



Cite this: DOI: 10.1039/d1sc00206f

All publication charges for this article have been paid for by the Royal Society of Chemistry

Dual Ag/Co cocatalyst synergism for the highly effective photocatalytic conversion of CO₂ by H₂O over Al-SrTiO₃[†]

Shuying Wang,^a Kentaro Teramura,^{ab} Takashi Hisatomi,^c Kazunari Domen,^{cd} Hiroyuki Asakura,^{ab} Saburo Hosokawa^{ab} and Tsunehiro Tanaka^{ab}

Loading Ag and Co dual cocatalysts on Al-doped SrTiO₃ (AgCo/Al-SrTiO₃) led to a significantly improved CO-formation rate and extremely high selectivity toward CO evolution (99.8%) using H₂O as an electron donor when irradiated with light at wavelengths above 300 nm. Furthermore, the CO-formation rate over AgCo/Al-SrTiO₃ (52.7 μmol h⁻¹) was a dozen times higher than that over Ag/Al-SrTiO₃ (4.7 μmol h⁻¹). The apparent quantum efficiency for CO evolution over AgCo/Al-SrTiO₃ was about 0.03% when photoirradiated at a wavelength at 365 nm, with a CO-evolution selectivity of 98.6% (7.4 μmol h⁻¹). The Ag and Co cocatalysts were found to function as reduction and oxidation sites for promoting the generation of CO and O₂, respectively, on the Al-SrTiO₃ surface.

Received 12th January 2021
Accepted 24th February 2021

DOI: 10.1039/d1sc00206f

rsc.li/chemical-science

Introduction

The photo-irradiative conversion of CO₂ into chemicals over semiconductor photocatalysts typically includes three steps: (1) light harvesting, (2) photoexcited electron-hole pair generation, separation, and transfer (charge transfer), and (3) surface catalytic reactions that include reduction by electrons and oxidation by holes.^{1–3} An urgent challenge involves preventing the rapid recombination of photogenerated electron-hole pairs within single photocatalyst particles, which results in lower photocatalytic performance of the semiconductor catalyst.^{4–6} Cocatalysts that provide essential functions during photocatalytic reactions not only promote the separation of photo-generated electron-hole pairs, but also reduce the activation potential and serve as active sites for the photocatalytic evolution of products.^{1,2,4,7–14} Photocatalyst surfaces decorated with Ag nanoparticles were reported in 2011 to be good cocatalysts that exhibit superior selectivities for photocatalytic products, such as CO from CO₂, when photoirradiated.¹² In addition, many kinds of material, including noble metals (such as Au,¹² Pt,^{15,16} and Pd¹⁷), non-noble metals (such as Cu^{15,18}), and metal oxides (such as MgO,¹⁶ NiO,¹⁹ and Cu₂O²⁰) have been studied as

cocatalysts for the photocatalytic conversion of CO₂ into chemicals such as CO and CH₄. Furthermore, dual cocatalysts have been investigated to overcome the shortcomings of single cocatalyst-loaded photocatalysts; *i.e.*, to improve CO₂ chemisorption,¹⁶ suppress the back reaction,²¹ and enhance the consumption of photogenerated holes.^{22,23} Dual metal-metal alloy cocatalysts (such as Pt-Cu,¹⁵ and Au-Cu^{24,25}) and metal-metal-oxide dual cocatalysts (such as Pt-MgO,¹⁶ Ni@NiO,²⁶ Ag@Cr,^{21,27} and Pt@Cu₂O²⁸) have been loaded onto the surfaces of photocatalysts to improve their photocatalytic performance for the conversion CO₂ into chemicals when photoirradiated. Specifically, a photocatalytic reaction contains two important components: (1) photoexcited electrons for the reduction reaction, and (2) photoexcited holes for the oxidation reaction, based on charge balance. Normally, attention is only paid to the photocatalytic CO₂ reduction side involving photoexcited electrons, while the H₂O oxidation process has largely been neglected. Improving the activity of the oxidation half-reaction should reduce photoexcited electron-hole recombination inside a single photocatalyst, which would enhance the activity of the reduction half reaction as larger numbers of electrons are transferred for the photocatalytic reaction on the surface of the photocatalyst.

The morphological structure of a semiconductor catalysts is known to clearly affect photocatalytic performance, as photocatalytic reactions occur on its surface. Recently, Yu *et al.* reported that the photocatalytic activity for CO₂ reduction over anatase TiO₂ depended on the {001} : {101} facet ratio.²⁹ In particular, the crystal structures of La-doped NaTaO₃,³⁰ anatase TiO₂,^{31,32} BiVO₄,^{11,33,34} 18-facet-SrTiO₃,³⁵ and Al-SrTiO₃ (ref. 36 and 37) have been studied as water-splitting photocatalysts, with photogenerated electrons and holes reportedly spatially

^aDepartment of Molecular Engineering, Graduate School of Engineering, Kyoto University, Kyoto 615-8510, Japan. E-mail: teramura@moleng.kyoto-u.ac.jp; tanakat@moleng.kyoto-u.ac.jp

^bElements Strategy Initiative for Catalysts and Batteries, Kyoto University, Kyoto 615-8510, Japan

^cResearch Initiative for Supra-Materials, Shinshu University, 4-17-1 Wakasato, Nagano 380-8553, Japan

^dThe University of Tokyo, 7-3-1 Hongo, Bunkyo-ku, Tokyo 113-8656, Japan

[†] Electronic supplementary information (ESI) available: Experimental details, and characterizations. See DOI: 10.1039/d1sc00206f

transferred to different exposed facets when photoirradiated, which is considered to enhance photocatalytic activity. Li *et al.* proposed that cocatalysts on the spatial facets of the BiVO₄ photocatalyst not only serve their traditional reaction roles, but also align the built-in electric field vectors of the photocatalyst particles, which maximizes the separation and transfer of photoexcited electrons and holes.¹¹ CoO_x and other Co species have recently been loaded on various photocatalysts as oxidation cocatalysts that enhance photocatalytic oxygen-evolution reaction (OER) activity.^{38,39} In particular, Domen *et al.* reported that Al-SrTiO₃ modified with the dual Rh/Cr₂O₃ cocatalyst and CoOOH on its various facets exhibited extremely high photocatalytic water-splitting activity without losses due to charge recombination, with a quantum efficiency of up to 96% at wavelengths of 350–360 nm.³⁷ Furthermore, the photocatalytic activity for CO₂ reduction has also been significantly improved by simultaneously loading dual cocatalysts (reduction and oxidation cocatalysts) onto different spatial facets of a photocatalyst.^{22,40–42} In our previous work, Al-SrTiO₃ fabricated using a flux method and modified with a Ag cocatalyst showed good efficiency and selectivity towards CO evolution in an aqueous solution when irradiated with light at wavelengths above 300 nm.^{43,44} In addition, the Ag cocatalyst on the surface of Al-SrTiO₃ prepared using a chemical reduction (CR) method was found to show good photocatalytic performance for the reduction of CO₂ compared to those prepared by impregnation (IMP) and photodeposition (PD) methods, because highly dispersed metallic Ag nanoparticles were produced.^{44–46} Herein, we report the effect of Ag and Co dual-cocatalyst loading on Al-SrTiO₃ (AgCo/Al-SrTiO₃) on the photocatalytic reactivity for the conversion of CO₂ into CO when photoirradiated, with water as the reductant.

