

Note

Characterization of Iron(II) Complexes with
 β -Diketones and Monothio- β -diketone

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β -Diketones and monothio- β -diketones are well-known and frequently used ligands. Change of a donor atom from oxygen to sulfur often gives rise to different properties in complexes formed. For example, the tris complexes of iron(III) with β -diketones are high-spin, but the tris complexes with some monothio- β -diketones exhibit an anomalous magnetic behavior ascribed to a spin-equilibrium between 5A_1 - 2T_2 states.^{1,2)} The ligands are capable of forming stable adducts with iron(II) in the presence of Lewis bases. Emmert *et al.* synthesized bis(acetylacetonato)dipyridineiron(II), Fe(II)(acac)₂Py₂, and bis(oxo-1, 3-diphenylprop-2-en-1-onato)dipyridineiron(II), Fe(II)(dibz)₂Py₂.^{3,4)} Uhlemann and Thomas isolated bis(thioxo-1, 3-diphenylprop-2-en-1-onato)dipyridineiron(II), Fe(II)(sdibz)₂Py₂.⁵⁾ These complexes are soluble in organic solvents. Therefore, it is possible to extract iron(II) into organic solvents containing these ligands synergistically and determine it spectrophotometrically.⁶⁻¹⁰⁾ The author characterized Fe(II)(acac)₂Py₂, Fe(II)(dibz)₂Py₂ and Fe(II)(sdibz)₂Py₂ through spectroscopic and thermogravimetric measurements and will report the results in this paper.

EXPERIMENTAL

Monothiodibenzoylmethane synthesized according to the literature was offered by Fukui.¹¹⁾ Acetylacetone, dibenzoylmethane, pyridine and Mohr's salt were obtained commercially and pyridine was purified by distillation. Fe(II)(acac)₂Py₂ was prepared by the addition of acetylacetone (1.0g) and pyridine (7ml) to an aqueous solution (15ml) of Mohr's salt (2.0g). The solution upon recrystallization from pyridine gave black crystals. Fe(II)(dibz)₂Py₂ and Fe(II)(sdibz)₂Py₂ were prepared by using dibenzoylmethane and monothiodibenzoylmethane instead of acetylacetone, respectively.

Optical absorption measurements were made on a Hitachi 323 automatic recording spectrophotometer. Pyridine (10%)—reagent grade benzene (90%) was used as solvent. Mössbauer spectra were obtained with a scanning velocity spectrometer in the time

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mode. The valocity scale was calibrated with metallic iron. Thermogravimetric curves were obtained with a Shimazu thermal balance.

RESULTS AND DISCUSSION

Table I is the data of elemental analysis for the complexes prepared. $\text{Fe(II)(dibz)}_2\text{-Py}_2$ and $\text{Fe(II)(sdibz)}_2\text{Py}_2$ are stable in solid, but the crystal of $\text{Fe(II)(acac)}_2\text{Py}_2$ changes color from black to red after a week. The color of the complexes in pyridine (10%)—benzene (90%) solvent changes completely after a day.

Table I. Data of Elemental Analysis

Complex	Color	C(%)		H(%)		N(%)		S(%)	
		Found	Calcd.	Found	Calcd.	Found	Calcd.	Found	Calcd.
monothiodibenzoylmethane	red	74.29	74.97	5.12	5.04			13.51	13.34
$\text{Fe(II)(acac)}_2\text{Py}_2$	black	57.38	58.25	5.85	5.87	6.63	6.80		
$\text{Fe(II)(dibz)}_2\text{Py}_2$	dark blue	72.34	72.70	4.70	4.93	4.37	4.24		
$\text{Fe(II)(sdibz)}_2\text{Py}_2$	dark green	69.53	69.36	4.77	4.66	4.34	4.05		

Table II. Mössbauer Data of Iron(II) Complexes

Complex	Temp.	I. S	Q. S.
$\text{Fe(II)(acac)}_2\text{Py}_2$	R. T.	1.08	2.40
	~110K	1.12	2.44
$\text{Fe(II)(dibz)}_2\text{Py}_2$	R. T.	1.00	2.25
	~110K	1.06	2.28
$\text{Fe(II)(sdibz)}_2\text{Py}_2$	R. T.	1.02	2.48
	~110K	1.13	2.50

unit: mm/s

The Mössbauer parameters are shown in Table II. The parameters are consistent with characteristic values observed in high-spin iron(II) complexes. In the present complexes, change of a donor atom from oxygen to sulfur does not give rise to a sharp difference in the electronic ground state of the central iron(II).

