilyl), **1o** (R¹ = CMe₂(OTBS)), **1p** (R¹ = *t*-Bu), and **1q** (R¹ = Cy) afforded the corresponding products (**2n–q**) in 88%, 55%, 57%, and 26% yields, respectively (entries 13–16). Propargyl acetate having phenyl ring (**1**: R¹ = Ph, R² = Me, R³ = H) afforded the product in 9% yield. Substrate bearing terminal alkyne moiety (**1**: R¹ = H, R² = Me, R³ = H) did not provide the carboxylated product.¹²

The TMS group of the products in Table 2 could easily be removed via protodesilylation¹³ in the presence of a suitable base. In the case of α,α -disubstituted carboxylic acids such as **2a**, reaction with tetrabutylammonium fluoride (TBAF, 1.0 M in THF) provided the corresponding carboxylic acid **2r** in 78% yield (Scheme 1a). In contrast, when **2a** or **2d** were treated with KOH (crushed), carboxylic acids bearing an allenyl moiety (**2s**, **2t**) were selectively obtained in

80% and 65% yields, respectively (Scheme 1b). A similar reaction of **2a** with K_2CO_3 resulted in the formation of a mixture of **2r** and **2s** (**2r**/**2s** = 1/3). On the other hand, α,α,α -trisubstituted carboxylic acids such as **2i** and **2j** reacted with K_2CO_3 to give **2u** and **2v** in 98% and 79% yields (Scheme 1c). Aryl and alkenyl carbons were introduced onto the terminal alkyne moiety of **2u** and **2v** by the Sonogashira coupling reaction^{14a} (**2w** and **2x**, Scheme 1d and 1e). Moreover, Au-catalyzed intra-molecular cyclization^{14b} of **2v** provided unsaturated γ -lactone **2y** smoothly (Scheme 1f). Thus these TMS moieties are very useful for the further derivatization.

When an optically pure (*S*)-**1a** was employed as the substrate in the present carboxylation, a racemic **2a** was obtained in 72% yield (eqn (1)).



A plausible reaction mechanism is shown in Scheme 2. Initially, the reduction of a Co(II) complex with manganese affords a Co(I) catalyst species (**A**). Then, oxidative addition of a propargyl acetate (**1**) takes place via C–O bond cleavage, giving a Co(III) intermediate (**B**) (step a). Subsequent reduction of propargyl Co(III) with manganese gives propargyl cobalt(II) species (**C**) (step b).^{3a,b,15} Then, the more nucleophilic^{3b} Co(II) species (**C**) reacts with CO_2 to give the carboxylatocobalt intermediate (**D**) (step c). Finally, the reduction of **D** with manganese affords the corresponding manganese carboxylate and the Co(I) catalytic species (**A**) regenerates (step d). Further studies about the reaction mechanism are now in progress.

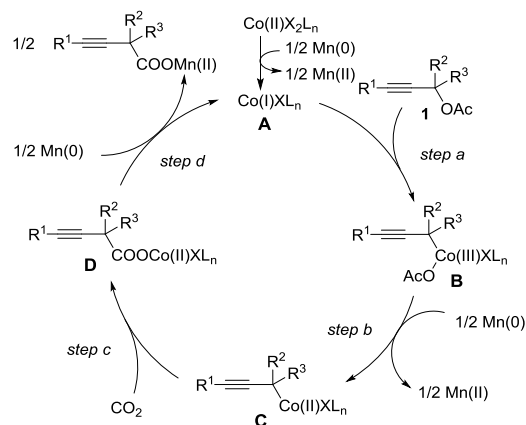
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Notes and references

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† Electronic Supplementary Information (ESI) available: Experimental procedures and characterization data for all new compounds. See DOI: 10.1039/c000000x/

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Scheme 2 Plausible catalyst cycle.

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- Recently, in the Ni-catalyzed carboxylation of benzylic chlorides with CO_2 ,^{3d} $C_6H_5C\equiv CCH_2Cl$ was used as a substrate and afforded $C_6H_5C\equiv CCH_2COOH$ in only 29% yield; other propargylic substrates were not examined in the reaction.
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- In those cases including entries 14–16 in Table 2, most substrates were consumed and some oligomerizations occurred judging from GPC (gel permeation chromatography) analysis of the resulting reaction mixtures.
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