Results and discussion

Fig. 1(b) and (c) show the XRD patterns of SrTiO₃ and Al-SrTiO₃, in which most peaks are assigned to the SrTiO₃ phase (Fig. 1(a)), with only a few small peaks attributable to the Sr₃Ti₂O₇ phase. Moreover, the impurity phase was absent after doping with Al using the flux method (Fig. 1(c)), with all peaks assigned to the Al-SrTiO₃ phase. Furthermore, Fig. 1(d) and (e) show SEM

Table 1 Photocatalytic conversion of CO₂ into CO by reduction with H₂O over various cocatalyst-loaded photocatalysts prepared by CR method^a

Photocatalyst	Product formation rate (μmol h ⁻¹)			Selectivity toward CO (%)	e ⁻ /h ⁺
	H ₂	O ₂	CO		
Al-SrTiO ₃	0.35	0.2	0.08	18.5	1.08
Ag/Al-SrTiO ₃	0.1	2.6	4.7	97.7	0.92
Co/Al-SrTiO ₃	41.6	19.5	0.1	0.2	1.07
AgCo/Al-SrTiO ₃ _CR	0.1	25.8	52.7	99.8	1.02
AgCo/Al-SrTiO ₃ _PD	14.1	17.5	20.5	59.2	0.99
AgCo/Al-SrTiO ₃ _IMP	0.2	2.8	4.1	95.0	0.77

^a Photocatalytic reaction conditions: amount of photocatalyst, 0.5 g; amount of Ag loaded, 1.7 mol%; amount of Co loaded, 0.85 mol%; volume of the reaction solution (H₂O), 1.0 L; additive, 0.1 M NaHCO₃; CO₂ flow rate, 30 mL min⁻¹; light source, 400 W high-pressure Hg lamp with a Pyrex® jacket (to cutoff light at λ < 300 nm).

images of SrTiO₃ and Al-SrTiO₃. The SrTiO₃ and Al-SrTiO₃ particles exhibit irregular polyhedral morphologies, with numerous nanosteps formed on the spatial surfaces of Al-SrTiO₃, as reported previously.⁴³ The morphologies and crystalline structures of the prepared samples are consistent with those in previous reports.^{36,43}

The photocatalytic performance of Al-SrTiO₃ modified with various cocatalysts, including single Ag and Co, and dual Ag and Co cocatalysts prepared by the CR method are summarized in Table 1. H₂, O₂, and CO were detected by GC as the gaseous products, and no liquid products were observed in the reaction solutions by using high performance liquid chromatography (HPLC). Importantly, the consumed e⁻/h⁺ ratios were found to be close to 1.0 over the SrTiO₃, Ag/SrTiO₃, Co/Al-SrTiO₃, and AgCo/Al-SrTiO₃ photocatalysts. Table S1† shows that Ag and Co were successfully loaded on the surface of the Al-SrTiO₃ photocatalyst. Only very small amounts of gaseous H₂, O₂, and CO were detected over bare Al-SrTiO₃, and the H₂-formation rate was much higher than the CO-formation rate with non-stoichiometric amounts of O₂ produced. The main reductive product was found to be gaseous CO (97.7%) over Ag/SrTiO₃, and a stoichiometric amount of O₂ was evolved, which indicates

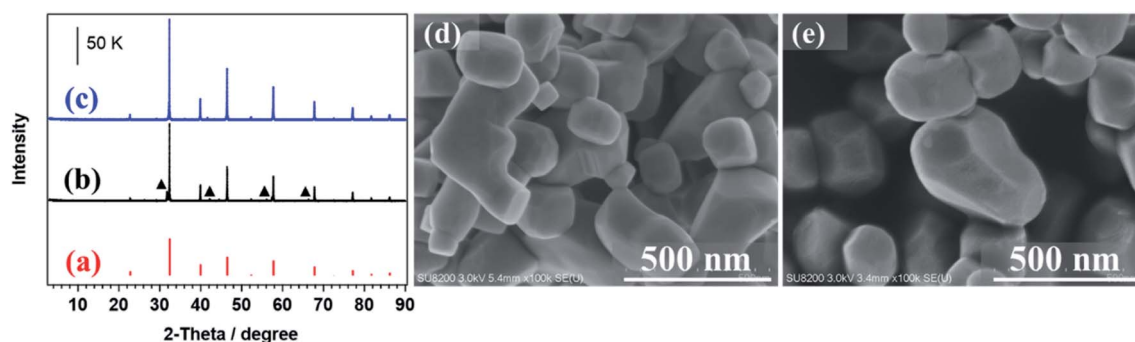


Fig. 1 XRD patterns of (a) the SrTiO₃ reference (ICSD no. 23076) and the synthesized (b) SrTiO₃ (▲ Sr₃Ti₂O₇) and (c) Al-SrTiO₃. SEM images of the synthesized (d) SrTiO₃ and (e) Al-SrTiO₃.



that H₂O functions as an electron donor in the photocatalytic reaction. The CO-formation rate was significantly enhanced (4.7 $\mu\text{mol h}^{-1}$) compared to the undoped catalyst. As mentioned in previous reports,^{12,13,47} Ag functions as a potential cocatalyst that enhances the photocatalytic activity for the reduction of CO₂ and is selective toward CO evolution. Furthermore, the Al-SrTiO₃ photocatalyst was also loaded with Co species, as Co has been described to be a good cocatalyst for the photocatalytic oxidation of water;^{38,39} while high H₂- and O₂-formation rates (41.6 and 19.5 $\mu\text{mol h}^{-1}$, respectively) were obtained over Co/Al-SrTiO₃, and that for CO formation was only 0.1 $\mu\text{mol h}^{-1}$, which indicates that overall only water splitting took place over the Co/Al-SrTiO₃ photocatalyst. Surprisingly, the highest CO-formation rate (52.7 $\mu\text{mol h}^{-1}$), which is 10-times higher than that observed for Ag/Al-SrTiO₃, was delivered by the Al-SrTiO₃ photocatalyst dual-modified with both Ag and Co by the CR method (AgCo/Al-SrTiO₃_CR). Meanwhile, the selectivity for CO evolution was further improved to 99.8%, with only a trace of H₂ evolved over AgCo/Al-SrTiO₃_CR. To the best of our knowledge, this is the highest selectivity for CO evolution reported for photocatalytic CO₂ conversion using H₂O as a reductant under aqueous conditions at wavelength (λ) above 300 nm.^{12,13,21,23,48–51} Consequently, we conclude that Co species in the photocatalyst significantly enhance photocatalytic performance for CO₂ conversion because they effectively promote the water-oxidation half reaction.