Figure 1 shows the absorption spectra and the band maxima are listed in Table III. Acetylacetonone does not exhibit any absorption band in the visible region. Therefore, the bands of $\text{Fe(II)(acac)}_2\text{Py}_2$ at 349 and 442 nm are probably classified to charge-transfer band. The absorption bands of $\text{Fe(II)(dibz)}_2\text{Py}_2$ except the band at 346 nm, which is ascribed to the $\pi\text{-}\pi^*$ transition of dibenzoylmethane, are probably assigned to charge-transfer bands as well. Akaiwa *et al.* proposed a method for the spectrophotometric determination of iron(II) extracted into benzene containing dibenzoylmethane and Lewis base.⁹⁾ The absorption spectrum of the pyridine system in the report is different from that in this study, because they employed the benzene used for the extraction of the iron(II)-free solution as a reference. They determined iron(II) spectrophotometrically at 408 nm, but the spectrum in Fig. 1 suggests that it seems good to use the charge-transfer band at 582 nm if the high sensitivity is not required. In $\text{Fe(II)(sdibz)}_2\text{-}$

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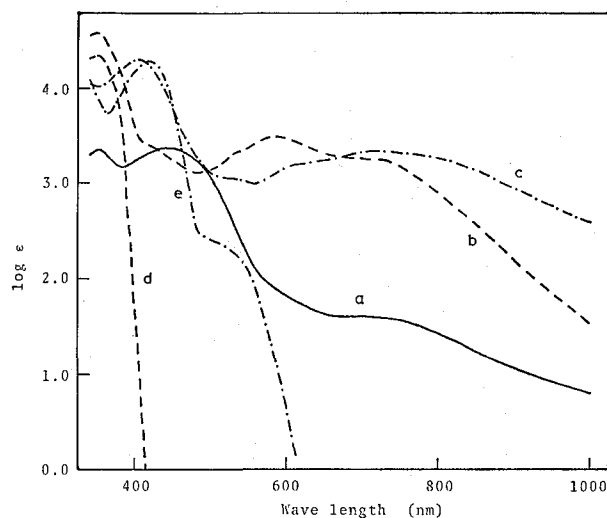


Fig. 1. Absorption spectra of iron(II)- β -diketone complexes, a: Fe(II) (acac)₂Py₂, b: Fe(II) (dibz)₂Py₂, c: Fe(II) (sdibz)₂Py₂, d: dibenzoylmethane, e: monothiodibenzoylmethane

Table III. Absorption Maxima of Iron(II) complexes

Complex	Absorption Maxima nm(ϵ)					
dibenzoylmethane	346(22, 100)					
monothiodibenzoylmethane	415(19, 100)	500 ^a (262)				
Fe(II) (acac) ₂ Py ₂	349(2, 249)	442 (2, 447)	673 (42. 4)			
Fe(II) (dibz) ₂ Py ₂	346(37, 270)	420 ^a (2, 520)	530 ^a (2, 030)	582 (3, 080)	680 (1, 900)	
Fe(II) (sd bz) ₂ Py ₂	403(20, 470)	529 (1, 120)	626 ^a (1, 600)	720(2, 160)	800 ^a (1, 830)	

a: shoulder

Py₂, the charge-transfer bands are observed at 620, 720 and 800 nm, and the π - π^* band of monothiodibenzoylmethane, at 404 nm. The absorption bands at 720 and 800 nm can be used for the photometric determination of iron(II) extracted.⁹⁾ Chikuma *et al.* and Honjo *et al.* proposed a more sensitive method for the determination of iron(II) with monothiothenoyltriuroacetone and pyridine, where iron(II) is determined at 810 nm ($\epsilon=5,590$).^{9,10)} They concluded that the species extracted is the mono-adduct with pyridine. The methods for the determination iron(II) with the β -diketone or the monothio- β -diketones are not so good as those with *o*-phenanthroline or bathophenanthroline in sensitivity. Thus, another method using trioctylphosphine oxide instead of pyridine was proposed in order to improve the sensitivity.⁸⁾

Figure 2 shows the thermogravimetric curves of Fe(II) (acac)₂Py₂ and Fe(II) (dibz)₂Py₂. 100% Weight loss in the vertical axis is consistent with the calculated decrease in weight when pyridine is perfectly released. Unstable Fe(II) (acac)₂Py₂ releases pyridine at a lower temperature than Fe(II) (dibz)₂Py₂ does. Two pyridine molecules are released at the same time in both complexes, suggesting that the mono-adducts are unstable.

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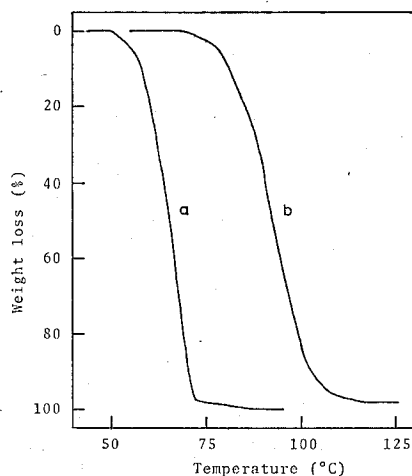


Fig. 2. Thermogravimetric curves of iron(II)- β -diketone complexes.
 a: $\text{Fe(II)(acac)}_2\text{Py}_2$, b: $\text{Fe(II)(dibz)}_2\text{Py}_2$

bauer measurements. The author is grateful to Mr. K. Fukui for supplying monothio-dibenzoylmethane.

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