The H₂-, O₂-, and CO-formation rates during the photocatalytic conversion of CO₂ by H₂O over AgCo/Al-SrTiO₃ prepared by the IMP, CR, and PD methods are also listed in Table 1. The amounts of Ag and Co loaded into AgCo/Al-SrTiO₃ prepared by these methods were determined by XRF spectroscopy (Table S1†), which suggested that the Ag and Co cocatalysts had been successfully loaded onto the surface of the Al-SrTiO₃ prepared using each method.

The selectivity toward CO evolution was very low (59.2%) over AgCo/Al-SrTiO₃_PD. In addition, AgCo/Al-SrTiO₃_IMP exhibited a very low CO-formation rate of 4.1 $\mu\text{mol h}^{-1}$. Undoubtedly, the highest CO-formation rate was delivered by AgCo/Al-SrTiO₃ prepared by the CR method, with exceptionally good selectivity toward CO evolution also observed. Compared with Ag/Al-SrTiO₃ and Co/Al-SrTiO₃, AgCo/Al-SrTiO₃ prepared by the PD and CR methods exhibited higher activities for photocatalytic CO₂ reduction and H₂O oxidation, which indicates that, apart from Co playing an important role in enhancing the water oxidation half reaction, the Ag and Co species function synergistically. Furthermore, the amounts of the Ag and Co cocatalysts loaded by the CR method were also determined by XRF spectroscopy, the results of which are summarized in Table S1.† The XRF-determined amount of Ag on the surface of the Al-SrTiO₃ photocatalyst was found to be similar to the amount of Ag precursor used in the reaction (Table S1†), whereas the amount of Co was determined to be lower than the amount of precursor used, particularly at Co precursor levels higher than 0.85%. The effect of the loaded-amount of the dual Ag and Co cocatalyst on the photocatalytic activity for CO₂ reduction are shown in Fig. S1–S3,† which reveal that photocatalytic performance depends on the Ag and Co loading. These results show

that Ag(1.7)Co(0.85)/Al-SrTiO₃ and Ag(1.7)Co(1.275)/Al-SrTiO₃ prepared by the CR method exhibit the highest CO-formation rates of 52.7 $\mu\text{mol h}^{-1}$ and 54.6 $\mu\text{mol h}^{-1}$, respectively, with good selectivities toward CO evolution. Ag(1.7)Co(0.85)/Al-SrTiO₃ is referred to as AgCo/Al-SrTiO₃_CR hereafter.

To determine apparent quantum efficiencies, photocatalytic CO₂ reductions over AgCo/Al-SrTiO₃_CR were also carried out in aqueous NaHCO₃ (0.1 M) using various light sources. Fig. 2(a) shows time courses for the evolution of CO, H₂, and O₂ during the photocatalytic conversion of CO₂ by H₂O over AgCo/Al-SrTiO₃_CR when irradiated at λ above 300 nm for 5 h under typical conditions. Importantly, the same reaction (photocatalytic reduction of CO₂ by H₂O) was also carried out under monochromatic light (λ = 365 nm) in NaHCO₃ (0.2 L, 0.1 M), the results of which are shown in Fig. 2(b); the CO-formation rate was approximately 7.4 $\mu\text{mol h}^{-1}$ with good selectivity toward CO evolution (98.6%). The apparent quantum efficiency (AQE) of 0.03% was calculated using eqn (1):

$$\text{AQE (\%)} = (\text{number of reacted electrons/number of incident photons}) \times 100 \quad (1)$$

Photocatalytic reaction conditions: amount of photocatalyst, (a) 0.5 g and (b) 0.1 g; amount of Ag loaded, 1.7 mol%; amount of Co loaded, 0.85 mol%; volume of the reaction solution (H₂O), (a) 1.0 L and (b) 0.2 L; additive, 0.1 M NaHCO₃; CO₂ flow rate; 30 mL min⁻¹.

Although this AQE value is much lower than desired, it is important to note that an AQE was obtained in this photocatalytic CO₂-reduction system without any sacrificial reagent under relatively high wavelength of photoirradiation. The amounts of Ag and Co in the AgCo/Al-SrTiO₃_CR photocatalyst were also determined by XRF spectroscopy after CO₂ had been photoreduced for 5 h, the results of which are summarized in Table S1,† which reveals that Ag and Co loadings are only slightly lower compared to those of the as-prepared sample. These results suggest that Ag and Co are durable under the reaction conditions, as their photocatalyst loadings were stable over the 5 h duration of the photocatalytic reaction.

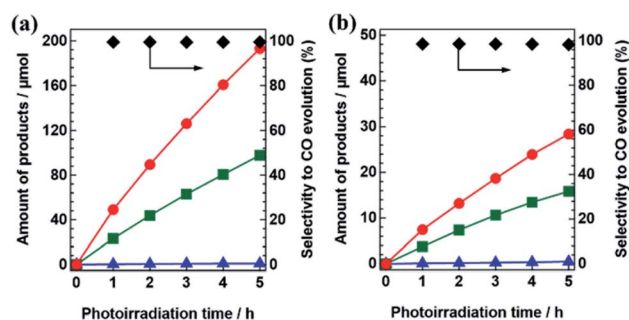


Fig. 2 Time courses for the evolution of H₂ (blue), O₂ (green), CO (red), and selectivity toward CO evolution (black) for the photocatalytic conversion of CO₂ by H₂O over AgCo/Al-SrTiO₃_CR using various light sources: (a) 400 W high pressure Hg lamp with a Pyrex® jacket at $\lambda > 300$ nm, (b) LED lamp at $\lambda = 365$ nm.

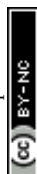


Fig. 3(a) shows Ag K-edge XANES spectra of Ag/Al-SrTiO₃ and AgCo/Al-SrTiO₃_CR, and those of Ag foil, Ag₂O, and Ag₃PO₄ as references. The absorption edges of the Ag species in Ag/Al-SrTiO₃ and AgCo/Al-SrTiO₃ prepared by the CR method are consistent with that of Ag foil, at 25 528 eV;^{21,23} in addition, the Ag K-edge XANES spectra of Ag/Al-SrTiO₃ and AgCo/Al-SrTiO₃_CR exhibit features that are similar to that of Ag foil. These results indicate that the Ag species in these photocatalysts are metallic, and the presence of Co does not influence the chemical state of the Ag species. Fig. 3(b) shows the Co K-edge XANES spectra of Co/Al-SrTiO₃ and AgCo/Al-SrTiO₃_CR, with those of Co₃O₄, CoOOH,^{52,53} Co₃(PO₄)₂, and CoO included for reference, which reveals that the Co K-edge XANES spectra of Co/Al-SrTiO₃ and AgCo/Al-SrTiO₃_CR do not match those of Co₃(PO₄)₂ and CoO; the AgCo/Al-SrTiO₃_CR absorption edge was closer to those of CoOOH and Co₃O₄. Furthermore, the white-line and peak maximum (7727.6 eV) in the Co K-edge XANES spectrum of AgCo/Al-SrTiO₃_CR are consistent with those of the CoOOH reference, even though the spectral features are slightly different.^{53,54} The differences between the Co K-edge XANES spectra of AgCo/Al-SrTiO₃_CR and CoOOH are possibly ascribable to highly dispersed Co species on the Al-SrTiO₃ surface, as XANES spectra are reportedly affected by nanoparticle size.^{55,56} In addition, the Co K-edge EXAFS oscillation of AgCo/Al-SrTiO₃_CR is approximately consistent with that of CoOOH rather than Co₃O₄ (Fig. S4†), which indicates that CoOOH exists on the AgCo/Al-SrTiO₃_CR surface. The XANES spectra of Co/Al-SrTiO₃ and AgCo/Al-SrTiO₃_CR also overlap, as shown in Fig. 3(c). The Co absorption edge of Co/Al-SrTiO₃ is slightly shifted to lower energy compared to that of AgCo/Al-SrTiO₃_CR, while the Co K-edge XANES features are also different. Consequently, the presence of Ag affects the chemical structure of the Co species when Ag and Co species are simultaneously loaded using the CR method.

Fig. 4 shows SEM images of the Ag/Al-SrTiO₃, Co/Al-SrTiO₃, and AgCo/Al-SrTiO₃_CR photocatalysts. Ag nanoparticles approximately 2–25 nm in size are uniformly dispersed on the

Al-SrTiO₃ surface, as shown in Fig. 4(a), while no obvious nanoparticles are observed in the SEM image of Co/Al-SrTiO₃ (Fig. 4(b)), although the XRF data indicate that Co had been successfully loaded on the Al-SrTiO₃ surface (Table S1†). In addition, the Co/Al-SrTiO₃ EDS map and Co 2p XPS spectrum (Fig. S5†) also show that Co species are present on the Al-SrTiO₃ surface. We believe that the Co species on the Al-SrTiO₃ are too small to be observed by SEM. AgCo/Al-SrTiO₃_CR exhibited a similar morphology to that of Ag/Al-SrTiO₃; the Al-SrTiO₃ particles are well covered by nanoparticles around 2–20 nm in size, with nanoparticles larger than 20 nm rarely observed on the photocatalyst surface. Fig. 4(d–f) show SEM images of Ag/Al-SrTiO₃, Co/Al-SrTiO₃, and AgCo/Al-SrTiO₃_CR after 5 h of photoirradiation, which reveal that Ag metal nanoparticles are only present on smooth Ag/Al-SrTiO₃ {100} facets, with no Ag nanoparticles present on other facets, whereas Co species are observed on nanostep Co/Al-SrTiO₃ {110} facets; therefore, the Ag metal and Co nanoparticles apparently moved to the smooth {100} and nanostep {110} facets during photoirradiation, respectively. Note that the positions of the Ag and Co species, which were highly dispersed on the surface in a random fashion, are different after photoirradiation. In addition, the Ag and Co nanoparticle cocatalysts are both observed on every facet of each Al-SrTiO₃ particle in Fig. 4(f) in the case of AgCo/Al-SrTiO₃_CR, which indicates that the smooth {100} and nanostep {110} facets prefer to be decorated with Ag metal nanoparticles and Co nanoparticles, respectively, when irradiated. Moreover, the SEM images of the AgCo/Al-SrTiO₃ photocatalysts prepared by the IMP and PD methods are shown in Fig. S6,† which reveals that the surface of AgCo/Al-SrTiO₃_IMP contains several large cocatalyst particles (Fig. S6(a)†). Numerous Ag nanoparticles were observed to have aggregated on {100} facets after 5 h of photocatalytic reaction, in addition to the large cocatalyst particles on the surfaces of the photocatalyst particles. The unusual size of the cocatalyst on AgCo/Al-SrTiO₃_IMP negatively affects photocatalytic performance. Fig. S6(b)† shows that the Ag and Co dual cocatalyst is spatially located on the {100} facets

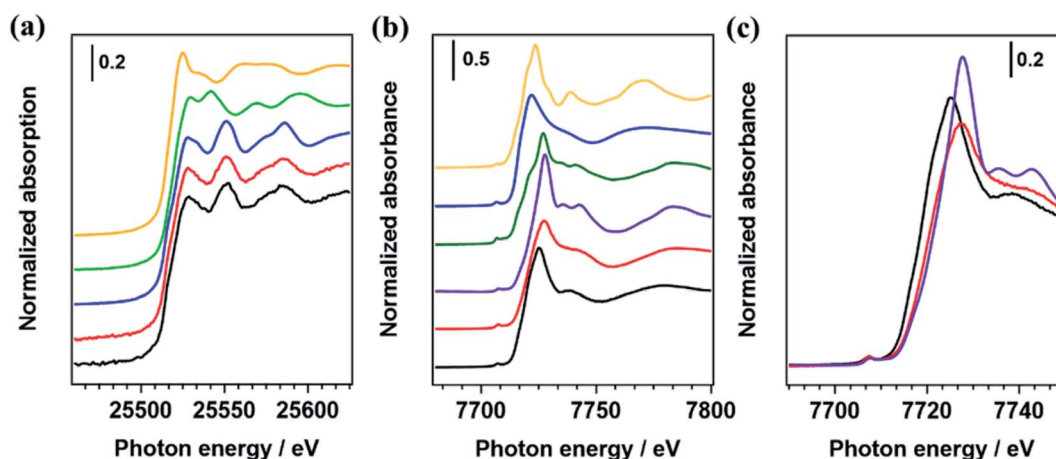


Fig. 3 (a) Ag K-edge XANES spectra of Ag/Al-SrTiO₃ (black), AgCo/Al-SrTiO₃_CR (red), Ag foil (blue), Ag₂O (green), and Ag₃PO₄ (yellow). (b) and (c) Co K-edge XANES spectra of AgCo/Al-SrTiO₃_CR (red), Co/Al-SrTiO₃ (black), CoOOH (purple), Co₃O₄ (green), Co₃(PO₄)₂ (blue), and (f) CoO (yellow).



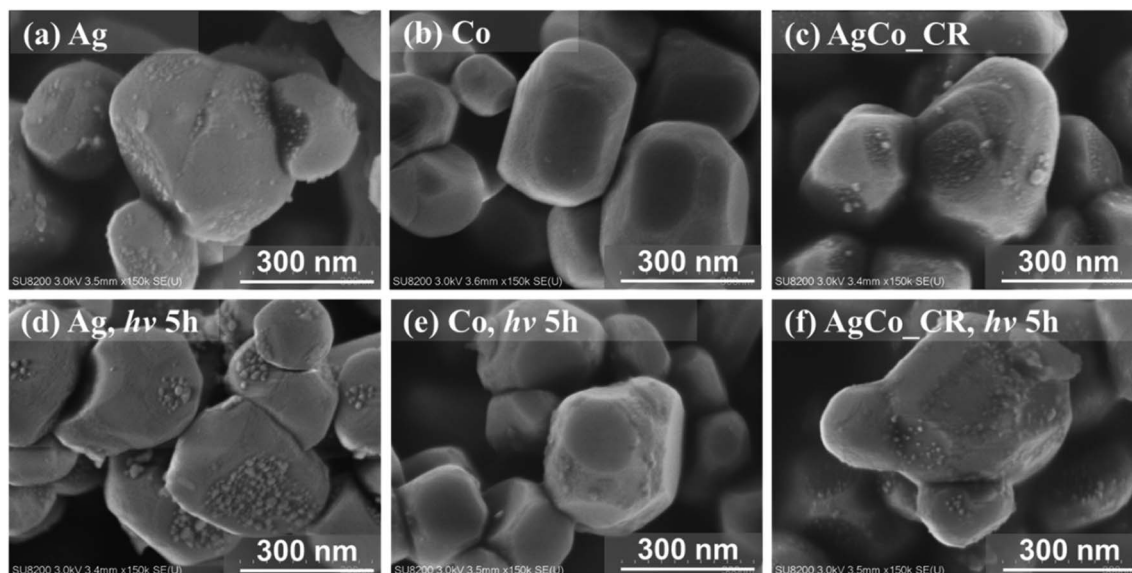


Fig. 4 SEM images of various photocatalysts: (a, d) Ag/Al-SrTiO₃ prepared using CR method, (b, e) Co/Al-SrTiO₃ prepared using CR method, and (c, f) AgCo/Al-SrTiO₃_CR; (a–c) show fresh samples and (d–f) show samples after 5 h of photoirradiation.

and {110} facets when the PD method was used. In addition, the Ag nanoparticles on the {100} facets are much larger than those obtained using the CR method. Thus, the main reason for the poor selectivity of AgCo/Al-SrTiO₃_PD is likely to be the oversized Ag nanoparticles on its surface, as it is known that the Ag cocatalyst plays an important role during the photocatalytic reduction of CO₂ to CO. BiVO₄,^{11,34} 18-facet-SrTiO₃,³⁵ Al-SrTiO₃,³⁷ and KTaO₃,⁴⁰ which have exposed anisotropic facets, were also reported to exhibit similar properties. Li *et al.* reported that metals (Ag, Pt, and Au) and metal oxides (CoO_x, MnO₂, and

PbO₂) are selectively loaded on various BiVO₄^{11,34} and 18-facet-SrTiO₃³⁵ facets when irradiated under aqueous conditions. In addition, the Rh/Cr hydrogen-evolution reaction (HER) cocatalyst and the CoOOH OER cocatalyst were also reported to be spatially photodeposited on the {100} and {110} facets of Al-SrTiO₃, respectively.³⁷ Selective photodeposition was proposed to be due to the charge-rectification effect inside individual semiconductor photocatalyst particles.^{37,57} Moreover, an anisotropically deposited cocatalyst was reported to play a positive charge-separation and transfer role inside individual

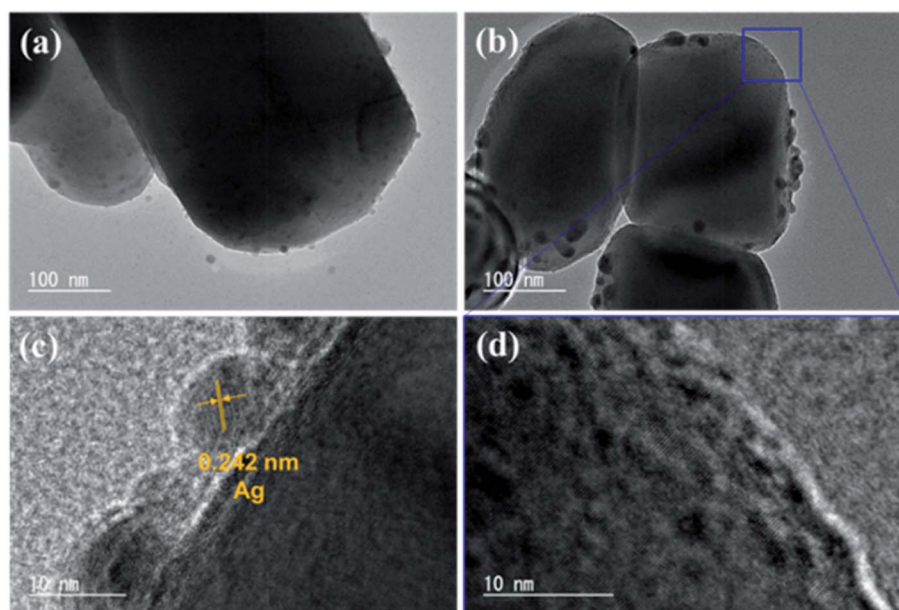


Fig. 5 TEM images of AgCo/Al-SrTiO₃_CR photocatalysts: (a) before reaction and (b–d) after 5 h of photocatalysis. (c) Ag nanoparticles on a smooth Al-SrTiO₃ facet and (d) enlarged region showing AgCo/Al-SrTiO₃_CR nanostep facet.



photocatalyst particles by aligning the electric fields built in the cocatalyst-loaded photocatalyst.¹¹ Therefore, Ag and Co species possibly appear on different Al-SrTiO₃ facets after photo-irradiation because the {100} and {110} facets of Al-SrTiO₃ exhibit different band structures and band-edge positions.^{29,33,37}

The TEM images of AgCo/Al-SrTiO₃-CR reveal that the Ag and Co cocatalysts were deposited on the Al-SrTiO₃ surface, as shown in Fig. 5, with nanoparticles around 2–20 nm in size; these nanoparticles exhibit a lattice distance of 0.242 nm, which is attributable to the Ag {111} facet.^{58,59} Ag nanoparticles are aggregated on the {100} facets of single photocatalyst particles and much larger particles were observed after 5 h of photo-irradiation, with the largest around 20 nm in size. In addition, Fig. 5(b) and (d) reveal the appearance of numerous small nanoparticles on the {110} facets of the photocatalyst. As mentioned earlier, the XANES spectrum (Fig. 3) suggests that highly dispersed CoOOH was generated on the Al-SrTiO₃ surface; however, there are no clear CoOOH fringe patterns visible in Fig. 5(d), although the fringe pattern of the {110} facet of metallic Ag is observed, which indicates that the main Co species in AgCo/Al-SrTiO₃-CR is amorphous CoOOH.

The UV-Vis DR spectra of the Al-SrTiO₃, Ag/Al-SrTiO₃, Co/Al-SrTiO₃, and AgCo/Al-SrTiO₃-CR cocatalysts prepared by the CR method are shown in Fig. 6. Al-SrTiO₃ exhibits an absorption edge at approximately 390 nm, which is consistent with that reported previously.^{36,43} However, broad absorption bands, which are assignable to surface plasmon resonance (SPR), are evident in the spectra of Ag/Al-SrTiO₃ and AgCo/Al-SrTiO₃-CR, but not Co/Al-SrTiO₃; these characteristic bands correspond to Ag nanoparticles.⁶⁰ In addition, spectra f and g in Fig. 6 reveal that the SPR bands are slightly redshifted after use, which indicates that much larger nanoparticles are generated during photocatalysis. We conclude that cocatalyst nanoparticles migrate and aggregate on the Al-SrTiO₃ surface during photo-irradiation, as shown in the SEM and TEM images (Fig. 4 and 5).⁶¹

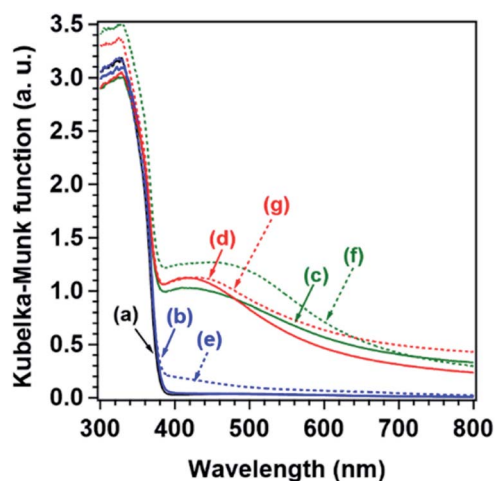
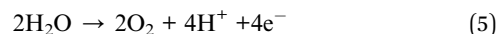
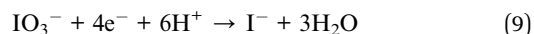


Fig. 6 UV-Vis DR spectra: (a) Al-SrTiO₃, (b, e) Co/Al-SrTiO₃, (c, f) Ag/Al-SrTiO₃, and (d, g) AgCo/Al-SrTiO₃-CR. (b–d) As-prepared samples and (e–g) samples after 5 h of photocatalysis.

To further study the anisotropic properties of Al-SrTiO₃, we investigated depositing various metals, such as Ag, Pt, Au, and metal oxides, such as MnO₂, PbO₂, and Co₃O₄, on the Al-SrTiO₃ surface by the PD method. Interestingly, metal nanoparticles (*i.e.*, Ag, Pt, and Au), which were observed by EDS (Fig. S7†), were spatially deposited on the {100} facets of Al-SrTiO₃, as shown in Fig. 7(a)–(c), which suggests that the metal cations in aqueous solutions of AgNO₃, H₂PtCl₆, and HAuCl₄ are reduced to Ag, Pt, and Au nanoparticles on the {100} facets by H₂O as the electron donor, as shown by the following equations:



where eqn (2)–(4) show the reductions of metal ions and eqn (5) shows the oxidation of water. Meanwhile, Fig. 7(d)–(f) reveal that metal oxide nanoparticles (*i.e.*, MnO₂, PbO₂, and Co₃O₄) are spatially located on the {110} facets when NaIO₃ was used as the hole donor, according to eqn (6)–(9):^{35,62}



where eqn (5)–(7) show metal-ion oxidation and eqn (8) corresponds to the reduction of iodate.

The abovementioned results suggest that photogenerated electrons and holes selectively transfer to different facets of the photocatalyst because the electric fields within single photocatalyst particles are aligned;^{11,37} therefore, reductive and oxidative sites exist on different Al-SrTiO₃ facets. Table S2† summarizes the photocatalytic activities of various single-cocatalyst-loaded Al-SrTiO₃ samples prepared by the PD method, which reveals that only the Ag metal nanoparticle cocatalyst exhibited good selectivity toward CO evolution (92.2%) by CO₂ reduction, even though the CO-formation rate was only 3.2 μmol h^{−1}. The Pt-, Au-, Co₃O₄-, MnO₂-, and PbO₂-loaded Al/SrTiO₃ only split water, with poor CO-evolution selectivities.

A possible mechanism for the photoreduction of CO₂ into CO over AgCo/Al-SrTiO₃-CR using H₂O as the reductant is depicted in Scheme 1. As mentioned above, different cocatalyst are spatially deposited on anisotropic facets, with Ag⁺, Au³⁺, and Pt⁴⁺ reduced and loaded onto smooth {100} facets, and Mn²⁺, Co²⁺, and Pb²⁺ oxidized and loaded on {110} facets. We propose that photogenerated electrons and holes are transferred to different Al-SrTiO₃ facets; therefore, reduction and oxidation occur at different facets.^{35,37} Even though the Ag and Co species were dispersed well on the surface of AgCo/Al-SrTiO₃-CR, these



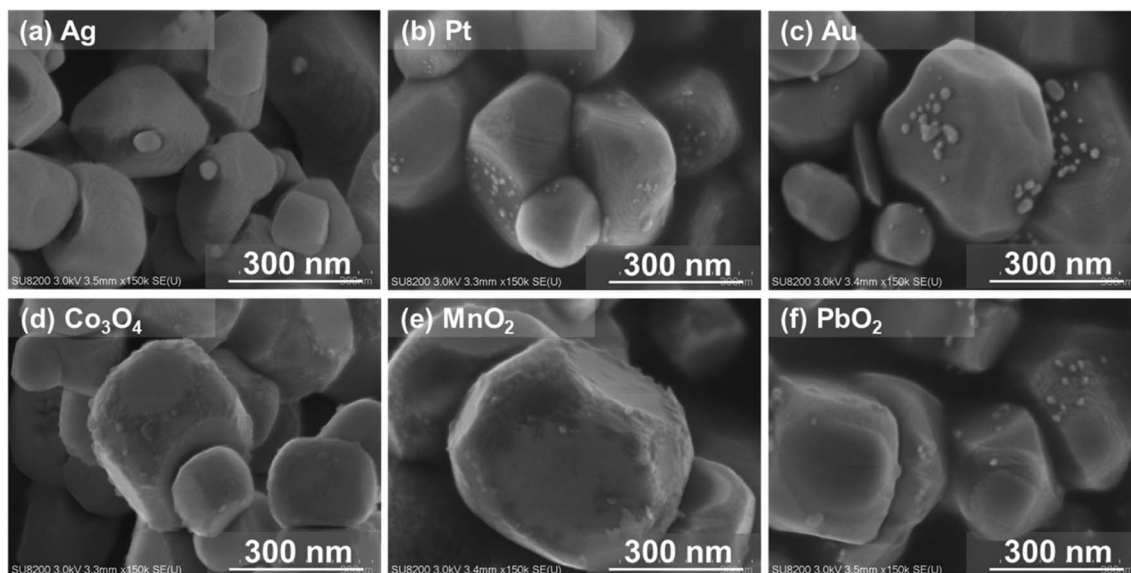
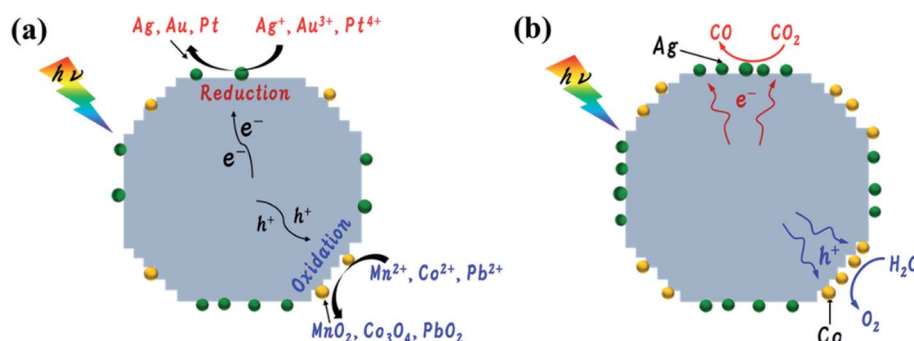


Fig. 7 SEM images of various of cocatalyst-loaded Al-SrTiO₃ samples prepared by the PD method: (a) Ag, (b) Pt, (c) Au, (d) Co₃O₄, (e) MnO₂, and (f) PbO₂.



Scheme 1 Plausible photoirradiation mechanisms: (a) cocatalyst loading by the PD method and (b) the photocatalytic conversion of CO₂ into CO over Al-SrTiO₃ modified with Ag and Co by the CR method using water as the reductant.

Ag and Co were redeposited onto different crystal facets of the Al-SrTiO₃ during the photocatalytic reaction under photoirradiation as shown in Fig. 4 and 5. Moreover, based on these results, CO₂ is photocatalytically converted into CO at Ag cocatalyst particles on the {100} facets of Al-SrTiO₃, whereas Co cocatalyst particles oxidize water on the {110} facets. Photocatalytic activity is significantly enhanced after dual cocatalyst loading because photogenerated electrons and holes move to different facets and are quickly captured by the Ag and Co cocatalysts, respectively.

Conclusions

Loading Ag and Co onto Al-SrTiO₃ significantly improved its activity for the photocatalytic conversion of CO₂ by H₂O (as the electron donor), with extremely high selectivity toward CO evolution (99.8%), in which Ag and Co enable CO₂ reduction and H₂O oxidation on the Al-SrTiO₃ surface, respectively. A CO-evolution rate of up to 52.7 μmol h⁻¹ was observed over AgCo/Al-SrTiO₃_CR when irradiated with light at wavelengths above

300 nm, which is ten-times higher than that observed for Ag/Al-SrTiO₃ (4.7 μmol h⁻¹). Furthermore, 7.4 μmol h⁻¹ of CO gas was evolved when irradiated with LED light (365 nm), with an apparent quantum efficiency of 0.03%. In addition, this study demonstrated that the reductive and oxidative sites are distributed on the {100} and {110} facets of Al-SrTiO₃, respectively. Therefore, photocatalytic CO₂ reduction and water oxidation occur separately on different Al-SrTiO₃ facets. Synergism between the Ag and Co dual cocatalysts effectively enhances the photocatalytic conversion of CO₂ over Al-SrTiO₃ with H₂O as the electron donor.

Conflicts of interest

The authors declare no competing financial interest.

Acknowledgements

This research was partially supported by the Program for Elements Strategy Initiative for Catalysts and Batteries,



commissioned by the Ministry of Education, Culture, Sports, Science, and Technology (MEXT) of Japan. This work was also supported by JSPS KAKENHI grant number 18H01788. Shuying Wang thanks the State Scholarship of the China Scholarship Council, which is affiliated with the Ministry of Education of the People's Republic of China. Takashi Hisatomi thanks the Japan Society for the Promotion of Science (JSPS) for a Grant-in-Aid for Scientific Research on Innovative Areas (grant no. 18H05156).

Notes and references

- 1 J. Yang, D. Wang, H. Han and C. Li, *Acc. Chem. Res.*, 2013, **46**, 1900–1909.
- 2 J. Ran, M. Jaroniec and S. Z. Qiao, *Adv. Mater.*, 2018, **30**, 1704649.
- 3 X. Chang, T. Wang and J. Gong, *Energy Environ. Sci.*, 2016, **9**, 2177–2196.
- 4 A. Kudo and Y. Miseki, *Chem. Soc. Rev.*, 2009, **38**, 253–278.
- 5 H. Abdullah, M. M. R. Khan, H. R. Ong and Z. Yaakob, *J. CO₂ Util.*, 2017, **22**, 15–32.
- 6 R. Daghrir, P. Drogui and D. Robert, *Ind. Eng. Chem. Res.*, 2013, **52**, 3581–3599.
- 7 S. Trasatti, *J. Electroanal. Chem.*, 1972, **39**, 163–184.
- 8 K. Maeda, K. Teramura, N. Saito, Y. Inoue and K. Domen, *J. Catal.*, 2006, **243**, 303–308.
- 9 H. Kato and A. Kudo, *J. Phys. Chem. B*, 2001, **105**, 4285–4292.
- 10 X. Zong, H. Yan, G. Wu, G. Ma, F. Wen, L. Wang and C. Li, *J. Am. Chem. Soc.*, 2008, **130**, 7176–7177.
- 11 J. Zhu, S. Pang, T. Dittrich, Y. Gao, W. Nie, J. Cui, R. Chen, H. An, F. Fan and C. Li, *Nano Lett.*, 2017, **17**, 6735–6741.
- 12 K. Iizuka, T. Wato, Y. Miseki, K. Saito and A. Kudo, *J. Am. Chem. Soc.*, 2011, **133**, 20863–20868.
- 13 H. Nakanishi, K. Iizuka, T. Takayama, A. Iwase and A. Kudo, *ChemSusChem*, 2017, **10**, 112–118.
- 14 T. Takayama, K. Tanabe, K. Saito, A. Iwase and A. Kudo, *Phys. Chem. Chem. Phys.*, 2014, **16**, 24417–24422.
- 15 O. K. Varghese, M. Paulose, T. J. LaTempa and C. A. Grimes, *Nano Lett.*, 2009, **9**, 731–737.
- 16 S. Xie, Y. Wang, Q. Zhang, W. Deng and Y. Wang, *ACS Catal.*, 2014, **4**, 3644–3653.
- 17 T. Yui, A. Kan, C. Saitoh, K. Koike, T. Ibusuki and O. Ishitani, *ACS Appl. Mater. Interfaces*, 2011, **3**, 2594–2600.
- 18 K. Sayama and H. Arakawa, *J. Phys. Chem.*, 1993, **97**, 531–533.
- 19 P.-W. Pan and Y.-W. Chen, *Catal. Commun.*, 2007, **8**, 1546–1549.
- 20 I.-H. Tseng, W.-C. Chang and J. C. Wu, *Appl. Catal., B*, 2002, **37**, 37–48.
- 21 R. Pang, K. Teramura, H. Tatsumi, H. Asakura, S. Hosokawa and T. Tanaka, *Chem. Commun.*, 2018, **54**, 1053–1056.
- 22 T. F. Xie, D. J. Wang, L. J. Zhu, T. J. Li and Y. J. Xu, *Mater. Chem. Phys.*, 2001, **70**, 103–106.
- 23 X. Zhu, A. Yamamoto, S. Imai, A. Tanaka, H. Kominami and H. Yoshida, *Chem. Commun.*, 2019, 13514–13517.
- 24 S. T. Neațu, J. A. Maciá-Agulló, P. Concepción and H. Garcia, *J. Am. Chem. Soc.*, 2014, **136**, 15969–15976.
- 25 Q. Kang, T. Wang, P. Li, L. Liu, K. Chang, M. Li and J. Ye, *Angew. Chem., Int. Ed.*, 2015, **54**, 841–845.
- 26 C.-W. Tsai, H. M. Chen, R.-S. Liu, K. Asakura and T.-S. Chan, *J. Phys. Chem. C*, 2011, **115**, 10180–10186.
- 27 R. Pang, K. Teramura, H. Asakura, S. Hosokawa and T. Tanaka, *ACS Sustainable Chem. Eng.*, 2018, **7**, 2083–2090.
- 28 Q. Zhai, S. Xie, W. Fan, Q. Zhang, Y. Wang, W. Deng and Y. Wang, *Angew. Chem., Int. Ed.*, 2013, **52**, 5776–5779.
- 29 J. Yu, J. Low, W. Xiao, P. Zhou and M. Jaroniec, *J. Am. Chem. Soc.*, 2014, **136**, 8839–8842.
- 30 H. Kato, K. Asakura and A. Kudo, *J. Am. Chem. Soc.*, 2003, **125**, 3082–3089.
- 31 T. Ohno, K. Sarukawa and M. Matsumura, *New J. Chem.*, 2002, **26**, 1167–1170.
- 32 N. Murakami, Y. Kurihara, T. Tsubota and T. Ohno, *J. Phys. Chem. C*, 2009, **113**, 3062–3069.
- 33 T. Tachikawa, T. Ochi and Y. Kobori, *ACS Catal.*, 2016, **6**, 2250–2256.
- 34 R. Li, F. Zhang, D. Wang, J. Yang, M. Li, J. Zhu, X. Zhou, H. Han and C. Li, *Nat. Commun.*, 2013, **4**, 1432.
- 35 L. C. Mu, Y. Zhao, A. L. Li, S. Y. Wang, Z. L. Wang, J. X. Yang, Y. Wang, T. F. Liu, R. T. Chen, J. Zhu, F. T. Fan, R. G. Li and C. Li, *Energy Environ. Sci.*, 2016, **9**, 2463–2469.
- 36 Y. Ham, T. Hisatomi, Y. Goto, Y. Moriya, Y. Sakata, A. Yamakata, J. Kubota and K. Domen, *J. Mater. Chem. A*, 2016, **4**, 3027–3033.
- 37 T. Takata, J. Jiang, Y. Sakata, M. Nakabayashi, N. Shibata, V. Nandal, K. Seki, T. Hisatomi and K. Domen, *Nature*, 2020, **581**, 411–414.
- 38 H. Zhang, C. Guo, J. Ren, J. Ning, Y. Zhong, Z. Zhang and Y. Hu, *Chem. Commun.*, 2019, **55**, 14050–14053.
- 39 D. Kong, J. Xie, Z. Guo, D. Yang and J. Tang, *ChemCatChem*, 2020, 2708–2712.
- 40 X. Zhu, A. Yamamoto, S. Imai, A. Tanaka, H. Kominami and H. Yoshida, *Appl. Catal., B*, 2020, 119085.
- 41 Y. Bai, L. Ye, L. Wang, X. Shi, P. Wang and W. Bai, *Environ. Sci.: Nano*, 2016, **3**, 902–909.
- 42 Q. Liu, Y. Zhou, J. Kou, X. Chen, Z. Tian, J. Gao, S. Yan and Z. Zou, *J. Am. Chem. Soc.*, 2010, **132**, 14385–14387.
- 43 S. Wang, K. Teramura, T. Hisatomi, K. Domen, H. Asakura, S. Hosokawa and T. Tanaka, *ACS Appl. Energy Mater.*, 2020, **3**, 1468–1475.
- 44 S. Wang, K. Teramura, T. Hisatomi, K. Domen, H. Asakura, S. Hosokawa and T. Tanaka, *ChemistrySelect*, 2020, **5**, 8779–8786.
- 45 Z. Wang, K. Teramura, S. Hosokawa and T. Tanaka, *J. Mater. Chem. A*, 2015, **3**, 11313–11319.
- 46 Z. Wang, K. Teramura, S. Hosokawa and T. Tanaka, *Appl. Catal., B*, 2015, **163**, 241–247.
- 47 K. Teramura and T. Tanaka, *Phys. Chem. Chem. Phys.*, 2018, **20**, 8423–8431.
- 48 S. Kikkawa, K. Teramura, H. Asakura, S. Hosokawa and T. Tanaka, *J. Phys. Chem. C*, 2018, **122**, 21132–21139.
- 49 Z. Huang, K. Teramura, H. Asakura, S. Hosokawa and T. Tanaka, *Catal. Today*, 2018, **300**, 173–182.
- 50 R. Pang, K. Teramura, H. Asakura, S. Hosokawa and T. Tanaka, *Appl. Catal., B*, 2017, **218**, 770–778.
- 51 A. Anzai, N. Fukuo, A. Yamamoto and H. Yoshida, *Catal. Commun.*, 2017, **100**, 134–138.



- 52 G. Amatuucci, J. Tarascon, D. Larcher and L. Klein, *Solid State Ionics*, 1996, **84**, 169–180.
- 53 M. W. Kanan, J. Yano, Y. Surendranath, M. Dinca, V. K. Yachandra and D. G. Nocera, *J. Am. Chem. Soc.*, 2010, **132**, 13692–13701.
- 54 M. Yoshida, T. Mineo, Y. Mitsutomi, F. Yamamoto, H. Kurosu, S. Takakusagi, K. Asakura and H. Kondoh, *Chem. Lett.*, 2016, **45**, 277–279.
- 55 D. Bazin, D. Sayers, J. Rehr and C. Mottet, *J. Phys. Chem. B*, 1997, **101**, 5332–5336.
- 56 K. Beppu, S. Hosokawa, T. Shibano, A. Demizu, K. Kato, K. Wada, H. Asakura, K. Teramura and T. Tanaka, *Phys. Chem. Chem. Phys.*, 2017, **19**, 14107–14113.
- 57 R. Chen, S. Pang, H. An, J. Zhu, S. Ye, Y. Gao, F. Fan and C. Li, *Nat. Energy*, 2018, **3**, 655–663.
- 58 R. Nazir, P. Fageria, M. Basu and S. Pande, *J. Phys. Chem. C*, 2017, **121**, 19548–19558.
- 59 S. Navaladian, B. Viswanathan, T. Varadarajan and R. Viswanath, *Nanotechnology*, 2008, **19**, 045603.
- 60 H. Sakai, T. Kanda, H. Shibata, T. Ohkubo and M. Abe, *J. Am. Chem. Soc.*, 2006, **128**, 4944–4945.
- 61 K.-C. Lee, S.-J. Lin, C.-H. Lin, C.-S. Tsai and Y.-J. Lu, *Surf. Coat. Technol.*, 2008, **202**, 5339–5342.
- 62 J. Menze, B. Mei, P. Weide and M. Muhler, *J. Mater. Chem. A*, 2017, **5**, 17248–17252.

